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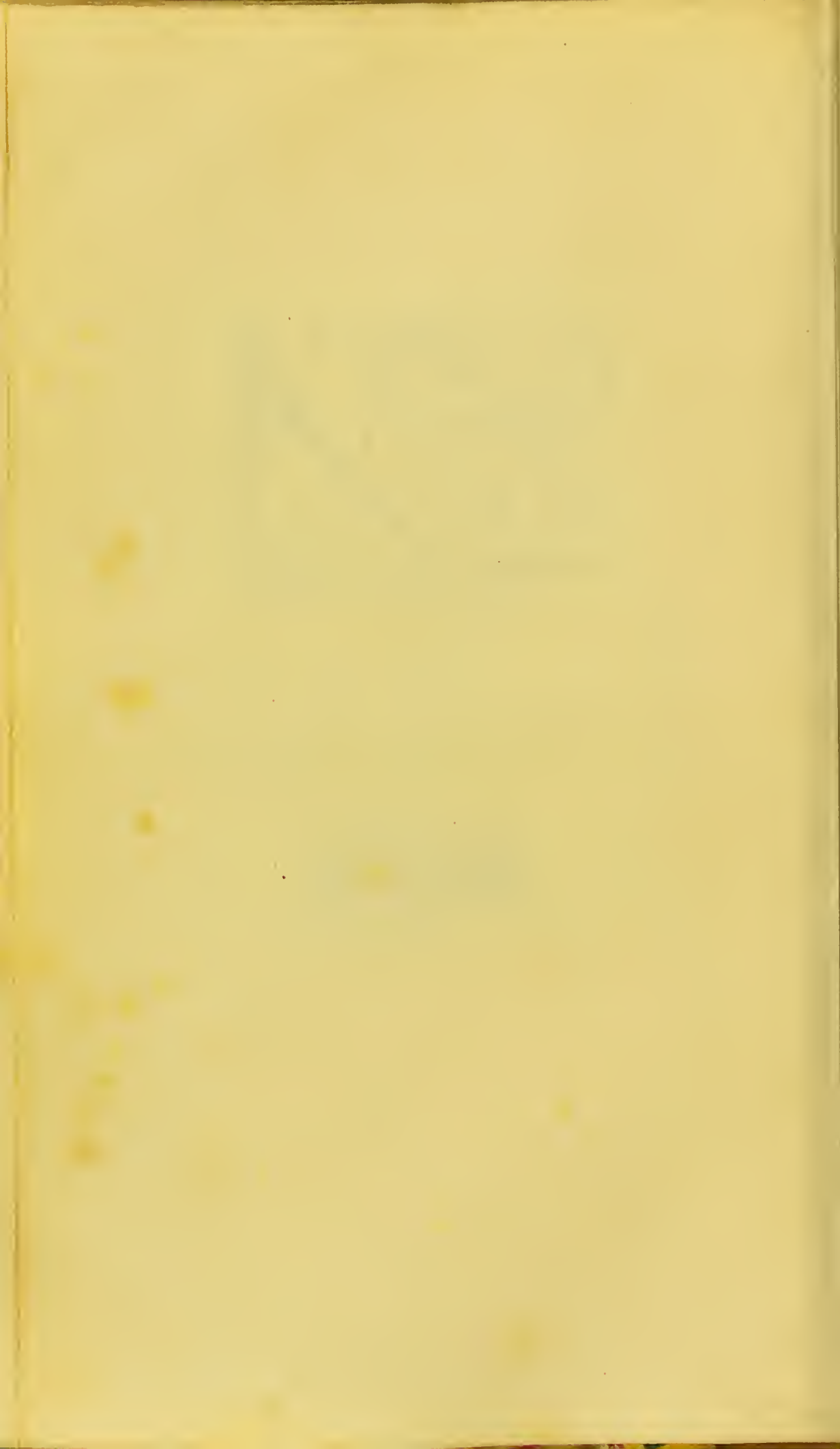
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A
SYSTEM
OF
CHEMISTRY,



IN FOUR VOLUMES.

BY THOMAS THOMSON, M. D.

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OF THE ROYAL ACADEMY OF SCIENCES OF NAPLES, OF THE MINERALOGICAL
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CONTENTS

OF

VOLUME FOURTH.

	Page
BOOK IV. Of Vegetables	1
CHAP. I. Of the Ingredients of Plants.....	2
SECT. 1. Of Acids	3
2. Of Sugar	11
3. Of Sarcocoll	31
4. Of Gum	34
5. Of Mucus	42
6. Of Jelly.....	44
7. Of Ulmin	45
8. Of Colouring Matters	47
9. Of Bitter Principles	49
10. Of Nicotin.....	55
11. Of Extractive	56
12. Of Emetin.....	61
13. Of Asparagin.....	63
14. Of Cerasin.....	64
15. Of Inulin	67
16. Of Starch	69
17. Of Indigo	79
18. Of Gluten	90
19. Of Pollenin	97
20. Of Fibrin	98
21. Of Oils	99
22. Of Wax.....	103
23. Of Camphor	109
24. Of Bird-Lime	118
25. Of Olivile	121
26. Of Resins	122
27. Of Guaiacum.....	147
28. Of Balsams	151

	Page
SECT. 29. Of Gum Resins	160
30. Of Caoutchouc	170
31. Of Cotton	176
32. Of Suber	178
33. Of Medullin.....	180
34. Of Wood or Lignin.....	ibid.
35. Of Fungin	184
36. Of the Foreign Substances found in Plants	186
37. Of Vegetable Principles in general	198
CHAP. II. Of the Parts of Plants	205
SECT. 1. Of the Sap of Plants.....	206
2. Of the Peculiar Juices	211
3. Of the Gases in Plants	216
4. Of the Wood and Roots.....	217
5. Of Barks	229
6. Of Leaves	239
7. Of Flowers	245
8. Of Pollen.....	251
9. Of Seeds	255
10. Of Fruits.....	274
11. Of Bulbs	276
12. Of Lichens	283
13. Of Mushrooms	287
14. Of Fuci	297
15. Of the Diseases of Plants	300
CHAP. III. Of Vegetation	302
SECT. 1. Of Germination	304
2. Of the Food of Plants.....	311
3. Of the Motion of the Sap.....	327
4. Of the Functions of the Leaves.....	338
5. Of the peculiar Juices of Plants	356
6. Of the Decay of Plants.....	361
CHAP. IV. Of the Decomposition of Vegetable Sub- stances.....	362
SECT. 1. Of the Vinous Fermentation.....	365
2. Of the Acetous Fermentation.....	385
3. Of Putrefaction	387
BOOK V. Of Animals	389
CHAP. I. Of Animal Substances.....	390
SECT. 1. Of Gelatin	ibid.
2. Of Albumen.....	398
3. Of Fibrin.....	408

	Page
SECT. 4. Of the Colouring Matter of Blood	412
5. Of Mucus	415
6. Of Osmazome	416
7. Of Pieromel	417
8. Of Urea or Nephriin	419
9. Of Saccharine Matter	423
10. Of Cantharadin	426
11. Of Cochenilin	427
12. Of Oils	431
13. Of Animal Resins	435
14. Of Sulphur and Phosphorus	439
15. Of Acids	ibid.
16. Of Alkalies, Earths, and Metals	443
17. Of Animal Principles in general	ibid.
CHAP. II. Of the Parts of Animals	445
SECT. 1. Of Bones, Shells, and Crusts	446
2. Of Horns, Nails, and Scales	460
3. Of the Muscles of Animals	463
4. Of the Skin	468
5. Of Membranes, Tendons, Ligaments, and Glands	473
6. Of the Brain and Nerves	474
7. Of Marrow	477
8. Of Hair and Feathers	480
9. Of Silk	483
10. Of Chyle	487
11. Of Blood	489
12. Of Milk	594
13. Of Eggs	504
14. Of Saliva	506
15. Of Bile	509
16. Of the Cerumen of the Ear	513
17. Of Tears and Mucus	514
18. Of the Liquor of the Pericardium	518
19. Of the Humours of the Eye	ibid.
20. Of Sinovia	521
21. Of Semen	524
22. Of the Liquor of the Amnios	527
23. Of Animal Poisons	532
24. Of the Air contained in the Swimming Blad- ders of Fishes	533
25. Of Sweat	536

	Page
SECT. 26. Of Urine.....	539
27. Of Feces.....	553
28. Of Morbid Concretions.....	559
29. Of Morbid Secretions	583
CHAP. III. Of the Functions of Animals	588
SECT. 1. Of Digestion	ibid.
2. Of Respiration	601
3. Of the Action of the Kidneys.....	620
4. Of Perspiration	621
5. Of Assimilation	625
CHAP. IV. Of the Decomposition of Animal Bodies....	631

APPENDIX.

Tables exhibiting the weight of the atom of all substances
hitherto accurately examined.

TABLE 1.	637
TABLE 2.	644
TABLE 3. Specific Gravity of Gases.....	648
INDEX	651

A . y

SYSTEM

OF

CHEMISTRY.

BOOK IV.

OF VEGETABLES.

VEGETABLES are too well known to require any definition. They are perhaps the most numerous class of bodies belonging to this globe of ours: the species already known amounting to no less than 30,000, and very considerable additions are daily making to the number. But it is not my intention in this place either to enumerate, to describe, or to classify plants. These tasks belong to the botanist, and have been successfully accomplished by the zeal, the singular address, and the indefatigable labour, of Linnæus and his followers.

It is the business of the chemist to analyze vegetables, to discover the substances of which they are composed, to examine the nature of these substances, to investigate the manner in which they are combined, to detect the processes by which they are formed, and to ascertain the chemical changes to which plants, after they have ceased to vegetate, are subject. Hence a chemical investigation of plants comprehends under it four particulars: 1. An account of the different *substances* found in the vegetable kingdom. 2. An account of the composition and properties of the different *parts* of vegetables; as the leaves, the bark, the seed, &c. These must be composed of combinations of the various substances described under the

Book IV. first head. 3. An account of the *vegetation* of plants, as far as it can be illustrated by chemistry. 4. An account of the *changes* which plants undergo after they cease to vegetate.

These different heads shall form the subject of the four following chapters.

CHAP. I.

OF THE INGREDIENTS OF PLANTS.

Chap. I.

THE substances hitherto found in the vegetable kingdom, all of them at least which have been examined with any degree of accuracy, may be reduced under five heads: I. Substances soluble in cold water. II. Substances insoluble in cold water, but partially soluble in hot water, at least in certain states. III. Substances which melt when heated (if not already fluid) and which burn like oils. They are insoluble in water; but most of them are soluble in alcohol. IV. Substances neither soluble in water, alcohol, nor ether, and which have a fibrous or woody texture. V. Substances which belong to the mineral kingdom, which occur only in small quantity in vegetables, and may therefore be looked upon as extraneous or foreign. The following table exhibits a view of the different vegetable substances hitherto discovered, arranged under their respective heads:

- | | |
|--------------------------|--------------------|
| I. 1. Acids. | 5. Indigo. |
| 2. Sugar. | 6. Gluten. |
| 3. Sarcocoll. | 7. Pollenin. |
| 4. Gum. | 8. Fibrin. |
| 5. Mucus. | III. 1. Fixed oil. |
| 6. Jelly. | 2. Wax. |
| 7. Ulmin. | 3. Volatile oil. |
| 8. Colouring principles. | 4. Camphor. |
| 9. Bitter principles. | 5. Birdlime. |
| 10. Nicotin. | 6. Olivile. |
| 11. Extractive. | 7. Resins. |
| 12. Emetin. | 8. Guaiacum. |
| II. 1. Asparagin. | 9. Balsams. |
| 2. Cerasin. | 10. Gum-resins. |
| 3. Inulin. | 11. Caoutchouc. |
| 4. Starch. | |

- IV. 1. Cotton.
2. Suber.
3. Medullin.
4. Fungin.
5. Lignin.

- V. 1. Acids.
2. Alkalies.
3. Earths.
4. Metals.

The properties of these different substances form the subject of the following sections:

SECT. I.

OF ACIDS.

THE combustible or vegetable acids found ready formed in plants are the following:

- | | |
|---------------|------------------|
| 1. Acetic. | 8. Meconic. |
| 2. Oxalic. | 9. Benzoic. |
| 3. Tartaric. | 10. Hydrocyanic. |
| 4. Citric. | 11. Kinic. |
| 5. Moroxylic. | 12. Gallic. |
| 6. Boletic. | 13. Tannin. |
| 7. Malic. | |

1. Acetic acid is easily detected by its peculiar odour, and ^{Acetic.} by the compounds which it forms with the alkalies and earths. It has been found by Vauquelin in the sap of different trees which he examined, and likewise in the acid juice of the *cicer arietinum*, which consists of oxalic acid, malic acid, and a little acetic acid.* Scheele obtained it from the *sambucus nigra*.†

It has been found also in the following plants:

Phœnix dactilifera.‡

Rhus typhinus.||

Gallium verum.§

2. Oxalic acid is easily detected and distinguished by the ^{Oxalic.} following properties: It decomposes all calcareous salts, and forms with lime a salt insoluble in water. It readily crystallizes. Its crystals are quadrilateral prisms. It is totally destroyed by heat.

Oxalic acid was first detected in vegetables by Scheele. It ^{Plants containing it.} exists in the state of binoxalate of potash in the leaves of the *oxalis acetosella*, *oxalis corniculata*, and different species of

* Jour. de Phys. lii. 362.

† Ann. de Chim. vi. 13.

‡ Fourcroy, as quoted by John, Chemische tabellen der Pflanzenanalyse, p. 53.

§ Schrader, Ibid.

|| John, Ibid.

Book IV.

rumex; and likewise in the *geranium acidum*. In the root of *rhubarb* Scheele found it in the state of oxalate of lime; and Deyeux, Dispan, and Vanquelin, have observed it uncombined in the liquid which exudes from the *cicer parietinum*. Scheele detected the presence of oxalate of lime in a variety of roots and barks besides *rhubarb*. He reduced the substances to be examined to powder or thin slices, digested them for two hours in diluted muriatic acid, and dropped ammonia in excess into the filtered solution. If the oxalate was present, the liquid became gradually milky, and deposited it. The following is his table of the roots and barks from which he obtained this salt:*

I. ROOTS.

Alkana.	Vincetoxicum.	Saponaria.
Apium.	Lapathum.	Scilla.
Bistorta.	Liquiritia.	Sigillum salomonis.
Carlina acaulis.	Mandragora.	Tormentilla.
Curcuma.	Ononis.	Valeriana.
Dictamnus albus.	Iris florentina.	Zedoaria.
Fæniculum.	Iris nostras.	Zingiber.
Gentiana rubra.		

II. BARKS.

Berberis.	Cassia caryophyllata.	Quassia.
Cassia fistularis.	China.	Quercus.
Canella alba.	Culilavan.	Simaruba.
Cinnamomum.	Frangula.	Lignum sanctum.
Cascarilla.	Fraxinus.	Ulmus.

Oxalic acid has been found likewise united to potash, lime, or magnesia, in

Spinacea oleracea. †
Lichnis dioica. †
Atropa belladonna. ‡

Tartaric.

3. Tartaric acid is known by the following properties: When a little potash is cautiously dropped into a solution containing it, common tartar is formed, and precipitates to the bottom. Tartaric acid does not decompose the sulphate, nitrate, or muriate of lime. Tartrate of lime is insoluble

* Crell's Annals, iii. 2. English Trans.

† Braconnot, John, Chemische tabellen der Pflanzenanalysen, p. 52.

‡ Melandre, John, ibid. p. 54.

in water. Tartaric acid crystallizes. Its crystals are long slender prisms. It is destroyed by heat. Chap. I.

Tartaric acid has been found in the following vegetable substances: Plants containing it.

Pulp of the tamarind.*
 The juice of grapes.
 Rhus typhinum.†
 Vaccinium oxycocos.‡
 Rheum raponticum.§
 Morus alba.||
 Rhus corearia.**
 Pinus sylvestris.†
 Pinus abies.†
 Pinus larix.†
 Sium sisarum.††
 Triticum repens.||
 Leontodon taraxicum.||

4. Citric acid is distinguished by the following properties: Citric.
 It does not form tartar when potash is added to it. With lime it forms a salt insoluble in water, which is decomposed by sulphuric, nitric, and muriatic acids. It readily crystallizes. It is destroyed by heat.

Citric acid has been found unmixed with other acids in the following vegetable substances: †† Plants containing it.

The juice of oranges and lemons.
 The berries of vaccinium oxycoccus, *cranberry*.
 ——— vitis idæa, *red whortleberry*.
 Prunus padus, *birdcherry*.
 Solanum dulcamara, *nightshade*.
 Rosa canina, *hip*.

It occurs mixed with other acids in many other fruits. Citrate of lime exists also in the onion.

5. Moroxylic acid has been hitherto found only in the bark of the *morus alba*, or white mulberry. Moroxylic. For its characteristic properties I refer to vol. ii. p. 152.

6. Boletic acid has been hitherto found only in the *boletus pseudo igniarius*. Boletic. For its characters I refer to vol. ii. p. 155.

* Vauquelin, Ann. de Chim. v. 92.

† John, ibid, p. 50.

‡ Bindheim, Crell's Annalen, 1787, p. 557.

§ Bindheim, ibid. 1788, ii. 42.

|| Hermbstadt.

** Trommsdorf and Gren, John, ibid. 50.

†† Margraaf, John, ibid. p. 54.

‡‡ Scheele, Crell's Journ. ii. 8. Eng. Trans.

Book IV.
Malic.

7. Malic acid is known by the following properties: it forms with lime a salt soluble in water, which is decomposed by citric acid. It does not form tartar with potash. It is incrustalizable. Heat destroys it.

Plants containing it.

Malic acid has been found by Scheele* in the fruits of the following plants, which contain no other acid:

Apples.

Berberis vulgaris, *barberry*.

Prunus domestica, *plum*.

—— *spinosa*, *sloe*.

Sambucus nigra, *elder*.

Sorbus aucuparia, *roan* or *service*.

Plants containing citric and malic acids.

In the following fruits he found nearly an equal quantity of malic and citric acids: †

Ribes grossularia, *gooseberry*.

—— *rubrum*, *currants*.

Vaccinium myrtyllus, *bleaberry*.

Crategus aria, *beam*.

Prunus cerasus, *cherry*.

Fragaria vesca, *strawberry*.

Rubus chamæmorus, *cloudberries*, *evrochs*.

—— *idæus*, *raspberry*.

Malic acid has also been found in the agave americana, ‡ and in the pulp of tamarinds. § In the first of these it is mixed with tartaric acid; in the second with tartaric and citric acids. Vauquelin has lately detected it in the state of malate of lime, in the following plants:

Sempervivum tectorum.

Sedem album, *acre*, *telephium*.

Arum maculatum.

Different species of *crassula* and *mesembrianthemum*. ||

It has been found likewise in the following plants chiefly combined with lime or potash:

Bromelia ananas.**

Delphinium elatum.

Ruta graveolens. ††

Portulaca oleracea.

* Scheele, Crell's Jour. ii. 8. Eng. Trans.

† Ibid.

‡ Hoffman of Weimar.

§ Vauquelin. Ann. de Chim. v. 92.

|| Vauquelin, Ann. de Chim. xxxv. 153.

** Adet, Ann. de Chim. xxv. 32.

†† Braconnot, Ann. de Chim. lxx. 277. To him likewise we are indebted for all the other plants contained in the table.

Eupatorium cannabinum.	Chærophyllum sylvestre.
Nicotiana rustica.	Actæa spicata.
Mirabilis Jalappa.	Brionia alba.
Tropæolum majus.	Syringa vulgaris.
Ricinus communis.	Valeriana officinalis.
Trifolium melilotus.	Spinacea oleracea.
Rubia tinctorum.	Staphylea pinnata.
Reseda luteola.	Sambucus ebulus.

8. Meconic acid has been hitherto found only in opium. ^{Meconic.} But it is not unlikely that it exists in the milky juice of the different species of *lactuca*, particularly the *virosa*. For its characters the reader is referred to Vol. ii. p. 158.

9. Benzoic acid is distinguished by its aromatic odour, and ^{Benzoic.} its volatility on the application of a very moderate heat. It has been found hitherto only in a few vegetable substances, to which the French chemists have confined the term *balsam*. These are, *benzoin*, *balsam of tolu*, *storax*, &c. It has been detected likewise in the following plants:

Origanum majorana.*	Fungi.
Salvia sclarea.†	Trifolium melilotus officinalis.
Epidendron vanilla.‡	Tonkin bean.**
Cicer arietinum.§	

10. The smell and the property of forming Prussian blue with ^{Hydrocyanic.} the oxygenized salts of iron are sufficient to characterize the hydrocyanic acid. Ever since Scheele pointed out the method of procuring this acid in a separate state, the similarity between its smell and the odour of bitter almonds, peach blossoms, and other vegetable substances, did not fail to be remarked: but it was not till about the year 1802 that Bohn observed a similarity between the taste of hydrocyanic acid and the water distilled from bitter almonds; and that when this water was mixed with potash, it acquired the property of precipitating iron solutions blue. These results were confirmed soon after by Schræder, who succeeded also in obtaining prussian-blue, by the same means, from the distilled waters of *laurocerasus* leaves and peach blossoms. When lime

* Götting, as quoted by John, Tabellen, p. 52. But he thinks that Götting mistook camphor for benzoic acid.

† Braconnot, Ann. de Chim. lxx. 285.

‡ John, Tabellen, p. 53.

§ Proust, Ann. de Chim. xxxvi. 273, and xlix. 181.

|| Ibid.

** Vogel, Annals of Philosophy. xvi. 227.

Book IV. is added to the concentrated waters of these substances, there is formed a prussiate of lime; for when the liquid is mixed with a salt of iron, prussian blue precipitates the instant that another acid is added. When these waters are distilled off pure potash, a prussiate of potash remains, and even crystallizes. The water that comes over contains no hydrocyanic acid. When distilled with ammonia, no prussian ley is obtained.*

This discovery was soon after verified by the experiments of Gehlen, Bucholz, and Vauquelin. Bucholz found that hydrocyanic acid was contained in the essential oil of bitter almonds, and that it could be separated by means of fixed alkalies, but not by ammonia.† Vauquelin obtained hydrocyanic acid by distilling water off the kernels of apricots in a water bath. The liquid, when mixed with sulphate of iron, and a little ammonia added, yields a precipitate only partially soluble in sulphuric acid. The residuum is green, but becomes blue when boiled. Hence Vauquelin supposes that the hydrocyanic acid in these bodies is in a peculiar state.‡ Schræder has since obtained hydrocyanic acid from the flowers of the sloe (*prunus spinosa*,) and the leaves of the bay-leaved willow (*salix pentandra*.)§ It cannot be doubted that almost all the bitter-tasted kernels contain this acid.

Kinic. 11. Kinic acid has been hitherto found only in the bark of the *cinchona officinalis*, where it occurs combined with lime. For its characters, see vol. ii. p. 179.

Gallic. 12. Gallic acid is known by the following properties: with the red oxide of iron it produces a black colour. It is crystallizable. Heat destroys it. It has been found in a great number of plants, chiefly in the bark. The following table, drawn up by Mr. Biggin,|| will serve to show the relative proportions of this acid in different plants:

Plants containing it.	Elm	7	Willow (trunk)	9
	Oak cut in winter	8	Sycamore	6
	Horse chesnut	6	Birch	4
	Beech	7	Cherry-tree	8
	Willow (boughs).	8	Sallow	8
	Elder	4	Mountain ash	3
	Plum-tree	8	Poplar	8

* Jour. de Phys. lvi. 224.

† Ann. de Chim. xlv. 206.

‡ Nicolson's Jour. iii. 394.

† Gehlen's Jour. i. 78.

§ Gehlen's Jour. i. 394.

Hazel	9	Oak cut in spring	10	} Chap. I.
Ash	10	Huntingdon or Leicester	} 10	
Spanish chesnut	10	willow		
Smooth oak	10	Sumach	14	

13. Tannin, like most other vegetable substances, seems to be susceptible of different modifications. The following are the different species of tannin which have been hitherto noticed. Species of tannin.

(1). Tannin from nutgalls. This is the common species described in this Work under the name of tannin. It precipitates iron black, and forms a firm insoluble brown precipitate with glue. The bark of oak, and most other astringent trees in this country, are supposed at present to contain this species of tannin. From nutgalls,

(2). The tannin which constitutes so large a proportion of catechu forms the second species. Its peculiar nature was first observed by Proust. It was afterwards more particularly examined by Davy. It forms with ~~an~~ iron an olive-coloured precipitate. Catechu,

(3). The tannin of *kino* constitutes a third species. This substance is obtained from different vegetables. It was originally imported, as is supposed from Africa; but at present the common kino of the shops is, according to Dr. Duncan, an extract from the *coccoloba urifera*, or *sea-side grape*, and is brought chiefly from Jamaica. But the finest kino is the product of different species of *eucalyptus*, particularly the *resinifera* or brown gum tree of Botany Bay.* It is an astringent substance of a dark red colour, and very brittle. It dissolves better in alcohol than water. The solution in the latter liquid is muddy; in the former transparent, and a fine crimson if sufficiently diluted. It throws down gelatine of a rose-colour, and forms with salts of iron a deep-green precipitate, not altered by exposure to the air. These properties, first observed by Dr. Duncan, and afterwards by Vauquelin, show that it contains abundance of tan, but in a particular state.† Kino,

(4). The fourth variety of tannin is contained in *sumach*. This is a powder obtained by drying and grinding the shoots of the *rhus coriaria*; a shrub cultivated in the southern parts of Europe. The tan, which it contains in abundance, yields a Sumach,

* Nicolson's Jour. vi. 234.

† Ann. de Chim. xlvi. 321.

Book IV. precipitate with gelatine, which subsides very slowly, and remains in the state of a white magma without consistence.

Fustic. (5). The fifth variety, according to Proust, is to be found in the wood of the *morus tinctoria*, or *old fustic* as the British dyers term it. This wood gives out an extract both to alcohol and water, which yields a precipitate with gelatine. A solution of common salt is sufficient to throw it down.*

Tannin affects particularly the bark of trees; but it exists also in the sap and in the wood of a considerable number, and even in the leaves of many. It is very seldom that it exudes spontaneously; yet this seems to be the case with a variety of kino from Botany Bay, which I have examined.†

It has been ascertained by Mr. Biggin,‡ that when the barks of trees are examined at different seasons they vary in the quantity of tan. The quantity varies also with the age and size of the trees.§ The greatest proportion of tan is contained in the inner bark.|| The epidermis usually contains none.

The following table exhibits the proportion of solid matter extracted by water from different vegetable substances, and the quantity of tan contained in that solid matter, as ascertained by the experiments of Davy.

Table of the quantity of tannin in plants.	One ounce of	Solid Matter.		Tan.
		Grains.	Grains.	Grains.
White inner bark of old oak	108	72
 young oak	111	77
 Spanish chesnut	89	63
 Leicester willow	117	79
Coloured or middle bark of oak	43	19
 Spanish chesnut	41	14
 Leicester willow	34	16
Entire bark of oak	61	29
 Spanish chesnut	35	21
 Leicester willow	71	33
 elm	—	13
 common willow	—	11
Sicilian sumach	165	78
Malaga sumach	156	79
Sonchong tea	—	48
Green tea	—	41

* Proust, Ann. de Chim. xlii. 94.

† Phil. Tran. 1798, p. 299.

|| Davy, *ibid.*

† Annals of Philosophy, i. 163.

§ Davy, *ibid.* 1803.

One ounce of	Solid Matter, Grains.	Tan. Grains.	Chap. I.
Bombay catechu.....	— 261	}
Bengal catechu	— 231	
Nutgalls	180 127	

SECT. II.

OF SUGAR.

SUGAR, which at present forms so important an article in ^{History.} our food, seems to have been known at a very early period to the inhabitants of India and China; but Europe probably owes its acquaintance with it to the conquests of Alexander the Great. The word *saccharon* occurs in Pliny* and Dioscorides.† They describe it as a white brittle solid, which exuded spontaneously from a species of reed. From the description, it does not appear that the sugar known to the ancients was similar to our sugar; but their account is so imperfect that little can be collected from it.‡ For ages after its introduction into the West, it was used only as a medicine; but its consumption gradually increased; and during the time of the Crusades, the Venetians, who brought it from the East, and distributed it to the northern parts of Europe, carried on a lucrative commerce with sugar. It was not till after the discovery of America, and the extensive cultivation of sugar in the West Indies that its use in Europe, as an article of food, became general.§

Sugar was formerly manufactured in the southern parts of Europe; but at present almost the whole of our sugar is produced in the East and West Indies. The plant from which it is produced is the *arundo saccharifera* or *sugar cane*. Other plants indeed contain it; but not in such abundance. In North America, however, it is extracted from the *acer saccharinum* or sugar maple, but in too small quantity for exportation. Attempts have been lately made to extract it from the beet.

* Lib. xii. c. 8.

† De Medica Materia, lib. ii. c. 104, or, in some copies, c. 74.

‡ See Beckmann's *Historia Sacchari*. *Commentationes Soc. Reg. Gottingensis*, 1782, vol. v. p. 56.

§ See Falconer's *Sketch of the History of Sugar*, *Manchester Memoirs*, iv. 291; and Mozeley's *History of Sugar*.

Book IV.
 Manufacture of sugar
 in Indostan.

1. The method of making sūgar practised in Indostan is exceedingly simple, and requires little or no expensive apparatus. The soil chosen is a rich vegetable mould, in such a situation that it can be easily watered from the river. About the end of May, when the soil is reduced to the state of soft mud, either by rain or artificial watering, slips of the cane, containing one or two joints, are planted in rows about four feet from row to row, and eighteen inches asunder in the rows. When they have grown to the height of two or three inches, the earth round them is loosened. In August small trenches are cut through the field to drain off the rain, if the season prove too rainy; and to water the plants, if the season prove too dry. From three to six canes spring from each slip set. When they are about three feet high, the lower leaves of each cane are carefully wrapped round it; and then the whole belonging to each slip are tied to a strong bamboo eight or ten feet high, and stuck into the earth in the middle of them. They are cut in January and February, about nine months after the time of planting. They have now reached the height of eight or ten feet, and the naked cane is from an inch to an inch and a quarter in diameter. They have not flowered. When this happens, the juice loses much of its sweetness. The newly cut canes are put through the rollers of a mill, and their juice collected into large iron boilers, where it is boiled down smartly to a proper consistence, the scum being carelessly taken off. The fire is then withdrawn, and the liquid by cooling becomes thick. It is then stirred about with sticks till it begins to take the form of sugar, when it is put in mats made of the leaves of the palmira tree (*borassus flabelliformis*), and the stirring continued till it is cold.* This process yields a *raw* or *powdered* sugar; but it is clammy, and apt to attract moisture from the atmosphere, because the acids in the juice have not been removed. By the addition of quicklime to the juice in the proportion of about three spoonfuls to every 14 gallons, the sugar loses this property. The impure sugar prepared by this method is called *jagary*. Every three quarts of juice, or every six pounds, yield about one pound of sugar. From an acre of ground about 5000 pounds of sugar, and consequently about 30,000 pounds of juice are obtained.

* See Dr. Roxburgh's Account of the Hindoo method of cultivating the Sugar Cane, Repertory, ii. 425, Second Series; and Tennant's Indian Recreations, ii. 31:

2. In the West India islands the raising of sugar is much more expensive, and the produce much less, owing to the high price of labour; or, which is the same thing, to the nature of the labourers, and to the inferiority of the soil. The juice extracted by passing the cane twice between iron rollers is received in a leaden bed, and thence conducted into a *receiver*. Here it cannot be allowed to stand above 20 minutes without beginning to ferment. Therefore, as soon as collected, it is run into a flat copper cauldron, called a *clarifier*, capable of holding 400 gallons or more. Here it is mixed with a quantity of lime. The maximum used is a pint of lime to every hundred gallons; but in general much less will serve. Fire is immediately applied, and the juice heated to the temperature of 140°. The fire is then extinguished. A thick viscid scum forms upon the top, which remains unbroken, and the clear liquid is drawn off from under it by a cock or syphon, and introduced into a large copper boiler. Here it is boiled briskly; the scum, as it forms, being continually removed by large scummers. When the bulk of the liquid is sufficiently diminished, it is introduced into a second boiler, and the boiling and scumming continued as before; lime-water being sometimes added, if the liquid does not appear clear. From the second copper it passes into a third; and from that to a fourth, where the very same process is continued. This last copper is called the *teache*. When by continued evaporation it has become sufficiently concentrated, which is judged of by its viscosity, it is withdrawn from the fire, and introduced into the *cooler*. This is a wooden vessel about 11 inches deep, seven feet long, and from five to six wide. As it cools it crystallizes or *grains*, forming an irregular mass separated from the uncrystallized syrup or *mélasses*. From the cooler it is taken and put into empty hogsheads, having holes in the bottom; through each hole the stalk of a plantain leaf is thrust, long enough to rise above the top of the hogshead. Through these holes the *mélasses* drain into a receiver. In about three weeks the sugar becomes tolerably dry and fair, and the process is finished. The sugar in this state is sent to Britain, under the name of *muscovado* or *raw sugar*.* A gallon of the raw juice yields at an average about a pound of raw sugar.†

In the French West India islands the process is somewhat

* Edwards's History of the West Indies, ii. 222.

† Ibid. p. 224.

Book IV. different. From the cooler the inspissated juice is poured into conical-shaped pots with the point undermost. As soon as it has become solid, a plug is withdrawn from the point of the cone, and the melasses suffered to drain. When no more will drop, a stratum of clay moistened with water is laid upon the top of the sugar. The water from this, slowly penetrating the sugar, carries down the remains of the melasses. Sugar treated in this way is called *clayed* sugar. The process is said to have been discovered by accident. A hen having her feet dirty had gone over a pot of sugar, and the sugar under her tread was found whiter than elsewhere. By claying, the sugar loses about one-third of its weight; but a considerable portion of the loss may be obtained by boiling down the juice that runs off. The process is not considered as so profitable in the British islands as the one which is usually followed.*

In North
America.

3. In North America the farmers procure sugar for their own use by a still simpler process, from the sap of the *acer saccharinum*, or sugar maple tree, which abounds in the woods. It reaches maturity in about 20 years, and is then from two to three feet in diameter. In February, March, and April, the tree is bored with an auger to the depth of about $\frac{3}{4}$ inch, and in an ascending direction. The hole is then deepened to two inches. A wooden spout is introduced into the hole to direct the flow. The sap flows from four to six weeks. When it ceases on the south side, the north side is bored. This process does not injure the tree; but on the contrary improves it. An ordinary tree yields in good seasons from 20 to 30 gallons of sap; from which are made from five to six pounds of sugar; or every forty pounds of sap yields about a pound of sugar; so that it is not one-sixth so rich as the East India sugar cane.

The sap ought never to be kept longer than twenty-four hours after it is procured from the tree. It is improved by straining through a cloth. It is put into large flat kettles, mixed usually with quicklime, white of egg, and new milk. A spoonful of slacked lime, the white of one egg, and a pint of new milk, is sufficient for fifteen gallons of sap. A little butter is added to prevent the sap from boiling over. When boiled down sufficiently, it is allowed to *grain*, or form into

* Edwards's History of the West Indies, ii. 232.

small crystals, which constitute raw sugar, and then purified in the usual manner.* Chap. I.

4. The raw sugar imported into Europe is still farther purified. It is dissolved in water, mixed with lime, clarified by means of bullock's blood, boiled down to a proper consistency, scumming off the impurities as they rise to the top, and then poured into unglazed conical earthen vessels, where it is allowed to grain. The point of the cone is undermost, and perforated to allow the impurities to separate. The base of the cone is covered with moist clay; the water of which gradually filters through the sugar, and displaces a quantity of impure liquid. The sugar thus purified is called *loaf sugar*. When redissolved and treated in the same way a second time, it is called *refined sugar*. This process appears to have been first practised by the Venetians. Refining of sugar.

5. In the East Indies, where they make very pure sugar, their process is remarkably simple and economical, but tedious. An account of this process, as practised about Aska and Barampour, has been given by Mr. Anderson, surgeon of the Madras establishment. The cane juice is first strained, and then boiled with a little lime to the proper consistency for graining. It is afterwards put into small earthen pots with mouths six inches wide, and allowed to remain at least a month: sometimes six or eight months. A hole is then made in the bottom of the pots, and the liquid syrup suffered to run off from the crystallized sugar. The sugar is put upon a cloth and squeezed, being occasionally moistened with water, the better to remove the remains of the syrup. It is then dissolved in water, and boiled a second time to the proper consistency; milk and some water being added from time to time in order to clarify it, and the scum being removed as it rises. It is poured into small wide-mouthed pots, suffered to crystallize, and a hole bored in the bottom of the pots to allow the remains of the syrup to run off. To whiten the sugar, the tops of the pots are covered with the leaves of the *creeping vine*; and these are renewed every day for five or six days. When pure *sugar candy* is wanted, the sugar thus obtained is dissolved in water again, and the same process of boiling with milk and scumming is repeated. When poured into pots, thin slices of bamboo are introduced, which prevents it from running into lumps, and induces it to form large crystals.† How refined in Indostan.

* See Dr. Rush's Account of the Sugar Maple Tree. Amer. Trans. iii. 64.

† Phil. Mag. xxi. 272.

Book IV.
Theory of
sugar-making.

6. According to Dr. Higgins, who went to the West Indies on purpose to examine the manufacture of sugar, the juice of the sugar cane contains a great number of bodies; but the most important of them (not reckoning sugar and water) are those which he calls *herbaceous matter* and *melasses acid*. Concerning the nature of these substances nothing very precise is known; but the great object of the manufacturer is to remove them, as they impede the crystallizing of the sugar. The *herbaceous matter** is partly held in solution by the water, partly by the carbonic acid. When the liquid is heated to 140°, the carbonic acid is disengaged, and the herbaceous matter separates in flakes of an olive-green colour. This produces the scum. Lime facilitates the separation, partly by abstracting the carbonic acid, and partly by forming an insoluble compound with the herbaceous matter. Hence its use in the first part of the process. It is called *temper* by the manufacturers. Too much is injurious in the first part of the process. It redissolves the herbaceous matter, or at least deepens the colour of the syrup.†

This nature of the *melasses acid* is not better known than that of the herbaceous matter; but it combines with the sugar, and forms an uncrystallizable syrup. Lime is necessary in the last part of the process to remove this acid; it combines with it, and the compound runs off with the syrup during the graining of the sugar.‡ This acid is no doubt that which exists in unripe canes, and the quantity of it of course diminishes as the cane improves in quality. I have been informed by my friend Dr. Clarke of Dominica, that the quantity of lime necessary is much less when the canes are rich and ripe than when green or lodged, or when they have been cut some time before they are bruised in the mill: and that he has known the juice of ratoon rich canes boiled down into permanent crystals of sugar without the assistance of any lime. He substituted potash instead of lime; but found that the sugar obtained by this method was liable to deliquesce. Soda made the grain firmer, but smaller than usual. These facts would lead to a suspicion that the *melasses acid* is either

* This is probably a different substance from any that we know at present. It seems to possess many properties in common with *gluten*. Dr. Higgins says that it agrees in its properties with the dregs of refined indigo.

† See Higgins' Observations, Phil. Mag. xxiv. 308.

‡ Ibid.

the *acetic* or the *citric*. Dr. Clarke's experiments should exclude the oxalic, tartaric, and malic acids.

These observations of Dr. Clarke agree very well with the account of Dutrone, who made a very great number of experiments on sugar cane juice, and was very well acquainted with its properties. According to him it chiefly consists of sugar, extractive, and two different kinds of fecula, which are precipitated by lime. He endeavours to prove that unaltered juice contains no acid, and his experiments are sufficient to convince us that the quantity of acid is at least very minute.*

Dutrone's experiments were performed before vegetable chemistry had made nearly the progress which it has now reached. Hence it is difficult, from his description, to recognise the substances which he found in sugar-cane juice. But this loss has been supplied by the experiments of Proust. Living in Spain where the sugar-cane is still cultivated, he had an opportunity of examining its juice. He found it to contain gluten or green fecula, gum, extractive, malic acid, sulphate of lime, and two species of sugar; namely, common sugar, and a species which is incapable of crystallizing, and can only be exhibited in the state of syrup, and which, for that reason, he calls *liquid sugar*. The quantity of acid was very minute, and probably in a warmer climate than Spain its quantity would be still smaller.†

According to Dutrone, sugar-canes give about half their weight of juice, the specific gravity of which varies from 1.033 to 1.106. It is at first opaque, but when the feculent matter has precipitated, it becomes transparent, and of a green colour. When allowed to stand for some time it runs into acidity, but if the feculent matter be removed it undergoes the vinous fermentation.

7. Sugar procured pure by the preceding methods has a Properties. very strong sweet taste, but no smell. Its colour is white; and when crystallized it is somewhat transparent. It has often a considerable degree of hardness; but it is always so brittle that it can be reduced without difficulty to a very fine powder. When two pieces of sugar are rubbed against each other in the dark, a strong phosphorescence is visible.

Sugar is not altered by exposure to the atmosphere, excepting only that in damp air it absorbs a little moisture.

* Histoire de la Canne. Premiere Partie, chap. vi. and vii.

† Ann. de Chim. lvii. 148.

Book IV.
Solubility
in water.

8. It is exceedingly soluble in water. At the temperature of 48° water, according to Mr. Wenzel, dissolves its own weight of sugar.* The solvent power of water increases with its temperature; when nearly at the boiling point, it is capable of dissolving any quantity of sugar whatever. Water thus saturated with sugar is known by the name of *syrup*.

Syrup.

Syrup is thick, ropy, and very adhesive; when spread thin upon paper, it soon dries, and forms a kind of varnish, which is easily removed by water. Its specific caloric, according to the experiments of Dr. Crawford, is 1.086. When syrup is sufficiently concentrated, the sugar which it contains may be obtained in crystals. The syrup is boiled down till it is thick, poured immediately into pans, and kept in a stove or room, heated so as not to be supported by animals for any length of time. The crystals concrete upon small threads placed in the pans for that purpose.† The primitive form of these crystals is a four-sided prism, whose base is a rhomb, the length of which is to its breadth as 10 to 7, and whose height is a mean proportion between the length and breadth of the base. The crystals are usually four or six-sided prisms, terminated by two-sided, and sometimes by three-sided summits.‡ From the experiments of Berzelius they are composed of

Crystals.

Real sugar	100
Water	5.6
	105.6 §

Specific
gravity.

9. The specific gravity of very white sugar, according to the experiments of Fahrenheit, is 1.6065; || according to Hassenfratz, it is 1.4045.** I have found it 1.5629.

10. Sugar, as far as is known, is not acted upon by oxygen gas. The effect of the simple combustibles on it has not been tried; but it does not appear to be great. Azotic gas nor the metals have no sensible action on it.

11. From the recent experiments of chemists it appears that sugar is capable of combining with and neutralizing both acids and alkaline bodies.

Action of
acids on.

Sulphuric acid and muriatic acid act upon it with consi-

* Verwandtschaft, p. 308.

† Lewis, Neuman's Chemistry, p. 329.

‡ Gillot, Ann. de Chim. xviii. 317.

§ Annals of Philosophy, v. 264.

|| Phil. Trans. 1724, vol. xxxiii. p. 114.

** Ann. de Chim. xxviii. 15.

derable energy, charcoal is evolved, and much of the sugar destroyed, the remainder loses the power of crystallizing.*

Nitric acid dissolves it with an effervescence, occasioned by the evolution of nitrous gas, and converts it into malic and oxalic acids. Four hundred and eighty grains of sugar, treated with six ounces of nitric acid, diluted with its own weight of water, and cautiously heated, separating the crystals as they formed, yielded 280 grains of oxalic acid. So that 100 parts of sugar yield by this treatment 58 parts of oxalic acid. † When liquid chlorine is poured upon sugar in powder, it is dissolved, and immediately converted into malic acid; and the chlorine is converted into common muriatic acid. ‡

Oxalic, acetic, and tartaric acid, deprive sugar of the property of crystallizing, or at least the crystals are soft and ill defined. §

12. When lime is added to a solution of sugar in water, and the mixture boiled for some time, a combination takes place. The liquid still indeed retains its sweet taste; but it has acquired also a bitter and astringent one. A little alcohol added to the solution produced a precipitate in white flakes, which appeared to be a compound of sugar and lime. Sulphuric acid precipitated the lime in the state of sulphate, and restored the original taste of the sugar. When the compound of sugar and lime was evaporated to dryness, a semitransparent tenacious syrup remained, which had a rough bitter taste, with a certain degree of sweetness. ||

It would appear from the observations of Mr. Daniell, that lime has a tendency to convert sugar into gum. When raw sugar from the West Indies is kept for some time in this country, it is apt to assume a clamminess, with the soft feel of flour. In this state it is called *weak* sugar. Mr. Daniell is of opinion that this change is occasioned by the action of the lime (always present in raw sugar) upon the sugar.**

From the experiments of Mr. William Ramsay, it appears that sugar facilitates and increases the solubility of lime and strontian, and forms a combination with them. But barytes seems to act with more energy, and to occasion a partial de-

* Vogel, Ann. de Chim. lxxi. 93.

† Cruickshanks, Rollo on Diabetes, p. 460.

‡ Chenevix, Ann. de Chim. xxviii. 193.

§ Vogel, Ann. de Chim. lxxi. 95.

|| Cruickshanks, Rollo on Diabetes, p. 460.

** Quarterly Jour. vi. 32.

Book IV. composition of the sugar. For on endeavouring to dissolve it in syrup it was constantly converted into a carbonate, and very little in consequence dissolved.*

Of alkalies. The fixed alkalies combine with sugar, and form compounds not unlike those which have been just described. Potash destroys the sweet taste of syrup more completely than lime; but when it is neutralized by sulphuric acid, and the sulphate precipitated by alcohol, the sweet taste is completely restored. When alcohol is agitated with the compound of sugar and potash dissolved in water, it refuses to unite with it, but swims on the top in a state of purity.†

When a solution of sugar is digested on oxide of lead, the oxide is gradually dissolved, but after a certain interval of time a light white powder makes its appearance. This powder is a compound of sugar and oxide of lead. It is white, light, tasteless, and insoluble in water. According to Berzelius, to whom we are indebted for our knowledge of it, ‡ its constituents are as follows:

Sugar	41·74	100
Oxide of lead	58·26	139·6
	<hr style="width: 10%; margin: 0 auto;"/>		
			100·00

Of alcohol. 13. Sugar is soluble in alcohol, but not in so large a proportion as in water. According to Wenzel, four parts of boiling alcohol dissolve one of sugar.§ But this proportion is surely too great. Lewis could only dissolve one part of sugar in 12 of boiling rectified spirits, and Margraff in 16 parts. When the solution is set aside for a few days, the sugar separates in elegant crystals.|| Sugar unites readily with oils, and renders them miscible with water. A moderate quantity of it prevents, or at least retards, the coagulation of milk; but Scheele discovered that a very large quantity of sugar causes milk to coagulate.**

Of sulphurets, &c. 14. The hydrosulphurets, sulphurets, and phosphurets of alkalies and alkaline earths, seem to have the property of decomposing sugar, and of bringing it to a state not very dif-

* Nicholson's Journal, xviii. 9.

† Cruickshanks, Rollo on Diabetes, p. 460.

‡ Annals of Philosophy, v. 263.

§ Verwandtschaft, p. 305.

|| Lewis, Neuman's Chemistry, p. 329. Margraff, Opusc. i. 217

** Scheele, ii. 32. Dijon Trans.

ferent from that of gum. Mr. Cruickshanks introduced a quantity of syrup into a jar standing over mercury, and then added about an equal quantity of phosphuret of lime. Phosphuretted hydrogen gas was immediately extricated. In eight days the syrup was withdrawn: it had lost its sweet taste, and acquired a bitter and astringent one.* From this solution alcohol threw down white flakes, very much resembling those of mucilage separated from water by the same liquid.† A little sugar was dissolved in alcohol, and phosphuret of lime added to it. No apparent action took place. The mixture, after standing in the open air for some days, was evaporated and water added. No gas was disengaged, as the phosphuret had been converted into a phosphate. The liquid being filtered and evaporated, a tenacious substance remained, much resembling gum arabic. Its taste was bitter, with a slight degree of sweetness. It did not seem soluble in alcohol. It burned like gum.‡

Similar experiments were made by this ingenious chemist with the sulphurets. The sweet taste of the sugar was destroyed; but on account of the solubility of the different products, the nature of the change could not be ascertained.

15. When heat is applied to sugar it melts, swells, becomes brownish-black, emits air bubbles, and exhales a peculiar smell, known in French by the name of *caromel*. At a red heat it instantly bursts into flames with a kind of explosion. The colour of the flame is white with blue edges.

Action of heat.

When sugar is distilled in a retort, there comes over a fluid which, at first, scarcely differs from pure water; by and by it is mixed with what was formerly called pyromucous acid, and is now known to be a compound of oil and acetic acid; § afterwards some empyreumatic oil makes its appearance, and a bulky charcoal remains in the retort. This charcoal very frequently contains lime, because lime is used in refining sugar; but if the sugar, before being submitted to distillation, be dissolved in water, and made to crystallize by evaporation in a temperature scarcely higher than that of the atmosphere, no lime whatever, nor any thing else, except pure charcoal,

* This is the taste of phosphuret of lime.

† Rollo on Diabetes, p. 452.

‡ Cruickshanks, *ibid.*

§ Schrickel, in his dissertation *De Salibus Saccharinis*, published in 1776, endeavoured to show that pyro-mucous acid was a mixture of vinegar, oxalic and tartaric acids. Fourcroy and Vauquelin have proved more lately that it is merely acetic acid united to a little oil.

Book IV.

will be found in the retort. During the distillation, there comes over a considerable quantity of carbonic acid and carburetted hydrogen gas.* Sugar therefore is decomposed by the action of heat; and the following compounds are formed from it: Water, acetic acid, oil, charcoal, carbonic acid, carburetted hydrogen gas. The quantity of oil in a separate state is inconsiderable; by far the most abundant product is pyro-mucous acid. Sugar indeed is very readily converted into pyro-mucous acid; for it makes its appearance always whenever syrup is raised to the boiling temperature.

We are indebted to Mr. Cruickshanks for the most precise set of experiments on the decomposition of sugar by heat. Four hundred and eighty grains of pure sugar were introduced into a coated retort, and heated gradually to redness. The products were,

Pyro-mucous acid with a drop or two of oil ..	270 grains.
Charcoal	120
Carburetted hydrogen, and carbonic acid gas	90
	480

The pyromucous acid required about 75 grains of a solution of potash to saturate it; and when thus neutralized, no ammonia was disengaged. Hence sugar contains no azote, unless we suppose a very minute portion to be present in the pyromucous acid; and even this is not likely. The charcoal burns away without leaving any residue. Hence sugar contains no earth nor fixed alkali. The proportion of the gaseous products was 119 ounce measures of carburetted hydrogen, and 41 ounce measures of carbonic acid gas.†

Constitu-
ents,

These experiments are sufficient to show us, that sugar is composed entirely of oxygen, carbon, and hydrogen. It is of course a vegetable oxide. Lavoisier endeavoured to determine its constituents experimentally; but at that time pneumatic chemistry had made too little progress to permit him to approach very near the truth. More lately Gay-Lussac and Thenard have analysed sugar by burning a determinate quantity of it with chlorate of potash, and ascertaining the proportion of carbonic acid gas formed.‡ Berzelius has repeated this analysis with some modifications, and with his accustomed

* Scopoli and Morveau, Encyc. Meth. Chim. i. 269.

† Rollo on Diabetes, p. 452.

‡ Recherches Physico-chimiques, ii. 288.

precision.* The following table exhibits the results of these experiments:

	Gay-Lussac and Thenard.		Berzelius.	
Oxygen	50.63	.. 51.47	.. 49.015	.. 49.083
Carbon	42.47	.. 41.48	.. 44.200	.. 44.115
Hydrogen ..	6.90	.. 7.05	.. 6.785	.. 6.802
	<hr/>	<hr/>	<hr/>	<hr/>
	100.00	100.00	100.000	100.000

If we consider the compound of sugar and oxide of lead formed by Berzelius as composed of one atom sugar and one atom oxide of lead, then the weight of an atom of sugar will be 10.03; for 139.6 : 100 :: 14 : 10.03. The proportions of constituents which will agree best with this weight and with the preceding analyses are

Five atoms oxygen	= 5 49.4
Six atoms carbon	= 4.5 44.5
Five atoms hydrogen ..	= 0.625 6.1
	<hr/>	<hr/>
Atom of sugar	= 10.125	100.0

If Berzelius' saccharate of lead be a compound of 2 atoms sugar and 1 atom oxide of lead, then the weight of an atom of sugar will be only 5. But as this weight could not be so nearly reconciled to the analyses above stated, I am disposed to consider the saccharate of lead as a compound of 1 atom of sugar and 1 atom of oxide of lead. On that supposition the weight of an atom of sugar must be 10.125.

16. From the experiments of different chemists, especially Species. of Proust and Goettling, it appears that there are different species of sugar found ready prepared in the vegetable kingdom, distinguished from each other by the figure of their crystals, and other variations in their properties. The species hitherto examined are six in number; namely, common sugar, liquid sugar, sugar of figs, sugar of grapes and of starch, Botany Bay sugar, and manna.

Common sugar is the substance which has been described Common in the preceding part of this Section. It is obtained from the sugar cane. As far as is known at present, there is no difference between the sugar of the maple and common sugar. The sugar of beet is also the same as common sugar.

17. Liquid sugar was first pointed out by Proust. He has Liquid.

Book IV.

shown that it exists in a variety of fruits and vegetable juices. It is distinguished from every other species of sugar by being incapable of crystallizing. It can only be exhibited in a liquid state. It is transparent and colourless when pure, and is more soluble in alcohol than common sugar. By means of that liquid it may be separated from common sugar when they happen to be mixed. It exists in the sugar cane juice, and constitutes, according to Proust, a considerable portion of the molasses. It exists also in grapes, peaches, apples, and other fruits.* From the experiments of Auarie it appears, that a liquid sugar may also be obtained from the stalks of the *Zea Mays*, or Indian corn; but no method tried was capable of inducing it to crystallize.† From the experiments of Vogel and Bouillon Lagrange there is reason to suspect that liquid sugar may be merely common sugar deprived of the power of crystallizing by being combined with an acid.

Sugar of figs.

18. The sugar of figs may be seen in a concrete state upon the outside of that fruit in the state in which it is usually sold in the shops in this country. If we dissolve it in boiling alcohol and set the liquid aside, we may easily obtain it in the state of crystals. These crystals have a form different from that of common sugar; nor do they appear to my taste quite so sweet. Hence they must constitute a distinct species.

Sugar of grapes.

19. That grapes contain abundance of sugar has been long known. The Duc de Bullion first extracted it from the juice of grapes, and Proust pointed out the difference between it and common sugar. The juice of grapes, according to him, yielded from 30 to 40 per cent. of this sugar.‡ The sugar of grapes is not so white as common sugar, but it crystallizes much more readily.§

Proust has published a long dissertation on the properties of this sugar, and the method of extracting it from grapes. It was of great importance on the continent during the last war, on account of the difficulty of obtaining sugar from the West Indies.

Verjuice.

Verjuice, or the liquid obtained from unripe grapes, contains tartar, sulphate of potash, sulphate of lime, much citric acid, a little malic acid, extractive, and water; but neither gum nor sugar. As the grapes advance to maturity, the citric

* Proust, Ann. de Chim. lvii. 131.

† Ann. de Chim. lx. 61.

‡ Jour. de Phys. xxix. 5, and lvi. 113.

§ Nicholson's Jour. xiv. 178.

acid gradually disappears, and gum and sugar appear in its place. Chap. I.

The juice of ripe grapes contains also gluten and fibrous matter, merely in a state of mixture, and therefore separable by the filter, or still better by boiling and scumming the liquid. The substances held in solution are chiefly sugar, syrup, gluten, gum, and extractive. When this juice is evaporated to dryness, it yields from a third to a fifth of solid matter, according to the species of grape employed, and the season of the year. Juice of ripe grapes.

To extract the sugar from this juice, Proust saturated the acids which it contains with potash, boiled it down to a half, and left it at rest. By this means several of the salts subsided. Its specific gravity was 1.215. It was then mixed with blood, heated, scummed, filtered, and boiled down to a syrup. It gradually becomes crystallized, and resembles the raw sugar from the West Indies. In this state its specific gravity is about 1.500. This raw sugar, according to Proust, is composed of Sugar of grapes.

Crystallizable sugar	75.00
Syrup, or uncrystallizable sugar . . .	24.44
Gum	0.31
Malate of lime	0.25
	100.00

Besides some extractive, the quantity of which cannot well be ascertained. The syrup holds in solution a considerable quantity, probably more than half its weight, of crystallizable sugar; but it is difficult to separate it.

The raw sugar thus obtained is not so sweet as that from the sugar cane, since four parts of the latter will go as far as five parts of the former. But it may be applied to all the purposes of common raw sugar.

This raw sugar may be refined precisely in the same way as that of the sugar-cane. It is then white, but inferior in consistence to common sugar. It is not so sweet, and has a striking resemblance to the sugar of honey. It does not crystallize, but assumes the form of sphericles. It is not so soluble as the sugar of canes, and is therefore more easily separated from the other substances in the juice of grapes.

Proust informs us that the raw sugar from grapes, when diluted sufficiently with water, ferments and is converted into wine.

Starch may be converted into a sugar possessing exactly the Starch sugar.

Book IV. properties of sugar of grapes, by mixing it with about four times its weight of water and about one hundredth part of its weight of sulphuric acid, boiling the mixture for thirty-six hours, supplying water as fast as it evaporates; then saturating the acid with lime, separating the sulphate of lime and concentrating the liquid by sufficient evaporation.

This curious fact was accidentally discovered by Kirchoff, a Russian chemist, as he was employed in a set of experiments to convert starch into gum. He conceived that the starch would be rendered soluble in water by boiling it with very dilute sulphuric acid, and by prolonging the boiling he gradually observed the conversion of the starch into sugar. Vogel ascertained that during the conversion of starch to sugar no gas whatever is extricated. Mr. Moore* and M. de Saussure† ascertained that the quantity of sulphuric acid was not diminished by the process. Saussure ascertained that 100 parts of starch when converted into sugar become 110·14 parts. Hence he drew as a conclusion that starch sugar is merely a compound of starch and water in the solid state.‡ According to his analysis the constituents of starch sugar are as follows:

Oxygen	55·87
Carbon	37·29
Hydrogen	6·84
		100·00

As we have no data for determining the equivalent number for starch sugar, it is impossible to know of how many atoms of oxygen, carbon, and hydrogen it is composed. Were we to consider it as composed of five atoms of each of its ingredients, its composition and weight would be as follows:

5 Atoms oxygen = 5 53·53
5 Atoms carbon = 3·75 40·00
5 Atoms hydrogen = 0·625 6·67
		100·00
Atoms of starch sugar	.. = 9·375	100·00

These proportions do not differ much from the result of the analysis of Saussure. Were we to suppose the above result correct, starch sugar would differ from common sugar by containing one atom of carbon less than it does.

* Phil. Mag. xl. 134.

† Annals of Philosophy, vi. 424.

‡ Annals of Phil. vi. 426.

Saussure's analysis of sugar of grapes approaches very nearly to the preceding analysis. He found the composition of sugar of grapes as follows:

Chap. I.

Oxygen	56.51
Carbon	36.71
Hydrogen	6.78
		100.00

This species of sugar melts at the temperature of boiling water; when dissolved in water and sufficiently diluted, it undergoes the vinous fermentations without its being necessary to add any yeast. Its crystals are small spheres, exactly similar to the crystals of honey. I believe that these spheres are composed of a congeries of small needles; though it is not always easy to observe this structure.

20. Some years ago specimens of a sweet substance were brought from Botany Bay. They were snow-white, in the form of tears, and had obviously dropped in a liquid state from some vegetable. I was informed that these tears were collected in a plain covered with wood; but of what species I could not learn. Some bushels of it might have been collected. These tears had a sweet and agreeable taste. They dissolved in much greater quantity in alcohol than common sugar, and when the alcohol cooled it deposited the sugar abundantly in needle-form crystals. The form of these crystals approached that of manna; but was not quite the same. Nor did it make so cooling an impression on the tongue. It was therefore a peculiar species of sugar; though approaching much nearer to manna than any of the preceding species. It is very much to be wished that more complete information could be obtained respecting this species of sugar, and that sufficient quantities of it could be sent over to this country to permit an accurate set of experiments on it to be made. I only had a very few grains of it in my possession.

Botany Bay
sugar.

21. M. Braconnot obtained from the juice of the *agaricus volvaceus*, by evaporation, a species of sugar, which differs from every other hitherto known. It crystallizes in four-sided prisms with square bases. It has such a disposition to crystallize, that when a very weak aqueous solution of it is put upon the surface of a vessel it is immediately sprinkled with small acicular crystals. When heated this sugar melts, swells and takes fire, giving out the odour of caramel. There re-

Mushroom
sugar.

Book IV. mains a small quantity of charcoal, which is destitute of alkali. Acids do not deprive this substance of the power of crystallizing as they do common sugar. When digested with nitric acid it produces abundance of oxalic acid; but no bitter principle. It is capable of undergoing the vinous fermentation.*

Manna. 22. Manna was long considered as a substance which fell from the heavens, till incontestable experiments demonstrated it to be an exudation from trees. It has the form of oblong globules or masses, of a yellowish-white colour, and some degree of transparency. It is the produce of various trees, but is chiefly procured from the *fraxinus ornus*, a species of ash, which grows abundantly in Sicily and Calabria. It partly exudes spontaneously during the summer months, and is partly obtained by incisions. The juice gradually concretes into a solid mass, or it is dried in the sun or in stoves.†

Pure manna is very light, and appears to consist of a congeries of fine capillary crystals. Its taste is sweet, and it leaves a nauseous impression on the tongue. It acts as a mild purgative. When dissolved in water it may be obtained again unaltered by a gentle evaporation. Alcohol dissolves it when assisted by heat; and the solution, when set aside, gradually deposits about $\frac{2}{5}$ ths of the manna, in the state of a fine white light spongy crystalline mass, bearing some resemblance to camphor. This deposit has an agreeable sweet taste, and instantly melts upon the tongue like snow in warm water. This may be considered as pure manna. It differs from sugar in the nature of its crystals, and in its more rapid solubility. By evaporating the solution, and setting it aside repeatedly, about $\frac{3}{4}$ ths more of the manna is deposited, but not so fine in colour or taste as the first precipitate. By evaporation to dryness, the remaining 8th is obtained in the state of a thick extract, which cannot easily be reduced to dryness. This may be considered as consisting chiefly of the foreign bodies to which manna owes its nauseous taste.‡ Manna differs from common sugar in several remarkable particulars. It dissolves very readily and abundantly in alcohol, and crystallizes on cooling. When digested in nitric acid it yields both oxalic and saccharic acids, whereas sugar only yields oxalic acid.§ It does not fer-

* Ann. de Chim. lxxix. 278.

† See Neuman's Chemistry, p. 325, from which all the accounts of manna to be found in chemical books have been copied.

‡ Neuman, p. 326.

§ Proust, Ann. de Chim. lviii. 144.

ment like sugar, and of course does not seem capable of furnishing alcohol.* In a set of experiments on the juice of the common onion (*allium cœpe*) Fourcroy and Vauquelin found that at a temperature, between 66° and 80°, it gradually underwent the acetous fermentation without emitting any gas; and that by this process a quantity of uncrystallizable sugar which it contained assumed most of the properties of manna. It was not precisely the same, however, with manna, for it did not yield sacclactic acid when treated with nitric acid.

The common manna of the shops, according to the experiments of Fourcroy and Vauquelin, consists of four different ingredients. 1. Pure manna, which constitutes at least $\frac{3}{4}$ ths of the whole. 2. A little common sugar, which makes it fermentable to a small extent. 3. A yellow matter with a nauseous odour, to which the purgative quality of manna seems owing. 4. A little mucilage convertible into sacclactic acid. This last ingredient seems hypothetical. Several substances by fermenting seem to be converted into manna. The sugar in onion juice has been already mentioned. Fourcroy and Vauquelin found likewise that fermented melon juice contained manna, though none could be detected in it before the fermentation. Manna appears sometimes to be formed and deposited by insects.†

23. The plants containing sugar are very numerous. Margraff first pointed out a method of separating it from them. The plant suspected to contain it is reduced to powder or pulp, and boiled with strong alcohol. The liquid is filtered while hot, and set aside in a close vessel. In a few days the sugar separates from the alcohol and crystallizes.‡

Plants containing sugar.

The following are the chief plants from which it has been actually extracted by chemists: §

- The flower of the rhododendron ponticum,
- The sap of the acer saccharinum,
- betula alba,
- aselepias syriaca,
- heracium sphondilium,
- cocos meifera,
- juglans alba,

* Du Puytren and Thenard, Ann. de Chim. lix. 51.

† See Klaproth, Gehlen's Jour. iv. 328.

‡ Margraff's Opusc. i. 216.

§ See Gren's Handbuch, ii. 123.

The sap of the agave *Americana*,*
 ————— *ficus saccharinus*,
 ————— *ficus carica*,
 ————— *ceratonia siliqua*, †
 The juice of *arundo saccharifera*,
 ————— *arundo bambos*, ‡
 ————— *zea mays*,
 The roots of *pastinaca sativa*, §
 ————— *sium sisarum*, §
 ————— *beta vulgaris* and *cicla*, §
 ————— *daucus carota*, §
 ————— *apium petroselinum*,
 The bulb of the *allium cœpe*.

It is proper to observe, however, that from the *daucus carota*, Margraff could obtain only an uncrystallizable syrup. The sugar from the sap of the agave *Americana* bore a greater resemblance to manna than to sugar. || It is very seldom that sugar exudes spontaneously from vegetables; sometimes, however, it does. Tears of a sweet substance were observed upon the *ceratonia siliqua*, or locust tree, some time ago in Naples. These tears were examined by Klaproth, and found to be sugar mixed with a little tannin and oxalate of potash.** The inspissated juice of the bamboo (*arundo bambos*) is known in India by the name of *sacar nambu*; a term which is supposed to be the origin of our word *sugar*, and constitutes a species of sugar celebrated as a medicine. How far it agrees with common sugar has not been ascertained.

Small crystals of sugar are found occasionally in the flower of the *rhododendron ponticum*. I have received some of these crystals from my friend Dr. Charles Mackenzie, but too small in quantity to admit of a rigid examination. They have no regular shape; but in other respects seem to agree with common sugar in their properties. The same sugar has been noticed by Fourcroy, Vauquelin, and Bosc. ††

The list of the saccharine plants would be greatly extended were we to add all sweet-tasted fruits, such as grapes, &c.

* Margraff, Opusc. i. 213.

† Klaproth, Gehlen's Jour. iv. 326.

‡ Tennant's Indian Recreation, ii. 238.

§ Margraff, Opusc. i. 213.

|| Margraff, i. 241.

** Gehlen's Jour. iv. 327.

†† Ann. de Chim. lxiii. 102.

which obviously contain sugar, and some of the mushroom tribe, from which Humboldt affirms he extracted it.* Chap. I.

24. Sugar has now become an essential part of the food of ^{Uses.} Europeans. It contains perhaps a greater proportion of nourishment than any other vegetable substance in the same bulk. It has the advantage of most other articles of food, in not being liable to be injured by time nor by the weather. If we believe Dr. Rush, the plentiful use of it is one of the best preventatives of the diseases occasioned by worms. It has been long supposed to have a tendency to injure the teeth; but this prejudice is now given up. It has the property of preserving other vegetable substances from putrefaction; and accordingly is often employed for that purpose, constituting the base of conserves, &c.

SECT. III.

OF SARCOCOLL.

THIS substance has hitherto been confounded with the gum resins, though its properties are very different. It is supposed to be the produce of the *peneta mucronata*; a shrub which is said by botanical writers to be indigenous in the north-eastern parts of Africa.† Nothing precise is known concerning the way in which it exudes.

Sarcocoll comes to this country from Persia, Turkey, and ^{State.} India, and is usually sold in the state of oblong globules, from the size of a pea to that of a particle of sand. Its colour is usually yellow; and it has the semitransparency and much of the appearance of gum arabic. But some of the grains are reddish-brown. Its smell is peculiar, and not unlike that of aniseed. When carefully examined, four different substances may be detected: The first, and by far the most abundant, is pure sarcocoll: The second consists of small woody fibres, and a soft yellowish-white substance, not unlike the covering of the seeds of some of the cruciform plants: The third is a reddish-brown substance apparently earthy: And the fourth

* The reader will find a much fuller list than I have given, in John's Tabellen, p. 11. I have not ventured to make use of his materials, because I have no means of determining the particular species of sugar which each vegetable contains.

† Dioscorides informs us that sarcocoll exudes from a tree which grows in Persia, and nothing farther respecting it is known at present.

Book IV. is only detected when the sarcocoll is dissolved in water or alcohol. It then appears in soft transparent tremulous masses like jelly.

Properties. The pure sarcocoll amounts to 0·8 of the whole. When the sarcocoll is dissolved in alcohol or water, and obtained again by evaporation, it loses its smell. It then assumes the form of semitransparent brittle-brown cakes very like gum. Its specific gravity is 1·2684.*

This substance has a sweet taste, but leaves an impression of bitterness. In the mouth it dissolves like gum.

It is almost equally soluble in water and alcohol; the solution is yellow. The watery solution has the appearance of mucilage, and may be used for the same purposes.

Cannot be made to crystallize. When heated softens, but does not melt. It emits a slight smell of caromel. When strongly heated, it blackens, and assumes the consistence of tar, emitting a white heavy smoke having an acrid odour. In a strong fire it scarcely leaves any residuum.

To determine the properties of sarcocoll more in detail, I dissolved 50 grains of it in five ounce measures of water. The solution had a light brown colour, and tasted very like a decoction of liquorice. It did not feel in the least glutinous, and could not be used like gum to paste pieces of paper together, or to stiffen linen.

The effects of reagents on it are as follows:

Infusion of nutgalls	A slight precipitate.
Tincture of nutgalls	{ A copious light-yellow precipitate. Liquor becomes colourless.
Gallic acid	No change.
Hydrocyanate of potash	No change.
Nitric acid	{ A white precipitate and some effervescence.
Sulphuric acid	Becomes dark coloured.
Potash	Becomes green.
Carbonate of ammonia	Becomes green.
Lime water	Becomes green.
Strontian water	Becomes green.
Silicated potash	{ Becomes green. No precipitate.
Oxalate of ammonia	Becomes slightly opal.

* Brisson.

Sulphate of alumina	No change.	} Chap. I.
Muriate of magnesia	No change.	
Sulphate of iron	} Colour becomes darker and more opaque. A precipitate falls after an interval of some hours.	
Nitrate of lead		No change.
Acetate of lead	Becomes milky.	
Subacetate of lead	A copious white precipitate.	
Corrosive sublimate	No change.	
Nitrate of silver	No change.	
Sulphate of zinc	No change.	
Sulphate of copper	No change.	
Muriate of gold	A slight precipitate.	

The most remarkable property pointed out by this table is that of being precipitated by tannin. This distinguishes sarcocoll from gum and mucilage.

Liquorice seems to approach in its nature to sarcocoll. Liquorice.
This substance is obtained from the root of the *glycyrrhiza glabra*; a plant cultivated in the south of Europe, and even in Britain. The roots are long, slender, and fibrous, of a yellow-colour, and when fresh very juicy. The juice of these roots is expressed and inspissated by boiling. The substance thus obtained is *liquorice* or *black sugar*. It comes to this country from Spain in cylindrical rolls covered with hay leaves. It is afterwards purified by a subsequent process, and sold in small cylinders about the size of a goose quill, under the name of *refined liquorice*. It is of a fine black colour and glossy. It is brittle, and its fracture is glassy.

Besides the sweet matter of liquorice, there is present in it about a third part of mucilaginous matter,* and even some charcoal. When dissolved in nitric acid, a quantity of tannin is formed, owing probably to this charcoal; for neither the saccharine nor gummy substances yield this substance to nitric acid.† When treated with sulphuric acid, it leaves a quantity of charcoal, amounting to $\frac{1}{4}$ th of its weight. During the treatment of this substance with sulphuric acid, no tannin was evolved in Mr. Hatchett's experiments.‡ We are indebted to Robiquet for an interesting analysis of liquorice.§

* Neuman, p. 425.

+ Hatchett, *ibid.*

‡ Third Series of Experiments on Artificial Tannin, *Phil. Trans.* 1806.

§ *Ann. de Chim.* lxxii. 143.

Book IV.

The saccharine matter which it contains approaches very closely to sarcocoll. It has a brown colour, does not crystallize, is sparingly soluble in cold, but very soluble in hot water. It dissolves readily in alcohol, and does not appear susceptible of fermentation.

SECT. IV.

OF GUM.

THERE is a thick transparent tasteless fluid which sometimes exudes from certain species of trees. It is very adhesive, and gradually hardens without losing its transparency; but easily softens again when moistened with water. This exudation is known by the name of *gum*. The gum most commonly used is that which exudes from different species of the *mimosa*, particularly the *nilotica*.* It is known by the name of *gum arabic*. Gum likewise exudes abundantly from the *prunus avium*, or common black cherry tree of this country.

Properties.

Gum is usually obtained in small pieces like tears, moderately hard, and somewhat brittle while cold, so that it can be reduced by pounding to a fine powder. When pure it is colourless, but it has commonly a yellowish tinge, and it is not destitute of lustre. It has no smell. Its taste is insipid. Its specific gravity varies from 1.3161 to 1.4317.

Mucilage.

1. Gum undergoes no change from being exposed to the atmosphere; but the light of the sun makes it assume a white colour. Water dissolves it in large quantities. The solution which is known by the name of *mucilage*,† is thick and adhe-

* Schousbae, Phil. Mag. v. 241.

† Hermbstadt uses this word in a different sense. He makes a distinction between gum and mucilage. The solution of gum in water is transparent and glutinous, and can be drawn out into threads; whereas that of mucilage is opaque, does not feel glutinous, but slippery, and cannot be drawn into threads. Gum may be separated from mucilage by the following process:

Let the gum, which is supposed to be mixed with mucilage, previously reduced to a dry mass, be dissolved in as small a quantity of water as possible, and into the solution drop at intervals diluted sulphuric acid. The mucilage coagulates, while the gum remains dissolved. When no more coagulation takes place, let the mixture remain at rest for some time, and the mucilage will precipitate to the bottom, and assume the consistence of jelly. Decant off the liquid part, and evaporate the mucilage to dryness by a gentle heat till it acquires the consistence of horn. Med. and Phys. Jour. iii. 370.

sive: it is often used as a paste, and to give stiffness and lustre to linen. When spread out thin it soon dries, and has the appearance of a varnish; but it readily attracts moisture, and becomes glutinous. Water washes it away entirely. When mucilage is evaporated the gum is obtained unaltered. This mucilaginous solution may be kept for years without undergoing putrefaction. Scarcely any vegetable substance is less liable to decomposition. At last, however, the odour of acetic acid becomes perceptible in it.

When gum is exposed to heat it softens and swells, but it does not melt; it emits air bubbles, blackens, and at last, when nearly reduced to charcoal, it emits a low blue flame. This flame appears sooner if a flaming substance be held just above the gum. After the gum is consumed, there remains a small quantity of white ashes, composed chiefly of the carbonates of lime and potash.

2. It does not appear that gum is acted upon by oxygen gas. A solution of gum in water, when exposed to the air, soon becomes mouldy on the surface, but undergoes no farther change for a long time. The action of the simple combustibles on gum has scarcely been examined. Azotic gas seems to have no action whatever.

Gum does not act upon metals; but it has the property of combining with several of the metallic oxides, and forming compounds; at least some of the metallic salts occasion precipitates when dropped into solutions of gum. The most curious effect is that produced by the permuriate of iron. When this salt, concentrated, is dropped into a very strong mucilage, the whole becomes a brown semitransparent jelly, which is not readily dissolved by water. When dried, the jelly becomes lighter coloured, and assumes nearly the appearance of gum.

Action of
metallic
salts.

Its taste is that of gum mixed with iron.

The following table exhibits the effects of different metallic salts on the solution of gum in water, as far as my experiments go. The mucilage I used was composed of eight parts of water to one part of gum. It was transparent and quite fluid, though a little thready.

SALINE SOLUTIONS.	EFFECTS.
1. Muriate of gold	No change.
2. Nitrate of silver	No precipitate nor change.
3. Corrosive sublimate	No change.
4. Supersulph. of mercury	No change.
5. Nitrate of mercury	A white coagulum: it disappears by agitation, but returns when the solution is much diluted.
6. Prussiate of mercury	Becomes slightly opal, but no precipitate.
7. Nitrate of copper	No change.
8. Sulphate of copper	No change.
9. Muriate of copper	No change; yet the muriate is precipitated by water.
10. Cuprate of ammonia	No change.
11. Sulphate of iron	No change.
12. Permuriate of iron	Little change; except the mucilage be more concentrated.
13. Pernitrate of iron	No change.
14. Muriate of tin	No change.
15. Permuriate of tin	No change.*
16. Smoking liquor of Libavius	No change.
17. Nitrate of lead	No change.
18. Subacetate of lead	A copious coagulum.
19. Acetate of lead	No change.
20. Sulphate of zinc	No change.
21. Nitrate of zinc	No change.
22. Muriate of arsenic	No change.
23. Oxide of arsenic	No change.
24. Tartar emetic	Becomes yellow, but no precipitate.

* When the permuriate is approaching the gelatinous state, a flaky precipitate appears; but the same precipitate is occasioned by pure water.

SALINE SOLUTIONS.	EFFECTS.
25. Nitrate of bismuth	No precipitate; yet the salt is precipitated by pure water.

These phenomena indicate an affinity between gum and the peroxides of mercury and iron. Copper, antimony, and bismuth, seem also to be acted on by it; for it prevents water from precipitating them in the state of subsalts.

My experiments on the action of the alkalies and earths on mucilage of the strength indicated above, are contained in the following table: Action of alkalies and earths.

SOLUTIONS.	EFFECTS.
Potash	No precipitate.
Ammonia	No precipitate.
Lime water	No precipitate.
Barytes water	No precipitate.
Strontian water	No precipitate.
Alum	No precipitate.
Sulphate of magnesia	No precipitate.
Silicated potash	A white flaky precipitate, though very much diluted. The liquid remains transparent.
Aluminated potash	No precipitate.

From this table it appears that silica alone forms with gum an insoluble precipitate. It is by far the most delicate test of gum which I have yet met with.

Liquid potash first converts gum into a substance not unlike curd, and then dissolves it. The solution is of a light amber colour, and transparent. When long kept, the gum again falls in the state of curd. Alcohol throws down the gum in white flakes still soluble in water; but it retains the potash obstinately, and is much more friable than before. Lime-water and ammonia likewise dissolve gum, and it may be afterwards separated little altered.

Charcoal powder, when mixed with a solution of gum in water, gives a black colour, which cannot be removed by filtration, unless a very great proportion of the powder be added. In that case the water passes clear; but the whole of the gum Of charcoal.

Book IV. is retained by the charcoal. Mr. Lowitz found, that not less than 30 lbs. of charcoal powder must be mixed with water containing an ounce of gum dissolved in it before the water is entirely deprived of the gum.*

of acids. The vegetable acids dissolve gum without alteration; the strong acids decompose it.

When thrown into sulphuric acid it blackens, and soon is resolved into other substances. The charcoal formed, according to Hatchett's experiments, amounts to 0.29, and some traces of artificial tannin may also be detected.† Water and acetic acid are said likewise to be formed.‡

When gum is dissolved in strong muriatic acid, a brown solution is obtained, which becomes perfectly transparent when diluted with water, while at the same time some charry matter falls. If the solution be now saturated with ammonia, evaporated to dryness, and the residue digested in alcohol, the alcohol assumes a deep brown colour, and dissolves the whole except a very little sal-ammoniac. The gum now bears some resemblance to sugar in its properties; at least when heated it melts, and gives out a very strong smell of caromel.

Chlorine converts gum into citric acid, according to the experiments of Vauquelin. He passed a current of chlorine gas through a diluted solution of gum in water. In a few days almost the whole of the gum was acidified; and he detected citric acid by the formation of supercitrate of lime, soluble in water, and decomposable by oxalic acid.§ If nitric acid be slightly heated upon gum till it has dissolved it, and till a little nitrous gas is exhaled, the solution on cooling deposits sac-lactic acid. Malic acid is formed at the same time; and if the heat be continued, the gum is at last changed into oxalic acid. Thus no less than three acids are developed by the action of nitric acid on gum.|| We are indebted to Mr. Cruikshanks for the most precise experiments on the quantity of oxalic acid obtainable from gum by nitric acid. By digesting 480 grains of it with six ounces of nitric acid, he obtained 210 grains of oxalic acid and six grains of oxalate of lime.** No

* Crell's Annals, ii. 167. Eng. Trans.

† Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

‡ Fourcroy, vii. 197. Engl. Trans.

§ Ann. de Chim. vi. 178. These properties alone were not sufficient to prove the presence of citric acid, since malic acid possessed them both.

|| Fourcroy, vii. 148.

** Rollo on Diabes, p. 452.

artificial tannin is formed during this process.* According to the experiments of Fourcroy and Vauquelin, the quantity of sacclactic acid yielded by gum, when treated with nitric acid, varies from 0·14 to 0·26.†

Chap. I.

Gum is insoluble in alcohol. When alcohol is poured into mucilage, the gum immediately precipitates; because the affinity between water and alcohol is greater than that between water and gum. The gum in this case is in the state of soft opaque white flakes. Neither is gum soluble in ether. It is not soluble in oils; but when triturated with a little oil it renders the oil miscible with water.

Or alcohol.

The action of the hydrosulphurets, sulphurets, phosphurets, and of most of the salts on gum, has not been examined with any attention.

Gum and sugar readily unite together by dissolving both in water. By gentle evaporation a perfectly transparent solid substance is obtained, which does not crystallize. When treated with alcohol it becomes white, opaque, and soft. The greater part of the sugar is dissolved, and the gum remains united to a small portion. It has a sweetish taste, and very much resembles in appearance the substance of which the nests of wasps are formed.

Combines with sugar.

3. When gum is distilled in a retort, the products are water impregnated with a considerable quantity of pyromucous acid or *acetic acid* combined with oil, a little empyreumatic oil, carbonic acid gas, and carburetted hydrogen gas. When the pyromucous acid obtained by this process is saturated with lime, a quantity of ammonia is disengaged, with which that acid had been combined. The charcoal which remained in the retort leaves behind it, after incineration, a little lime and phosphate of lime. Mr. Cruickshanks, to whom we are indebted for these facts, gradually heated 480 grains of gum arabic to redness in a coated glass retort. The products were,

Destructive distillation.

Pyromucous acid mixed with some oil	210 gr.
Charcoal	96
Lime and a little phosphate of lime	10
Carburetted hydrogen and carbonic acid gas . .	164
	<hr/>
	480

* Hatchett, Additional Experiments on Artificial Tannin, Phil. Trans. 1805.

† Fourcroy, vii. 199. Engl. Trans.

Book IV.

The pyromucous acid liquid contained less than what was obtained from an equal weight of sugar in the proportion of 118 to 150. The gases consisted of 93 ounce measures of carbonic acid and 180 of carburetted hydrogen, composed of 5 parts charcoal to 1 of hydrogen. When the pyromucous acid was saturated with lime, ammonia was disengaged.*

From the experiments of Vauquelin, it appears that gum also contains traces of iron. A hundred parts of gum arabic left after incineration three parts of white ashes. These were composed chiefly of carbonate of lime; but contained also some phosphate of lime and iron, without any traces of an alkali or of sulphur. He conjectures, that in the gum the lime is usually combined with acetic acid, and perhaps sometimes with malic acid.†

Composi-
tion.

It is only of late that any notions approaching to precision respecting the composition of gum have been acquired. Gay-Lussac and Thenard analysed it, by burning it along with chlorate of potash, and found it composed of

Oxygen	50·84
Carbon	42·23
Hydrogen	6·93
	100·00 †

Berzelius analysed it by a process of the same kind but somewhat modified. He found its constituents as follows:

Oxygen	51·306
Carbon	41·906
Hydrogen	6·788
	100·000 §

To determine the equivalent number for the weight of an atom of gum, Berzelius mixed caustic ammonia with a boiling solution of gum, and then mixed the liquid with a solution of subnitrate of lead, likewise boiling hot. A precipitate fell, which when washed and dried was a compound of gum and oxide of lead. Its constituents, according to the analysis of Berzelius, were,

* Rollo on Diabetes, p. 452.

† Ann. de Chim. liv. 312.

‡ Recherches Physico-chimiques, ii. 290.

§ Annals of Philosophy, v. 270.

Gum	61·75	100
Oxide of lead	38·25	62·105
	<hr/>		
	100·00	*	

Chap. I.

If we suppose this compound to consist of an atom of gum united to an atom of oxide of lead, the weight of an atom of gum will be 22·542. But if the gummate contain 2 atoms of gum united to 1 atom of oxide of lead, (which may be the case,) then an atom of gum will weigh 11·271. I am rather disposed to adopt this last number for the weight of an atom of gum; because it seems to approach very nearly to sugar in its properties, and seems capable of being converted into sugar with great facility in the processes of vegetation. On that account it seems likely that the composition of gum and sugar do not differ very much from each other. But if the weight of an atom of gum be 11·250, (which differs immaterially from the preceding number) its composition, as indicated by the preceding analysis, must be as follows:

Six atoms oxygen = 6	53·33
Six atoms carbon = 4·5	40·00
Six atoms hydrogen .. = 0·75	6·67
	<hr/>	
	11·25	100·00

Sugar then, it would seem, differs from gum merely in containing an atom of water less. Hence an atom of sugar + an atom of water constitute an atom of gum.

4. The species of gum at present known are considerable in number; and it is likely that a more rigid examination of the vegetable kingdom will discover a greater number. The most remarkable are *gum arabic*, *gum senegal*, *gum kuteera*.

Gum arabic exudes from the *mimosa nilotica*, and other species of mimosa. It is the species described in the preceding part of this section.

5. Gum senegal, brought from the island of that name in the coast of Africa, commonly supplies the place of gum arabic in the shops, and is the sort chiefly employed by the calico printers. It is generally in larger masses than the arabic, and its colour is darker,† but in other respects its properties are the same, excepting that it does not go so far in thickening water as gum arabic.

* Annals of Philosophy, v. 270.

† Lewis, Neuman's Chem. p. 282.

Book IV.
Gum kuteera.

6. Gum kuteera is the produce of the *sterculia urens*,* a tree which grows in Indostan; and having a considerable resemblance to gum tragacanth, great quantities were imported into this country, but did not answer as a substitute. The only account of this gum with which I am acquainted was given by Mr. Cowie, in one of the volumes of the Society of Arts.† It is in loose wrinkled drops or pieces, without smell or taste, and mostly transparent. In water it slowly forms a pulp or jelly like gum tragacanth; but if pounded well in a mortar, and then boiled in water for fifteen minutes with constant agitation, it is completely dissolved. A tea spoonful of its powder gives to water the consistence of capillaire. In India it enters into the composition of some varnishes; it is used in calico-printing, and is one of the ingredients of a famous medicine for horses among them.‡

Plants containing it.

7. It has been remarked by Mr. Barrow, and probably also by others, that all the plants which yield gum have an astringent bark.§ Almost all the trees known to yield gum have been enumerated in the preceding part of this section.

Uses.

8. Gum is a nutritive food, though seldom employed for that purpose, except when in the state of mucilage, and though from M. Magendie's experiments it does not seem capable of supporting life by itself. It is used frequently as a paste, and to give stiffness or lustre to linen. The calico-printers use it in great quantities to give their colours such a degree of consistency as prevents them from running upon the cloth. It forms an ingredient in ink for a similar reason. In medicine it forms the base of many mixtures.

SECT. V.

OF MUCUS.

THE substances to which I give the name of *mucus* have been hitherto considered as varieties of gum. But the experiments on them by Dr. Bostock have shown that their properties differ so much from those of gum, as to entitle them to a separate place as peculiar vegetable principles. They

* Roxburgh, Nicholson's Journal, xxvii. 70.

† See Nicholson's Jour. vii. 301. ‡ Cowie, Nicholson's Jour. vii. 301.

§ Travels in South Africa.

are extremely numerous, existing in the roots, leaves, and seeds of a great variety of plants. They seldom or never separate spontaneously, but may be obtained artificially in a state of tolerable purity. Only a few species have been hitherto examined, so that we are uncertain how far their properties agree, and only include them all under one head from analogy.

Chap. 1.

Linseed yields one of the purest. Dr. Bostock obtained it by infusing the seeds in ten times their weight of water. A fluid is obtained of the consistence of the white of an egg, which has the adhesive qualities of mucilage of gum arabic. Its taste and appearance resemble those of that liquid. When mixed with alcohol the mucus is precipitated in white flocks, but the liquid does not become opaque and milky like mucilage of gum arabic when mixed with alcohol. Acetate of lead throws down a copious dense precipitate. Super-acetate of lead, and permuriate of tin render the liquid opaque, and also throw down a precipitate. The nitrate of mercury occasions a very slight precipitate, while muriate of gold, persulphate of iron, and silicated potash, produce no sensible effect whatever. No change is produced by the infusion of nutgalls.*

Linseed mucus.

From quince seeds, and the root of the hyacinth, Dr. Bostock obtained solutions of mucus which exhibited the above properties with some shades of difference. Thus quince seed mucus was coagulated by acids, and by most metallic salts, and the mucus from the hyacinth was precipitated by infusion of nutgalls. These differences were probably owing to foreign bodies; such as starch, gluten, &c. with which the mucus happened to be mixed in these cases, and from which we know no method of separating it completely.

Other varieties.

Mucus is contained in the roots and leaves of a vast number of plants. Almost all the bulbous roots and fleshy leaves yield it. For example, the roots of the *hyacinthus nonscriptus* and the *althea officinalis*; the leaves of the *althea*, of the *malva sylvestris*, of many of the *fuci*, and of the greater number of the lichens; the seeds of flax, quinces, fenugrec, &c.

The bulbs of the hyacinth contain so much mucus, that when dried they may be employed as a substitute for gum arabic. This was first made known to the public by Mr.

* Bostock, Nicholson's Jour. xviii. 31.

Book IV.

Thomas Willis.* He afterwards found that the roots of the vernal squill, the white lily, comfrey, and salop, might be employed also for similar purposes.† A mucus may be extracted from most of the stringy lichens, which likewise answers all the purposes of a solution of gum. This was first ascertained by Georgi, who published a dissertation on the subject in the Memoirs of the Petersburg Academy for 1779.‡ His experiments were repeated and varied by Amoreux, who published the result of his trials in 1787.§

Hoffmann also found abundance of gum in many lichens.|| Lord Dundonald first thought of applying this discovery to practical purposes. During the last war, when the price of gum was high, lichens, at his Lordship's suggestion, were very much employed as a substitute by the calico-printers in this country.

The mucilaginous quality of the fuci is apparently still greater than that of the lichens, though the mucilage obtained from them does not answer so well.

SECT. VI.

OF JELLY.

How obtained.

IF we press out the juice of ripe blackberries, currents, and many other fruits, and allow it to remain for some time in a state of rest, it [partly coagulates into a tremulous soft substance, well known by the name of *jelly*. If we pour off the uncoagulated part, and wash the coagulum with a small quantity of water, we obtain *jelly* approaching to a state of purity.

Properties.

In this state it is nearly colourless, unless tinged by the peculiar colouring matter of the fruit; it has a pleasant taste and a tremulous consistency. It is scarcely soluble in cold water, but very soluble in hot water; and when the solution

* Phil. Mag. xv. 1803.

† Nicholson's Jour. ix. 233.

‡ The lichens which he tried were the *farinaceus*, *glaucus*, *physodes*, and *pulmonarius*.§ The academy of Lyons had proposed a prize dissertation on the lichens in 1786. Three memoirs were approved of, and published by them in 1787, written respectively by Willemet, Amoreux, and Hoffmann. The lichens examined by Amoreux, and found to yield a gum, were the *pulmonarius*, *prunastris*, *islandicus*, *fraxineus*, *canius*, *caperatus*. See his Memoir, p. 95.|| See his Memoir, p. 23, and *passim*.

cools, it again coagulates into the form of a jelly.* When long boiled, it loses the property of gelatinizing by cooling, and becomes analogous to mucilage.† This is the reason that in making current jelly, or any other jelly, when the quantity of sugar added is not sufficient to absorb all the watery parts of the fruit, and consequently it is necessary to concentrate the liquid by long boiling, the mixture often loses the property of coagulating, and the jelly of course is spoiled. ‡

Jelly combines readily with alkalis; nitric acid converts it into oxalic acid, without separating any azotic gas. § When dried it becomes transparent. || When distilled it affords a great deal of pyro-mucous acid, a small quantity of oil, and scarcely any ammonia.**

Jelly exists in all acid fruits, as oranges, lemons, gooseberries, &c. If the juices of these fruits be allowed to gelatinize, and then poured upon a sarsee, the acid gradually filters through, and leaves the other; which may be washed with a little cold water, and allowed to dry. Its bulk gradually diminishes, and it concretes into a hard transparent brittle mass, which possesses most of the properties of gum. Perhaps, then, jelly is merely gum combined with vegetable acid.

SECT. VII.

OF ULMIN.

This substance was first noticed in 1804, by Klaproth, who found it in a spontaneous exudation from the *elm*.†† It was noticed in 1810 by Berzelius in his experiments on the bark of the pine, and on Jesuits bark.‡‡ On the tenth of December, 1812, a set of experiments on the exudation from the elm, which had been examined by Klaproth, was read to the Royal Society. These experiments were made by Mr. Smithson, well known for his precision. §§ Two or three days before the tenth of December, I had printed in the first number of the *Annals of Philosophy*, a set of experiments on an exudation from the common elm which had been collected at

* Vauquelin, Ann. de Chim. vi. 282.

† Ibid. v. 100.

‡ Ibid. v. 10.

§ Ibid. vi. 282.

|| Ibid. v. 100.

** Ibid. vi. 286.

†† Gehlen's Jour. iv. 329.

‡‡ Afhandlingar, iii. 347.

§§ Phil. Trans. 1813, p. 64.

Book IV. Plymouth by Mr. Coulson.* The result of my experiments corresponded very nearly with those of Mr. Smithson; and showed that the extraordinary properties of ulmin, announced by Klaproth, did not belong to it. I afterwards examined varieties of ulmin from the oak, the hornbeam, and the horse-chesnut; and was so fortunaté as to receive a specimen from Dr. Leach in a state of purity.†

From the observations of Berzelius it appears that ulmin forms a constituent of the bark of almost all trees. If the bark be first digested in alcohol, and afterwards in cold water, we may obtain the ulmin nearly pure, by digesting the bark thus treated in hot-water, holding some carbonate of potash in solution. ‡ The properties of ulmin are as follows.

Properties.

1. Its colour is dark-brown or almost black. It has scarcely any taste or smell.

2. It dissolves slowly, and only in small quantity, in water and in alcohol. The solution is pale-brown and tasteless. Hardly any metalline salt precipitates the aqueous solution of ulmin, except pernitrate of mercury and acetate of lead, both of which cause brown flocks to fall down. The alcohol solution is precipitated dark-brown by sulphate of iron and sulphate of copper.

3. When a weak solution of carbonate of potash is poured upon ulmin a dark-brown solution is immediately effected. The ulmin as it exudes from the elm is almost always combined with a quantity of carbonate of potash. Hence the reason that it dissolves so readily in water.

This solution (supposing the carbonate of potash to amount to $\frac{1}{5}$ th of the weight of the ulmin) has a dark brown colour, it is not in the least mucilaginous or ropy, and cannot, like the solution of gum in water, be employed to paste pieces of paper together. It is not precipitated by gelatine, tincture of nut-galls, or hydrocyanate of potash. Sulphate of iron occasions a copious precipitate, which in some cases is brown, and in some cases green. Nitric acid occasions a reddish-brown precipitate by saturating the potash which held the ulmin in solution. Alcohol occasions no precipitate. But all the acids precipitate the ulmin. When thus obtained it is buff-coloured, and after being dried is not sensibly soluble either in water or alcohol.

* Annals of Philosophy, i. 23.

† Ibid. ii. 11, 395, 396.

‡ Ibid. ii. 314.

4. Ulmin when heated swells and catches fire; but does not melt. The smell which is given out has not the least resemblance to that of caromel. Chap. I.

Ulmin seems hitherto to have been confounded with *extractive* by chemists. In many cases it has passed for tannin.

SECT. VIII.

OF COLOURING MATTERS.

THE colouring principles of vegetables are so numerous and so various in their properties that they will hereafter fill a very extensive department of vegetable chemistry. But hitherto so few of them have been obtained in a separate state and subjected to a chemical examination, that it would be in vain to attempt a particular description of each. I shall therefore merely notice the following, which have been honoured with a more particular investigation of their properties

I. POLYCHROITE.

This substance constitutes the colouring matter of *crocus sativus* or *safron*, and is the substance which Hermbstaedt announced as pure extractive. Its properties were investigated in 1811, by Bouillon Lagrange and Vogel,* who gave it the name of polychroite on account of the different colours which it is capable of assuming.

To obtain polychroite, form an infusion of safron in water, evaporate the infusion to the consistency of honey, digest the residue in alcohol of the specific gravity 0.800, filter the solution and evaporate to dryness. The dry mass is polychroite in a state of purity. It possesses the following properties. How obtained.

Its colour is intensely yellow. When exposed to the air it absorbs moisture and becomes a viscid liquid. It is very soluble in water and in alcohol. But insoluble, or scarcely soluble in sulphuric ether. Properties.

When the aqueous solution is exposed to the rays of the sun it becomes colourless, and the yellow colour cannot be again restored. Hence it appears that the polychroite is destroyed, or at least its nature altered by the action of the solar ray.

When sulphuric acid is dropped into the aqueous solution of polychroite the liquid assumes a deep indigo-blue colour,

* Ann. de Chim. lxxxii. 188.

Book IV. which gradually becomes lilac. The same change is produced by sulphuric acid on the alcoholic solution.

Nitric acid causes these liquids to assume a green colour. The addition of a little water causes these colours to disappear, and when the proportion of acid is increased, other shades make their appearance.

Chlorine destroys the colour of these liquids altogether.

Sulphate of iron throws down a dark-brown precipitate, lime-water a yellow precipitate, barytes water a red precipitate. Subacetate of lead throws down a safron coloured precipitate. Acetate of lead produces no effect; but nitrate of mercury throws down a red powder.

Polychroite is insoluble in oils.

When this colouring matter is distilled, there comes over an acid water, and an oil of a lemon-yellow colour, which is succeeded by an oil of a darker colour, and almost brown. At the same time carbonic acid and carburetted hydrogen gases are disengaged. When the eliareoal that remains is calcined with potash a portion of hydrocyanic acid is formed.

II. HEMATIN.

This substance constitutes the colouring matter of the *hæmatoxylon campechianum* or *logwood* so universally known as a dye stuff. Its properties were investigated by Chevreul in 1812, who first recognised it as a peculiar substance, and gave it the name prefixed to this article.*

How obtained.

Hematin may be obtained by the following process. Digest for several hours logwood in powder in water of the temperature 125°. Filter the liquid, evaporate it to dryness, and digest the residue for a whole day in alcohol of the specific gravity 0.837. Filter the alcohol, concentrate it by evaporation, then add a little water, evaporate a little farther, and leave it to itself. Crystals of hematin are deposited in abundance. These, when washed with alcohol and dried, may be considered as hematin in a state of purity.

Properties.

Hematin thus prepared is in small brilliant crystals, of a reddish-white colour, and a slightly astringent, bitter, and acrid taste.

Boiling water dissolves it with facility, and acquires an orange-red colour; which becomes yellow when the liquid cools; but heat again restores the original colour. When this

* Ann. de Chim. lxxxi. 128.

liquid is evaporated the hematin crystallizes. The addition of an acid renders it first yellow and then red. Sulphurous acid however gradually destroys the colouring matter altogether if it be left long enough in contact with it. Potash, soda, and ammonia, give it a purplish-red colour, and when a great excess of these alkalies is added the colour becomes violet-blue; then reddish-brown, and at last yellowish-brown. The hematin is then decomposed as its colour cannot be again restored by the addition of acids. Barytes, strontian, and lime-water produce the same effects: only they gradually precipitate hematin from its solution.

When a current of sulphuretted hydrogen gas is passed through an aqueous solution of hematin it assumes a yellow colour, which disappears in the course of a few days. The oxides of lead, tin, iron, copper, nickel, zinc, antimony, and bismuth, unite with hematin, and give it a blue colour with a shade of violet. The peroxide of tin acts upon it as the mineral acids do. Gelatine throws it down in reddish flocks.

Besides these colouring matters there are several others, such as *madder*, *safflower*, *litmus*, that would deserve a particular description. But as these substances have not yet been analyzed, and their colouring ingredient obtained in a separate state, I think it better to defer the account of them to the ensuing chapter. Indigo, which has been sufficiently examined to ascertain its peculiar properties seems entitled, from its great importance, to rank as a peculiar vegetable principle.

Other species.

SECT. IX.

OF BITTER PRINCIPLES.

MANY vegetable substances have an intensely bitter taste, and on that account are employed in medicine, by brewers, &c. This is the case with the wood of the *quassia amara* and *excelsa*, the common *quassia* of the shops; with the roots of the *gentiana lutea*, common *gentian*; the leaves of the *humulus lupulus* or *hop*; the bark and wood of the *spartium scoparium* or common broom; the flowers and leaves of the *anthemis nobilis* or *chamomile*; and many other substances. These bodies owe their bitter taste to the presence of peculiar vegetable substances differing from every other, which may be distinguished by the name of *bitter principles*.

Plants containing bitter principles.

Book IV.

Of late years several of these peculiar substances have been examined and characterized; and doubtless there are others with which we are not yet acquainted. I shall give an account in this section of the different bitter tasted vegetable substances that appear to possess distinct qualities and to differ from each other by well characterized properties. I have little doubt that several of them will be found to possess alkaline properties.

I. *Quassin, or Bitter Principle of Quassia.*

Experiments on the infusion of quassia.

I. When water is digested over *quassia* for some time, it acquires an intensely bitter taste and a yellow colour, but no smell. When water thus impregnated is evaporated to dryness in a low heat, it leaves a brownish-yellow substance, which retains a certain degree of transparency. It continues ductile for some time, but at last becomes brittle. This substance I shall consider as *quassia* in a state of purity. If it contain any foreign body, it must be in a very minute proportion. This substance I find to possess the following properties:

1. Its taste is intensely bitter. Colour brownish-yellow.
2. When heated, softens, and swells, and blackens; then burns away without flaming much, and leaves a small quantity of ashes.
3. Very soluble in water and alcohol.
4. Does not alter the colour of infusion of litmus.
5. Lime-water, barytes-water, and strontian-water, occasion no precipitate. Neither is any precipitate thrown down by silicated potash, aluminated potash, or sulphate of magnesia.
6. The alkalies occasion no change in the diluted solution of the bitter principle.
7. Oxalate of ammonia occasions no precipitate.
8. Nitrate of silver renders the solution muddy, and a very soft flaky yellow precipitate falls slowly to the bottom.
9. Neither corrosive sublimate nor nitrate of mercury occasion any precipitate.
10. Nitrate of copper, and the ammoniacal solution of copper, produce no change; but muriate of copper gives the white precipitate, which falls when this liquid salt is dropped into water.
11. Sulphate and permuriate of iron occasion no change.
12. Muriate of tin renders the solution muddy, but occa-

sions no precipitate, unless the solution be concentrated; in that case a copious precipitate falls. Chap. I.

13. Acetate of lead occasions a very copious white precipitate; but the nitrate of lead produces no change.

14. Muriate of zinc occasions no change.

15. Nitrate of bismuth produces no change, though when the salt is dropped into pure water a copious white precipitate appears.

16. Tartar emetic produces no change; but when the muriate of antimony is used, the white precipitate appears, which always falls when this salt is dropped into pure water.

17. Muriate and arseniate of cobalt occasion no change.

18. Arseniate of potash produces no effect.

19. Tincture of nutgalls, infusion of nutgalls, gallic acid, occasion no effect.

These properties are sufficient to convince us that the bitter principle is a substance differing considerably from all the other vegetable principles. The little effect of the different reagents is remarkable. Nitrate of silver and acetate of lead are the only two bodies which throw it down. This precipitation cannot be ascribed to the presence of muriatic acid; for if muriatic acid were present, nitrate of lead would also be thrown down. Besides, the flakes produced by nitrate of silver are too light, and indeed have no resemblance whatever to muriate of silver. The precipitate by acetate of lead is very copious. This salt is therefore the best substance for detecting the presence of the bitter principle, when we are certain that no other substance is present which throws down lead.

Vauquelin has examined the bitter principle in the fruit of the colocynth (*cucumis colocynthis*), and in the root of the briony (*brionia alba*), two vegetable bodies which have been employed in medicine, and possess strong lethargic powers. The bitter principle in these substances, as far as his experiments went, possesses exactly the properties of the bitter principle of quassia.* It is probable also that hops, gentian, and broom, contain the same bitter principle. Exists in other plants.

When alcohol is digested for some months upon wheat flour, and then distilled, it leaves a yellowish-brown substance behind it, which possesses (as far as I have examined it) all the properties of the bitter principle of quassia.

* Vauquelin, *Essais Analytiques des Racines d'Ellebore d'Hiver et] de Brione*. Ann. de Mus. d'Hist. Nat. No. xliii. 91.

II. *Scillitin*, or *Bitter Principle of Squills*.

Book IV.

How obtained.

This is the peculiar substance to which the bulb of the *scilla maritima*, or squill, owes its medicinal qualities. Its properties were investigated by Vogel in 1812, who pointed out its peculiar nature, and gave it the name by which it is distinguished.* It may be obtained by the following process. The juice of the bulbs being expressed was boiled for a few minutes, and a quantity of citrate of lime which appeared was separated. It was then evaporated to dryness, and the dry residue digested in alcohol as long as that liquid took up any thing. The alcoholic solution was evaporated to dryness, and the residue (which consisted of scillitin and tannin) was redissolved in water. Acetate of lead was dropped in to throw down the tannin. The liquid was filtered, and the excess of lead which had been added was separated by means of a current of sulphuretted hydrogen gas. The liquid being again filtered, was evaporated to dryness to drive off the acetic acid from the acetate. The dry mass was scillitin, mixed with a little sugar, from which it was not in Vogel's power to separate it.

Properties.

Scillitin thus obtained is white and transparent, and breaks with a resinous fracture. It was easily reduced to powder. It absorbs water rapidly from the atmosphere, and becomes at first a viscid mass, and at last quite fluid. Its taste is intensely bitter, leaving a slight impression of sweet from the sugar with which it is mixed. It dissolves readily in water, and gives it a mucilaginous consistency like gum. It dissolves likewise very readily in cold alcohol of 0.817, and still more readily in that liquid while hot. When heated, it swells prodigiously, and gives out a distinct odour of caramel. When treated with nitric acid no saccharic acid is formed.

III. *Caffein*, or *Bitter Principle of Coffee*.

How obtained.

This species is distinguished from the preceding by the property which it has of striking a green colour with iron, and of precipitating that metal from concentrated solutions. Mr. Chenevix separated a portion of it from coffee by the following process: he digested unburnt coffee in water, and filtered the liquid. It was then treated with muriate of tin. The precipitate wasedulcorated, mixed with water, and treated

* Ann. de Chim. lxxxiii. 247.

with sulphuretted hydrogen gas. The tin was thus precipitated, and the substance with which it had been combined was dissolved by the water. The liquid was then evaporated to dryness. The substance thus obtained possessed the following properties :

1. Semitransparent like horn, and of a yellow colour. * Properties.
2. When exposed to the air, it does not attract moisture.
3. Soluble in water and in alcohol. The solution in water is semitransparent, and has a pleasant bitter taste. When the alkaline solutions are dropped into it, the colour becomes garnet-red.
4. It is not precipitated from water by the alkaline carbonates. Sulphuric acid renders the solution brown, but produces no further change. Neither muriatic acid, nor phosphoric acid, nor the vegetable acids, produce any change on this solution.
5. The muriates of gold, platinum, and copper, occasion no change.
6. With solutions of iron it forms a fine green-coloured liquid; and when concentrated, iron throws down a green-coloured precipitate. Indeed it is almost as delicate a test of iron as tan and gallic acid.
7. Muriate of tin throws down a copious yellow precipitate. This precipitate, and that by iron, are soluble in all acids, but they lose their colour.
8. Neither lime nor strontian water occasion any precipitate in the aqueous solutions of this substance; but barytes water occasions a brown precipitate.
9. Gelatine occasions no precipitate.*

From the experiments of Bouillon la Grange, it appears that a substance similar to this species of bitter principle exists in the flowers of the *arnica montana*.† There is reason for believing that the bitter principle which characterizes the following plants belongs also to this species; *absinthium vulgare* (wormwood,) *juniperus sabinus* (sabine,) *ruta graveolens* (rue,) *anthemis nobilis* (chamomile,) *achillea millefolium* (milfoil.)

IV. *Daphnin*, or *Bitter Principle of Daphne Alpina*.

This species was obtained by Vauquelin from the bark of *daphne alpina*. The bark was digested in alcohol, and the liquid was concentrated till the resin separated. The residual Preparation.

* Phil. Mag. 1802, May, p. 350.

† Ann. de Chim. lv. 35.

Book IV.

liquid, being diluted with water and filtered, was mixed with acetate of lead. A copious yellow precipitate falls. This precipitate being separated by the filter, is diluted with water, and is treated with sulphuretted hydrogen gas to get rid of the lead. Being filtered and sufficiently concentrated, the daphnin separated in small transparent crystals. It is hard, has a greyish colour, and a very bitter taste. When heated, it evaporates rapidly in sharp vapours. It is sparingly soluble in cold water, much more soluble in boiling water, and separates in crystals as the solution cools. Its solution is not precipitated by acetate of lead. When heated in a retort, it melts, blackens, and emits an acid.*

Properties.

V. *Artificial Bitter Principle.*

History.

This species has been formed by the action of nitric acid on various vegetable and animal substances. It was first obtained by Haussman while examining indigo, but he mistook its nature. Welther afterwards formed it by digesting silk in nitric acid; ascertained its properties, and gave it the name of *yellow bitter principle*; † he is therefore to be considered as the real discoverer. Bartholdi afterwards procured it by treating the *white* willow with nitric acid. ‡ Mr. Hatchett obtained it during his experiments on artificial tannin, by treating indigo with nitric acid; § and about the same time Fourcroy and Vauquelin procured it by the same means, and examined its properties in detail. || Mr. Chevreul is of opinion that it is a compound of nitric acid and a peculiar vegetable substance, probably of an oily nature.*** It possesses the following properties:

Properties.

Its colour is a deep yellow, its taste intensely bitter. It is soluble both in water and alcohol, and has the property of dyeing silk, woollen cloth, and cotton, of a durable yellow colour. It crystallizes in elongated plates, and possesses many of the characters of an acid, combining readily with alkaline substances, and forming crystallizable salts. When potash is dropped into a concentrated solution of it, small yellow prismatic crystals are gradually deposited, consisting of bitter principle combined with potash. These crystals were examined by Welther, but it was Fourcroy and Vauquelin that

* Ann. de Chim. lxxxiv. 174.

† Ibid. xxxii. 187.

|| Phil. Mag. xxiii. 255.

† Ibid. xxix. 301.

§ Phil. Trans. 1805.

** Ann. de Chim. lxxii. 113.

ascertained their composition. They have a bitter taste, are not altered by exposure to the air, and are less soluble than pure bitter principle. When thrown upon hot charcoal they burn like gunpowder, and detonate very loudly when struck upon an anvil, emitting a purple light.* Ammonia dropped into the solution of bitter principle deepens its colour, and occasions a copious deposition of fine yellow spicular crystals. These are a combination of bitter principle and ammonia.† It combines likewise with the alkaline earths, and with metallic oxides, as has been shown by the experiments of Chevreul.

VI. *Artificial Tannin.*

This itself may be considered as approaching the bitter principle in many of its properties. Its taste is always intensely bitter, and the colour of the precipitates which it throws down from the metals is similar to what takes place when artificial bitter principle is present. It is indeed possible that the bitter taste may be owing not to the tannin, but to a portion of artificial bitter principle which may be always formed along with the tannin; but this has not been ascertained. It is well known that the bitter taste very easily overpowers and conceals all other tastes.

SECT. X.

OF NICOTIN.

THIS substance exists in the leaves of the *nicotiana latifolia*, or *tobacco*, and gives that plant its peculiar properties. It was obtained in 1809 by M. Vauquelin, to whom we are indebted for all the facts respecting it yet known.‡ It may be obtained from the juice of tobacco leaves by the following process.

Evaporate the juice to one fourth of its volume, and allow Preparation. it to cool. A quantity of malate of lime will be deposited under the form of a gritty powder. Separate this matter, and continue to concentrate the liquid as long as any saline matter can be obtained from it. When it is so much inspissated that nothing farther can fall, digest it in alcohol, which will take up the uncombined malic and acetic acids that it contains, the nicotin and a portion of sal ammoniac. Distil the alcohol off

* Welther, Fourcroy, and Vauquelin.

† Hatchett, Phil. Trans. 1805.

‡ Ann. de Chim. lxxi. 139.

Book IV.

in a retort, and concentrate the residual matter nearly to dryness, taking care not to injure it by the application of too much heat. Dissolve it a second time in more concentrated alcohol, in order to separate a portion of animal matter which had been dissolved in the first alcohol solution in consequence of the water present. Distil off this second portion of alcohol. Dissolve the residue in water, saturate the acetic and malic acids with potash; and distil the water off to dryness. This water, though colourless, contains in solution a quantity of nicotin. By dissolving the residual matter in water, and distilling a second time, an additional quantity of nicotin may be obtained. And by repeating this process, a sufficient number of times, the whole of the nicotin may be obtained in solution in water. By a very cautious evaporation, the water may be driven off, and the nicotin obtained in a state of purity.

Properties.

This substance is colourless. It has an acrid taste, and the peculiar smell which distinguishes tobacco. It occasions violent sneezing.

It is soluble in water and in alcohol. Both solutions are colourless, and distinguished by the peculiar taste and smell of nicotin.

It is precipitated from its solutions by tincture of nutgalls. It is volatile, and possesses poisonous qualities. And according to Vauquelin approaches the volatile oils in its properties.

SECT. XI.

OF EXTRACTIVE.

History.

The word *extract* was at first applied to all those substances which were extracted from plants by means of water, and which remained behind in the state of a dry mass when the water was evaporated; consequently it included gum, jelly, and several other bodies. But of late it has been confined by many to a substance which is supposed to exist in many plants, and which may be obtained tolerably pure from the bark of the *cinchona officinalis* according to the experiments of Schrader. But as the word *extract* occurs even in modern authors in its original sense, I shall rather denote this substance by the phrase *extractive principle*, to prevent ambiguity.

Indeed it is becoming more and more probable every day, that no substance entitled to the name of extractive exists in plants.* Chap. I.

The difficulty of obtaining the extractive principle in a separate state, and the facility with which it alters its nature, have hitherto prevented chemists from examining it with that attention to which it is entitled. It was first particularly attended to by Rouelle; but it is to Fourcroy and Vauquelin that we are chiefly indebted for ascertaining its characters. The dissertation of Vauquelin in the *Journal de Pharmacie*, is by far the best account of extractive matter which has hitherto appeared. Many valuable facts and curious observations were published by Schrader in his dissertations on *extractive*.† But unfortunately the term has not been always taken by chemists in the same acceptation. Parmentier has for example published a dissertation on the *extracts* of vegetables taken in the loose and general sense of the word, which contains much information.‡ There cannot I think be any doubt that under the term *extractive*, chemists have hitherto confounded together a variety of different vegetable principles. Some of these, as *ulmin*, *polychroit*, &c. I have distinguished by peculiar names, and perhaps it would be wisest to lay aside the term *extractive* altogether. But in the present state of obscurity in which this department of vegetable chemistry continues, I am unwilling to withhold the facts collected in this section, which though imperfect and probably partly inaccurate, will yet be useful to those who apply themselves to the investigation of this subject. Uncertainty.

The extractive principle possesses the following properties:

1. Soluble in water, and the solution is always coloured. Properties.
When the water is slowly evaporated, the extractive matter is obtained in a solid state and transparent; but when the evaporation is rapid the matter is opaque.
2. The taste of extractive is always strong; but it is very different according to the plant from which it is obtained.
3. Insoluble in absolute alcohol and in ether; but soluble in alcohol when it contains water.
4. By repeated solutions and evaporations, the extractive

* The reader is referred to a valuable paper by Braconnot, on this subject, in the *Jour. de Phys.* lxxxiv. 267.

† Gehlen's *Jour.* Second Series, viii. 548, and Schweigger's *Jour.* ix. 139.

‡ *Ann. de Chim.* xliii. 19.

Book IV.

matter acquires a deeper colour, and becomes insoluble in water. This change is considered as the consequence of the absorption of the oxygen of the atmosphere, for which the extractive principle has a strong affinity. But if the solution be left to itself, exposed to the atmosphere, the extract is totally destroyed in consequence of a kind of putrefaction which speedily commences.

5. When chlorine is poured into a solution containing extractive, a very copious dark-yellow precipitate is thrown down, and the liquid retains but a light lemon-colour. These flakes are considered as oxygenized *extractive*. It is now insoluble in water; but hot alcohol dissolves it.

6. The extractive principle unites with alumina, and forms with it an insoluble compound. Accordingly, if sulphate or muriate of alumina be mixed with a solution of extractive, a flaky insoluble precipitate appears, at least when the liquid is boiled; but if an excess of acid be present, the precipitate does not always appear.

7. It is precipitated from water by concentrated sulphuric acid, muriatic acid, and probably by several other acids. When the experiment is made with sulphuric acid, the fumes of vinegar generally become sensible.

8. Alkalies readily unite with extractive, and form compounds which are insoluble in water.

9. The greater number of metallic oxides form insoluble compounds with extractive. Hence many of them, when thrown into its solution, are capable of separating it from water. Hence also the metallic salts mostly precipitate extractive. Muriate of tin possesses this property in an eminent degree. It throws down a brown powder perfectly insoluble, composed of the oxide of tin and vegetable matter.

10. If wool, cotton, or thread, be impregnated with alum, and then plunged into a solution of extractive, they are dyed of a fawn-brown colour, and the liquid loses much of its extractive matter. This colour is permanent. The same effect is produced if muriate of tin be employed instead of alum. This effect is still more complete if the cloth be soaked in chlorine, and then dipped into the infusion of extractive. Hence we see that the extractive matter requires no other mordant than oxygen to fix it on cloth.

11. When distilled, extractive yields an acid liquid impregnated with ammonia.*

* See Vauquelin.

It cannot be doubted that there are many different species of extractive matter; though the difficulty of obtaining each separately has prevented chemists from ascertaining its nature with precision. Extracts are usually obtained by treating the vegetable substance from which they are to be procured with water, and then evaporating the watery solution slowly to dryness. All extracts obtained by this method have an acid taste, and redden the infusion of litmus. They all yield a precipitate while liquid if they are mixed with ammonia. This precipitate is a compound of lime and insoluble extractive. Lime always causes them to exhale the odour of ammonia. It has been ascertained that the extractive principle is more abundant in plants that have grown to maturity than in young plants.

Properties of
extracts.

As the *extracts* of vegetables prepared by apothecaries for medical purposes, besides the extractive principle, always contain other bodies, frequently to the number of eight or more, and as the greater number of them are still but imperfectly examined, I shall satisfy myself at present with pointing out some of those vegetable substances which have been ascertained to contain extractive principle, and stating the constituents of such as have been analysed.

1. Extractive principle is not an uncommon ingredient in the sap of trees. Indeed Deyeux and Vauquelin found it in almost all those which they examined. It is usually thrown down when the sap is mixed with chlorine, and it precipitates in brown flakes while the sap is evaporating on a sand bath.

Extractive
of sap.

2. It forms a constituent of the bark of all trees hitherto examined. This was evidently the case with all the barks which Davy subjected to experiment; namely, those of oak, Leicester willow, Spanish chesnut, elm, common willow, and undoubtedly all barks which have an astringent taste; for tan and extractive seem scarcely ever to be found separate.

Of bark,

3. The infusion of catechu contains an extractive principle, united chiefly to tan. It may be obtained in a state of purity by washing the catechu in powder repeatedly with water till the fluids obtained cease to precipitate gelatine. What remains is extractive. It is of a pale reddish-brown colour, and a slightly astringent taste, leaving in the mouth a sensation of sweetness. It has no smell. Its solution in water is at first yellow-brown; but it acquires a tint of red when left exposed to the air. The solution in alcohol is of a dirty brown. It does not affect vegetable blues. Alkalies brighten

Of catechu,

Book IV.

its colour; but neither these bodies nor the alkaline earths precipitate it from water. Nitrate of alumina and muriate of tin render the solution slightly turbid. Nitrate of lead throws down a dense light-brown precipitate. It renders persulphate of iron green, and throws down a green precipitate, becoming black by exposure to the air. Linen, when boiled in the solution, extracts almost the whole of the extractive, and acquires a light red-brown colour. When this extractive is exposed to heat, it softens, and its colour becomes darker, but it does not melt. When distilled, it yields carbonic acid and carburetted hydrogen gas, weak acetic acid, and a little unaltered extractive. A porous charcoal remains.*

Of senna,

4. The infusion of senna contains a matter of a very peculiar nature, but which may be considered as a species of *extractive*. The senna of commerce consists of the dried leaves of the *cassia senna*, a shrub-like annual cultivated in Egypt. Water, at the ordinary temperature of the atmosphere dissolves nearly one-half of the substance of these leaves. The infusion obtained has a brown colour, a bitter taste, and a peculiar aromatic odour. It contains a considerable proportion of carbonate of lime, sulphate of potash, and carbonate of magnesia, besides a little silica; but the most curious of its constituents is the extractive. When common air is made to pass for some hours through the infusion, a yellow-coloured precipitate appears: the same substance is thrown down immediately by muriatic acid and chlorine. It appears also when a current of chlorine gas is made to traverse the infusion. This substance is the extractive altered by its combination with oxygen. It has a slight bitter taste. It is no longer soluble in water. Alcohol dissolves it; but lets it fall when diluted. The alkalies dissolve it, and form a deep-brown solution. On burning coals it emits a thick smoke, exhales an aromatic odour, and leaves a spongy charcoal. These properties indicate a very decided approach to the resinous state.†

Of quinquina.

5. The infusion of Peruvian bark likewise yields an extractive matter of a peculiar nature, which assumes a fine red-colour when united to oxygen; and, like the extractive of senna, acquires nearly the properties of a resin. It was obtained by Fourcroy from the quinquina of St. Domingo. Water, boiled on it till it refused to dissolve any thing more,

* Davy on Astringent Vegetables, p. 25.

† Bouillon La Grange, Ann. de Chim. xxiv. 4.

was slowly evaporated, and the extract obtained was dissolved in alcohol. The alcohol, by evaporation, deposited the peculiar extractive. Its colour was brown, its taste bitter, insoluble in cold, but very soluble in hot water. It was soluble in alcohol. When dry it was black, brittle, and broke with a polished fracture. Lime-water precipitated it in the state of a red powder: chlorine threw it down in the state of a fine red powder, neither soluble in water nor alcohol, but capable of uniting with alkalis. A stronger dose of chlorine rendered it yellow.*

6. Steinacher has shown that distilled vinegar contains a quantity of extractive.†

The resemblance between extractive bodies and the colouring matter of plants is sufficiently striking. It is more than probable, that when this last set of bodies have been examined with more precision by chemists, they will be found to belong to the same class.

SECT. XII.

OF EMETIN.

This remarkable substance was extracted from ipecacuanha by MM. Magendie and Pelletier, and called by them *emetin*, because it constitutes the substance in that root, which occasions vomiting.

They obtained it by the following method: Ipecacuanha was digested first in sulphuric ether, then in alcohol. The alcoholic solution was evaporated to dryness, re-dissolved in water and acetate of lead dropped into the solution. A copious precipitate fell, which beingedulcorated and diffused through water was exposed to a current of sulphuretted hydrogen gas. The lead was precipitated in the state of sulphuret, and the emetin dissolved in water. The liquid being filtered and evaporated to dryness, the residue consisted of emetin in a state of purity. Thus obtained it possessed the following properties:

It is in the form of transparent scales which have a brownish red colour. It has no smell. Its taste is bitter and a little acrid; but not in the least nauseous. It is not altered by a

* Fourcroy, Ann. de Chim. viii. 120.

† Jour. de Phys. lxxv. 327.

Book IV. degree of heat inferior to that of boiling water. When exposed to a stronger heat it does not melt; but swells, blackens, and is decomposed, furnishing water, carbonic acid, a little oil and acetic acid. A very spongy and light coal remains. No ammonia can be discovered in these products, indicating that azote does not enter into the composition of emetin.

Emetin when exposed to the air undergoes no change, unless the air be very damp, in which case it becomes moist. It dissolves very readily in water; but cannot be made to crystallize. It is soluble in alcohol; but insoluble in sulphuric ether.

Diluted sulphuric acid produces no effect on it; but concentrated sulphuric acid chars and destroys it. Nitric acid, when applied either cold or hot, dissolves it and forms a fine red-coloured solution, which gradually passes into yellow while nitrous gas exhales, and crystals of oxalic acid are formed, but no yellow bitter principle. Muriatic and phosphoric acid dissolve it without alteration, and let it fall again when they are saturated with an alkali, acetic acid is one of the best solvents of it. Gallic acid on the contrary throws it down from its solution in the state of a dirty white precipitate. It is precipitated equally by infusion of nutgalls, and a slight precipitate is produced by the infusion of cinchona.

The alkalies dissolve the compound of emetin and gallic acid. When iodine is poured into the tincture of emetin a red precipitate falls which is a compound of emetin and iodine.

Acetate of lead throws down emetin from its solution in abundance. Protonitrate of mercury gradually throws down a slight precipitate. Corrosive sublimate occasions rather a greater one. Tartar emetic has no action on it whatever.

Neither vegetable nor animal substances as far as they have been tried have any action on emetin.

Half a grain of emetin when swallowed occasions violent vomiting followed by sleep, and the animal awakes in a state of health. Twelve or even six grains of emetin produce vomiting and sleep, followed by death. A violent inflammation of the lungs and intestinal canal appears to be the proximate cause of the death which in this case ensues.*

* Ann. de Chim. et Phys. iv. 172.

SECT. XIII.

OF ASPARAGIN.

I GIVE this name to a substance discovered in the juice of asparagus by Vauquelin and Robiquet.* It was procured by the following process. Chap. I.

The juice of asparagus is expressed in the usual way, filtered, evaporated to the consistence of a syrup, and then set aside. Various crystals gradually make their appearance, and among others crystals of asparagin, easily separated from the rest on account of their colour and figure. This substance possesses the following properties: Preparation,

1. The crystals are white and transparent, and have the figure of rhomboidal prisms. The greater angle of the rhomboidal base, as ascertained by Haüy, is 130° . The edges of this base, and the two angles at the extremity of its greater diagonal, are truncated. Properties.

2. It is hard and brittle. Its taste is cool and slightly nauseous, so as to occasion a secretion of saliva.

3. It dissolves readily in hot water; but in cold water, only sparingly. Alcohol does not dissolve it.

4. The aqueous solution does not affect vegetable blues. Neither infusion of nutgalls, acetate of lead, oxalate of ammonia, muriate of barytes, nor the hydrosulphuret of potash, occasion any change in it.

5. When triturated with potash no ammonia is disengaged. The potash appears to render it more soluble in water.

6. When heated it swells, and emits penetrating vapours, affecting the eyes and nose like the smoke of wood. It leaves a large portion of insipid charcoal, which, when incinerated, leaves scarcely a trace of residue.

7. Nitric acid dissolves it with the evolution of nitrous gas. The solution has a yellow colour and a bitter taste, like that of animal substances in the same acid. Lime disengages from it a considerable quantity of ammonia.

Such are the properties of this substance ascertained by Vauquelin and Robiquet. They distinguish it sufficiently from all other vegetable substances known.

* Ann. de Chim. lv. 167. Nicholson's Jour. xv. 242.

SECT. XIV.

OF CERASIN.

Book IV.

I ADOPT this name from Dr. John,* to denote a particular vegetable principle which has been long known; but which hitherto has been confounded with gum. It may be distinguished by the following characters.

Properties.

1. It is a solid substance possessing the appearance of gum and distinguished by a similar taste. It is usually harder than gum and not so easily reduced to powder; though this is not always the case, especially when it is mixed with gum which it frequently is.

2. When put into water it imbibes that liquid, swells up very considerably, and becomes semitransparent and gelatinous, but is not in the least soluble. It dissolves in boiling water, but again precipitates when the liquid cools, and remains in the state of a jelly. This gelatinous mass may be used, like gum, to paste together pieces of paper.

3. Cerasin is insoluble in alcohol and ether.

4. Cold water acidulated with nitric, sulphuric, or muriatic acid dissolves a small portion of it, and when heat is applied, the whole of the cerasin dissolves in these liquids. When alkali is dropped into the nitric acid solution, a portion only of the cerasin is precipitated. If the liquid be evaporated, the remainder is obtained converted into a peculiar bitter tasted substance.

Tragacanth.

Gum tragacanth may be considered as pure cerasin. It has been longest known and is in the most general use. Perhaps on that account it might have been as well to have given to cerasin the name of tragacanth.

Tragacanth is the produce of the *astragalus tragacantha*, a thorny shrub which grows in Candia and other islands of the Levant. The gum is said to exude about the end of June from the stem and larger branches, and soon dries in the sun. It is in the state of whitish vermiform pieces, not nearly so transparent as gum arabic. From the two preceding species it differs extremely in many of its properties. When put into water it slowly imbibes a large quantity of the liquid, swells into a large volume, and forms a soft, but not fluid, mucilage. If the quantity of water be more than the gum can imbibe, the

* More lately he has proposed to substitute the term *prunin*, because the substance which exudes from the fruit of the *prunus myrobola* when punctured by insects is almost pure cerasin. See *Chemische Untersuchungen*, 1. 17.

mucilage forms an irregular mass, which does not unite with the rest of the liquid. By agitation, indeed, an apparent solution is effected, and the whole assumes a wheyish appearance; but on standing, the mucilage again subsides, as at first, and the water becomes transparent.* When a solution of gum arabic is poured into this solution, the mucilaginous gum tragacanth separates much sooner than usual, and forms no union with the gum arabic.† When treated with nitric acid, it yields abundance of saccharic acid, malic acid, and oxalic acid; but not the slightest trace of artificial tannin.‡ Sulphuric acid develops some traces of artificial tannin. The quantity of charcoal which it yielded, when digested in sulphuric acid, was 0.22, or considerably less than gum arabic yielded in the same circumstances.§ When Mr. Cruickshanks distilled 480 grains of it in a glass retort, he obtained the following products:

Pyromucous acid	245 gr.
Charcoal	93
Lime with some phosphate	12
Carbonic acid and carburetted hydrogen gas	130
	480

When the pyromucous acid was saturated with lime, a considerably greater proportion of ammonia was disengaged than from the pyromucous acid of gum arabic. The gases were 78 ounce measures of carbonic acid, and 91 of carburetted hydrogen.|| This gum yielded to Vauquelin 0.03 of ashes. They dissolved in muriatic acid with effervescence, while the odour of sulphuretted hydrogen was perceptible. They consisted chiefly of carbonate of lime, with a little iron and phosphate of lime, and contained likewise some potash.**

These experiments show us, that gum tragacanth contains more azote and lime, and perhaps more oxygen and less carbon, than gum arabic.

Dr. Bostock by digesting tragacanth in water till it became gelatinous, and then triturating it with pure water in a mortar,

* Lewis, Neumann's Chem, p. 282.

† Lewis, *ibid.*

‡ Hatchett, Phil. Trans. 1805.

§ Hatchett, Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

|| Cruickshanks, Rollo on Diabetes, p. 452. ** Ann. de Chim. liv. 512.

Book IV.

formed a homogeneous mucilage, consisting of 100 parts of water and one part of tragacanth. Subacetate of lead formed a copious precipitate with this mucilage. Acetate of lead produced a light precipitate, which increased by standing, though on mucilage of gum arabic it produced no effect. The permuriate of tin likewise throws down a firm coagulum, though it does not alter mucilage of gum arabic. Persulphate of iron produced no effect. Nitrate of mercury threw down a slight precipitate of a reddish tinge. Silicated potash produced no effect. The permuriate of tin likewise throws down a slight precipitate of a reddish tinge.

I suppose the gum bassora of Vauquelin to be a variety of tragacanth.

Cerasin constitutes a portion of the gummy matter that exudes from the *prunus cerasus*, *prunus avium*, *prunus domestica*, *xanthorea hastilis*. It was on that account, I presume, that Dr. John distinguished it by the name of *cerasin*. It may be worth while to state here the result of my experiments on cherry tree gum.

Cherry tree gum.

The *prunus avium*, the common cherry and plum trees, and the almond and apricot likewise, yield a gum which exudes in great abundance from natural and artificial openings in the stem. It is of a reddish-brown colour, in large masses, at first much softer than gum arabic, but by keeping it becomes much harder than any other species of gum which I have seen, except tragacanth. When put into water it gradually swells, and is converted into a semitransparent reddish-brown jelly. A portion dissolves in the water, but a considerable portion remains in the gelatinous state, and does not dissolve even though boiled for some hours in a very large portion of water. In this respect cherry-tree gum differs from every other species which I have tried. The portion dissolved in water possesses the common appearance of mucilage of gum arabic, excepting that its colour is darker, and that it is much weaker, and of course less adhesive and thready. The gum is not precipitated by alcohol, nor is any effect produced by silicated potash. Acetate of lead produces no immediate effect; but on standing the whole becomes opaque, and a precipitate at last subsides. Fuming liquid of libavius causes the liquid to gelatinize immediately. The acetate of lead, and the nitrate of mercury, produce no effect. No precipitate was produced by any of the

* Nicholson's Jour. lviii. 30.

other metallic salts tried. Dr. Bostock found, that when treated with nitric acid it yielded a portion of sacclactic acid. Chap. I.

The gentlemen constituting the late unfortunate expedition to the Congo, while travelling along the banks of that river, picked up a large piece of gum, which bore a very close resemblance to the cherry-tree gum of this country. Its colour was equally dark and its hardness as great. On examination, I found that it consisted entirely of cerasin darkened by a colouring matter, which was partly soluble in water and possessed the characters of extractive. But the cerasin still continued dark coloured even after all the extractive which water was capable of taking up was removed. Congo gum

Some time ago, I received from Dr. Clarke of Dominica, specimens of gum very different in appearance from any that I had previously seen. It was in large masses like stalactites, four or five inches long, and above an inch thick, and composed of a number of irregularly cylindrical masses, some of them hollow, cemented together, or running into each other. I have learned nothing respecting the plant from which this gum exudes. It is rather brittle, light yellowish-brown, and translucent. I found it a compound of about 3 parts cerasin and 1 part gum. Dominica gum.

SECT. XV.

OF INULIN.

I GIVE this name to a new vegetable principle, discovered by Rose in the roots of the *inula helenium* (*elecampane*.)

When the roots of this vegetable are boiled in water, the decoction, after standing some hours, deposits the *inulin* in the form of a white powder like starch. Its properties are as follows :

1. It is insoluble in cold water. By trituration the inulin is uniformly diffused, and gives the liquid a milky appearance; but it soon falls down in the state of a white powder, leaving the liquid quite transparent. Properties.

2. It dissolves readily in hot water. One part of inulin in four parts of boiling water formed a solution which passed readily through the filter, though its consistence was somewhat mucilaginous, and it was not quite transparent. After some hours the greater part of the inulin precipitates from the water Action of water,

Book IV. in the form of a white powder. An equally concentrated solution of gum arabic is much thicker and more adhesive.

Alcohol, 3. When the aqueous solution of inulin is mixed with an equal bulk of alcohol, no change takes place for some time, but the inulin soon separates, and falls to the bottom in the state of a bulky white powder. A solution of gum arabic, when treated in this manner, remains milk-white for days without any precipitate falling.

Heat, 4. When thrown upon burning coals it melts as readily as sugar, and emits a thick white smoke, not unpleasantly pungent, and similar in odour to that of burning sugar. The residue, which is but small, sinks into the coal. Starch emits a similar smoke, but leaves a more bulky residue. When heated in an iron spoon inulin melts, and emits a thick smoke, with the same odour; and as soon as the spoon becomes red-hot, burns with a vivid flame, and leaves a very small coaly residue.

5. When distilled it yields a brown acid liquid, having the smell of the pyromucous acid, but not a trace of oil.

Nitric acid. 6. When treated with nitric acid it yields malic and oxalic acids, or acetic acid if too much nitric acid be employed; but no saccharic acid is formed, as happens with the gums, neither is any of the waxy matter separated which is formed when starch is digested with the same acid.

Such are the properties of inulin, as far as they have been investigated by Rose.* When the root of elecampane is distilled, it yields a portion of a white solid matter, which seems to be intermediate between camphor and volatile oil in its nature.† No other vegetable substance has been observed to yield a similar product. How far it is connected with inulin remains to be determined.

Inulin has been found, by Dr. John, in the roots of several other plants. Besides the inula helenium in which it had been already detected by Rose, he found it in the roots of the anthemis pyrethrum, in nutgalls, in angelica archangelica. He suspects its presence, likewise, in the *hyacinthus nonscriptus*, and in several other plants.‡

* Gehlen's Jour. iii. 217.

† Neumann's Chem. p. 420.

‡ John's Tabellen der Pflanzen Analysen, p. 17.

SECT. XVI.

OF STARCH.

IF a quantity of wheat flour be formed into a paste, and then held under a very small stream of water, kneading continually till the water runs off from it colourless, the flour by this process is divided into two distinct constituents. A tough substance of a dirty white colour, called *gluten*, remains in the hand; the water is at first milky, but soon deposits a white powder, which is known by the name of *starch*. A sweet-tasted mucilaginous substance remains dissolved in the water.

Chap. I.
Analysis of
wheat flour.

The starch obtained by this process is not altogether free from gluten; hence its colour is not very white, and it has not that fine crystallized appearance which distinguishes the starch of commerce.* Manufacturers employ a more economical and more efficacious process. Good wheat is allowed to steep in cold water till it becomes soft, and yields a milky juice when squeezed. It is then taken out of the water; put into coarse linen sacks, which are subjected to pressure in a vat filled with water; a milky juice containing abundance of starch exudes, and mixes with the water of the vat. This process is repeated as long as the wheat yields any milky juice. The sack and its contents are then removed. The starch soon falls to the bottom of the vat; and the water which covers it gradually ferments, in consequence of the substances which it holds in solution. Alcohol and vinegar are formed in it, partly, no doubt, at the expense of the starch. The vinegar, thus evolved, dissolves all the impurities, and leaves nothing behind but starch. It is then poured off, and the starchedulcorated with water. It is afterwards dried by a moderate heat. During the drying it usually splits into small columnar masses, which have a considerable degree of regularity. The water which has stood over the starch was analysed by Vauquelin. It contains a considerable portion of alcohol and of acetic acid. The acid holds in solution gluten somewhat altered, phosphate of lime, and ammonia.†

Method of
making
starch.

* Its colour may be much improved by digesting it in an alkaline ley of moderate strength.

† Ann. de Chim. xxxviii. 248. See La Fabrique de l'Amidon, by Du Hamel de Monceau. See also Gmelin's Handbuch der Technischen Chemie, ii. 737. The reader will find a description of the process followed by our manufacturers in making starch in Phil. Mag. xxix. 166.

Book IV.

Starch was well known to the ancients. Pliny informs us, that the method of obtaining it was first invented by the inhabitants of the island of Chio.*

Properties.

1. Starch has a fine white colour, and is usually concreted in longish masses; it has scarcely any smell, and very little taste. When kept dry, it continues for a long time uninjured though exposed to the air.

Action of water,

2. Starch does not dissolve in cold water, but very soon falls to powder, and forms with it a kind of emulsion. It combines with boiling water, and forms with it a kind of jelly, which may be diffused through boiling water; but when the mixture is allowed to stand a sufficient time, the starch slowly precipitates to the bottom. This subsidence takes place even when 90 parts of water are employed to dissolve one of starch; but in that case, at least a month elapses before the starch begins to precipitate. This solution (if it be entitled to the name) is glutinous in proportion to the quantity of starch. If the quantity be considerable, linen dipped into it, and suddenly dried, acquires (as is well known) a great degree of stiffness. When the solution is evaporated to dryness, a brittle opaque substance is obtained, differing in appearance from common starch, but exhibiting nearly the same properties with reagents. Hence the apparent difference is probably owing to a portion of water remaining united to the boiled starch. When the solution of starch is left exposed to damp air, it soon loses its consistency, acquires an acid taste, and its surface is covered with mould.

Alcohol,

3. Starch is so far from dissolving in alcohol, even when assisted by heat, that it does not even fall to powder. Neither does starch dissolve in ether.

4. The action of oxygen gas and of the simple combustibles and azote upon starch has not been examined, but it is probable that it is not remarkable.

Metallic solutions,

5. The metals seem to have but little action on starch; neither does it combine with the metallic oxides; at least none of the metallic salts which I have tried, have the property of forming a precipitate when dropped into a solution of starch in water. The following were the metallic salts which I tried. The decoction used was made by boiling together one part of starch and 90 parts of water.

* Lib. xviii. cap. 7.

Nitro-muriate of gold, of platinum.
 Nitrate of silver, of mercury, of lead.
 Muriate of tin; acetate of lead.
 Sulphate, muriate, and nitrate of copper.
 Sulphate, muriate, and nitrate of iron.
 Sulphate, muriate, and nitrate of zinc.
 Ammoniated nickel and cobalt.

Dr. Bostock found that sub-acetate of lead throws down a copious precipitate from the solution of starch in water. He obtained likewise a precipitate when permuriate of tin was used.* Hence it is obvious that his starch infusion must have contained a much greater proportion of starch than mine, which formed no precipitate with muriate of tin.

6. When lime-water is mixed with the same decoction of starch, no change is produced, neither is any perceptible alteration occasioned by strontian water; but barytes water throws down a copious white flaky precipitate. This precipitate is redissolved by muriatic acid; but appears again on standing, unless a considerable excess of acid be employed. Yet muriate of barytes occasions no change in the decoction of starch. Silicated potash occasions no precipitate in this decoction.

7. When starch is triturated with the hot infusion of nutgalls, a complete solution is effected. The solution is transparent, and rather lighter coloured than the infusion of nutgalls. When this solution cools it becomes opaque, and a copious curdy precipitate falls. A heat of 120° redissolves this precipitate, and renders the solution transparent, but the precipitate is again deposited when the solution cools. This property I consider as characteristic of starch. The infusion of nutgalls throws it down from every solution, but the precipitate is redissolved by heating the liquid to 120° .

This precipitate is a compound of tannin and starch, and the constituents are capable of uniting in various proportions; but it is least soluble when deposited from the solution of 24 grains of starch in a half-ounce measure of the infusion of nutgalls, made by boiling together two parts of water and one of nutgalls. In that case it is composed of about three parts starch and two tannin. The whole of it does not precipitate; about $\frac{1}{4}$ th remains in solution, seemingly containing

* Nicholson's Jour. xviii. 33.

Book IV.

an excess of tannin. This compound of starch and tannin is of a light-brownish colour, semitransparent, and brittle, and has a good deal of resemblance to common sarcocol. Its taste is astringent. It feels glutinous between the teeth like gum. In cold water it dissolves very imperfectly, but very copiously and readily in hot water. Alcohol, when digested on it, acquires a brown colour, but is incapable of separating the whole of the tannin from the starch. When this compound is heated it froths, swells, and melts, and then burns with a clear flame, leaving behind it a small portion of white ashes.

Action of
potash,

8. When potash is triturated with starch, and a little water added, the whole assumes, on standing, the appearance of a semitransparent jelly. On adding water, an opal coloured solution is obtained, from which the starch is readily thrown down by an acid. When muriatic acid is employed a peculiar aromatic odour is exhaled. When the infusion of galls is dropped into the solution of starch in potash, a yellowish-white precipitate appears, but is immediately redissolved, and the liquid remains opaque, and of a dark brown colour. On adding muriatic acid a copious precipitate appears, similar to the compound of starch and tannin; but nitric acid occasions no precipitate, neither does ammonia.

The decoction of starch is neither altered by potash, carbonate of potash, nor ammonia.

Of acids,

9. When starch is thrown into any of the mineral acids, at first no apparent change is visible: But if an attempt is made to reduce the larger pieces while in acids to powder, they resist it, and feel exceedingly tough and adhesive. Sulphuric acid dissolves it slowly, and at the same time a smell of sulphurous acid is emitted; and such a quantity of charcoal is evolved, that the vessel containing the mixture may be inverted without spilling any of it. Indeed if the quantity of starch be sufficient, the mixture becomes perfectly solid. Diluted sulphuric acid does not act sensibly on starch while cold; but when heated it dissolves it completely, apparently without decomposition, as the starch may be again thrown down by means of alcohol. Sulphurous acid has no effect upon starch.

Diluted nitric acid reduces starch to powder, and then slowly dissolves it without the assistance of heat. The acid assumes a green colour, and a small portion of white matter swims on the surface, on which the acid does not act. Alcohol

throws down the starch from this solution.* Concentrated nitric acid dissolves starch pretty rapidly, assuming a green colour, and emitting nitrous gas. The solution is never complete, nor do crystals of oxalic appear unless heat be applied. In this respect starch differs from sugar, which yields oxalic acid with nitric acid, even at the temperature of the atmosphere. When heat is applied to the solution of starch in nitric acid, both oxalic and malic acid are formed, but the undissolved substance still remains. When separated by filtration, and afterwardsedulcorated, this substance has the appearance of a thick oil, not unlike tallow; but it dissolves readily in alcohol. When distilled it yields acetic acid, and an oil having the smell and the consistence of tallow.†

Strong muriatic acid dissolves starch slowly and without effervescence. When the starch does not exceed $\frac{1}{10}$ th of the acid, the solution is colourless and transparent; but if we continue to add starch, a brown colour soon appears, and the acid loses a portion of its liquidity. Its peculiar smell is destroyed, and replaced by the odour which may be perceived in corn mills.

Acetic acid does not dissolve starch. I have not tried the action of any of the other acids.

10. Alcohol separates starch in part from its decoction. A solution of potash in alcohol occasions a copious white precipitate, which is redissolved on adding a sufficient quantity of water. A solution of sulphuret of potash in alcohol occasions a flaky precipitate in the decoction of starch. This precipitate has sometimes an orange colour.

11. When starch is thrown upon a hot iron, it melts, blackens, froths, swells, and burns with a bright flame like sugar, emitting at the same time a great deal of smoke; but it does not explode, nor has it the caramel smell which distinguishes burning sugar. When distilled it yields water impregnated with an acid, supposed to be the pyromucous, a little empyreumatic oil, and a great deal of carbonic acid and carburetted hydrogen gas. The charcoal which remains is easily dissipated when set on fire in the open air; a proof that it contains very little earth. When starch is torrifed in a flat

* Vauquelin has applied this property of nitric acid to detect starch in vegetables. He digested the substance under examination in diluted nitric acid for two days, and then poured into the solution alcohol. The starch precipitated. See Ann. de Chim. lv. 28.

† Scheel, Crell's Annals, ii. 14. English Trans.

Book IV. vessel till it assumes a grey colour it becomes partly soluble in water, and if the torrefaction be continued till the starch becomes brown it is wholly soluble in cold water. The solution is glutinous, and may be applied to most of the purposes of gum. But tannin still continues to precipitate it; a proof that it has not been converted into gum; but merely undergone an alteration which renders it soluble.*

Constituents
of starch.

12. Three sets of experiments have been made in order to determine the constituents of starch. Gay-Lussac and The-
nard burnt a quantity of wheat starch made up into a paste with chlorate of potash, and ascertained the substances formed. According to their experiments 100 parts of wheat starch are composed of

Hydrogen	6.77
Carbon	43.55
Oxygen	49.68
		100.00†

M. Theodore de Saussure analysed wheat starch by burning it in oxygen gas, ascertaining the proportion of oxygen gas consumed, and the quantity of carbonic acid gas formed. According to him 100 parts of starch are composed of

Azote	0.40
Hydrogen	5.90
Carbon	45.39
Oxygen	48.31
		100.00‡

Professor Berzelius analysed a quantity of potatoe starch by burning it along with chlorate of potash, and using various precautions to ensure an accurate result. According to his experiments 100 parts of potatoe starch are composed of

Hydrogen	7.064
Carbon	43.481
Oxygen	49.455
		100.000§

* This is farther confirmed by the recent experiments of Lassaigne. He found that roasted starch when treated with nitric acid yielded oxalic acid without any trace of saccharic acid. See *Annals of Philosophy*, xvi. 87.

† *Recherches Physico-chimiques*, ii. 291.

‡ *Annals of Philosophy*, vi. 427.

§ *Ibid.* v. 273.

To be able from these experiments to determine the constitution of starch, it is requisite to know the equivalent number, denoting the weight of it which enters into combinations. Berzelius united it with oxide of lead by digesting an excess of subnitrate of lead in a boiling solution of starch in water. The precipitate was washed repeatedly in boiling water, in order to remove any excess of subnitrate of lead that might be contained in it. The amylate of lead thus formed was composed of

Starch	72	100	36
Oxide of lead ..	28	38.89	14

If this compound consisted of 1 atom of starch, united with 1 atom of oxide of lead, the weight of an atom of starch will be 36. But it is more probable, from the process employed to form it, that it is a binamylate of lead, or a compound of 2 atoms starch and 1 atom protoxide of lead. On this supposition an atom of starch must weigh 18. The nearest number to this formed of the requisite proportion of atoms to correspond with the preceding analyses is 17.75. If we consider this as the weight of an atom of starch its constituents will be

10 atoms hydrogen	=	1.25	7.04
10 atoms carbon	=	7.50	42.25
9 atoms oxygen	=	9.00	50.71
		17.75		100.00

13. M. Theodore de Saussure has published an elaborate set of experiments on the spontaneous changes which a mixture of starch and water undergo when left to itself, either in vacuo or with the access of air. Nearly one fourth of the whole starch disappeared, and the remaining three fourths were converted into the following substances :

1. Starch sugar.
2. Gum.
3. Amylin.
4. Starchy lignin.
5. Lignin mixed with charcoal.
6. Undecomposed starch.

The starch sugar amounted to nearly half the weight of the starch. The gum ought rather to have been termed roasted starch, as it possessed the characters of that substance.

The *amylin* (called *amidine* by Saussure) is a substance intermediate between gum and starch. It is obtained from the

Book IV.

residuum of the starch after every thing soluble in cold water is removed. Boiling water dissolves the amylin, and leaves it in a state of purity when evaporated to dryness. It is semi-transparent and brittle, and very soluble in water, of the temperature of 144°. This solution assumes a blue colour, with iodine. It is coagulated by subtracetate of lead, and copiously precipitated by barytes water; but not by lime water nor infusion of nutgalls. It dissolves in the aqueous solution of potash, and the solution has no viscosity. From this solution it is thrown down by acids and by alcohol.

The substance called *starchy lignin* by Saussure, was obtained by digesting the insoluble residue of decomposed starch in ten times its weight of an alkaline ley, containing $\frac{1}{12}$ th of potash. A brown coloured solution is obtained, from which dilute sulphuric acid throws down the starchy lignin, under the form of a light brown combustible substance, having the lustre of jet. It gives a blue colour to the aqueous solution of iodine.*

Plants containing it.

14. Starch is contained in a great variety of vegetable substances; most commonly in their seeds or bulbous roots, but sometimes also in other parts. Mr. Parmentier, whose experiments have greatly contributed towards an accurate knowledge of starch, has given us the following list of the plants from the roots of which it may be extracted:

Arctium lappa.	Imperatoria ostrutheum.
Atropa belladonna.	Hyoscyamus niger.
Polygonum bistorta.	Rumex obtusifolius.
Bryonia alba.	———— acutus.
Colchicum autumnale.	———— aquaticus.
Spiræa filipendula.	Arum maculatum.
Ranunculus bulbosus.	Orchis mascula.
Scrophularia nodosa.	Iris pseudacorus.
Sambucus ebulus.	—— foetidissima.
———— nigra.	Orobis tuberosus.
Orchis morio.	Bunium bulbocastanum.

It is found also in the following seeds:

Oats.	Millet.	Peas.
Rice	Chesnut.	Beans.
Maze.	Horse-chesnut.	Acorn.

Indeed the greater number, if not the whole, of the vege-

* Phil. Trans. 1819, p. 29.

table seeds employed by man as an article of food, consists chiefly of starch. But that substance is always combined with some other which serves to disguise its properties; such as sugar, oil, extractive, &c. It is only by processes similar to those described in the beginning of this Section that it is extracted from these substances in a state of tolerable purity. The following substances, which may be considered as varieties of starch, deserve particular attention.

Chap. I.

1. *Potatoe Starch.*

When the potatoe is grated down to a pulp, and placed on a fine searce, if water be poured on it, a great deal of starch passes through the meshes of the searce, and may be collected in proper vessels. When washed with water and dried, it assumes a fine white colour, and possesses all the essential properties of starch. Indeed it goes much further; a smaller quantity being sufficient to form a thick paste with water than is required of wheat starch. It has a very perceptible crystallized appearance, and is much heavier apparently than common starch. It is not likely therefore that it could be employed with the same advantage as a hair powder.

Potatoe starch.

2. *Sago.*

This substance is extracted from the pith of several species of palm in the Moluccas, Philippines, and other East Indian islands. The palm is cut into pieces of five or six feet in length: the woody part is cut off one side, exposing the pith lying, as it were, in the hollow of a canoe. Cold water is poured in, and the pith well stirred; by which means the starch is separated from the fibrous part, and passes through with the water when the whole is thrown on a searce. The sago, thus separated, is allowed to settle; the water is poured off; and when it is half dry it is granulated, by being forced through a kind of funnel. It is said to acquire its grey colour while dried in an artificial heat. This substance is employed as an article of food, and its nourishing properties are well known.*

Sago.

3. *Salop.*

This substance comes from Persia; but is said also to be manufactured in Europe. It is supposed to be the prepared roots of different species of *orchis*, as the *morio*, *mascula*, *bifo-*

Salop.

* Forest's Voyage, p. 39.

Book IV. *lia, pyramidalis*. According to Mout, the bulbous roots of these plants are deprived of their cuticle, baked in an oven for 10 or 12 minutes, which gives them the semitransparency, and then fully dried in a moderate heat.* Like sago, salop is used only as a nourishing article of food. It is said by Dr. Percival to have the property of concealing the taste of salt water.†

4. *Cassava* or *Tapioca*

Cassava. As it is called, is prepared from the roots of the *jatropha manihot*, an American plant. They are peeled and subjected to pressure in a kind of bag made of rushes. The juice that is forced out is a deadly poison, and is employed by the Indians to poison their arrows; but it deposits gradually a white starch, which when properly washed is innocent. What remains in the bag consists chiefly of the same starch. It is dried in smoke, and afterwards passed through a kind of sieve. Of this substance the cassava bread is made.

5. *Sowans*.

Sowans. This very nutritious article of food is made in this country from the husk of oats, by a process not unlike that by which common starch is made. The husk of the oat (called *seeds*) is separated from oat-meal by the sieve. It still retains a considerable portion of farinaceous matter. It is mixed with water, and allowed to remain for some days till the water has become sour. The whole is then thrown upon a sieve. The milky water passes through; but all the husk remains behind. The water thus obtained is loaded with starchy matter, which soon subsides to the bottom. The sour liquor is decanted off, and about an equal quantity of fresh water added. This mixture when boiled forms a very nourishing article of food; and the portion of the sour water which still adheres to the starch gives the whole a pleasant acidity.

It is curious enough that the starch-maker's water, notwithstanding the great quantity of acid which it contains, likewise the still sourer water of sowans, are swallowed greedily by hogs. They fatten upon it.

* Phil. Trans. lix. 2.

† Phil. Mag. xviii. 161.

SECT. XVII.

OF INDIGO.

THIS valuable pigment, one of the capital manufactures of America, is obtained from the leaves of different species of plants; the *indigofera argentea* or *wild indigo*, which is said to yield the finest pigment, but in smallest quantity; the *indigofera disperma* or *Guatemala indigo*; and the *indigofera tinctoria* or *French indigo*, which yields the greatest quantity of indigo, and is therefore preferred by the planter, though its quality be inferior to the indigo obtained from the other species.* For the first accurate account of the mode of preparing indigo from these plants, we are indebted to Labat.† Since his time some improvements have been introduced into the manufacture; but the mode of rearing the plants still continues the same. In the West Indies the seeds are sown in March, in trenches about a foot asunder, and the plant comes into blossom, and is fit for cutting down in May; but in South America about six months elapse before it can be cut. In the West Indies four cuttings are often obtained in the same plant in the course of a year; but in America never more than two, and often only one. The produce continually diminishes after the first cutting; so that it is necessary to renew the plants from seed every year.

Manufac-
ture.

The plants are cut down with sickles, and laid in strata in the *steep*, till it is about three parts full. This is a large cistern of mason work or wood about 16 feet square. Here they are pressed down with planks, and loaded to prevent them from swimming, and covered with water to the height of four or five inches. Here they ferment; and the utmost attention is required to the process. If they be allowed to remain too long, the pigment is spoiled; and if the water be drawn off too soon, much of the indigo is lost. This fermentation, according to Le Blond, goes on best at the temperature of 80°. The water soon becomes opaque, and assumes a green colour; a smell resembling that of volatile alkali is exhaled, and bubbles of carbonic acid are emitted.‡ When the fermentation has conti-

* Edward's History of the West Indies, ii. 275; and Berthollet, Elemens de la Teinture, ii. 37.

† See his account detailed by Dr. Lewis in Neumann's Chemistry, p. 435.

‡ Jour. de Phys. xxxviii. 141.

Book IV.

nued long enough, the liquor is let out into a second cistern, placed lower than the first. This cistern is called the *battery*, and is commonly about 12 feet square and four and a half deep. Here it is agitated for 15 or 20 minutes by means of levers moved by machinery, till the flocculi, beginning to separate, give it a curdled appearance. A quantity of lime water is now poured in, and the blue flocculi are allowed to subside. The lime water is supposed to prevent putrefaction, which otherwise might endanger the pigment; and according to Le Blond, it absorbs carbonic acid, which existed in the liquid, and prevented the indigo from separating. The water is drawn off, and the pigment put to be drained in small linen bags; after which it is put into little square boxes, and allowed to dry in the shade.*

Europe was long supplied with indigo from the West Indies and America; but of late the manufacture has been introduced into Indostan, and carried on with such spirit and success as promises fair to secure to that country in future the supply of the European market.†

The leaves of the *indigofera* yield a green infusion to hot water, and a green powder may be precipitated from it; but unless a fermentation has taken place, neither the colour nor properties of it have any resemblance to indigo. Le Blond affirms that the contact of air is necessary to the fermentation of indigo; but his proofs are insufficient. The separation of carbonic acid is a sufficient proof that chemical decompositions are going on during the fermentation; but of what nature we are ignorant.

The *indigofera* is not the only genus of plants from which indigo may be obtained. Dr. Roxburgh first drew the attention of manufacturers to the *nerium tinctorum*, a tree very common in Indostan, from the leaves of which it may be extracted in abundance. His process is shorter. The leaves are kept in a copper full of water, supported at the temperature of 160°, till they assume a yellowish hue, and the liquid acquires a deep green colour. The liquid is then drawn off, agitated in the usual manner, and the indigo thrown down by lime-water.‡

Indigo may be obtained also from the *isatis tinctoria*, or

* Edward's History of the West Indies, ii. 278.

† See Tennant's Indian Recreations, ii. 142.

‡ Bancroft on Permanent Colours, p. 423.

woad; a plant commonly enough cultivated in Britain, and even found wild in England.* When arrived at maturity, this plant is cut down, washed, dried hastily in the sun, ground in a mill, placed in heaps, and allowed to ferment for a fortnight. It is then well mixed, and made up into balls, which are piled upon each other, and exposed to the wind and sun. In this state they become hot, and exhale a putrid ammoniacal smell. The fermentation is promoted, if necessary, by sprinkling the balls with water. When it has continued for a sufficient time, the woad is allowed to fall to a coarse powder; in which state it is sold as a dye-stuff. By treating woad nearly in the same manner with the *indigofera*, indigo has been obtained from it by different chemists. Justi † was one of the first of these. Borth published a paper on the subject in 1754, ‡ Kulencamp § in 1755, and Edel || a third in 1756. And several experiments on the subject have been published in Germany.**

Chevreul obtained, by digesting *woad* in alcohol, and evaporating the solution, *white* crystalline grains, which gradually became blue when exposed to the atmosphere, and assumed the usual appearance of indigo. These white grains he considers as indigo destitute of oxygen. Their crystallized state is remarkable.††

The method of preparing indigo, and of applying it to the purposes of dyeing, seems to have been very early known in India; but in Europe, though it had been occasionally used as a paint, ‡‡ its importance as a dye-stuff was not understood before the middle of the 16th century. It is not even mentioned in Plietho, a celebrated work on dyeing published in Italy in 1548. The Dutch were the people who first imported it from India, and made its importance known in Europe. Even as late as the middle of the seventeenth century, its use was restricted in different countries. It was prohibited in England during the reign of Queen Elizabeth; and the prohibition was not taken off till the reign of Charles II. It was prohibited also in Saxony. In the edict it is spoken of

* Some think that this is the plant with which the ancient Britons stained their bodies, to make them appear terrible to their enemies.

† Schreber's Sammlungen, i. 68.

‡ In the Halle Newspaper.

§ Schreber's Samml. viii. 443.

|| De Ind. Germ. lit. et. G. sto.

** See Saltzer, Schweigger's Jour. iii. 417; Gehlen, ibid. v. 191, and vi. 1, and viii. 136.

†† Jour. de Phys. lxxvi. 373.

‡‡ Plinii, Lib. xxxv. c. 6.

Book IV. as a corrosive substance, and called *food for the Devil*. Colbert restricted the French dyers to a certain quantity of it.* Soon after its importance began to be generally understood, it was cultivated in Mexico and the West Indies with such success that the indigo from these countries was preferred to every other; but within these few years the East Indian indigo, owing entirely to the enlightened exertions of some of our own countrymen, has recovered its character, and is now imported in considerable quantities to Britain.

Properties. Indigo is a fine light friable substance of a deep blue colour. Its texture is very compact, and the shade of its surface varies according to the manner in which it has been prepared. The principal tints are copper, violet, and blue; the lightest indigo is the best: but it is always more or less mixed with foreign substances, partly owing, doubtless to the carelessness of the preparation, and partly to the bodies which the plant containing indigo yields to water. From the analysis of Bergman, to whom we are indebted for one of the most complete treatises on the properties of indigo which has yet appeared, the purest indigo which he could procure was composed of the following constituents:

47	pure indigo
12	gum
6	resin
22	earth
13	oxide of iron
<hr style="width: 10%; margin: 0 auto;"/>	
100	†

The earth consisted of

10·2	barytes
10·0	lime
1·8	silica
<hr style="width: 10%; margin: 0 auto;"/>	
22·0	†

But in all probability the earth differs in different specimens; for Proust found magnesia in considerable quantity in the specimens which he examined. § The 47 parts of blue pigment are alone entitled to the name of indigo; and to them therefore we shall confine our attention in this section.

* See Bancroft on Permanent Colours, and Berthollet on Dyeing.

† Bergman, v. 36.

‡ Ibid. v. 33.

§ Nicholson's Journal, iii. 325.

Chevreul has lately published the result of his experiments on Guatemala indigo, which shows that the indigo of commerce is a still more complex substance than Bergman found it. Chap. I.

Hot water digested on indigo became yellow. This liquid being distilled gave out ammonia, and let fall a greenish matter which became blue upon exposure to the air, and possessed the properties of indigo. After this matter has separated, flocks precipitated of a green colour, to which Chevreul has given the name of *green matter*. It is scarcely soluble in water, but dissolves in alkaline leys and becomes reddish. When the alkali is saturated by an acid the colour becomes green, and when the liquid is concentrated the green matter precipitates in flocks. Alcohol dissolves it and becomes reddish; but when spread thin upon any substance it still looks green. Alcohol being poured into the water from which the green matter had precipitated, a substance fell of a brownish-yellow colour and a bitter and astringent taste, and which burnt on coals, emitting the odour of empyreumatic vinegar. Constituents
of common
indigo.

The indigo thus treated was digested in alcohol. The alcohol dissolved; 1. *green matter*; 2. *red matter*; 3. *indigo*. The *red matter* agrees with the *green* in all its properties except colour. The colour does not change by solution in alkaline leys and alcohol.

The indigo in the third place was treated with muriatic acid which dissolved, 1. *iron*; 2. *carbonate of lime*; 3. *red matter*.

The following table exhibits the substances taken from Guatemala indigo by these processes, and of consequence its composition :

Water	{	ammonia disoxygenized indigo 12 green matter bitter matter
Alcohol	{	green matter 30 red matter indigo
Muriatic acid.	{	red matter 6 carbonate of lime 2 oxide of iron and alumina 2 silica 3 pure indigo 45
Total		100

Book IV.

Action of
water.

All indigos do not yield the same substances nor the same proportions. In some the *green matter* has a brown colour.*

Indigo is a soft powder, of a deep-blue inclining to purple, without either taste or smell. It undergoes no change though kept exposed to the air. Water does not dissolve any part of it, nor produce any change upon it. Bergman however found, that indigo, when kept long under water, underwent a kind of putrefaction, or at least exhaled a fetid odour.† When heat is applied to indigo it emits a bluish-red smoke, and at last burns away with a very faint white flame, leaving behind it the earthy parts in the state of ashes.‡ The purple smoke, as Chevreul ascertained, crystallizes in needles, and is indigo unaltered. Hence the easiest method of obtaining pure indigo is to sublime it in a platinum crucible.

Neither oxygen nor the simple combustibles have any effect upon indigo, except it be in a state of solution; and the same remark applies to the metallic bodies.

Of alkalies.

The fixed alkaline solutions have no action on indigo, except it be newly precipitated from a state of solution. In that case they dissolve it with facility. The solution has at first a green colour, which gradually disappears, and the natural colour of the indigo cannot be again restored. Hence we see, that the alkalies when concentrated decompose indigo. § Pure liquid ammonia acts in the same way. Even carbonate of ammonia dissolves precipitated indigo, and destroys its colour; but the fixed alkaline carbonates have no such effect.||

Earths.

Lime-water has scarcely any effect upon indigo in its usual state; but it readily dissolves precipitated indigo. The solution is at first green, but becomes gradually yellow. When the solution is exposed to the air, a slight green colour returns, as happens to the solution of indigo in ammonia; but it soon disappears.** The effect of the other alkaline earths upon indigo has not hitherto been tried; but it cannot be doubted that they would act nearly as lime-water, but with more energy. The other earths seem to have but little action on indigo in any state.

Acids.

The action of the acids upon indigo has been examined with most attention, and it certainly exhibits the most important phenomena.

* Chevreul, Jour. de Phys. lxx. 309.

† Opusc. v. 29.

‡ Ibid. p. 31.

§ Ibid. 24.

|| Ibid.

** Ibid. 25.

When diluted sulphuric acid is digested over indigo, it produces no effect except that of dissolving the impurities; but concentrated sulphuric acid dissolves it readily. One part of indigo, when mixed with eight parts of sulphuric acid, evolves heat, and is dissolved in about 24 hours. The solution of indigo is well known in this country by the name of *liquid blue*. Bancroft calls it *sulphate of indigo*. While concentrated, it is opaque and black; but when diluted, it assumes a fine deep-blue colour, and its intensity is such, that a single drop of the concentrated sulphate is sufficient to give a blue colour to many pounds of water.* Bergman ascertained the effect of different reagents on this solution with great precision. His experiments threw light, not only on the properties of indigo, but upon the phenomena that take place when it is used as a dye-stuff. The following is the sum of these experiments:

Chap. I.
Sulphuric
acid.

Effect of re-
agents on
the sulphate.

1. Dropped into diluted water. No change.
2. Dropped into sulphurous acid. Colour at first blue, then green, and very speedily destroyed.
3. In weak nitric acid it experiences no change.
4. In muriatic acid. No change.
5. In weak tartaric acid. Becomes gradually green, and in 144 hours had assumed a very pale yellow colour. Colour not restored by alkalies.
6. In vinegar. Becomes green, and in four weeks the colour disappeared.
7. In weak potash. Becomes green and then colourless.
8. In weak carbonate of potash. The same changes, but more slowly. If the solution be very weak, the colour of the indigo is not destroyed.
9. Soda. The same as potash.
10. In ammonia and its carbonate. Colour becomes green, and then disappears.
11. In a weak solution of sulphate of soda. Colour after some weeks becomes green.
12. In nitre, common salt, and sal ammoniac. No change.
13. In tartrate of potash. Became green and then colourless.
14. In a solution of sugar. Became green and at last yellowish.
15. In alum. No change.
16. In sulphate of copper. No change.

* Bergman, v. 7.

Book IV.

17. In sulphate of iron. Colour became green, and in three weeks disappeared.

18. In the sulphurets. Colour destroyed in a few hours.

19. Realgar, white oxide of arsenic, and orpiment, produced no change.

20. Black oxide of manganese destroyed the colour completely.

21. In the infusion of madder. Colour became green, and at last yellow.

22. In the infusion of woad. The same changes, but more speedily.*

Blue colour
of indigo
owing to
oxygen.

From these experiments, it is obvious that all those substances which have a very strong affinity for oxygen give a green colour to indigo, and at last destroy it. Hence it is extremely probable, that indigo becomes green by giving out oxygen. Of course it owes its blue colour to that principle. This theory was first suggested by Mr. Haussman, and still farther confirmed by Berthollet.† Now it is only when green that it is in a state capable of being held in solution by lime, alkalies, &c. in which state it is applied as a dye to cloth. The cloth, when dipped into the vat containing it thus dissolved, combines with it, and the blue colour is restored by exposure to the atmosphere. It may be restored equally by plunging the cloth into chlorine dissolved in water. Hence the restoration is probably owing to oxygen. Hence then the reason that sulphurous acid, the vegetable acids, sulphate of iron, give sulphate of indigo a green colour.

From these experiments, we see also that the colour of indigo is destroyed by the addition of those substances which part with oxygen very readily, as the black oxide of manganese. In that case the indigo is destroyed, for its colour cannot be again restored. When the sulphate of indigo is poured into boiling water, it affords a green coloured solution; but with cold water a deep blue solution. What is called smoking sulphuric acid dissolves indigo much more readily than the pure acid, and evolves much more heat during the solution. Bucholz has shown, that by boiling sulphur in pure sulphuric acid, it acquires the property of dissolving indigo as readily as the smoking acid.‡

* Bergman, v 8.

† See Jour. de Phys. 1788; Berthollet's Notes on Poerner sur l'Art de la Teinture, p. 173; and Berthollet on Dyeing.

‡ Gehlen's Journal, iii. 1

Nitric acid attacks indigo with great violence; the evolution of abundance of heat and nitrous gas. When of the specific gravity 1.52, it even sets fire to indigo. This fact was discovered by Mr. Woulfe. He communicated it to Rouelle, who exhibited it in his lectures. Sage first published an account of it.* When the acid is diluted, the action is still violent, unless the proportion of water be considerable. Mr. Hatchett poured upon 100 grains of indigo an ounce of nitric acid diluted with an equal quantity of water. The action was so rapid, that he found it necessary to add another ounce of water. When the effervescence had nearly subsided, the liquid was placed on a sand bath for some days, and evaporated to dryness. Water poured upon the residuum dissolved a considerable portion of it, and formed a beautiful deep-yellow solution of an intense bitter taste. This solution contains only a very small portion of oxalic acid; but with a solution of isinglass it forms a copious yellow insoluble precipitate, and of course contains a portion of artificial tannin. With ammonia, crystals precipitate, consisting of *bitter principle* combined with ammonia.†

Chap. I.
Action of
nitric acid.

Forms
tannin.

When four parts of nitric acid are boiled upon one part of indigo, the pigment soon loses its colour, and is dissolved. The solution becomes yellow, and a thin layer of a resinous matter appears on the surface. If the process be now stopped, the resinous matter becomes firm by cooling. If this matter be removed, and the solution evaporated to the consistence of honey, redissolved in hot water, and filtered, potash throws down yellow spicular crystals, consisting of *bitter principle* combined with potash. These crystals have the curious property of detonating with a purple light when wrapped up in paper and struck with a hammer; the resin, by treating it with fresh nitric acid, may be converted into the same bitter principle. If the process be stopped sooner than the point mentioned above, yellow crystals are obtained, which are more soluble in water, and which sublime in white needles, having all the properties of benzoic acid. These curious facts were ascertained by Fourcroy and Vauquelin.‡

And benzoic
acid.

* Proust, Nicholson's Journal, iii. 328.

† Hatchett's Additional Experiments on Artificial Tannin, Phil. Trans. 1805.

‡ Phil. Mag. xxiii. 256.

Book IV.

Thus it appears, that by treating indigo with nitric acid, it is converted into tannin, oxalic acid, benzoic acid, and *bitter principle*.

Action of other acids.

Muriatic acid does not act upon indigo in its common state, but it readily dissolves indigo precipitated from the sulphate, and forms a blue coloured solution. The same phenomena are exhibited by the phosphoric, acetic, tartaric acids, and probably by all except the acid supporters.*

Chlorine destroys the colour of indigo as readily as nitric acid, and obviously for the same reason.†

Alcohol dissolves a small proportion of indigo, but it gradually precipitates again unless *red matter* be present, in which case the solution is permanent. ‡

Indigo is not acted upon by ether or oils, at least if the experiments of Bergman be accurate. §

Fermentation.

When indigo is mixed up with bran, woad, and other similar substances which readily undergo fermentation, it assumes a green colour during the fermentation, and is then easily dissolved by lime or potash. It is by this process that it is usually rendered proper for dyeing.

Action of heat.

When indigo is distilled, it yields products different from any other vegetable substance, if the accuracy of Bergman is to be trusted. He distilled 576 grains in a small retort connected with a pneumatic apparatus. He obtained the following products:

19 grains carbonic acid gas.
 173 ——— of a yellow acid liquid containing ammonia.
 53 ——— oil.
 331 ——— charcoal.
 ———
 576 ||

He observed no other gas but carbonic acid. The experiment ought to be repeated: if it prove true, no other vegetable substance agrees with it in this respect except gallic acid. The oil was black, and had the smell of tobacco. It dissolved readily in alcohol. The quantity of charcoal appears enormous; but Bergman's experiment was made with common indigo: therefore 158 grains of earth and iron are to be

* Bergman, v. 22.

† Chevreul, Jour. de Phys. lxx. 309.

‡ Bergman, v. 35.

† Ibid.

§ Opusc. v. 35.

subtracted from the sum. This decomposition indicates the presence of oxygen, carbon, hydrogen, and azote in indigo. The proportion of the two first constituents is uncommonly great.

Chevreul subjected common indigo to distillation. The products obtained were, 1. ammoniacal water. 2. Sulphur, probably combined with oily hydrogen. 3. A thick brown oil containing carbonate and acetate of ammonia. 4. Prussiate and hydroguretted sulphuret of ammonia. 5. A purple substance crystallized in silky threads on the dome of the retort. It was pure indigo. 6. A bulky charcoal yielding prussic acid when coloured with potash. 7. Gases not examined.

To ascertain the constituents of indigo I had recourse to the indigo vat of the calico-printer, in which the indigo is deprived of its blue colour by means of sulphate of iron, and is held in solution by means of lime-water. The colour of the liquid in this vat is yellow, but it becomes immediately blue, and the indigo precipitates when the liquid is exposed to the air. After digesting the indigo, thus regenerated in muriatic acid and drying it sufficiently, I considered it as pure, and exposed it to a red heat, mixed with a sufficient quantity of peroxide of copper. One grain of indigo yielded

Composition of indigo.

Cubic Inches.

Carbonic acid gas 3.23
Azotic gas 0.46

Now $0.46 \times 7 = 3.22$. Hence we see that indigo contains 7 atoms of carbon to 1 atom of azote. The weight of the carbon and azote in these gases is,

Carbon	0.408
Azote	0.136
	.544
Loss456
	1.000

This loss must be oxygen. It is equivalent to 6 atoms. Thus the constituents of indigo are as follows :

7 atoms carbon	= 5.25
6 atoms oxygen	= 6.00
1 atom azote	= 1.75
	13.00

Book IV.

So that an atom of indigo weighs 13. It is possible that the azote which I detected may have been owing to the presence of some impurity in the indigo; but I have never succeeded in obtaining any indigo which did not yield azote:

By exposing the yellow solution of indigo to a given bulk of oxygen gas till the indigo was precipitated blue, and noting the bulk of gas absorbed, and the weight of indigo present, I ascertained that indigo becomes soluble in alkalies, and loses its blue colour when deprived of a single atom of oxygen. Hence the soluble basis of indigo is composed of

7 atoms carbon	= 5.25
5 atoms oxygen	= 5.00
1 atom azote	= 1.75
	<hr style="width: 10%; margin: 0 auto;"/>
	12.00 *

Such are the properties of indigo, as far as they have been ascertained. To descant on its importance is quite unnecessary. It is perhaps the most valuable, and certainly the most permanent, of all the dye-stuffs.

SECT. XVIII.

OF GLUTEN.

Extraction. IF wheat flour be kneaded into a paste with a little water, it forms a tenacious, elastic, soft, ductile mass. This is to be washed cautiously, by kneading it under a small jet of water till the water no longer carries off any thing, but runs off colourless; what remains behind is called *gluten*. It was discovered in 1742 by Beccaria, an Italian philosopher, to whom we are indebted for the first analysis of wheat flour.†

Properties. 1. Gluten, when thus obtained, is of a grey colour, exceedingly tenacious, ductile, and elastic, and may be extended to twenty times its original length. When very thin, it is of a whitish colour, and has a good deal of resemblance to animal tendon or membrane. In this state it adheres very tenaciously to other bodies, and has been often used to cement together broken pieces of porcelain. Its smell is peculiar. It has scarcely any taste, and does not lose its tenacity in the mouth. In the air, it assumes a brown colour, and becomes as it were covered with a coat of oil.

* Annals of Philosophy, xv. 466.

† Collect. Academ. xiv. 1.

When exposed to the air, it gradually dries; and when completely dry, it is pretty hard, brittle, slightly transparent, of a dark brown colour, and has some resemblance to *glue*. It breaks like a piece of glass, and the edges of the fracture resemble in smoothness those of broken glass; that is to say, it breaks with a *vitreous* fracture.

Fresh gluten imbibes water, and retains a certain quantity of it with great obstinacy. To this water it owes its elasticity and tenacity. When boiled in water it loses both these properties. Soluble in water.

2. M. Taddey, an Italian chemist, has lately ascertained that fresh gluten of wheat may be decomposed into two principles, which he has distinguished by the names of *gliadin* (from $\gamma λ ι α$, gluten,) and *zimome* (from $\xi ζ υ μ η$, ferment). They are obtained in a separate state by kneading the fresh gluten in successive portions of alcohol, as long as that liquid continues to become milky when diluted with water. The alcohol solutions being set aside gradually deposite a white matter, consisting of small filaments of gluten, and become perfectly transparent. Being now left to slow evaporation, the *gliadin* remains behind of the consistence of honey, and mixed with a little yellow resinous matter, from which it may be freed by digestion in sulphuric ether, in which gliadin is not sensibly soluble. The portion of gluten not dissolved by the alcohol is the *zimome*.

Gliadin when dry has a straw yellow colour, is slightly transparent, in thin plates, brittle, with a slight honeycomb smell, and when slightly heated gives out an odour similar to that of boiled apples. In the mouth it becomes adhesive, and has a sweetish and balsamic taste. It is pretty soluble in boiling alcohol, but the solution loses its transparency in proportion as it cools, and when cold it retains only a small quantity in solution. It forms a kind of varnish on those bodies to which it is applied. It softens, but does not dissolve in distilled water. At a boiling heat it is converted into froth, and the liquid remains slightly milky. It is specifically heavier than water. Gliadin.

The alcoholic solution of gliadin becomes milky when mixed with water, and is precipitated in white flocks by the alkaline carbonates. It is scarcely affected by the mineral and vegetable acids. Dry gliadin dissolves in caustic alkalies and in acids. It swells upon red hot coals, and then contracts in the manner of animal substances. It burns with a pretty

Book IV. lively flame, and leaves behind it a light spongy charcoal difficult to incinerate. It is very sensibly affected by the infusion of nutgalls. It is capable of undergoing a slow fermentation, and produces fermentation in saccharine solutions.

Zimome.

The gluten thus treated with alcohol is reduced to about the third of its former bulk. This diminution is owing not merely to the separation of the gliadin, but likewise of the water which it formerly contained. The zimome thus obtained is in small globules, or constitutes a shapeless mass which is hard, tough, destitute of cohesion, and of an ash-white colour. When washed in water it recovers part of its viscosity, and becomes speedily brown when left in contact of the air. It is specifically heavier than water. Its mode of fermenting is no longer that of gluten; for when it putrefies it exhales a fetid urinous odour. It dissolves completely in vinegar, and in the mineral acids at a boiling temperature. With caustic potash, it combines and forms a kind of soap. When put into lime water or into the solutions of the alkaline carbonates, it becomes harder and assumes a new appearance without dissolving. When thrown upon red hot coals, it exhales an odour similar to that of burning hair or hoofs, and burns with flame. It produces various kinds of fermentation, according to the nature of the substance with which it comes in contact.*

3. When fresh gluten is macerated for a considerable time in cold water the liquid becomes opaque, and contains small films suspended, which do not soon subside. By repeated filtrations it becomes transparent; but it holds in solution a little gluten, which renders it frothy, and gives it the property of precipitating when mixed with chlorine or the infusion of nutgalls. Thus gluten is to a certain extent soluble in cold water. When the water is heated, the gluten separates in the state of yellow flakes.†

When kept moist, it very soon begins to decompose, and to undergo a species of fermentation. It swells, and emits air-bubbles, which Proust has ascertained to consist of hydrogen and carbonic acid gases.‡ It emits also a very offensive odour, similar to what is emitted by putrefying animal bodies. Cadet kept gluten in a vessel for a week in a damp room. Its surface became covered with byssi, the fermenta-

* Annals of Philosophy, xv. 390.

† Fourcroy and Vanquelin, Annales de Mus. d'Hist. Nat. No. xxxvii. p. 3.

‡ Jour. de Phys. lvi. 108.

tion just mentioned had commenced, and the odour was distinctly acid. In 24 days, on removing the upper crust, the gluten was found converted into a kind of paste, of a greyish white colour, not unlike bird-lime. In that state he gave it the name of *fermented gluten*.* If the gluten be still left to itself, it gradually acquires the smell and the taste of *cheese*. This curious fact was first ascertained by Rouelle junior. In that state it is full of holes, and contains the very same juices which distinguish some kinds of cheese. Proust ascertained that it contains ammonia and vinegar; bodies which Vauquelin detected in cheese: and ammonia robs both equally of their smell and flavour.†

Cb. p. 1.

Fermented
Gluten.

Fourcroy and Vauquelin have examined the changes which gluten sustains when left for some months to putrefy under water. When left in that liquor, at the temperature of 54°, it swells, becomes sour and fetid, and rises to the surface, giving out carbonic acid. The water acquires the property of giving out a red colour to paper stained with litmus. The fixed alkalis disengage ammonia, and chlorine and infusion of nutgalls throw down a precipitate, indicating the presence of gluten in solution. When sugar is dissolved in this water, the liquid assumes the properties of vinegar, without fermentation or contact of the atmosphere. If the gluten be put again into water at the same temperature, the fermentation goes on, carbonic acid is disengaged, and the liquid slightly reddens litmus paper. The reagents indicate the presence of gluten and ammonia. The gluten, which forms a pellicle on the surface of the liquid, becomes first purple-coloured, and then blackish-grey, and exhales the fetid odour of putrefying mucous membranes. The water is now blackish; it gives a brown colour to nitrate of silver, blackens nitrate of mercury, becomes milky, and loses its smell when mixed with chlorine, and is not precipitated by infusion of nutgalls. After three months putrefaction, the gluten was greatly diminished; it had a brown colour, and exhaled only a weak smell. When dried it formed lumps, which emitted the odour of the earth from a burial ground. They softened under the fingers, and burned like fat. In alcohol it partly dissolved, leaving behind it a substance which resembled woody fibre in its properties.‡ Thus by putrefaction gluten is converted into *car-*

* Ann. de Chim. xli. 315.

† Proust, Jour. de Phys. lvi. 100.

‡ Ann. de Mus. d'Hist. Nat. No. xxxvii. 4.

Book IV. *lonic acid, ammonia, a fatty matter, and a substance analogous to woody fibre.*

Action of alcohol.

4. Fresh gluten does not sensibly dissolve in alcohol, which even throws it down from water; yet in certain cases this liquid forms a solution of gluten in very small proportion. I poured two ounces of alcohol, of the specific gravity 0.837, upon 100 grains of pearl barley ground to a coarse powder, and allowed it to stand untouched for two years. On evaporating the liquid, it left behind it a yellow matter, which partly resembled an oil. Alcohol digested upon this substance took up the yellow matter and left *gluten*, easily recognizable by its properties. On trying the same experiment with wheat flour, no gluten was taken up, but only a yellow matter, which possessed the properties of *bitter principle*.

When the *fermented* gluten of Cadet is triturated with a little alcohol into a mucilage, and then mixed with a sufficient quantity of that liquid, a portion of it is dissolved. This solution constitutes an excellent varnish, possessed of considerable elasticity. It may be spread over paper or wood; and when dry resists other bodies, as well as most varnishes. In this state, too, it may be employed to cement china; and triturated with paints, especially vegetable colours, it forms a very good ground. When this solution is mixed with a sufficient quantity of lime, it forms a very good lute; and bits of linen dipped in it adhere very strongly to other bodies.*

5. Ether does not sensibly dissolve gluten.

6. Acids act upon gluten differently according to the peculiar properties of each.

Acids.

Concentrated acetic acid dissolves it readily in considerable quantity, and without altering its nature. The solution is muddy, but permanent; and the gluten may be thrown down by means of alkalies. This acid dissolves the fermented gluten of Cadet; and the solution may be substituted for the solution in alcohol as a varnish; but it does not answer to mix it with colours.†

Concentrated sulphuric acid renders it violet-coloured, and at last black; inflammable air escapes, and charcoal, water, and a portion of ammonia, are formed.‡ When nitric acid is poured on it, and heat applied, there is a quantity of azotic gas emitted, as Berthollet discovered; and by continuing the

* Cadet, Ann. de Chim. xli. 316.

† Ibid.

‡ Fourcroy, vii. 304.

heat, some little oxalic acid is formed,* and likewise malic acid, while a number of yellow-coloured oily flakes make their appearance in the solution.

Muriatic acid dissolves gluten with facility when its action is assisted by heat. When gluten is placed in chlorine it softens, and seems to dissolve, but soon coagulates again in yellow-coloured flakes, which become transparent and greenish-coloured by drying. When heated, they exhale chlorine, and assume the appearance of common gluten.† Chlorine has the property of precipitating gluten from water in the state of yellowish-white flakes.

7. Alkalies dissolve gluten when they are assisted by heat.‡ Alkalies. The solution is never perfectly transparent. Acids precipitate the gluten from alkalies, but it is destitute of its elasticity.§ Alkalies, when much concentrated, form with it a kind of soap, converting it into oil and ammonia; which last is dissipated during the trituration.

8. Gluten is precipitated from water, and from some of its other solutions, by the infusion of nutgalls. The colour of the precipitate is usually yellowish-brown, and it does not dissolve though the solution be heated. Nutgalls.

9. The action of the metallic oxides and their salts upon gluten has not been tried.

It has a strong affinity for the colouring matter of vegetables, and likewise for resinous bodies.

10. When moist gluten is suddenly dried, it swells amazingly. Dry gluten, when exposed to heat, cracks, swells, melts, blackens, exhales a fetid odour, and burns precisely like feathers or horn. When distilled, there come over water impregnated with ammonia and an empyreumatic oil; the charcoal which remains is with difficulty reduced to ashes. Heat.

11. The properties of this substance clearly point out a resemblance between it and animal matter; and the phenomena of its fermentation and destructive distillation show us that oxygen, hydrogen, carbon, and azote, are constituents of it. Proust first observed that the vapour which it emits, while fermenting, blackens silver and lead, and of course contains sulphur.

12. Like all other vegetable principles, gluten is susceptible

* Vauquelin, Ann. de Chim. vi. 278.

† Fourcroy and Vauquelin, Ann. de Mus. d'Hist. Nat. No. xxxvii. 3.

‡ By digesting starch in weak alkaline leys, it may be completely freed from gluten.

§ Fourcroy.

Book IV.

of various shades of properties, which constitute so many species. In wheat flour it occurs in the greatest abundance, and from it we can extract it with the greatest ease: but the sagacity and industry of Rouelle and Proust have detected its presence in many other vegetable substances. Rouelle found it in the leaves of all the vegetable substances which he examined. The exactness of this opinion was called in question by Fourcroy, who treated the experiments of Rouelle with contempt; but it has been examined, and confirmed by very decisive experiments of Proust.

Green fecula contains gluten.

When the juice of cabbage leaves, cresses, scurvy-grass, and other similar plants, is extracted by pressure and passed through a cloth, it still continues far from transparent. Its muddiness is owing to a fine soft silky-green powder suspended in it, which subsides to the bottom so slowly as to take at least a week before it is deposited. This green powder has been distinguished by the name of the *green fecula* of plants. Rouelle first examined it with attention, and ascertained its properties; and the subject has been carried still farther by Proust. The slowness with which it subsides shows that its specific gravity does not differ much from that of water. When once it has fallen, it is insoluble. This substance consists chiefly of three principles: 1. A green matter to which it owes its colour, separated by digestion in alcohol, and which possesses the properties of a *resin*. 2. A substance which consists chiefly of *woody fibres*, and which is left behind when the fecula is digested in potash. 3. A species of gluten, which constitutes the greatest part of it, and to which it owes its characteristic properties.

When the juice of the plants is exposed to a heat of about 130°, the green fecula undergoes a kind of coagulation, concreting into large flakes, which subside very quickly. At this temperature albumen is not altered by heat. This is the method commonly taken to *clarify* these juices. We see from it, that the fecula was united to the water by a very small force, which the addition of heat weakened sufficiently to enable the gluten to cohere. This coagulation by heat takes place how diluted soever the juices are with water; which is by no means the case with albumen. It is thrown down also by the addition of a little alcohol, by all acids, by ammonia, by sulphuretted hydrogen gas, or by throwing into the liquid crystals of carbonate of potash, magnesia, common salt, muriate of potash, nitre, sal ammoniac, &c.

When separated from water it soon dries, and becomes

elastic, and somewhat of the appearance of horn: and in that state is scarcely softened by hot water. When treated like gluten, it gradually acquires the cheesy taste and smell. When kept under water it very soon begins to putrify, and exhales a gas which blackens silver and solutions of lead. This speedy putrefaction in stagnant water takes place when flax and hemp are steeped. These substances contain green fecula in their rind, and the putrefaction occasions the separation of the whole, which could not otherwise be accomplished. The water which has been allowed to remain for a whole year over green fecula contains sulphuretted hydrogen, carbonate of ammonia, and gluten seemingly held in solution by the ammonia. The stench of putrefaction still continues even after the water has been boiled.*

13. The number of plants containing gluten is very considerable. Proust found it in acorns, chesnuts, horse-chesnuts, rice, barley, rye, peas, and beans; and in apples and quinces. He found it also in the leaves of cabbage, sedums, cress, hemlock, borage, saffron, &c.; in the berries of the elder, the grape, &c.; in the petals of the rose, &c.† It occurs also in several roots: Proust could find none in the potatoe.

Plants containing gluten.

14. Gluten must be considered as one of the most useful of the vegetable principles. It constitutes an essential ingredient in wheat, and is the substance which renders flour of wheat so fit for forming bread. It seems also to constitute the essential part of yeast. Its uses as a varnish, a ground for paint, &c., pointed out by Cadet, likewise deserve attention.

Albumen and *gluten* have been shown by Link to resemble each other so closely that they cannot be considered as distinct principles.‡

SECT. XIX.

OF POLLENIN.

This is a name given by Dr. John to a peculiar substance found in the pollen of the *pinus abies*, *pinus sylvestris*, *lycopodium clavatum*, and supposed by him to constitute the characteristic constituent of every species of pollen. It seems to have been first recognised in 1806, by Bucholz, in the pollen

* Proust, Jour. de Phys. lvi. 97.

† Ibid.

‡ Annals of Philosophy, vii. 455.

Book IV.

of the lycopodium.* From the experiments of Foureroy and Vanquelin, it appears also to constitute a portion of the pollen of the phœnix daetlifera. John's analysis of the pinus sylvestris, in which he gives the characters of pollenin, as they had been previously given by Bucholz, may be found in the fourth volume of his *Chemical Researches*.

Properties.

Pollenin has a yellow colour, and is destitute of taste and smell. It is not altered by exposure to the air.

It is insoluble in water, alcohol, ether, fat and volatile oils, and likewise in petroleum.

According to John it dissolves in alkaline leys, and the solution has a brown colour. But in the trials of Bueholz this solution did not take place.

When exposed to the heat of a candle it burns with rapidity and flame, as is the case with the pollen of the lycopodium elavatum.

When exposed to the air it speedily assumes the smell and taste of cheese, and soon runs into the putrifaective fermentation, containing a considerable quantity of ammonia. When distilled it yields a good deal of ammonia.

The remaining properties of this substance have not been investigated. It is obvious, however, from the preceding facts that pollenin approaches very nearly to gluten.

SECT. XX.

OF FIBRIN.

THAT peculiar substance which constitutes the fibrous part of the muscles of animals has been called *fibrin* by chemists. A substance resembling it, as it exists in the blood has been detected by Vauquelin in the juice of the papaw tree. Fibrin then must be ranked among vegetable substances.

Found in
papaw.

The *papaw tree*, the *carica papaya* of botanists, grows in Peru, &c.; and in the Isle of France, where the milky juice that exudes from it is said to be employed with efficacy against the *Guinea-worm*. Two specimens of this juice were brought from that island to Paris by Charpentier de Cossigny. In the one, the juice had been evaporated to dryness, and was in the state of an extract; in the other, the juice was preserved by being mixed with an equal bulk of rum. Both were subjected

* Gehlen's Jour. vi: 599.

to a chemical analysis by Vauquelin. The first was of a yellowish-white colour, and semitransparent. Its taste was sweetish. It had no smell, and was pretty solid; but attracted moisture when kept in a damp place. The second was reddish-brown, and had the smell and taste of boiled beef. When the first specimen was macerated in cold water, the greatest part of it dissolved. The solution frothed with soap. The addition of nitric acid coagulated it, and rendered it white; and when boiled, it threw down abundance of white flakes.*

When the juice of the papaw is treated with water, the greatest part dissolves; but there remains a substance insoluble, which has a greasy appearance. It softens in the air, and becomes viscid, brown, and semitransparent. When thrown on burning coals it melted, let drops of grease exude, emitted the noise of meat roasting, and produced a smoke which had the odour of fat volatilized. It left behind it no residuc. This substance was the *fibrin*. The resemblance between the juice of the papaw and animal matter is so close, that one would be tempted to suspect some imposition, were not the evidence that it is really the juice of a tree quite unexceptionable.†

The properties of fibrin are the following:

1. It is tasteless, fibrous, elastic, and resembles gluten.
2. It is insoluble in water and in alcohol.
3. It is not dissolved by diluted alkalies.
4. But acids dissolve it without difficulty.
5. With nitric acid it gives out much azotic gas.
6. When distilled it yields much carbonate of ammonia and oil.
7. It soon putrifies when kept moist, becomes green; but does not acquire any resemblance to cheese.

Properties.

A fuller description of this substance must be reserved for the next book, as it belongs rather to the animal than vegetable kingdom.

SECT. XXI.

OF OILS.

THERE are two species of oils; namely, *fixed* and *volatile*; both of which are found abundantly in plants.

1. Fixed oil is only found in the seeds of plants, and is

* Ann. de Chim. xliii. 267.

† Ibid.

Book IV.

almost entirely confined to those which have two cotyledons; as linseed, almonds, beech-nut, poppy seed, rape seed, &c. Sometimes, though rarely, it is found in the pulp which surrounds the stone of certain fruits. This is the case with the olive, which yields the most abundant and most valuable species of fixed oil. The bicotyledinous seeds, besides oil, contain also a mucilaginous substance; and they have all the character of forming, when bruised in water, a milky liquid, known by the name of *emulsion*.

The following is a list of the plants yielding the fixed oils which usually occur in commerce.*

List of fixed oils.

1. <i>Linum usitatissimum et perenne</i>	..	Linseed oil.
2. <i>Corylus avellana</i>	}	Nut oil.
3. <i>Juglans regia</i>		
4. <i>Papaver somniferum</i>	Poppy oil.
5. <i>Cannabis sativa</i>	Hemp oil.
6. <i>Sesamum orientale</i>	Oil of Scsamum.
7. <i>Olea Europea</i>	Olive oil.
8. <i>Amygdalus communis</i>	Almond oil.
9. <i>Guilandina Mohringa</i>	Oil of behen.
10. <i>Curcubita pepo et melapepo</i>	Cucumber oil.
11. <i>Fagus sylvatica</i>	Beech oil.
12. <i>Sinapis nigra et arvensis</i>	Oil of mustard.
13. <i>Helianthus annuus et perennis</i>	Oil of sunflower.
14. <i>Brassica napus et campestris</i>	Rape seed oil.
15. <i>Ricinus communis</i>	Castor oil.
16. <i>Nicotiana tabacum et rustica</i>	Tobacco seed oil.
17. <i>Prunus domestica</i>	Plum kernel oil.
18. <i>Vitis vinifera</i>	Grape seed oil.
19. <i>Theobroma cacao</i>	Butter of cacao.
20. <i>Laurus nobilis</i>	Laurel oil.
21. <i>Arachis hypogæ</i>	Ground nut oil.†

2. Volatile oils are found in every part of plants except the cotyledons of the seeds, where they never occur. The root, the stem, the leaves, the flower, the rind or pulp of the fruit of a variety of plants, are loaded with volatile oils, from which they are extracted by expression or by distillation. The number of these oils is so great that it baffles all description. Almost every plant which is distinguished by a peculiar

* See Gren's Handbuch, ii. 166.

† Nicholson's Jour. vi. 224.

odour contains a volatile oil, to which it is indebted for that odour.

Chap. I.

The following table contains a pretty copious list of plants which yield volatile oils. The part of the plant from which it is extracted, and the English name of the oil, are added in separate columns.*

Plants.	Parts.	Oil of	Colour.	List of volatile oils.
1. <i>Artemisia absinthium</i>	Leaves	Wormwood	Green	
2. <i>Acorus calamus</i>	Root	Sweet flag	Yellow	
3. <i>Myrtus Pimenta</i>	Fruit	Jamaica pep. §	Yellow	
4. <i>Anethum graveolens</i>	Seeds	Dill	Yellow	
5. <i>Angelica archangelica</i>	Root	Angelica		
6. <i>Pimpinella anisum</i>	Seeds	Anise	White	
7. <i>Illicium anisatum</i>	Seeds	Stellat. anise	Brown	
8. <i>Artemisia vulgaris</i>	Leaves	Mugwort		
9. <i>Citrus aurantium</i>	Rind of the fruit	Bergamotte	Yellow	
10. <i>Meloleuca leucodendra</i>	Leaves	Cajeputt †	Green	
11. <i>Eugenia caryophyllata</i>	Capsules	Cloves §	Yellow	
12. <i>Carum carvi</i>	Seeds	Caraways	Yellow	
13. <i>Amomum cardamomum</i>	Seeds	Card. seeds	Yellow	
14. <i>Carlina acaulis</i>	Roots		White	
15. <i>Scandix chaerifolium</i>	Leaves	Chervil	Sulph. yel.	
16. <i>Matricaria chamomilla</i>	Petals	Chamomile	Blue	
17. <i>Laurus cinnamomum</i>	Bark	Cinnamon §	Yellow	
18. <i>Citrus medica</i>	Rind of the fruit	Lemons	Yellow	
19. <i>Cochlearia officinalis</i>	Leaves	Scurvy grass	Yellow	
20. <i>Copaifera officinalis</i>	Extract	Copaiba	White	
21. <i>Coriandrum sativum</i>	Seeds	Coriand. seed	White	
22. <i>Crocus sativus</i>	Pistils	Saffron §	Yellow	
23. <i>Piper cubeba</i>	Seeds	Cubeb pep.	Yellow	
24. <i>Laurus culilaban</i>	Bark	Culilaban	Brown yel.	
25. <i>Cuminum cyminum</i>	Seeds	Cumin	Yellow	
26. <i>Inula helenium</i>	Roots	Elecampane	White	
27. <i>Anethum fœniculum</i>	Seeds	Fennel	White	
28. <i>Croton eleutheria</i>	Bark	Cascarilla	Yellow	
29. <i>Maranta galanga</i>	Roots	Galanga	Yellow	
30. <i>Hyssopus officinalis</i>	Leaves	Hyssop	Yellow	
31. <i>Juniperus communis</i>	Seeds	Juniper	Green	
32. <i>Lavendula spica</i>	Flowers	Lavender	Yellow	

* See Gren's Handbuch, ii. 204.

§ The oils marked § sink in water.

† I find the specific gravity of this oil 0.9066.

Book IV.

Plants.	Parts.	Oil of	Colour.
33. <i>Laurus nobilis</i>	Berries	Laurel	Brownish
34. <i>Prunus laurocerasus</i>	Leaves	Lauroceras. §	
35. <i>Levisticum ligusticum</i>	Roots	Lovage	Yellow
36. <i>Myristica moschata</i>	Seeds *	Mace	Yellow
37. <i>Origanum majorana</i>	Leaves	Majorum	Yellow
38. <i>Pistacia lentiscus</i>	Resin	Mastich	Yellow
39. <i>Matricaria parthenium</i>	Plant	Motherwort	Blue
40. <i>Melissa officinalis</i>	Leaves	Balm	White
41. <i>Mentha crispa</i>	Leaves		White
42. ——— <i>piperitis</i>	Leaves	Peppermint	Yellow
43. <i>Achillea millefolium</i>	Flowers	Millefoil	{ Blue and green
44. <i>Citrus aurantium</i>	Petals	Neroli	Orange
45. <i>Origanum creticum</i>	Flowers	Spanish hop	Brown
46. <i>Apium petroselinum</i>	Roots	Parsley	Yellow
47. <i>Pinus sylvestris et abies</i>	{ Wood & resin.	{ Turpentine	Colourless
48. <i>Piper nigrum</i>	Seeds	Pepper	Yellow
49. <i>Rosmarinus officinalis</i>	Plant	Rosemary	Colourless
50. <i>Mentha pulegium</i>	Flowers	Pennyroyal	Yellow
51. <i>Genista canariensis</i>	Root	Rhodium	Yellow
52. <i>Rosa centifolia</i>	Petals	Roses	Colourless
53. <i>Ruta graveolens</i>	Leaves	Rue	Yellow
54. <i>Juniperus sabina</i>	Leaves	Savine	Yellow
55. <i>Salvia officinalis</i>	Leaves	Sage	Green
56. <i>Santalum album</i>	Wood	Santalum §	Yellow
57. <i>Laurus sassafras</i>	Root	Sassafras	Yellow
58. <i>Satureia hortensis</i>	Leaves	Satureia	Yellow
59. <i>Thymus serpyllum</i>	{ Leaves and flower	{ Thyme	Yellow
60. <i>Valeriana officinalis</i>	Root	Valerian	Green
61. <i>Kæmpferia rotunda</i>	Root	Zedoary	{ Greenish blue
62. <i>Amomum Zinziber</i>	Root	Ginger	Yellow
63. <i>Andropogon schænan- thum</i>	{	Sira	Brown

Several of the gum resins, as *myrrh* and *gallbanum*, yield likewise an essential oil, and likewise the balsams, as *ben-zoin*, &c.

§ The oils marked § sink in water.

* They yield also a fixed oil.

SECT. XXII.

OF WAX.

THE upper surface of the leaves of many trees is covered with a varnish, which may be separated and obtained in a state of purity by the following process:

Chap. I.

Digest the bruised leaves, first in water and then in alcohol, till every part of them which is soluble in these liquids be extracted. Then mix the residuum with six times its weight of a solution of pure ammonia, and, after sufficient maceration, decant off the solution, filter it, and drop into it, while it is incessantly stirred, diluted sulphuric acid, till more be added than is sufficient to saturate the alkali. The varnish precipitates in the form of a yellow powder. It should be carefully washed with water, and then melted over a gentle fire.

How obtained from plants.

Mr. Tingry first discovered that this varnish possessed all the properties of *bees wax*:† Wax, then, is a vegetable product. Several plants contain wax in such abundance as to make it worth while to extract it from them. But let us, in the first place, consider the properties of bees wax, the most common and important species. This substance, as Huber has demonstrated, contrary to the generally received opinion, is prepared by the bees from honey or sugar, the latter yielding the greatest proportion of it.‡

1. Wax, when pure, is of a whitish colour; it is destitute of taste, and has scarcely any smell. Bees wax indeed has a pretty strong aromatic smell; but this seems chiefly owing to some substance with which it is mixed; for it disappears almost completely by exposing the wax, drawn out into thin ribands, for some time to the atmosphere. By this process, which is called *bleaching*, the yellow colour of the wax disappears, and it becomes very white. Bleached wax is not affected by the air.§

Properties of bees wax.

2. The specific gravity of unbleached wax varies from 0.9600|| to 0.9650;*** that of white wax from 0.8203 to 0.9662.**

* Hermbstadt, Med. and Phys. Jour. iii. 372.

† Enc. Meth. Forets et Bois, i. 100.

‡ Nicholson's Jour. ix. 182.

§ Sennebier, Ann. de Chim. xii. 60. and Jour. de Phys. xxxviii. 56.

|| Bostock, Nicholson's Jour. iv. 130.

** Fabroni, Crell's Annals, 1797, ii. 125.

Book IV.

3. Wax is insoluble in water; nor are its properties altered though kept under that liquid.

4. When heat is applied to wax it becomes soft; and at the temperature of 142° , if unbleached, or of 155° if bleached,* it melts into a colourless transparent fluid, which concretes again, and resumes its former appearance as the temperature diminishes. If the heat be still farther increased, the wax boils and evaporates; and if a red heat be applied to the vapour, it takes fire and burns with a bright flame. It is this property which renders wax so useful for making candles.

Action of alcohol.

5. Wax is scarcely acted on by alcohol when cold, but boiling alcohol dissolves it. This was known to Neumann,† and has been lately verified by Dr. Bostock.‡ Rather more than 20 parts of alcohol are necessary to dissolve 1 part of wax; and as the solution cools, the greater part of the wax precipitates, and the remainder is thrown down by water.§

6. Ether has but little action on wax while cold; but when assisted by heat, it takes up about $\frac{1}{20}$ th of its weight of it, and lets the greatest part precipitate on cooling.||

Dr. John ascertained that when bees wax, or the wax of the *myrica cordifolia* is digested in a sufficient quantity of boiling alcohol, it is divided into two distinct substances. One of these substances is taken up by the boiling alcohol while the other remains undissolved. To the first of these bodies he has given the name of *cerin*, to the second that of *myricin*.

Cerin.

Cerin he distinguishes by the following properties. It is soluble in fixed and volatile oils: from the last it precipitates in greasy looking grains. Insoluble in water, cold alcohol, and ether; soluble in hot alcohol and ether, and precipitating again as the liquids cool. Melts between 108° and 140° . Of the specific gravity 1.000. Of the consistence of wax. Unites with caustic alkalies, and forms a soap.

Myricin.

The characters of *myricin* are as follows. Insoluble in water, ether, and alcohol, both hot and cold. Soluble in fixed and volatile oils: does not precipitate from the former of these. Melts between 100° and 140° . Somewhat glutinous; but of the consistence of wax. Specific gravity 0.900.**

Chevreul has given the name of *cerin* to a substance which

* Bostock, Nicholson's Jour. i. 71.

† Neumann's Chemistry, p. 331.

‡ Nicholson's Jour. iv. 133.

§ Nicholson's Jour. iv. 133.

|| Id. ibid.

** John's Chemische tabellen der Pflanzenanalysen, p. 9, and Chemische Untersuchungen, iii. 42.

be extracted from cork, and which possessed the following properties. It is white and in small needles. It does not melt in boiling water; but becomes soft and sinks to the bottom of that liquid; while wax melts at 145° and swims upon the surface of water. When heated or distilled, it undergoes nearly the same changes as wax. It is rather more soluble in alcohol than wax. Nitric acid gradually dissolves it, and converts a portion of it into oxalic acid. It does not dissolve in an alcoholic solution of potash.*

7. Wax combines readily with fixed oils when assisted by heat, and forms with them a substance of greater or less consistency according to the quantity of oil. This composition, which is known by the name of *cerate*, is much employed by surgeons. Cerate.

The volatile oils also dissolve it when heated. This is well known, at least, to be the case with oil of turpentine. A part of the wax precipitates usually as the solution cools, but of a much softer consistence than usual, and therefore containing oil.

8. The fixed alkalies combine with it, and form a compound which possesses all the properties of common soap. When boiled with a solution of fixed alkalies in water, the liquid becomes turbid, and after some time the soap separates and swims on the surface. It is precipitated from the alkali by acids in the state of flakes, which are the wax very little altered in its properties.† Punic wax, which the ancients employed in painting in encausto, is a soap composed of 20 parts of wax and one of soda.‡ Its composition was ascertained by Mr. Lorgna.§ Action of alkalies.

When boiled with liquid ammonia, it forms a kind of soapy emulsion. As the mixture cools, the greatest part of the compound rises to the surface in the state of white flakes. This soap is scarcely soluble in water.||

9. The acids have but little action on wax; even chlorine, which acts so violently on most bodies, produces no other change on it than that of rendering it white. This property which wax possesses, of resisting the action of acids, renders it very useful as a lute to confine acids properly in vessels, or to prevent them from injuring a common cork.

* Annals of Philosophy, ix. 53.

† Nicholson's Jour. iv. 132.

‡ Pliny, lib. xxi. c. 14.

§ Jour. de Phys. Nov. 1735.

|| Bostock, Nicholson's Jour. iv. 134.

Book IV.

10. Lavoisier endeavoured to ascertain the composition of bees wax, by burning a quantity of it in oxygen gas, and determining the proportions of carbonic acid, and water formed. But pneumatic chemistry had not at that time made sufficient progress to enable him to give his experiments the requisite degree of precision. Gay-Lussac and Thenard analysed it by burning a determinate quantity along with chlorate of potash in a proper vessel, and ascertaining the quantity of carbonic acid formed.* According to their experiments 100 parts of wax are composed of

Composi-
tion.

Oxygen	5.544
Hydrogen	12.672
Carbon	81.784

As we are ignorant of the number which represents the weight of an atom of wax, we cannot from this analysis determine the number of atoms of each of these bodies that enter into its composition. But the smallest number of atoms, which will nearly correspond with the proportions obtained by Gay Lussac and Thenard, are,

1 Atom oxygen	=	1	5.48
20 Atoms carbon	=	15	82.19
18 Atoms hydrogen	=	2.25	12.33
—				
39				100.00
		18.25		

According to this supposition wax is a compound of 39 atoms, and the weight of an integrant particle of it is 18.25. From the experiments of Chevreul on soaps, there is some reason for believing that the weight of an atom of wax is nearly double the number here assigned, or 36. If this number were accurate its true constituents would be,

2 atoms oxygen	=	2
40 atoms carbon	=	30
32 atoms hydrogen	=	4
		36

11. Wax possesses all the essential properties of fixed oil. We must therefore consider it as a fixed oil rendered concrete. Now that species of fixed oils, distinguished by the epithet *fat*,

* Recherches Physico-chimiques, ii. 316.

have the property of becoming concrete, and assuming the appearance of wax, when exposed long to the air, in consequence, it is supposed, of the absorption of oxygen. Hence probably the difference between wax and fixed oils consists in the oxygen which it contains as the component part. The wax at its first formation was in all probability in the state of a fixed oil; but by the absorption of oxygen it gradually concreted into wax. Wax, then, may be considered as a fixed oil saturated with oxygen.

12. It is natural to suppose, if this theory be just, that fixed oil will occur in plants in various states of hardness: and this accordingly is the case. Sometimes it is of the consistency of butter, and this is denominated a *butter*; thus we have the butter of *cacao*, the butter of *coco*, the butter of *galam*. Sometimes it is of a greater consistency, and then is denominated *tallow*; thus we have the tallow of the *croton*, extracted by boiling water from the fruit of the *croton schifera*. When its consistency is as great as possible, it then takes the appellation of *wax*. Thus we have the *myrtle wax* of America, extracted from the seeds of the *myrica cerifera*, and the *pela* of the Chinese. The species of wax, then, which exist in the vegetable kingdom, may possibly be as numerous as the fixed oils. Let us take a view of some of the most remarkable.

Varieties of
wax.

Bees wax is the species whose properties have been described in the former part of this section.

13. The myrtle wax of North America is obtained from the *myrica cerifera*. We are indebted to Dr. Bostock * and Mr. Cadet † for a very exact account of its properties and extraction. The *myrica cerifera* is a shrub which grows abundantly in Louisiana and other parts of North America. It produces a berry about the size of a pepper corn. A very fertile shrub yields nearly seven pounds. The berries are picked off, thrown into a kettle, and covered with water to the depth of about half a foot. The kettle is then hoiled, and the berries stirred and squeezed against the sides of the vessel. The wax which they contain is melted out and swims on the surface. It is skimmed off, passed through a cloth, dried, melted again, and cast into cakes. From the observations of Cadet, we learn that the wax forms the outer covering of the berries. The wax thus obtained is of a pale green colour. Its specific gravity is 1.0150. It melts at the temperature of 109°: when strongly heated

Myrtle wax.

* Nicholson's Jour. iv. 130.

† Ann. de Chim. xiv. 160.

Book IV.

it burns with a white flame, produces little smoke, and during the combustion emits an agreeable aromatic odour. Water does not act upon it. Alcohol, when hot, dissolves $\frac{1}{10}$ th of its weight, but lets most of it fall again on cooling. Hot ether dissolves about $\frac{1}{4}$ th of its weight; and when slowly cooled, deposits it in crystalline plates, like spermaceti. The ether acquires a green colour, but the wax becomes nearly white. Oil of turpentine, when assisted by heat, dissolves it sparingly. Alkalies act upon it nearly as on bees wax. The same remark applies to acids. Sulphuric acid, when assisted by heat, dissolves about $\frac{1}{12}$ th of its weight, and is converted into a thick dark-brown mass.* Mr. Hatchett has detected a substance precisely similar to myrtle wax in *lac*.† It probably exists in many vegetables.

The wax from the *myrica cordifolia*, a shrub which grows at the Cape of Good Hope, has been examined by Dr. John, and from the characters which he assigns to it, we may consider it as very nearly the same with the myrtle wax of North America.‡

Brazil wax.

14. Mr. Brande in 1811, published a chemical examination of wax from Brazil, the produce of an unknown tree in that country. It had a greenish colour, was insoluble in water, but soluble in alcohol, ether, and oils. Its specific gravity was 0.980. It melted at the temperature of 206°. It could not be made to form a soap with the fixed alkalies; but answered remarkably well for candles.§

Wax from rose and lavender oil.

15. Dr. Macculloch found a waxy substance in the oil of roses and of lavender. It was precipitated from both by cold, and likewise by the addition of alcohol. Its colour is white; it has a greasy feel, swims on sulphuric ether, and crystallizes in shining scales. It melts at 96°. It dissolves in alcohol and ether, and seems to approach nearer spermaceti in its properties than wax.||

16. The Chinese extract a wax from various vegetables, which they manufacture into candles, and of which they form many of the delicate ornaments which are brought to Europe.**

* Bostock, Nicholson's Jour. iv. 130

† Analytical Experiments on Lac, Phil. Trans. 1804.

‡ Chemische Untersuchungen, iii. 38.

§ Nicholson's Jour. xxxi. 14.

|| Phil. Mag. xxxix. 123.

** See Phil. Trans. 1753, p. 253; in which Father d'Incarville gives some account of the tallow-tree, and mode of extracting the tallow or wax.

17. Wax has been detected by Proust in the green fecula of many plants; the cabbage for instance. He thinks that it forms a constituent of the pollen of all flowers, and that the bees collect it along with the gluten of the pollen, which, according to him, serves them for food.* But there is undoubted evidence that wax is actually formed by bees. The observations of Mr. John Hunter led him to conclude that it comes from under the scaly rings which cover their hinder parts.

Chap. I.
In the green
fecula of
plants.

Proust has also detected wax covering the rind of resins, plums, oranges, and other similar fruits.† In short, it seems to be one of the most abundant of vegetable principles.

SECTION XXIII.

OF CAMPHOR.

THE substance called *camphor*, though unknown to the Greeks and Romans, seems to have been long known in the East. When it was first brought to Europe does not appear, though it seems to have been introduced by the Arabians. *Ætius* is the first person who takes notice of it. It is mentioned in the writings of *Paracelsus*, in such a manner as to lead to the supposition that it was in common use in his time. Chemists sometimes arranged it among resins, sometimes among volatile oils, † till *Neumann* published his dissertation on it in 1725.§ He there examined its properties in detail, and demonstrated that they differed from those of every other substance; and concluded in consequence, that camphor must be considered as a peculiar vegetable principle. To *Neumann*, then, we are indebted for the first correct experiments on the properties of camphor.

History.

It comes to Europe chiefly from Japan. It is obtained from the *laurus camphora*, a tree common in the East, by distilling the wood along with water in large iron pots, on which are fitted earthen heads stuffed with straw. The camphor sublimes, and concretes upon the straw in the form of a grey powder.|| It is afterwards refined in Europe by a second sublimation. The vessels are of glass, and somewhat of the

How pro-
cured.

* Jour. de Phys. lvi. 87, and 111.

† Ibid.

‡ Hoffmann, Observ. Phys. Chym. Selec. p. 44.

§ Phil. Trans. xxxiii. 321.

|| Neumann's Chem. p. 319.

Book IV.

shape of a turnip, with a small mouth above loosely covered with paper. According to Ferber, about $\frac{1}{4}$ th of pounded chalk is mixed with the crude camphor;* but others assure us that there is no addition whatever employed. According to Lewis, nothing more is necessary than a proper regulation of the fire;† and Professor Robison, who witnessed the process as well as Neumann, informs us, that the camphor in the subliming vessel is in a liquid state, which it could scarcely be if quick-lime were employed, at least in any considerable quantity.‡

Properties.

1. Camphor thus refined is a white brittle substance, having a peculiar aromatic odour, and a strong hot acrid taste. Its specific gravity is 0.9887.§

It melts at the temperature of 288, and boils, when heated to 400° according to my experiments.||

It is not altered by atmospheric air; but it is so volatile, that if it be exposed during warm weather in an open vessel, it evaporates completely. When sublimed in close vessels it crystallizes in hexagonal plates or pyramids.**

2. It is insoluble in water; but it communicates to that liquid a certain portion of its peculiar odour.††

Action of alcohol.

3. It dissolves readily in alcohol, and is precipitated again by water. According to Neumann, well rectified alcohol dissolves three-fourths of its weight of camphor. By distillation the alcohol passes over first, and leaves the camphor. This property affords an easy method of purifying camphor. Dissolve the camphor in alcohol, distil off the spirit, and melt the camphor into a cake in a glass vessel.‡‡ If the alcoholic

* Gren's Handbuch, ii. 219.

† Neumann, *ibid.*‡ Black's Lectures, ii. 351. See Gay-Lussac *Ann. de Chim. et de Phys.* viii. 75.§ According to Brisson. Dr. Shaw states it at 0.996, (*Shaw's Boyle*, ii. 346;) and this has been copied into most of the recent chemical books. Probably it varies in its density considerably.|| Gay-Lussac says, that it melts at 349°, and boils at 399° (*Ann. de Chim. et de Phys.* viii. 78). This statement does not agree with my experience on the subject.

** Romieu.

†† From the experiments of Cadet it appears, that a French pint of water dissolves about 16 grains of camphor, and that the camphor may be precipitated by pure potash. *Ann. de Chim.* lxii. 132.‡‡ This process, proposed by Lewis (*Neumann's Chem.* p. 320,) is surely preferable to that of Trommsdorf, who precipitates the camphor by water from the alcohol, and then melts it into a cake. See *Gren's-Handbuch*, ii. 220.

solution of camphor be diluted with water as much as possible, without causing the camphor to precipitate, small crystals of camphor resembling feathers gradually form.*

4. Camphor is soluble also in oils, both fixed and volatile. Oils. If the solution be made by means of heat, as it cools part of the camphor precipitates, and assumes the form of plumose or feather-like crystals.†

5. Camphor is not acted on by alkalies, either pure or in the state of carbonates. Alkalies. Pure alkalies indeed seem to dissolve a little camphor; but the quantity is too small to be perceptible by any other quality than its odour.‡ Neither is it acted on by any of the neutral salts which have hitherto been tried.

6. Acids dissolve camphor without effervescence, and in general it may be precipitated unaltered from the recent solution. Acids.

To Mr. Hatchett we are indebted for an accurate investigation of the action of sulphuric acid on camphor. Sulphuric acid. Upon a hundred grains of powdered camphor he poured an ounce of sulphuric acid. The camphor immediately became yellow, and gradually dissolved while the acid changed, first to brownish-red, and afterwards to brown. In about an hour the liquid became blackish-brown, and began to emit abundance of sulphurous acid gas. In four hours the whole appeared like a thick black liquid, and no other smell except that of sulphurous acid could be distinguished in it. As during two days no farther alteration took place, the alembic containing the solution was put upon a sand bath moderately warm, by which means an additional quantity of sulphurous acid gas was driven off. After two days the liquid was slowly mixed with six ounces of water. It became reddish-brown, a considerable coagulum of the same colour subsided, the sulphurous acid smell became imperceptible, and was succeeded by a smell similar to that of a mixture of the oils of lavender and peppermint. The whole was now slowly distilled. The water which came over had the same smell as the original liquid, and there floated over it a little yellowish oil. A blackish-brown mass remained behind, not acted on by water; but alcohol extracted a resinous substance, and acquired a blackish-

* Romieu, Mem. Par. 1756, p. 41.

† Romieu, *ibid.*

‡ Bouillon la Grange. Ann. de Chim. xxiii. 154.

Book IV.

brown colour. What remained was charcoal. Thus, by the action of sulphuric acid, the camphor was decomposed and converted into oil, blackish-brown resin, and charcoal. The proportion of each was as follows:

Yellow oil	3 grains.
Charcoal	53
Resinous substance	49

 105

Making an increase of five grains, either from water which was retained by the resinous-like substance, or from oxygen with which it had combined.

This resinous-like substance, thus obtained, was in reality a species of artificial tannin. It was very brittle, had the odour of caramel, and an astringent taste. It dissolved in cold water, and the solution precipitated iron, lead, tin, and lime, dark brown. It precipitated gold in the metallic state, threw down isinglass in the state of a blackish insoluble substance, and had the property of converting skin into leather. A small quantity of nitric acid converted it into the artificial tannin obtained from charcoal by nitric acid.*

From the farther researches of Chevreul on this subject, it appears that when sulphuric acid is distilled off camphor there is formed, 1. A volatile oil which has the odour of camphor: 2. A charry matter which is a combination of sulphuric acid and a hydroguretted charcoal. 3. An astringent substance which is likewise a compound of sulphuric acid and a charcoal still more hydroguretted than the preceding. The charry residuum is not sensibly soluble in water. When distilled it gives sulphurous acid, carbonic acid, and leaves a residue which is a compound of carbon and sulphur.†

Nitric acid.

Nitric acid dissolves camphor readily, and in great abundance. The solution separates into two portions; that which contains the camphor, and most of the acid, floats upon the top of the other in the form of a very pale yellow oil. This solution is known by the name of *oil of camphor*. Water and several metals precipitate the camphor unaltered.‡ Alcohol combines with the oil of camphor. When this solution of

* Hatchett's additional Experiments on artificial Tannin, Phil. Trans. 1805.

† Ann. de Chim. lxxiii. 167.

‡ Neumann's Chemistry, p. 321.

camphor in nitric acid is long kept, a portion of the camphor separates in crystals, and swims on the surface, and a small portion is converted into camphoric acid.* Chap. I.

Muriatic, sulphurous, and fluoric acids, in the state of gas, dissolve camphor. When water is added, the camphor appears unaltered in flakes, which swim on the surface of the water. † It is dissolved also by water impregnated with carbonic acid gas, ‡ by acetic acid, § and probably by all acids. Other acids.

7. When heat is applied to camphor it is volatilized. It catches fire very readily, and emits a great deal of flame as it burns, but it leaves no residuum. It is so inflammable that it continues to burn even on the surface of water. When camphor is set on fire in a large glass globe filled with oxygen gas, and containing a little water, it burns with a very bright flame, and produces a great deal of heat. The inner surface of the glass is soon covered with a black powder which has all the properties of charcoal; a quantity of carbonic acid gas is evolved; the water in the globe acquires a strong smell, and is impregnated with carbonic acid and camphoric acid. || Action of heat.

If two parts of alumina and one of camphor be formed into a paste with water, and distilled in a glass retort, there comes over into the receiver (which should contain a little water, and communicate with a pneumatic apparatus) a volatile oil of a golden yellow colour, a little camphoric acid, which dissolves in the water, and a quantity of carbonic acid gas and carburated hydrogen gas, which may be collected by means of a pneumatic apparatus. There remains in the retort a substance of a deep black colour, composed of alumina and charcoal. By this process, from 122·284 parts of camphor M. Bouillon la Grange, to whom we are indebted for the whole of the analysis of camphor, obtained 45·856 parts of volatile oil and 30·571 parts of charcoal. The proportion of the other products was not ascertained.**

From this analysis, M. Bouillon la Grange concludes that camphor is composed of volatile oil and charcoal or carbon combined together. We learn from his experiments, that the

* Planche found these changes in a vial of oil of camphor, which had been kept unopened for 14 years. *Ann. de Chim.* xliii. 346.

† Fourcroy.

‡ *Jour. de Phys.* lii. 67.

§ *Phil. Mag.* xv. 156. Aromatic vinegar consists chiefly of this compound.

|| Bouillon la Grange, *Ann. de Chim.* xxiii. 168.

** *Ibid.* 157.

Book IV. ultimate ingredients of camphor are carbon and hydrogen, and that the proportion of carbon is much greater than in oils.

To determine its composition I passed a grain of camphor through a copper tube filled with red hot peroxide of copper. By this process, I obtained 5·837 cubic inches of carbonic acid gas, and 1·3 grain of water = 0·14 grain hydrogen = 6·8237 cubic inches of hydrogen gas. Now the carbon in 5·837 cubic inches of carbonic acid gas amounts to

	0·738 gr.
The hydrogen to	0·144
	<hr style="width: 50px; margin-left: auto; margin-right: 0;"/>
	·882
Loss of weight	= ·118
	<hr style="width: 50px; margin-left: auto; margin-right: 0;"/>
	1·000

and this deficiency must be owing to the oxygen which the camphor contained. But 0·118 gr. of oxygen gas = 0·348 of a cubic inch. It appears from this experiment that the constituents of camphor, supposing them in the gaseous state would be as follows :

Hydrogen	6·8237 cubic inches.
Carbon	5·837
Oxygen	0·348

Now this is equivalent very nearly to

10 atoms hydrogen	= 1·250 or 14·49	
8½ atoms carbon	= 6·375	73·91
1 atom oxygen	= 1·000	11·60
	<hr style="width: 50px; margin-left: auto; margin-right: 0;"/>	<hr style="width: 50px; margin-left: auto; margin-right: 0;"/>
	8·625	100·00

Species of camphor;

8. There are several species of camphor which have been examined by chemists, and which differ considerably from each other in their properties. The most remarkable are *common camphor*, the *camphor of volatile oils*, and the *camphor* obtained by treating *oil of turpentine* with *muratic acid*.

1. Common camphor;

9. Common camphor, obtained by distillation from the *laurus camphora*, is the substance which has been described in the preceding part of this Section. In Borneo and Sumatra camphor is procured from the *laurus sumatrensis*; but as none of this camphor is brought to Europe, we do not know how far it agrees with common camphor in its properties. The *laurus cinnamomum* likewise yields camphor.

10. The second species of camphor seems to exist in a great variety of plants, and is held in solution by the volatile oils extracted from them. Neumann obtained it from oils of thyme, marjoram, cardomum;* Hermann, from oils extracted from various species of mint.† Cartheuser obtained it from the roots of the *maranta galanga*, *kæmpferia rotunda*, *amomum zinziber*, *laurus cassia*,‡ and rendered it probable that it is contained in almost all the labiated plants. It has been supposed to exist in these plants combined with volatile oil. Proust has shown how it may be extracted, in considerable quantity, from many volatile oils.§ All that is necessary, is to expose the oil to the open air at a temperature between 22° and 54°; the oil gradually evaporates, and the camphor crystallizes. By this process he obtained from the oil of rosemary $\frac{1}{10}$ th of its weight of camphor, from that of sweet marjoram $\frac{1}{10}$ th, from oil of sage $\frac{1}{8}$ th, and from oil of lavender more than $\frac{1}{4}$ th. By distilling the oil in a water-bath, some degrees under the boiling point, till one-third of the oil comes over, part of the camphor is found crystallized in the still vessel. This being removed, and the oil again distilled as before, more camphor crystallizes. By three such processes the whole of the camphor is obtained from the oil. The camphor is then purified by mixing it with a little dry lime, and subliming it.

Chap. I.
2. Camphor
from oils;

From the observations of Mr. John Brown, there is reason to believe that the camphor from oil of thyme differs from common camphor in several respects. It does not appear to form a liquid solution either with nitric or sulphuric acid; nor is it precipitated from nitric acid in powder like common camphor, but in a glutinous mass.||

11. The artificial camphor yielded by oil of turpentine, when saturated with muriatic acid gas, was discovered by Mr. Kind, apothecary in Eutin, while employed in making a medicine called the *liquor arthriticus Pottii*. He put a quantity of oil of turpentine into a Woulfe's bottle, and caused a current of muriatic acid gas, separated from common salt by sulphuric acid, to pass through it. The salt used was of the same weight with the oil of turpentine. At first the oil became yellow, then brown, and at last became almost solid, from the

3. Artificial
camphor.

How pro-
cured.

* Phil. Trans. 1725, p. 321.

† Junker, Conspectus Chæmiæ, ii. 100.

‡ De Genericis quibusdam Plantarum Princip. 1754; as quoted by Gren, Handbuch, ii. 221.

§ Ann. de Chim. iv. 179.

|| Phil. Trans. 1725, xxxiii. 361.

Book IV.

formation of a great number of crystals in it, which possessed the properties of camphor.* This curious experiment was repeated by Trommsdorf with a similar result. He examined the camphoric crystals obtained, and found them in most respects to agree with common camphor, though in some they differed.† The society of apothecaries at Paris, on being informed of these particulars, appointed Messrs. Cluzel, Chomet, and Boullay, to examine the subject. The result of their labours has been given to the world by Boullay.‡ They ascertained the proportion of camphor yielded by oil of turpentine, the best dose of muriatic acid gas, examined the properties of the camphoric production, and endeavoured to explain the way in which it had been formed. The experiment of Kind was repeated by Hagen in 1804;§ and more lately, a set of experiments on the same subject was made by Gehlen, Schuster, and Pesth, chiefly in order to elucidate the theory of the process.|| Thenard has repeated them with the same object.**

The proportion of muriatic gas found to answer best, is what can be separated by sulphuric acid and heat from a quantity of common salt equal in weight to the oil of turpentine employed. The camphor produced amounts nearly to one-half of the oil of turpentine. The French apothecaries put into a Woulfe's bottle four pounds of the oil of turpentine, and into a tubulated retort four pounds of common salt, and two pounds of sulphuric acid, and drove off the muriatic acid gas by means of heat. It was totally absorbed by the oil, the bulk of which was augmented by $\frac{1}{10}$ th, while it became at the same time very hot. Its colour became brown; it retained its transparency; and drops of oil gradually formed at the bottom of it. In 24 hours it deposited regular crystals; which, being separated from the brown liquid with which they were mixed, and dried on blotting paper, weighed 24 ounces. The liquor being left for some days in a cellar, deposited four ounces more of the same crystals; and by exposure to an artificial cold of about 14° , two ounces more were obtained, making a total 30 ounces, or $1\frac{2}{3}$ lb. from the four pounds of oil of turpentine. No additional crystals were obtained by increasing the quantity of gas; but the French chemists did not find that

* Trommsdorf's Jour. de Pharmacie, xi. 132; as quoted in Gehlen's Jour. vi. 459.

† Ibid.

‡ Ann. de Chim. li. 270.

§ Gehlen's Jour. ii. 237.

|| Ibid. vi. 470.

** Mem. d'Arcueil, ii. 29.

the produce of camphor was diminished by this augmentation, as had been concluded by Trommsdorf from his trials. Chap. I.

The camphor thus procured was very white; it had a peculiar odour, in which that of the oil of turpentine could be distinguished. When washed with water, it became beautifully white, and gave no longer signs of containing an acid, but still had the smell of oil of turpentine. Water containing some carbonate of potash deprived it of part of this odour, but not the whole. When mixed with its own weight of charcoal powder, wood-ashes, quicklime, or porcelain clay, and sublimed, it was obtained in a state of purity. Properties.

Its smell when pure resembles that of common camphor, but is not so strong. Its taste also resembles that of camphor. It swims on water, to which it communicates its taste, and burns upon its surface. It dissolves completely in alcohol, and is precipitated by water. Nitric acid, of the specific gravity 1.261, had no action on it, though it readily dissolves common camphor; but concentrated nitric acid dissolves it with the disengagement of nitrous gas; and water does not precipitate it from its solution as it does common camphor. Acetic acid does not dissolve it. When heated it sublimes without decomposition; and when set on fire burns like camphor.

Such are the properties of this artificial camphor, as far as they have been investigated. We see from them that it resembles common camphor, except as far as regards the action of acids. In the experiment of Thenard, the residual liquid was not brown but colourless. He ascribes the brown colour of it, in preceding experiments, to the oil of turpentine used containing some resin mixed with it, whereas his had been purified by distilling it. According to him the oil of turpentine is not decomposed by this process. The residual liquid, he thinks, may be owing to the presence of some foreign oil in oil of turpentine. When this residual liquid is exposed to the air for some days it loses its acid properties, though it still contains abundance of acid. But the acid is neutralized by the oil.

Action of
muriatic
acid gas on
oil of tur-
pentine.

Gehlen ascertained that neither fluoric acid gas, sulphurous acid gas, nor ammoniacal gas, produce any such decomposition in oil of turpentine. Neither is any such change produced by the action of chlorine gas;* but when oil of turpentine is distilled slowly at a heat of 122°, a number of crystals of cam-

* Boullay, Ann. de Chim. li. 283.

Book IV. phor sublime from it.* No other oils tried yielded camphor when impregnated with muriatic acid gas.†

12. Camphor is much used in medicine. It is a powerful stimulant; it is considered as peculiarly efficacious in diseases of the urinary organs; it is often serviceable in mania, and procures sleep when every other medicine fails.

SECT. XXIV.

OF BIRD-LIME.

Discovery. THE vegetable principle to which I give the name of *bird-lime*, was first examined by Vauquelin, who found it possessed of properties different from every other. It was found collected on the epidermis of a plant brought to Europe by Michaud, and called *robinia viscosa* by Cels; constituting a viscid substance, which made the fingers adhere to the young twigs.‡ From the analysis of *bird-lime* by Bouillon la Grange, it is obvious that it owes its peculiar properties to the presence of an analogous substance, which indeed constitutes the essential part of that composition.§ Hence the reason why I have given the name of *bird-lime* to the principle itself.

**1. Natural
bird-lime.**

1. Natural bird-lime (or that which exudes spontaneously from plants) possesses the following properties :

Its colour is green; it has no sensible taste or smell; is extremely adhesive; softens by the heat of the fingers, and sticks to them with great obstinacy. When heated it melts, swells up, and burns with a considerable flame, leaving a bulky charcoal behind it. It does not dissolve in water; alcohol has but little action on it, especially when cold. By the assistance of heat it dissolves a portion of it; but on cooling, allows the greatest part to precipitate again. When exposed to the air it continues glutinous, never becoming hard and brittle like the resins.

It combines readily with oils. Ether is its true solvent, dissolving it readily without the assistance of heat. The solution is of a deep green colour. The alkalies do not combine with it; the effect of the acids was not tried.|| These properties

* Boullay, Ann. de Chim. li. 283.

† Ann. de Chim. xxviii. 223.

|| Vauquelin, Ann. de Chim. xxviii. 254.

‡ Ibid.

§ Nicholson's Journ. xiii. 144.

are sufficient to distinguish bird-lime from every other vegetable principle. Chap. I.

It is not unlikely that the viscid substance which covers the stem of the *lychnis viscaria*, *saxifraga tridactylites*, and a few other plants common in this country, consists of this vegetable principle. It has not hitherto, as far as I know, been examined by any chemist.

2. Artificial bird-lime is prepared from different substances in different countries. 2. Artificial
bird-lime. The berries of the mistletoe are said to have been formerly employed. They were pounded, boiled in water, and the hot water poured off. At present bird-lime is usually prepared from the middle bark of the holly. The process followed in England, as described by Geoffroy, is as follows: the bark is boiled in water seven or eight hours till it becomes soft. It is then laid in quantities in the earth, covered with stones, and left to ferment or rot for a fortnight or three weeks. By this fermentation, it changes to a mucilaginous consistency. It is then taken from the pits, pounded in mortars to a paste, and well washed with river water. Bouillon la Grange informs us, that at Nogent le Rotrou bird-lime is made by cutting the middle bark of the holly into small pieces, fermenting them in a cool place for a fortnight, and then boiling them in water, which is afterwards evaporated. At Commerci various other plants are used.*

Bouillon la Grange made bird-lime for the purpose of analysis by the following process. He bruised a sufficient quantity of the middle bark of the holly, boiled it in water for four or five hours, and then deposited it in pits placed in earthen pans, where it continued, being moistened occasionally with water till it became viscous. Lastly, it was freed from all heterogeneous substances by washing it with pure water. Thus prepared, it resembled the bird-lime of Commerci very exactly.†

Its colour is greenish, its flavour sour, and its consistence Properties. gluey, stringy, and tenacious. Its smell is similar to that of linseed oil. When spread on a glass plate and exposed to the air and light it dries, becomes brown, loses its viscosity, and may be reduced to powder; but when water is added to it, the glutinous property returns. It reddens vegetable blues.

When gently heated it melts and swells, and emits an odour

* Nicholson's Jour. xiii. 145.

† Bouillon la Grange, Nicholson's Jour. xiii. 145.

Book IV.

like that of animal oils. When heated on red-hot coals, it burns with a lively flame, and gives out a great deal of smoke, leaving a white ash composed of carbonate of lime, alumina, iron, sulphate, and muriate of potash.

Water has little action on bird-lime. When boiled in water the bird-lime becomes more liquid, but recovers its original properties when the water cools. The water, by this treatment, acquires the property of reddening vegetable blues, and when evaporated leaves a mucilaginous substance, which may be likewise separated by alcohol.

A concentrated solution of potash forms with bird-lime a whitish magma, which becomes brown by evaporation, while ammonia separates. The compound thus formed is less viscid than bird-lime, and in smell and taste resembles soap. In alcohol and water it dissolves almost completely, and possesses properties similar to those of soap.

Weak acids soften bird-lime and partly dissolve it; strong acids act with more violence. Sulphuric acid renders it black; and when lime is added to the solution, acetic acid and ammonia separate. Nitric acid cold has little effect; but when assisted by heat it dissolves the bird-lime; and the solution, when evaporated, leaves behind it a hard brittle mass. By treating this mass with nitric acid, a new solution may be obtained, which by evaporation yields malic and oxalic acids, and a yellow matter which possessed several of the properties of wax. Cold muriatic acid does not act on bird-lime; hot muriatic acid renders it black.

Bird-lime, when treated with chlorine, becomes white, and is divided into hard compact masses, having unaltered bird-lime in their centre. This white substance may be pulverised; it is insoluble in water, does not melt when heated; and when treated with nitric acid, it neither becomes yellow, nor does it yield resin.

Acetic acid softens bird-lime, and dissolves a certain portion of it. The liquid acquires a yellow colour. Its taste is insipid. When carbonate of potash is dropped into this solution, no precipitate falls. By evaporation it yields a resinous-like substance.

Some of the metallic oxides are reduced when heated with bird-lime. Litharge combines with it, and forms a kind of plaster.

Alcohol of the specific gravity 0.817 dissolves bird-lime at a boiling heat. On cooling it lets fall a yellow matter similar

to wax. The filtered liquid is bitter, nauseous, and acid. Chap. I.
 Water precipitates a substance similar to resin.

Sulphuric ether dissolves bird-lime readily, and in great abundance. The solution is greenish. When mixed with water, an oily substance separates, which has some resemblance to linseed oil. When evaporated a greasy substance is obtained, having a yellow colour and the softness of wax. Oil of turpentine dissolves bird-lime readily.

Such are the properties of artificial bird-lime, as far as they have been investigated by Bouillon la Grange.* It is obvious that it contained acetic acid, mucilage, and several alkaline and earthy bodies, which must be considered as foreign substances. When these are removed, the close analogy between artificial and natural bird-lime is sufficiently obvious.

SECT. XXV.

OF OLIVILE.

This is a name which M. Pelletier has given to a peculiar substance, which constitutes one of the ingredients of a gummy-looking substance, which exudes from the olive tree. It is easily obtained by dissolving the *gummy substance* in rectified alcohol and setting the alcohol aside for spontaneous evaporation. It gradually deposits crystals in flattened needles of a yellowish white colour. These crystals are the *olivile* in a state of purity.

It has the form of a white, brilliant starchy powder, or of flattened needles. Properties. It has no smell; but its taste is quite peculiar, being at the same time bitter and sweet, and having something aromatic. It melts when heated to the temperature of 158° , and when cooled has the appearance of a transparent resin of a slightly yellowish colour. In this state it becomes electric by friction; but does not appear to have undergone any chemical change. When thrown upon a red-hot coal it takes fire with difficulty, and burns with the emission of much smoke. When distilled it furnishes water, acetic acid, and oil; but no ammonia.

It is scarcely soluble in cold water. Thirty-two parts of boiling water dissolve one part of it. The solution is trans-

* Nicholson's Jour. xiii. 146.

Book IV.

parent and colourless but becomes milky on cooling from the separation of the olivile, which remains long suspended in the liquid. The alkalies facilitate the solution of olivile in water without altering its nature when they are not concentrated.

Alcohol is the best solvent of olivile. When hot it appears to dissolve it in any quantity whatever. Sulphuric ether has no action on it; neither have the fixed or the volatile oils, But these last bodies, when hot, dissolve a small portion of it which they allow to precipitate as they cool.

Nitric acid dissolves it without the assistance of heat, acquiring a deep red colour. When heat is applied the action increases very much; the red colour disappears and the liquid becomes yellow on cooling. Much oxalic acid is deposited and some yellow bitter matter is formed. Diluted sulphuric acid has no action on olivile; but concentrated sulphuric acid immediately chars it. Concentrated acetic acid dissolves it very readily, either cold or hot. The solution is colourless, and no precipitate falls when it is mixed with water; unless the olivile has been mixed with resin, in which case the resin is precipitated in the state of a yellowish powder.

The acetates of lead are the only salts which act upon solutions of olivile; they precipitate its aqueous solution in very white flocks, which are soluble in acetic acid.*

SECT. XXVI.

OF RESINS.

It is at present the opinion of chemists, that *resins* stand in the same relation to the *volatile* oils that *wax* does to the *fixed*. Wax is considered as a fixed oil saturated with oxygen; resins, as volatile oils saturated with the same principle.

The resins are very numerous and on account of the various purposes to which they are applied, and the peculiarity of their properties, constitute one of the most important genera of vegetable substances. Till lately they have been very much overlooked by chemists, who satisfied themselves with picking up doubtful information from artists and manufacturers. Many erroneous opinions concerning them have of course been admitted into every system of chemistry. The subject has engaged the attention of Mr. Hatchett, whose happy talent for

* Annals of Philosophy, xii. 33.

observation peculiarly fitted him for the task. To that gentleman I am indebted for some of the most important facts in this section.

Chap. I.

Resins often exude spontaneously from trees; they often flow from artificial wounds, and not uncommonly are combined at first with volatile oil, from which they are separated by distillation. The reader can be at no loss to form a notion of what is meant by *resin*, when he is informed that common *rosin* furnishes a very perfect example of a resin, and that it is from this substance that the whole genus derived their name: for rosin is frequently denominated *resin*.

Origin.

I. Resins may be distinguished by the following properties: Properties.

1. They are solid substances, naturally brittle; have a certain degree of transparency, and a colour most commonly inclining to yellow. Their taste is more or less acrid, and not unlike that of volatile oils; but they have no smell unless they happen to contain some foreign body. They are all heavier than water. They are all non-conductors of electricity; and when excited by friction, their electricity is negative.

2. Their specific gravity varies considerably. The following table exhibits a view of the specific gravity of such of the resins as have been hydrostatically examined.* Specific gravity.

Elemi	1·0182
Anime	1·0284
Highgate resin	1·046 †
Copal	1·069 †
Tacamahac	1·0463
Rosin	1·080 †
Mastich	1·0742
Sandarich	1·060 †
Lac	1·1390
Labdanum	1·1862

3. When exposed to heat they melt; and if the heat be increased they take fire, and burn with a strong yellow flame, emitting at the same time a vast quantity of smoke. Action of heat.

4. They are all insoluble in water whether cold or hot; but Water.

* The specific gravities in the text were ascertained by Brisson and by myself. The older writers differ so much from each other, that there is reason to suspect that the substances tried were not always those to which we at present give the same names.

† By my trials.

Book IV. when they are melted along with water, or mixed with volatile oil, and then distilled with water, they seem to unite with a portion of that liquid; for they become opaque, and lose much of their brittleness. This at least is the case with common rosin.

Alcohol and ether.

5. They are all, with a few exceptions, soluble in alcohol, especially when assisted by heat. The solution is usually transparent; and when the alcohol is evaporated, the resin is obtained unaltered in its properties. Alcohol, according to Tingry, never takes up more than one-third of its weight of resin. When the solution is mixed with water, it becomes milky, and the resin falls in the state of a white powder. They are soluble also in sulphuric ether.

Oils.

6. Several of them are soluble in fixed oils, especially in the drying oils. The greater number are soluble in the volatile oils; at least in oil of turpentine, the one commonly employed.

Alkalies.

7. Mr. Hatchett first examined the action of fixed alkalies on resins, and ascertained, contrary to the received opinion of chemists, that alkaline leys dissolve them with facility. He reduced a quantity of common rosin to powder, and gradually adding it to a boiling lixivium of carbonate of potash, a perfect solution was obtained of a clear yellow colour, which continued permanent after long exposure to the air. The experiment succeeded equally with carbonate of soda, and with solutions of pure potash or soda. Every other resin tried was dissolved as well as *rosin*. Nothing can afford a more striking proof than this, of the necessity of repeating the experiments of our predecessors before we put implicit confidence in their assertions. Mr. Hatchett's discovery must lead to very important consequences. The well-known fact, that the soap-makers in this country constantly mix *rosin* with their soap; that it owes its yellow colour, its odour, and its easy solubility in water to this addition, ought to have led chemists to have suspected the solubility of resins in alkalies. No such consequence, however, was drawn from this notorious fact.

These alkaline solutions of resins have the properties of soap, and may be employed as detergents. When mixed with an acid, the resin is separated in flakes, usually of a yellow colour, and not much altered in its nature.

Ammonia acts but imperfectly upon resins, and does not form a complete solution of any of those bodies hitherto tried.

Acids.

8. It was the received opinion of chemists that acids do not act upon resins. This opinion seems to have been founded on

the known effects of nitric acid upon oils, and on the old theory derived from that action, that resins are compounds of an oil and an acid.* Mr. Hatchett first ascertained this opinion also to be erroneous, and showed that most of the acids dissolve resins with facility, producing different phenomena according to circumstances.

When sulphuric acid is poured upon any of the resins in powder it dissolves them in a few minutes. At first the solution is transparent, of a yellowish-brown colour, and of the consistency of a viscid oil, and the resin may be precipitated nearly unaltered by the addition of water. If the solution be placed on a sand bath, its colour becomes deeper, sulphurous acid gas is emitted, and it becomes very thick, and of an intense black †. If the solution, some time after it is completed, and before it has acquired the deep black colour, be edulcorated, and the residuum digested in alcohol, and the alcohol afterwards separated by distillation, the residue is in part soluble in cold water, and the portion dissolved possesses the properties of artificial tannin.‡ Thus it appears, that sulphuric acid dissolves resins, but gradually acts upon them after the solution is completed, converting them first into artificial tannin, and afterwards reducing them to charcoal: for the last black state of the solution is owing to the evolution of this substance. The charcoal thus formed is dense, and burns rather like mineral than vegetable coal. Its quantity is also very considerable. The following table exhibits a view of the proportion of charcoal obtained by Mr. Hatchett, by digesting different resins in sulphuric acid, edulcorating the residue, and separating the tannin by means of alcohol and water.§ The quantity of resin employed was always 100 grains.

Copal	67 grains.
Mastich	66
Elemi	63
Tacamahac	62
Amber	56
Rosin	43

Charcoal
evolved.

The quantity of charcoal formed by this process is very

* See Hoffmann, *Observ. Phys. Chim. Select.* p. 55.

† Hatchett on an Artificial Tanning Substance, *Phil. Trans.* 1805.

‡ Hatchett's Third Series of Experiments, *ibid.* 1806.

§ Hatchett, *ibid.*

Book IV. remarkable. The same substances, when charred in the usual way by exposure to a red heat in close vessels, yield very little charcoal. The following table exhibits the quantity obtained by Mr. Hatchett from 100 grains of several of them.*

Mastich	4.50 grains.
Amber	3.50
Rosin	0.65

Nitric acid. Nitric acid likewise dissolves the resins with facility, but not without changing their nature. Mr. Hatchett was first led to examine the action of this acid on resins, by observing that resins are thrown down by acids from their solutions in alkalies in the state of a curdy precipitate; but when nitric acid is added in excess, the whole of the precipitate is redissolved in a boiling heat. This remarkable fact, which did not hold when sulphuric or muriatic acids were used, led him to try whether the resins were soluble in nitric acid. He poured nitric acid of the specific gravity 1.38, on powdered rosin in a tubulated retort; and by repeated distillation formed a complete solution of a brownish-yellow colour. The solution took place much sooner in an open matrass than in close vessels. The solution continues permanent, though left exposed to the air. It becomes turbid when water is added; but when the mixture is boiled, the whole is redissolved. When Mr. Hatchett collected the precipitate thrown down by water by filtration, he found that it still possessed several of the properties of resin. The resin is thrown down from nitric acid by potash, soda, and ammonia; but an excess of these alkalies redissolves the precipitate, and forms brownish orange-coloured liquids. When Mr. Hatchett dissolved resin in boiling nitric acid, the solution was attended with a copious discharge of nitrous gas; and when the powdered resin was thrown into cold nitric acid, a considerable effervescence soon took place, and a porous mass was formed, commonly of a deep orange colour.

When the digestion of nitric acid upon a resinous substance is continued long enough, and the quantity of acid is sufficient, the dissolved resin is completely changed; it is not precipitated by water; and by evaporation, a viscid substance of a deep-yellow colour is obtained, equally soluble in water and

* Hatchett's Third Series of Experiments, Phil. Trans. 1806.

alcohol, and seemingly intermediate between resin and extractive.* If the abstraction of nitric acid be repeated, this substance gradually assumes the properties of artificial tannin.† Thus it appears that nitric acid gradually alters the nature of resin, producing a suite of changes which terminate in artificial tannin, upon which nitric acid has no action.

Chap. I.
Tannin
formed.

Muriatic acid and acetic acid dissolve resin slowly, and it may be precipitated again from them unaltered. Mr. Hatchett, to whom we are indebted for the knowledge of all these solutions, recommends acetic acid as an excellent solvent of resins for vegetable analyses. He employed it himself with much address in his analysis of the different varieties of lac.‡

Acetic acid.

9. When resins are subjected to destructive distillation, we obtain, according to Gren, carburetted hydrogen and carbonic acid gas, a very small portion of acidulous water, and much empyreumatic oil. The charcoal is light and brilliant, and contains no alkali.§

Distillation.

10. When volatile oils are exposed for some time to the action of the atmosphere, they acquire consistency, and assume the properties of resins. During this change they absorb a quantity of oxygen from the air. Westrumb put 30 grains of oil of turpentine into 40 cubic inches of chlorine gas. Heat was evolved: the oil gradually evaporated, and assumed the form of yellow resin.¶ Mr. Proust observed, that when volatile oil is exposed to the air, it is partly converted into a resin, and partly into a crystallized acid; usually the benzoic or the camphoric. Hence we see that the oil is converted into two distinct substances. During this change oxygen is absorbed; and Fourcroy has observed that a portion of water is also formed.** It is probable, from these facts, that resin is volatile oil deprived of a portion of its hydrogen, and combined with oxygen. The experiments of Mr. Hatchett have added much to the plausibility of this theory.

Supposed to
be volatile
oils united
to oxygen.

No doubt the different resins contain different proportions of oxygen, carbon, and hydrogen, which are the only constituents. Hitherto only three resins have been analyzed with the requisite precision to enable us to estimate the number of atoms which enter into their composition. These are com-

* Hatchett on an Artificial Tanning Substance, Phil. Trans. 1805.

† Hatchett's Additional Experiments on Artificial Tannin, Phil. Trans. 1805.

‡ Phil. Trans. 1804.

§ Handbuch, ii. 140.

|| Crell's Annals, i. 1790.

** Fourcroy, viii. 16.

Book IV. *mon rosin, sandarach, and copal.* Rosin, according to the analysis of Gay-Lussac and Thenard, is composed of

Oxygen	13.337
Carbon	75.944
Hydrogen	10.719
	100.000 *

The smallest number of atoms which correspond with this analysis are the following :

2 atoms oxygen =	2 13.44
15 atoms carbon =	11.25 75.63
13 atoms hydrogen .. =	1.625 10.93
	30	14.875 100.00

Whether this table exhibits its real constitution cannot be known till the weight of an atom of rosin be determined by combining it with substances with which it enters into combination in definite proportions. But we are sure that its weight is either 14.875, or some multiple of that number. From this composition, we see that it may by heat be totally converted into

Water	2.25
Olefiant gas	9.625
Carbon	3.000
	14.875

It is very probable that different resinous substances are confounded together under the name of rosin; at least my experiments on the composition of rosin do not accord with those of Gay-Lussac and Thenard. By passing a grain of resin through a tube filled with red-hot peroxide of copper, I converted it into carbonic acid and water. I obtained 5 cubic inches of carbonic acid, and 1.05 gr. of water = 0.1166 gr. hydrogen. The carbon in 5 cubic inches of carbonic acid gas weighs

	Grains.
	0.6324
Hydrogen	0.1166
Total .. =	0.7490
Deficiency ...	0.2510

* Recherches Physico-chimiques, ii. 313.

This I consider as oxygen. Now these weights being converted into volumes, give us Chap. I.

	Cubic Inches.		Atoms.		Atoms.
Carbon	5	or	5	or	10
Hydrogen	5½	5½	11
Oxygen	0.75	1½	3
10 atoms carbon	=		7.5		
11 atoms hydrogen ..	=		1.375		
3 atoms oxygen	=		3		
			11.875		

When resin is heated to the temperature of 276°, it becomes quite fluid, and if it be kept at that temperature, it loses a quantity of water, and probably also of oil. When allowed to cool it concretes into a reddish-yellow rosin, much darker than before, and composed, according to my experiments, of

4 atoms carbon	=	3
1 atom hydrogen	=	0.125
3 atoms oxygen	=	3
		6.125*

When a grain of sandarach is passed through a tube filled with red hot peroxide of copper, we obtain 5.609 cubic inches of carbonic acid gas, and 1 grain of water. Now the weight of the carbon and hydrogen in these two substances, is as follows:

	Grains.
Carbon	0.708
Hydrogen	0.111
	0.819
Deficiency	0.181
	1.000

This deficiency I ascribe to oxygen. Now if we suppose these three constituents in the state of gas, their bulk would be as follows:

* Annals of Philosophy, xv. 468.

Book IV.

	Cubic Inches.	Volumes.
Carbon	5·609	or 14
Hydrogen	5·24 13
Oxygen	0·534 0·5

This is equivalent to

14 atoms carbon	=	10·500
13 atoms hydrogen	=	1·625
1 atom oxygen	=	1·000
		<hr/>
		13·125

When sandarach has been kept for some time in a heat of 308° in a state of fusion, its appearance and properties are considerably altered. On analyzing it in that state, I found it a compound of

13 atoms carbon	=	9·75
16 atoms hydrogen	=	2·00
4 atoms oxygen	=	4·00
		<hr/>
		15·75

Thus it appears that the proportion of hydrogen and oxygen is increased, while that of carbon is diminished by exposure to heat.

Copal, according to the experiments of Gay-Lussac and Thenard, is composed of

Oxygen	10·606
Carbon	76·811
Hydrogen	12·583
	<hr/>
	100·000 *

The smallest number of atoms which correspond with this analysis, are

2 atoms oxygen	=	2	10·81
19 atoms carbon	=	14·25	77·02
18 atoms hydrogen	=	2·25	12·17
<hr/>		<hr/>		<hr/>
39		18·5		100·00

Copal, according to this analysis, contains 5 atoms of hydrogen and 4 of carbon more than rosin. It may be resolved into

* Recherches Physico-chimiques, ii. 314.

Water	2·25
Olephant gas	14·00
Carbon	2·25
	<hr/>
	18·50

Chap. I.

11. Hermsbstadt affirms, that to know whether any vegetable substance contains resin, we have only to pour some sulphuric ether upon it in powder, and expose the infusion to the light. If any resin be present, the ether will assume a brown colour.

12. Having now described the general properties of resinous bodies, it will be proper to take a more particular view of those of them which are of the most importance, that we may ascertain how far each possesses the general characters of resins, and by what peculiarities it is distinguished from the rest. The most distinguished of the resins are the following:

1. *Rosin*.—This substance is obtained from different species of *fir*; as the *pinus abies*, *sylvestris*, *larix*, *balsamea*. It is well known that a resinous juice exudes from the *pinus sylvestris*, or common Scotch fir, which hardens into tears. The same exudation appears in the *pinus abies*, or spruce fir. These tears constitute the substance called *thus*, or common frankincense. When a portion of bark is stripped off these trees, a liquid juice flows out, which gradually hardens. This juice has obtained different names, according to the plant from which it comes. The *pinus sylvestris* yields *common turpentine*; the *larix*, *Venice turpentine*; the *balsamea*, *balsam of Canada*, &c. All these juices, which are commonly distinguished by the name of turpentine, are considered as composed of two ingredients; namely, oil of turpentine and rosin. When the turpentine is distilled, the oil comes over, and the rosin remains behind. When the distillation is continued to dryness, the residuum is known by the name of common rosin, or colophonium; but when water is mixed with it while yet fluid, and incorporated by violent agitation, the mass is called *yellow rosin*. During winter the wounds made in the fir trees become encrusted with a white brittle substance called *barras*, or *galipot*, consisting of rosin united to a small portion of oil. The yellow rosin made by melting and agitating this substance in water is preferred for most purposes; because it is more ductile, owing probably to its still containing some oil.* Alcohol of the specific gravity 0·835 dissolves the

* See Neumann's Chemistry, p. 265.

Book IV.

eighth of its weight of rosin at the temperature of 60°. The solution is yellowish and perfectly transparent. When the rosin is obtained from the alcohol by evaporation, it has a dark-brown colour, and is exactly similar to rosin which has been kept for some time in a state of fusion. Rosin, when heated to 156°, becomes viscid and of the consistence of common turpentine. At 276° it becomes quite fluid. When kept for some time at that temperature, it undergoes a considerable loss of weight.*

Mastich.

2. *Mastich*.—This resin is obtained from the *pistacea lentiscus*, a tree which grows in the Levant, particularly in the island of Chios. When transverse incisions are made into this tree, a fluid exudes, which soon concretes into yellowish semitransparent brittle grains.† In this state it is sold under the name of mastich. It softens when kept in the mouth, but imparts very little taste. This has induced surgeons to employ it to fill up the cavities of carious teeth, which it does tolerably well. When heated, it melts and exhales a fragrant odour. Its taste is slight, but not unpleasant. In Turkey great quantities of it are said still to be chewed for sweetening the breath, and strengthening the gums. It is to this use of the resin as a masticatory that it is supposed to owe its name.‡ Mastich does not dissolve completely in alcohol; a soft elastic substance separates during the solution. This substance, in Neumann's experiments, amounted to about $\frac{1}{10}$ th of the mastich; § but in those made more lately by Mr. Matthews, it was nearly a fifth. || The nature of this insoluble portion was first examined by Kind,** who found it possessed of all the properties of caoutchouc. These experiments have lately been repeated by Mr. Matthews with a similar result. || Mr. Brande, however, has observed, that when this insoluble substance is dried it becomes brittle, in which respect it differs from caoutchouc. He has observed also, that by passing a current of chlorine gas through the alcoholic solution of mastich, a tough elastic substance is thrown down, precisely similar to the original insoluble portion. †† When mastich is distilled either with water or

Contains caoutchouc.

* Annals of Philosophy, xv. 468.

† See Neumann's Chem. p. 297.

‡ Ibid. p. 298.

§ Ibid.

|| Nicholson's Jour. x. 247.

** Crell's Annals, 1794, ii. 185.

†† Phil. Mag. xxv. 111. From these experiments, together with those of Dr. Wollaston on this substance, there can be little doubt that it is a peculiar vegetable principle. It may be characterized thus: brittle, semitransparent, fusible like resins, insoluble in water and alcohol, soluble in ether.

alcohol, nothing considerable comes over with these liquids; but Hoffmann ascertained, that if the resin be first mixed in a mortar with its own weight of carbonate of potash, and afterwards distilled with alcohol, the liquid comes over both with the smell and taste of mastich.* With the other agents it exhibits the phenomena described in the former part of this section.

Chap. I.

3. *Sandarach*.—This resin is obtained from the *thuya articulata*, which grows in Barbary.† It exudes spontaneously, and is usually in the state of small round tears of a brown colour, and semitransparent, not unlike mastich, but rather more transparent and brittle. When chewed it does not soften as mastich does, but crumbles to powder. Mr. Matthews found it almost completely soluble in eight times its weight of alcohol. The residue was extraneous matter.‡ It does not dissolve in tallow or oil, as common resin does.§ Mr. Hatchett found it soluble in alkalis and acids with the same phenomena as common resin.|| It melts at 308°, and froths, giving out an agreeable smell. Sandarach.

4. *Elemi*.—This resin is obtained from the *amyris elemifera*; a tree which grows in Canada and Spanish America. Incisions are made in the bark during dry weather, and the resinous juice which exudes is left to harden in the sun. It comes to this country in long roundish cakes wrapped in flag leaves. It is of a pale yellow colour, semitransparent; at first softish, but it hardens by keeping. Its smell is at first strong and fragrant, but it gradually diminishes. Neumann found that alcohol dissolved $\frac{1}{6}$ th of this substance; the remainder consisted chiefly of impurities, and was partly taken up by water. Both water and alcohol, when distilled with it, come over strongly impregnated with its flavour. Along with the water there comes over a fragrant volatile oil, which amounts to about $\frac{1}{10}$ th of the resin employed.** Elemi.

* Hoffmann, *Observ. Phys. Chim. Select.* p. 68.

† Thenard, *Traité de Chimie*, iii. 225.

‡ Nicholson's *Jour.* x. 246.

§ *Ibid.*

|| Giese, an apothecary of Augsburg, announced in Scherer's *Journal* for 1802, p. 536, that sandarach does not dissolve completely in alcohol; that the residue amounts to about one-fifth of the whole. This residue he considers as a peculiar substance. Mr. Matthews proved, by a set of experiments, that the whole of this statement is inaccurate. I have no doubt that Giese employed mastich instead of sandarach, and that the insoluble substance detected was caoutchouc.

** Neumann's *Chem.* p. 296.

Book IV.
Tacama-
bac.

5. *Tacamahac*.—This resin is obtained from the *sagara octandra*, and likewise, it is supposed, from the *populus balsamifera*. It comes from America in large oblong masses wrapt in flag leaves. It is of a light brown colour, very brittle, and easily melted when heated. When pure it has an aromatic smell between that of lavender and musk. When distilled along with water or alcohol, nothing comes over with these liquids. When pure it dissolves completely in alcohol, and water has no action on it.*

Animé.

6. *Animé*.—This resin is obtained from the *hymenæa courbaril*, or locust tree, which is a native of North America. Animé resembles copal very much in its appearance; but it is easily soluble in alcohol, which copal is not; this readily distinguishes them. It is said to be very frequently employed in the making of varnishes. Alcohol dissolves it completely. Water, according to the experiments of Neumann, dissolves about $\frac{1}{16}$ th of it; and when the decoction is evaporated, it leaves an unctuous mass, which makes the fingers oily. Alcohol distilled over it acquires both the smell and taste of animé. Water distilled from it shows on its surface a small quantity of volatile oil.†

Ladanum.

7. *Ladanum* or *labdanum*.—This resin is obtained from the *cystus creticus*, a shrub which grows in Syria and the Grecian Islands. The surface of this shrub is covered with a viscid juice, which, when concreted, forms ladanum. It is collected while moist by drawing over it a kind of rake with thongs fixed to it. From these thongs it is afterwards scraped with a knife. It is always mixed with dust and sand, sometimes in great abundance. The best is in dark coloured masses, almost black, and very soft, having a fragrant odour and a bitterish taste. The impurities, even in the best kinds, amount to about $\frac{1}{4}$ th. Water dissolves rather more than $\frac{1}{12}$ th of the pure portion, and the matter taken up is said to possess gummy properties. When distilled with water, a small quantity of volatile oil rises. Alcohol likewise comes over impregnated with the taste and smell of labdanum.‡

Botany Bay
resin.

8. *Botany Bay resin*.—This resin is said to be the produce of the *acarois resinifera*; a tree which grows abundantly in New Holland, especially near Botany Bay. Specimens of it were brought to London about the year 1799, where it was tried as a medicine. Some account was given of it in Gover-

* Neumann's Chem. p. 295.

† Ibid. p. 297.

‡ Ibid. p. 295.

nor Philips's Voyage,* and in White's Journal of a Voyage to New South Wales;† but it is to Professor Lichtenstein that we are indebted for an account of its chemical properties. He obtained specimens from London, and published the result of his experiments in Crell's Journal.‡

The resin exudes spontaneously from the trunk of the singular tree which yields it, especially if the bark be wounded. It is at first fluid, but becomes gradually solid when dried in the sun. According to Governor Philips, it is collected usually in the soil which surrounds the tree, having doubtless run down spontaneously to the ground. It consists of pieces of various sizes of a yellow colour, unless when covered with a greenish grey crust. It is firm, yet brittle; and when pounded, does not stick to the mortar nor cake. In the mouth it is easily reduced to powder without sticking to the teeth. It communicates merely a slight sweetish astringent taste. When moderately heated, it melts; on hot coals it burns to a coal, emitting a white smoke, which has a fragrant odour somewhat like storax. When thrown into the fire, it increases the flame like pitch. It communicates to water the flavour of storax, but is insoluble in that liquid. When digested in alcohol, two-thirds dissolve: the remaining third consists of one part of extractive matter, soluble in water, and having an astringent taste; and two parts of woody fibre and other impurities, perfectly tasteless and insoluble. The solution has a brown colour, and exhibits the appearance and the smell of a solution of benzoin. Water throws it down unaltered. When distilled, the products were water, and empyrenmatic oil, and charcoal; but it gives no traces of any acid, alkali, or salt, not even when distilled with water.

Properties.

Twelve parts were boiled in a solution of pure soda in water. § Two parts of the resin were dissolved; the remaining ten parts were floating on the solution, cohering together in clots. No crystals were obtained by evaporating part of the solution; and when sulphuric acid was dropped into another portion, resin separated unaltered. When mixed with twice its weight of nitric acid, the resin swims unaltered on the surface; but when heat is applied, a considerable effervescence takes place. The digestion was continued till the effervescence stopped, and

§ Action of soda,

And nitric acid.

* Duncan's New Dispensary, p. 60.

† Appendix, p. 245.

‡ 1799, ii. 242.

§ By pure sodasalze, it is impossible to say whether Lichtenstein means the carbonate of soda or pure soda.

Book IV.

the resin swam on the surface of the liquid collected together in clots. It was then separated by filtration. It had lost $\frac{1}{2}$ th of its weight. The resin thus treated had acquired a bitterish taste, was not so easily melted as before, and alcohol was capable of dissolving only one-half of it. The solution was brown, tasted like bitter almonds; and when mixed with water, let fall a yellow resinous precipitate of a very bitter taste. The insoluble portion mixed with water, but formed a turbid liquid, which passed through the filter. The nitric acid solution separated from the resin by filtration was transparent; its colour was yellow; its taste bitter; and it tinged substances dipped into it of a yellow colour. By evaporation it yielded oxalic acid, and deposited a yellow earthy-like powder. This last substance was insoluble in water, and scarcely soluble in alcohol. Its taste was exquisitely bitter, like quassia. It mixed with the saliva, and readily stained the skin and paper yellow. The residuum continued bitter and yellow; but yielded no precipitate with potash and nitrate of lime.*

I have been thus particular in my account of Lichtenstein's experiments, on account of the very curious results which he obtained. Had he employed stronger reagents and greater heat, he would have most probably anticipated the curious discoveries of Hatchett. The bitter substance, into which he converted this resin by nitric acid, deserves particular attention. He suspects that it is capable of producing the same changes on all the resins: a conjecture which he verified with regard to *colophonium*, as he found it to yield equally a yellow bitter substance. We now know from the experiments of Hatchett, that the action of nitric acid on the resins terminates in the formation of artificial tannin, which possesses a very bitter taste.

Black poplar resin.

9. *Black poplar resin*.—In the year 1770, the black poplar was pointed out, in a German periodical work, as a tree from which abundance of wax could be obtained. It was even said, that a manufactory of candles from the wax of this tree had been established in Italy. This account having been revived in 1804, Schræder was induced to make a set of experiments on the subject. He found, that when the buds of this tree are boiled in water and properly pressed, they yield about $\frac{1}{8}$ th of their weight of a yellowish-white substance, which possesses the properties of a resin, and resembles, according to him, the

* Lichtenstein, Crell's Jour. 1799, ii. 242.

yellow resin of Botany Bay. When digested in water, a coloured solution is obtained, which reddens litmus paper, which becomes muddy by cooling, and which when evaporated deposits small crystals.*

Chap. I.

10. *The green resin*, which constitutes the colouring matter of the leaves of trees, and almost all vegetables, is insoluble in water, and soluble in alcohol. From the experiments of Proust we learn, that when treated with chlorine it assumes the colour of a withered leaf, and acquires the resinous properties in greater perfection.†

Green resin.

11. *Copal*.—This substance, which deserves particular attention from its importance as a varnish, and which at first sight seems to belong to a distinct class from the resins, is obtained, it is said, from the *thus copallinum*, a tree which is a native of North America; but the best sort of copal is said to come from Spanish America, and to be the produce of different trees. No less than eight species are enumerated by Hernandez.‡

Copal.

Copal is a beautiful white resinous substance, with a slight tint of brown. It is sometimes opaque, and sometimes almost perfectly transparent. When heated it melts like other resins; but it differs from them in not being soluble in alcohol, nor in oil of turpentine without peculiar management. Neither does it dissolve in the fixed oils with the same ease as the other resins. It resembles gum animé a little in appearance; but is easily distinguished by the solubility of this last in alcohol, and by its being brittle between the teeth, whereas animé softens in the mouth.§ The specific gravity of copal varies, according to Brisson, from 1.045 to 1.139. I found it 1.069. Mr. Hatchett found it soluble in alkalies and nitric acid with the usual phenomena; so that in this respect it agrees with the other resins. The solution of copal in alkalies he found indeed opalescent, but it is nevertheless permanent. It deserves attention, that he found rosin, when dissolved in nitric acid, and then thrown down by an alkali, to acquire a smell resembling that of copal.

When copal is dissolved in any volatile liquid, and spread thin upon wood, metal, paper, &c. so that the volatile men-

Forms the best varnish.

* Schrader, Gehlen's Jour. vi. 598.

† Jour. de Phys. lvi. 106.

‡ Lewis, Neumann's Chem. p. 299. From Dr. Roxburgh's account it appears that the Indian tree *Valeria indica* yields a resin possessing properties intermediate between those of copal and amber. Nicholson's Journal, xxvii. 72.

§ Ibid.

Book IV.

strum may evaporate, the copal remains perfectly transparent, and forms one of the most beautiful and perfect varnishes that can well be conceived. The varnish thus formed is called *copal varnish*, from the chief ingredient in it. This varnish was first discovered in France, and was long known by the name of *vernix martin*. The method of preparing it is concealed; but different processes for dissolving copal in volatile menstrua have been from time to time made public. The following are the most remarkable of these:

Solution in
linseed oil.

When copal is kept melted till a sour smelling aromatic odour has ceased to proceed from it, and then mixed with an equal quantity of linseed oil, which has been deprived of all colour by exposure to the sun, it unites with the oil, and forms a varnish which must be dried in the sun.*

Japanners'
varnish.

I have been informed by a very ingenious japan manufacturer in Glasgow, that the copal varnish used by the English japanners is made as follows: four parts by weight of copal in powder are put into a glass matrass and melted. The liquid is kept boiling till the fumes, condensed upon the point of a tube thrust into the matrass, drop to the bottom of the liquid without occasioning any hissing noise, as water does. This is a proof that all the water is dissipated, and that the copal has been long enough melted. One part of boiling hot linseed oil (previously boiled in a retort without any litharge) is now poured into it, and well mixed. The matrass is then taken off the fire, and the liquid, while still hot, is mixed with about its own weight of oil of turpentine. The varnish thus made is transparent, but it has a tint of yellow, which the japanners endeavour to conceal by giving the white ground on which they apply it a shade of blue. It is with this varnish that the dial plates of clocks are covered after having been painted white.

Solution in
oil of tur-
pentine.

When copal is treated with oil of turpentine in close vessels, the vapour being prevented from escaping, exerts a greater pressure, and the heat rises above the boiling point. This additional heat is said to enable the oil to dissolve the copal. The solution, mixed with a little poppy oil, forms a varnish which is distinguished from the *vernix martin* merely in having a very slight tinge of brown.†

The method of dissolving copal in oil of turpentine, published by Mr. Shel Drake, seems to depend upon the same prin-

* Dr. Black's Lectures, ii. 359.

† Ibid.

ciple with the last solution. On two ounces of copal, broken into small pieces, is poured a mixture of four ounces of ammonia with a pint of oil of turpentine. The whole is kept boiling very gently, so that the bubbles may be counted as they rise. If the heat be allowed to diminish, or if it be raised too high, the process stops, and cannot be again resumed. The matrass, in which the mixture is boiled, is stopped with a cork, secured in its place by a brass wire, and perforated by a pin. When the copal is nearly dissolved, the process is stopped, and the whole allowed to cool before uncorking the matrass. This varnish has a deep colour; but when spread thin and allowed to dry, it becomes colourless. Its defect is the difficulty with which it dries. This defect Mr. Sheldrake remedies by throwing the solution into its own weight of nut oil, rendered drying by white lead, and agitating till the turpentine is separated.

Chap. I.
Solution by means of ammonia,

To dissolve copal in alcohol, Mr. Sheldrake dissolves half an ounce of camphor in a pint of that liquid, and pours the solution on four ounces of copal. The matrass is placed in a sand bath, and the process is conducted exactly as the one last described. The solution thus formed contains a great deal of copal, and forms a varnish which is perfectly colourless; but considerable heat is necessary to drive off the camphor.

And camphor.

Mr. Sheldrake has lately favoured the public with another and easier method of dissolving copal. This method is as follows:

“Provide a strong vessel made of tin or other metal; it should be shaped like a wine bottle, and capable of holding two quarts; it will be convenient to have a handle strongly rivetted to the neck; the neck should be long and have a cork fitted to the mouth, but a notch or small hole should be made in the cork, that when the spirit is expanded by heat, a small portion may force its way through the hole, and thus prevent the vessel from bursting.

Sheldrake's process.

“Dissolve half an ounce of camphor in a quart of spirit of turpentine, and put it into the vessel; take a piece of copal the size of a large walnut, reduce it to a coarse powder or very small pieces; put them into the tin bottle, fasten the cork down with a wire, and set it as quick as possible upon a fire so brisk as to make the spirit boil almost immediately; then keep it boiling very gently for about an hour, when so much of the copal will be dissolved as will make a very good varnish; or, if the operation has been properly begun, but enough of copal has not been dissolved, it may be again put on the fire, and by

Book IV. boiling it slowly for a longer time, it may be at last brought to the consistence desired.”*

Demmenie's process. Van Mons relates another process much simpler than any of the above, which he says was taught him by Mr. Demmenie, a Dutch artist. It consists in exposing copal to the action of the steam of alcohol. A long-necked matrass is filled one-fourth full of strong alcohol, and a piece of copal is suspended above the surface of the liquid at some little distance; the top of the matrass is covered with a condensator; the alcohol is kept boiling: the copal softens, and drops down into the alcohol like oil. When these drops no longer dissolve, the process must be stopped. The solution thus obtained is perfectly colourless. Copal may be dissolved in oil of turpentine by the same process.†

Lenormand's process. The following method of making copal varnish has been recommended by Professor Lenormand. Drop upon the pieces of copal pure essential oil of rosemary. Those pieces that are softened by the oil are fit for the purpose, the others not. Reduce them to a fine powder, put this powder into a glass vessel not thicker than a finger breadth, pour oil of rosemary over it, and stir it about with a glass rod. In a short time the whole is converted into a very thick liquid. Pour alcohol on this liquid by little at a time, incorporating it, by gently agitating the vessel, till it is of the requisite thinness for use.‡

12. *Highgate resin*.—This substance was dug up at Highgate, near London, during the late attempt to run a tunnel through the hill. It was in small amorphous masses of different sizes. Colour a muddy-yellowish light-brown. Semitransparent. Lustre resinous, and surface smooth, as if it had been long agitated in water. It was less easily broken than common rosin, but much more easily than copal. Softer than copal. Has a resinous and aromatic smell, especially when heated. When heated it melts and may be rendered as liquid as water, without altering its colour. When in lumps it is insoluble in water, alcohol, potash ley, acetic acid. But ether renders it opaque and white, and quite tender; it loses its cohesion and crumbles into powder upon the least pressure between the fingers. The ether at the same time dissolves a portion, which it deposits when mixed with water. Nitric acid partly dissolves and partly converts it into a red coloured substance. Water

* Nicholson's Jour. ix. 157.

† Ibid. xxiv. 67.

‡ Jour. de Chim. iii. 213.

throws down the dissolved portion in the state of bitter-tasted white-flocks. Sulphuric acid readily chars it when assisted by heat. When in a state of a fine powder alcohol dissolves a small portion of it. It is insoluble in potash ley.*

13. *Lac*.—This is a substance deposited on different species of trees in the East Indies, by an insect called *chremes lacca*, constituting a kind of comb or nidus. It has been imported into Europe, and extensively used from time immemorial; but it is only of late years that correct information concerning it has been obtained. For what relates to the natural history of the insect, and the mode of forming the lac, we are indebted to Mr. Ker,† Mr. Saunders,‡ and Dr. Roxburgh.§ Though very often employed in the arts, it was neglected by chemists. Geoffroy, junior, indeed, published a dissertation on it, but it contains few chemical experiments. He merely subjected it to distillation, and obtained products which he thought similar to those given by wax in the same circumstances. || This led him to consider it as a species of wax, an opinion followed by Neumann; ** but Junker, †† and most of the subsequent chemical writers, place it among the resins. Mr. Hatchett has examined it with his usual address, and ascertained its composition and properties. †† To him we are indebted for almost every thing which we know respecting its chemical nature.

There are various kinds of lac distinguished in commerce. *Stick lac* is the substance in its natural state, encrusting small twigs. When broken off and boiled in water it loses its red colour, and is called *seed lac*. When melted and reduced to the state of a thin crust, it is called *shell lac*. *Stick lac* is of a deep red colour, and yields to water a substance which is used as a red dye. The other two varieties are brown.

Water dissolves the greatest part of the colouring matter of lac, which varies from 15 to $\frac{1}{2}$ per cent. Alcohol dissolves the greatest part of the resin, which constitutes the chief ingredient in the composition of lac. Ether acts more feebly. Sulphuric acid dissolves and gradually chars lac; nitric acid dissolves, and then produces the same changes on it as on other resinous bodies. Muriatic and acetic acids likewise act as sol-

* Annals of Philosophy, iii. 9.

† Phil. Trans. 1781, p. 376.

‡ Phil. Trans. 1789, p. 107.

§ Ibid. 1791, p. 228.

|| Mem. Par. 1714, p. 121; and Martine's translation of the Memoires of the French Academy, v. 4.

** Chemistry, p. 334.

†† Conspectus Chemicæ, ii. 70.

‡‡ Analytical Experiments on Lac, Phil. Trans. 1804.

Book IV.

vents. A solution of borax in water readily dissolves lac. The best proportions are 20 grains of borax, 100 grains of lac, and four ounces of water. This solution, mixed with lamp black, constitutes Indian ink; and may indeed be employed for many of the purposes of varnish. The fixed alkalies readily dissolve lac, but not the volatile. When placed on a hot iron it melts, and emits a thick smoke with an odour rather pleasant, leaving a spongy coal. When distilled, it yields water slightly acidulous, and a thick butyraceous oil. The gases emitted are a mixture of carbonic acid and carburetted hydrogen. Stick lac yields also some carbonate of ammonia; but the other two varieties none. The following table exhibits the constituents of the different varieties of lac, according to the analysis of Mr. Hatchett:

	Stick Lac.	Seed Lac.	Shell Lac.
Resin	68	88·5	90·9
Colouring matter ..	10	2·5	0·5
Wax	6	4·5	4·0
Gluten	5·5	2·0	2·8
Foreign bodies	6·5	—	—
Loss	4·0	2·5	1·8
	100·0	100·0	100·0

The resin is less brittle than those bodies usually are. The colouring matter possesses the properties of extractive; the wax is analogous to myrtle wax, and the gluten closely resembles the gluten of wheat.*

Dr. John published a set of observations on stick lac in 1810, and drew as a conclusion, that it consists of three distinct substances; viz. a colouring matter, a peculiar body to which he gave the name of *laccin* and resin.†

In the year 1816, a very elaborate analysis of stick lac appeared by the same chemist. According to him, 120 grains of stick lac are composed of the following substances:

* Hatchett, Phil. Trans. 1804.

† Chemische Untersuchungen, i. 52.

Resin, insoluble in ether	80
Laccin	20
Cochenillin	4.5
Extractive	3
Yellow extract	0.5
Laccic acid	0.75
Cochenil coloured coverings of insects ..	2.5
Waxy tallow	2
Laccate of potash	} 1.25
Sulphate of potash ..	
Muriate of potash	
Phosphate of lime	
A salt of iron	
Earthy matter	0.75
Loss	4.75
	120.00

The *laccin* was the matter that remained after the lac had been repeatedly digested in alcohol and water, till nothing further could be removed. It is distinguished by the following characters.

It is hard and brittle, has a yellow colour, and a certain degree of transparency. It is insoluble in cold water; but in hot water, though it does not dissolve, it becomes soft. In cold alcohol it softens, increases in bulk, and acquires a slippery feel. Even hot alcohol is unable to dissolve it. In ether and essential oils it swells a little, and becomes quite transparent, but does not dissolve. Laccin.

It dissolves very readily in potash ley, and the solution has a light-brown colour. Muriatic acid renders the solution milky, and the laccin slowly precipitates. Concentrated sulphuric acid dissolves it very rapidly; the solution has an amethyst colour, and becomes muddy when mixed with water. When long boiled with water, containing from $\frac{1}{10}$ th to $\frac{1}{8}$ th of its weight of sulphuric acid, only a small portion of it is dissolved. This portion is obtained in the state of a yellowish gum, when the acid is separated by means of lime, and the solution evaporated. This property distinguishes it readily from cerasin. Concentrated nitric acid dissolves it slowly when assisted by heat. The solution is clear, and has a yellow colour without any bitter taste. It gradually deposits some crystals of oxalic acid. Diluted nitric acid has no effect on it either cold or hot.

Book IV.

When heated, laccin gives out an aromatic odour, and becomes soft. It does not melt, but is gradually charred. When distilled, it gives out water, an acid, which when saturated with soda, throws down muriate of iron white, and a yellow and brown oil. No ammonia can be detected in the products of the distillation.*

Uses.

The uses to which lac is applied in India are very numerous. In Europe it forms the basis of sealing-wax; which is made by melting lac with different proportions of turpentine and some colouring matter, as ivory-black, for black sealing-wax; vermilion for red, &c. It constitutes also the basis of many varnishes and lackers. This last name indeed is derived from the word *lac*.†

Amber.

14. *Amber*.—This substance is undoubtedly of vegetable origin; and though it differs from resins in some of its properties, yet it agrees with them in so many others, that it may, without impropriety, be referred to them. For the chemical investigation of the properties of this substance, we are chiefly indebted to the labours of Hoffmann,‡ Bourdelin,§ Stockard de Neuforn,|| Heyer,** and more lately it has occupied the attention of Mr. Hatchett. The best account of amber varnish which I have seen is by Nils Nystrom in the Stockholm Transactions for 1797.††

Properties.

Amber is a brittle, light, hard substance, usually nearly transparent; sometimes almost colourless, but commonly yellow or even deep brown. It has considerable lustre. Its specific gravity is 1.065. It is tasteless, and without smell, except when pounded or heated, when it emits a fragrant odour. When heated it softens; but, as far as is known, cannot be melted without losing some of its weight, and altering its appearance. In a strong heat it burns, leaving a small

* Chemische Untersuchungen, iv. 12.

† A detailed account of the method of making several of these is given by Dr. Lewis in his Philosophical Commerce of the Arts, p. 223.

‡ Obs. Phys. Chem. p. 60 and 198.

§ Sur le Succin. Mem. Par. 1742, p. 192.

|| Specimen. Chem. Med. Inaugur. de Succino in Genere et speciatim de Succino Fossili Wisbolzensi, 1760. This tract contains a very copious set of experiments; which, however, do not always agree with those of other chemists. Wasserberg's Treatise on Amber is chiefly an abridgment of Stockard's.

** Chemische versuche mit Bernstein, 1787.

†† I have seen the paper only through the medium of Crell's Journal. It is inserted in Crell's Annals, 1799, ii. 171 and 253.

quantity of ashes, the nature of which has not yet been ascertained. Water has no action on it; but alcohol, by long digestion, dissolves about one-eighth of the amber, and forms a coloured solution, which when concentrated becomes milky when mixed with water. The precipitate possesses the properties of a resin. The residuum of the amber is not acted on by alcohol. Though amber be roasted before the action of the alcohol, the tincture is still formed. Hence we learn that the resinous part of amber is not expelled by a melting heat.* When amber is treated with a boiling fixed alkali, it is almost completely dissolved, according to Hoffmann, and the compound possesses the qualities of soap; for it is soluble in water and alcohol, and not thrown down by water. Mr. Hatchett found that the alkalis act only partially on amber, extracting a yellow tincture. Probably this ingenious chemist did not continue the process long enough; for I have accidentally ascertained, that a weak solution of potash is capable of dissolving amber completely without the assistance of heat, provided it be allowed to act for a sufficient time. I had formed a weak solution of potash (I believe carbonate) as nearly as possible of the specific gravity of amber, and I had put into it some amber powder, to show the supposed currents of Count Rumford during the heat of the liquid. On examining the infusion about a month after, I found the amber all lying at the bottom of the phial. I added more alkali to restore the equilibrium. Some time after the amber was again at the bottom, and it was necessary to add more alkali. By this time the solution had acquired a yellow colour. I therefore explained the sinking of the amber, by supposing that the potash had dissolved a portion of it, and that this had altered the specific gravity of the solution. Not knowing at the time that any experiments had been made on the subject, I put aside the phial to ascertain the result. Three years after only two or three particles of the amber at most could be detected, the rest having dissolved completely.

Action of
alkalies.

The weaker acids have no action on amber. Sulphuric acid converts it into a black resinous mass. Nitric acid acts upon it; when assisted by heat, nitrous gas is emitted. The amber is first converted into a light resinous substance, and at last dissolves completely. Heyer, who first made this experiment, could obtain neither oxalic nor acetic acid by the

of acids.

* Heyer.

Book IV. action of nitric acid on amber. That nitric acid is really capable of dissolving amber has been lately verified by the experiments of Hatchett, who found it soluble with the same phenomena as resins in general.

Amber varnish.

Neither fixed nor volatile oils have any action on amber unless it has been previously roasted or exposed to a melting heat. When thus treated, it combines with oils, and the solution forms *amber varnish*. The process recommended by Nystrom is this: amber is to be spread on a flat-bottomed iron pan, and placed on an equal coal fire till it melt; it is then to be withdrawn, covered with a plate of copper and iron, and allowed to cool. If the process be properly conducted, the amber will have lost half of its weight. If the fire be too strong, the amber will be scorched and rendered useless. If it be too low, the amber will not melt, but be reduced to a brown crust, which answers well enough for a varnish, provided it be exposed to heat till it is reduced to one-half of the original weight. One part of this roasted amber is to be mixed with three parts of linseed oil (rendered drying by litharge and white vitriol), and the mixture exposed to a gentle heat till the amber is dissolved: it is then to be withdrawn from the fire, and when nearly cold four parts of oil of turpentine are to be added. The whole is then allowed to settle, and the clear portion is passed through a linen cloth.

When amber is distilled, there comes over carburetted hydrogen and carbonic acid gas, an acidulous water, then an oil, at first thin and transparent, but becoming gradually darker and thicker. Towards the end of the process the succinic acid sublimes. From the observations of Vogel, it appears that a portion of the yellow matter which is obtained by this distillation possesses the properties of wax very nearly.*

Resin from bitumen.

15. *Resin from bitumen*. Mr. Hatchett ascertained, that when asphalt, mineral caoutchouc, and other similar bodies, are digested a sufficient time in nitric acid, they furnish a considerable portion of a brown substance, which possesses many of the properties of resins. Kilkenny coal, and every other variety of coal which contains no bitumen, is converted by nitric acid into artificial tannin, but yields no resin; while common coal furnishes both. Thus it appears that bitumen, when treated with nitric acid, has the property of yielding a species of resin.

* Gehlen's Jour. v. 272.

This substance has a pale brown-colour like that of Spanish snuff; its internal fracture is dark-brown, and it has a resinous lustre. When heated it does not readily melt; but when inflamed emits a resinous odour, mixed with that of fat oils, and leaves a coal much more bulky than the original substance. Alcohol dissolves it. Water throws down a part from the solution; but a portion remains undissolved, which acts on reagents like *extractive*. The taste is bitter. Hence this substance appears to be intermediate between extractive and resin. When digested with nitric acid it is readily converted into artificial tannin; and when digested with sulphuric acid it is converted into charcoal.*

I have now enumerated the most remarkable of the resinous bodies with which we are acquainted. There are indeed many of the substances employed in medicine, which undoubtedly contains resins; for example, *aloes*, *jalap*, *snake-root*, *arnica*, &c.; but as these have not yet been chemically examined, it would be in vain to dwell upon them.

SECT. XXVII.

OF GUAIIACUM.

THIS substance is obtained from the *guaiacum officinale*, a Original. tree which is a native of the West Indies, and yields a very hard heavy wood. The resin exudes spontaneously, and is also driven out artificially by heating one end of the wood in billets previously bored longitudinally; the melted resin runs out at the extremity farthest from the fire. This substance has been used in medicine for a considerable time, having been originally recommended in venereal diseases. Nothing is known concerning its original introduction into Europe.

It was considered by chemists as a resin, till Mr. Hatchett observed, that when treated with nitric acid it yielded products very different from those of resinous bodies.† This induced Mr. William Brande to examine its chemical properties in detail.‡ To his valuable paper we are indebted for almost

* Hatchett's Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

† Second Series of Experiments on Artificial Tannin, Phil. Trans. 1805.

‡ Chemical Experiments on Guaiacum, Phil. Trans. 1806, and Phil. Mag. xxv. 105.

Book IV. all the accurate information which we possess respecting its chemical nature.

Properties. 1. Guaiacum is a solid substance, resembling a resin in appearance. Its colour differs considerably, being partly brownish, partly reddish, and partly greenish; and it always becomes green when left exposed to the light in the open air.* It has a certain degree of transparency, and breaks with a vitreous fracture. When pounded it emits a pleasant balsamic smell, but has scarcely any taste, although when swallowed it excites a burning sensation in the throat. When heated it melts, and diffuses at the same time a pretty strong fragrant odour. Its specific gravity is 1.2289.†

Action of water. 2. When guaiacum is digested in water a portion of it is dissolved, the water acquiring a greenish-brown colour and a sweetish taste. The liquid, when evaporated, leaves a brown substance, which possesses the properties of *extractive*; being soluble in hot water and alcohol, but scarcely in sulphuric ether, and forming precipitates with muriates of alumina, tin, and silver. This extractive amounts to about nine parts in the hundred of guaiacum.‡

Alcohol. 3. Alcohol dissolves guaiacum with facility, and forms a deep brown-coloured solution. Water renders this solution milky by separating the resin. Muriatic acid throws down the guaiacum of an ash-grey, and sulphuric acid of a pale-green colour. Acetic acid and the alkalies occasion no precipitate. Liquid chlorine throws it down of a fine pale-blue, which does not change when dried. Diluted nitric acid occasions no change at first; but after some hours the liquid becomes green, then blue, and at last brown, and at that period a brown coloured precipitate falls down. If water be mixed with the liquid when it has assumed a green or a blue-colour, green and blue precipitates may be respectively obtained.§

Ether. 4. Sulphuric ether does not act so powerfully on guaiacum as alcohol. The solution obtained by means of it exhibits the same properties when treated with reagents as that in alcohol.||

Alkalies. 5. The alkaline solutions, both pure and in the state of carbonates, dissolve guaiacum with facility. Two ounces of a

* Wollaston, Nicholson's Jour. viii, 294.

† Brande, Phil. Mag. xxv, 105.

§ Ibid. p. 106.

‡ Ibid.

|| Ibid. p. 106.

saturated solution of potash dissolved about 65 grains of guaiacum; the same quantity of ammonia only 25 grains; or guaiacum dissolves in about 15 parts of potash and 38 parts of ammonia. Nitric acid threw down from these solutions a brown precipitate, similar to what is obtained when the alcoholic solution is mixed with the same acid. Muriatic acid and diluted sulphuric acid throw down a flesh-coloured curdy precipitate, which in its properties approaches the nature of extractive.*

6. Most of the acids act upon guaiacum with considerable energy. Action of acids.

Sulphuric acid dissolves it, and forms a deep-red liquid, which deposits while fresh a lilac-coloured precipitate when mixed with water. When heat is applied the guaiacum is charred.

Nitric acid dissolves guaiacum completely without the assistance of heat, and with a strong effervescence. When the solution is evaporated, it yields a very large quantity of oxalic acid.† No artificial tannin appears to be formed, but rather a substance possessing the properties of extractive. Diluted nitric acid converts guaiacum into a brown substance, similar to the precipitate obtained by nitric acid from the alcoholic solution of guaiacum. This brown matter possesses the properties of a resin.‡

Muriatic acid acts but slightly, as the guaiacum soon melts into a blackish mass, which is not acted upon.§

7. When guaiacum is distilled, 100 parts of it yielded to Distillation. Mr. Brande the following products:

Acidulous water	5·5
Thick brown oil	24·5
Thin empyreumatic oil	30·0
Charcoal	30·5
Gases, consisting of carbonic acid and carburetted hydrogen	9·0
Loss	0·5
	100·0

The coal when incinerated left three grains of lime, but no alkaline substance.||

* Brande, Phil. Mag. xxv. 109.

† Hatchett, Second Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

‡ Brande, Phil. Mag. xxv. 107.

§ Ibid.

|| Ibid.

Book IV.

8. Such are the properties of guaiacum, as far as they have been hitherto ascertained. From the preceding detail, it is obvious that guaiacum in many respects coincides with the resins: but it differs from them in three particulars so remarkable, that we cannot avoid considering it as a distinct substance.

Peculiar
properties.

The first of these is the great quantity of charcoal which it leaves when distilled in close vessels. This Mr. Brande found to amount to above 30 per cent.; while the resins, in like circumstances, hardly ever leave more than 15 per cent. of charcoal, and often not nearly so much. It is possible, however, and indeed not improbable, that this difference is to be ascribed to the different degrees of heat employed.

The second peculiarity is the action of nitric acid on guaiacum. This acid does not dissolve the resins without the assistance of heat, but converts them into a brown brittle mass; whereas it dissolves guaiacum completely. The action of this acid on the resins terminates in the formation of artificial tannin, whereas its action on guaiacum terminates in the formation of oxalic acid. This striking difference is alone sufficient to warrant a separation of guaiacum from the resins.

Forms
oxides.

The third peculiarity is the remarkable suite of changes of colour which guaiacum undergoes when its solutions are treated with nitric acid and chlorine. Dr. Wollaston first observed that guaiacum becomes green when exposed to light, provided air have access to it; and that the colour is again removed by the application of heat.* Hence it is probable that oxygen occasions the change. This opinion is much strengthened by the experiments of Mr. Brande. When guaiacum was put in contact with oxygen gas, it became green sooner than in the open air. When put into chlorine gas it became first green, then blue, and lastly brown; and ammonia, when left in contact with it, restored again its green colour. In like manner, by treating the alcoholic solution of guaiacum with nitric acid, green, blue, and brown precipitates are obtained, according to the length of time that the acid is allowed to act upon it. These facts give considerable plausibility to the opinion of Mr. Brande, that the changes of colour are owing to the combinations of oxygen with the guaiacum; that the green contains the least, and the

* Nicholson's Jour. viii. 295.

brown the most oxygen, while the blue is intermediate.* Thus guaiacum, in its changes of colour, bears some resemblance to indigo. Mr. Brande has remarked a coincidence also between guaiacum and the green resin of the leaves of plants.

Chap. I.

SECT. XXVIII.

OF BALSAMS.

THE term *balsam* or *balm* was originally confined to a thick fragrant juice obtained from the *amyris Gileadensis*, and afterwards applied by chemists to all substances which possessed the same degree of consistence and a strong smell, whether natural or artificial. Bucquét restricted the term to those resinous-like substances which yield benzoic acid when heated. This new meaning of the word, which has been adopted by chemists in general, has introduced into the class of balsams several substances which were formerly considered as resins. The word *balsam* originally implied a substance possessing a certain degree of fluidity; but now there are two classes of balsams; the one fluid, and the other solid and brittle.

A balsam, then, is a substance which possesses the general properties of a resin; but which, when heated or digested in acids, yields a portion of benzoic acid. Chemists, in general, have considered them as combinations of a resin with benzoic acid; but Mr. Hatchett has made it probable, that the acid is formed at the time of its separation.†

They are insoluble in water; but when boiled in that liquid often give out a portion of benzoic acid. Alcohol and ether dissolve them readily. The strong acids likewise dissolve them; and during the solution a portion of benzoic acid is separated. Nitric acid, in some cases, evolves likewise traces of prussic acid. The alkalies act upon them nearly as on the resins. They may be divided into two classes; namely, *liquid* and *solid* balsams.

I. LIQUID BALSAMS.

The liquid balsams at present known are five in number; namely,

* Brande, Phil. Mag. xxv. 107.

† Second Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

Book IV.

- | | |
|-----------------|------------|
| 1. Opobalsamum. | 4. Peru. |
| 2. Copaiva. | 5. Styrax. |
| 3. Tolu. | |

Opobalsamum.

1. *Opobalsamum* or *balm of Gilead*.—This balsam is obtained from the *amyris Gileadensis*, a tree which grows in Arabia, especially near Mecca. It is so much valued by the Turks that it is seldom or never imported into Europe. We are of course ignorant of its composition. It is said to be at first turbid and white, and of a strong aromatic smell, and bitter, acrid, astringent taste; but by keeping, it becomes limpid and thin, and its colour changes first to green, then to yellow, and at last it assumes the colour of honey, and the consistence of turpentine.*

Copaiva.

2. *Copaiva*.—This balsam is obtained from the *copaifera officinalis*; a tree which grows in South America, and some of the West Indian islands. It exudes from incisions made in the trunk of the tree. The juice thus obtained is transparent, of a yellowish colour, an agreeable smell, a pungent taste, at first of the consistence of oil, but it gradually becomes as thick as honey. Its specific gravity is 0.950.† When mixed with water and distilled there comes over with the water a very large portion of volatile oil. Lewis obtained half the original weight of this oil.‡ Schonberg, from eight ounces of copaiva, obtained $3\frac{1}{4}$ of oil by this process.§ It was colourless, very limpid, of the specific gravity 0.900; had the taste and smell of copaiva, but rather stronger. It dissolved in eight parts of alcohol; but the copaiva itself is a good deal more soluble.|| The oil ceases to come over before all the water has passed into the receiver. The residuum, of course, consists of two substances; namely, the watery portion, and a greyish-yellow substance, lying at the bottom of the vessel, which, on exposure to the air, dries, and becomes brittle and transparent. When heated it melts, and possesses the characters of a resin. When distilled it yielded a yellowish thick oil, some acidulous water, and a gas; one-sixth of which was carbonic acid, and the remainder seemed to possess the characters of olefiant gas.** From these facts, which have been long known, it was

Yields oil

And resin,

* This is the account of Professor Alpinus, as quoted by Lewis, Neumann's Chem. p. 284.

† Schonberg, Gehlen's Jour. vi. 494.

‡ Neumann's Chem. p. 285.

§ Schonberg, Gehlen's Jour. vi. 494.

§ Gehlen's Jour. vi. 494.

** Schonberg, ibid. 497.

concluded, that copaiva is a compound of a resin and a volatile oil, which passes over at a heat inferior to that of boiling water; but the experiments of Schonberg have rendered it much more probable, that the balsam is decomposed when distilled along with water, and that both the oil and resin are new products.

Chap. I.

But is decomposed.

When distilled on a water bath, nothing comes over but a few drops of water, and one or two drops of oil.* If the vessel be kept at a temperature between 234° and 257° , scarcely any thing more is obtained than when the distillation is conducted over a water bath. At the temperature of 504° the balsam begins to boil gently, a gas is extricated, and drops begin to pass more rapidly into the receiver. At 550° it boils briskly, and the distillation goes on rapidly. There passes into the receiver a limpid yellowish oil, occasionally mixed with a drop or two of water. As the distillation proceeds the oil becomes more and more yellow. At this period the balsam is as liquid as water, and boils without any frothing or swelling. After this period the oil becomes yellow, and then brownish-red; but still continues pretty thin. The first oil obtained by Schonberg from eight ounces of copaiva, by this process, was $3\frac{1}{16}$ ounces. The water had a sour taste, and reddened litmus paper. The gas amounted to 81 ounce measures; $\frac{1}{9}$ th of it was carbonic acid, the rest resembled olefiant gas.†

Distillation

Nitric acid acts upon this balsam with considerable energy. When one part of the balsam is mixed with four parts of nitric acid and two parts of water, and heated, a yellowish solution is formed, similar to the original balsam, but darker. When distilled, there comes over with the liquid that passes into the receiver an apple-green oil, which lines the helm of the retort. The nature of the residue was not examined. Doubtless it would have been found to contain artificial tannin, provided a sufficient quantity of acid was employed.‡

Action of nitric acid.

Similar to pentin.

When treated with sulphuric acid, it yields a portion of artificial tannin.§

Whether this balsam yields benzoic acid has not been ascertained. Its properties are rather against the probability of its doing so. Indeed it bears a striking resemblance to tur-

* Schonberg, Gehlen's Jour. vi. p. 495.

+ Id. ibid.

† Id. ibid. p. 499.

§ Hatchett, Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

Book IV. pentine in many respects; and ought, along with it, to constitute a class of bodies intermediate between volatile oils and resins, to which the name of *turpentine*s might be given.

Balsam of tolu. 3. *Balsam of Tolu*.—This substance is obtained from the *toluifera balsamum*, a tree which grows in South America. The balsam flows from incisions made in the bark. It comes to Europe in small gourd shells. It is of a reddish brown colour and considerable consistence; and when exposed to the air, it becomes solid and brittle. Its smell is fragrant, and continues so even after the balsam has become thick by age. When distilled with water, it yields very little volatile oil, but impregnates the water strongly with its taste and smell. A quantity of benzoic acid sublimes, if the distillation be continued.*

Action of alkalis,

Mr. Hatchett found it soluble in the alkalis like the rest of the balsams. When he dissolved it in the smallest possible quantity of lixivium of potash, it completely loses its own odour, and assumes a most fragrant smell, somewhat resembling that of the clove pink. "This smell," Mr. Hatchett observes, "is not fugitive, for it is still retained by a solution which was prepared in June, and has remained in an open glass during four months."

And acids.

When digested in sulphuric acid, a considerable quantity of pure benzoic acid sublimes. When the solution of it in this acid is evaporated to dryness, and the residue treated with alcohol, a portion of artificial tannin is obtained; the residual charcoal amounts to 0.54 of the original balsam.†

Mr. Hatchett found that it dissolved in nitric acid with nearly the same phenomena as the resins; but it assumes the odour of bitter almonds, which leads him to suspect the formation of prussic acid. During the solution in nitric acid, a portion of benzoic acid sublimes. By repeated digestions it is converted into artificial tannin.‡

Balsam of Peru.

4. *Balsam of Peru*.—This substance is obtained from the *myroxylon peruiferum*, which grows in the warm parts of South America. The tree is full of resin, and the balsam is obtained by boiling the twigs in water. It has the consistency of honey, a brown colour, an agreeable smell, and a hot

* Lewis, Neumann's Chem. p. 285.

† Hatchett, Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

‡ Ibid.

acid taste. When boiled with water for some time, the liquid separated by the filter reddens vegetable blues, and deposits crystals of benzoic acid on cooling. The water contains no other substance.* When distilled with water, it yields a very small quantity of reddish limpid oil. Hoffmann obtained only one part of oil from 16 of balsam.† Lichtenberg mixed two ounces of balsam with eight of water, and distilled. The first two ounces of water came over colourless, and had only a slight odour of the balsam. The next three were milky, smelt strongly of the balsam, and at the bottom of the receiver were some drops of colourless oil. The next 1½ ounce were similar, but contained more oil. Between 50 and 60 grains of benzoic acid had sublimed into the neck of the retort. By increasing the heat 3½ drachms of yellow oil came over, and a quantity of benzoic acid. A black shining coal remained in the retort.‡

When this balsam is exposed to the heat of a water bath, only a drop or two of water and a few drops of oil can be obtained.§ When placed in a sand bath, and exposed to a temperature gradually raised, nothing more comes over till the balsam is heated to 300°, when a portion of benzoic acid sublimes; and at 324°, drops of water and oil began to come over. At 550° the balsam begins to boil, and some gas is disengaged. At 594° the oil (mixed with a little water) comes over pretty fast. At 617° it comes over still more rapidly. Lichtenberg, to whom we owe these experiments, kept four ounces of balsam at that temperature for two hours, and obtained two ounces of a yellowish oil, and a crystallized mass of benzoic acid, which, together with the water, weighed 6½ drachms. The gas obtained amounted to 58 ounce measures; of these, 38 were carbonic acid. The rest burnt like olefiant gas. By increasing the heat a brownish oil came over, and at last a black oil of the consistence of pitch, and 49 ounce measures of gas were extricated. Of these, six were carbonic acid; the residue burnt with a bluish white flame.||

A saturated solution of carbonate of soda forms with this balsam a thick mass. When diluted with water and heated, a portion is dissolved. The solution, when saturated with sulphuric acid, deposits crystals of benzoic acid. One part of the

* Lichtenberg, Gehlen's Jour. vi. 489.

† Observ. Phys. Chym. Select. p. 71.

§ Lichtenberg, Gehlen's Jour. vi. 485.

‡ Gehlen's Jour. vi. 485.

|| Ibid. p. 487.

Book IV. } balsam, treated with one part of potash dissolved in four parts of water, formed an opaque solution, which gradually separated into two portions: the uppermost, a clear oil with some grey flakes at its lower surface; the undermost, a dark brownish-red opaque solution. This last solution, when saturated with sulphuric acid, lets fall a resinous-like substance, dissolved by boiling, while benzoic acid crystallized.*

Of acids. Nitric acid acts upon the balsam with energy, and gives it an orange-yellow colour when assisted by heat. When distilled with a sufficient quantity of this acid diluted, the liquid in the receiver smells of bitter almonds. When supersaturated with carbonate of potash, and mixed with a solution of iron, a precipitate falls, which, when treated with muriatic acid, leaves prussian blue, and indicates the presence of prussic acid. During the distillation benzoic acid sublimes. The residue in the retort has a crystalline appearance, is light yellow, dissolves sparingly in boiling water, and precipitates on cooling in the state of a yellow powder.† Such was the result of Lichtenberg's trials. Mr. Hatchett observed that the residue possessed the properties of artificial tannin.

When this balsam is treated with sulphuric acid, artificial tannin is also formed, and the residual charcoal amounts to no less than 0.64 of the original weight of the balsam.‡

Styrax. 5. *Styrax*.—This is a semifluid juice, said to be obtained from the *liquidambar styraciflua*, a tree which grows in Virginia, Mexico, and some other parts of America. § It is prepared, according to Mr. Petiver, in the island Cobross in the Red Sea, from the bark of a tree called *rosa mallos* by the natives, and considered by botanists as the same with the American species. The bark of this tree is boiled in salt water to the consistence of bird-lime, and then put into casks.|| Bouillon la Grange has published an account of its properties.** Its colour is greenish, its taste aromatic, and its smell agreeable. It is easily volatilized by heat. When treated with water, benzoic acid is dissolved. It is totally soluble in alcohol except the impurities. When exposed to the air it becomes harder, and absorbs oxygen. When distilled, it yields an acidulous

* Lichtenberg, Gehlen's Jour. vi. 487.

† Gehlen's Jour. vi. 491.

‡ Hatchett, Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

§ Lewis, Neumann's Chemistry, p. 291.

|| Phil. Trans. 1708, vol. xxvi. p. 44.

** Ann. de Chim. xxvi. 203.

water, having the odour of benzoic acid, a limpid colourless hot oil, a solid coloured oil, benzoic acid, and a mixture of carbonic acid and carburetted hydrogen. The charcoal is light and contains some oil.

II. SOLID BALSAMS.

The solid balsams at present known are only three in number; namely,

1. Benzoin.
2. Storax.
3. Dragon's blood.

1. *Benzoin*.—This substance is the produce of the *styrax* Benzoin. *benzoe*, a tree which grows in Sumatra, &c. and which has been described by Mr. Dryander.* Benzoin is obtained from this tree by incision; a tree yielding three or four pounds. It is a solid brittle substance, sometimes in the form of yellowish white tears joined together by a brown substance, and sometimes in the form of a brown substance not unlike common rosin. It has a very agreeable smell, which is increased by heating the benzoin. It has little taste. Its specific gravity is 1.092. This substance has been used in medicine for ages, and various processes have been pointed out by chemists for extracting benzoic acid from it: but the only person who has examined its properties in detail is Mr. Brande.†

Cold water has very little effect on benzoin, but boiling water takes up a portion of benzoic acid.

Alcohol dissolves it when assisted by a gentle heat, and forms a deep yellow solution inclining to reddish-brown. When this solution is diluted with water, the benzoin precipitates in the form of a white powder. It is precipitated also by muriatic and acetic acids, but not by the alkalies. A few drops of sulphuric acid likewise precipitate the benzoin; but an additional quantity redissolves it, and forms a liquid of the colour of port wine. When equal quantities of the alcoholic solution of benzoin and sulphuric acid are mixed, a dark-pink precipitate falls. The liquid assumes a pink colour, which becomes lilac when diluted with water. Nitric acid occasions a strong effervescence, and forms a dark-red fluid with the alcoholic solution, but throws down no precipitate.‡

Action of alcohol.

* Phil. Trans. 1787, p. 307.

† Nicholson's Jour. x. 82.

‡ Brande, Nicholson's Jour. x. 82.

Book IV.
Ether.
Acids.

Ether dissolves benzoin with facility, and the solution with reagents exhibits the same phenomena as the alcoholic.*

Nitric acid acts with violence on benzoin, and converts it into an orange coloured mass. When assisted by heat the acid dissolves the benzoin; and as the solution cools, crystals of benzoic acid gradually separate. Mr. Hatchett ascertained that by this process a quantity of artificial tannin is formed.

Sulphuric acid dissolves benzoin, while benzoic acid (as Hatchett discovered) sublimes; the solution is at first a deep red. By continuing the digestion, a portion of artificial tannin is formed, and the charcoal evolved amounts to 0.48 of the benzoin dissolved.†

Acetic acid dissolves benzoin without the assistance of heat. When heat is applied, the solution, as it cools, becomes turbid; owing to the separation of benzoic acid.‡

Alkalies.

Benzoin is dissolved by a boiling lixivium of the fixed alkalies; a dark-brown solution is formed, which becomes turbid after some days' exposure to the air. Ammonia likewise dissolves benzoin sparingly.§

Distillation.

When Mr. Brande exposed 100 grains of benzoin in a retort to a heat gradually raised to redness, the products were,

Benzoic acid	9.0
Acidulous water	5.5
Butyraceous and empyreumatic oil	60.0
Charcoal	22.0
Carburetted hydrogen and carbonic acid .	3.5
	100.0

Bucholz subjected 1500 grains of benzoin to a chemical analysis. He obtained the following substances,

Resin	1250
Benzoic acid	187
Substance similar to balsam of Peru	25
Aromatic substance soluble in water and alcohol	8
Woody fibres and impurities	30
	1500

* Brande, Nicholson's Jour. x. 82.

† Hatchett, Third Series of Experiments on Artificial Tannin.

‡ Brande, Nicholson's Jour. x. 85.

§ Ibid. p. 86.

|| Ann. de Chim. lxxxiv. 319.

2. *Storax*.—This is the most fragrant of all the balsams, and is obtained from the *styrax officinalis*, a tree which grows in the Levant, and it is said also in Italy. Sometimes it is in the state of red tears; and this is said to be the state in which it is obtained from the tree. But common storax is in large cakes; brittle, but soft to the touch, and of a reddish-brown colour. This is more fragrant than the other sort, though it contains a considerable mixture of saw-dust. It dissolves in alcohol. When distilled with alcohol or with water, scarcely any oil is obtained. When distilled by the naked fire, it seems, from the experiments of Neumann, to yield the same products as benzoin.*

Chap. I.
Storax.

3. *Dragon's blood*.—This is a brittle substance of a dark-red colour, which comes from the East Indies. There are two sorts of it; one in small oval drops or tears of a fine deep-red, which becomes crimson when the tears are reduced to powder; the other is in larger masses, some of which are pale-red, and others dark. It is probably obtained from different kinds of trees; the *calamus draco* is said to furnish most of what comes from India. The *dracæna draco* and the *ptero-carpus draco* are also said to furnish it.

Dragon's
blood.

Dragon's blood is brittle and tasteless, and has no sensible smell. Water does not act upon it, but alcohol dissolves the greatest part, leaving a whitish-red substance, partially acted upon by water. The solution has a fine deep-red colour, which stains marble, and the stain penetrates the deeper the hotter the marble is. It dissolves likewise in oils, and gives them a deep-red colour also. When heated it melts, catches flame, and emits an acid fume similar to that of benzoic acid.† When digested with lime, a portion of it becomes soluble in water, and it acquires a balsamic odour. On adding muriatic acid to the solution, a red resinous substance is precipitated, and slight traces of benzoic acid only become perceptible.‡ Nitric acid acts upon it with energy, changes it to a deep-yellow, a portion of benzoic acid is sublimed, and a brown mass remains soluble in water and possessing the properties of artificial tannin.§ When treated with sulphuric acid no perceptible portion of benzoic acid sublimes; but it is con-

Properties.

* Neumann's Chem. p. 290.

† Lewis, Neumann's Chem. p. 299.

‡ Hatchett, Second Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

§ Hatchett, *ibid*.

Book IV.

verted partly into artificial tannin, while a quantity of charcoal is evolved, amounting to 0.48 of the original dragon's blood employed.*

SECT. XXIX.

OF GUM RESINS.

THIS class of vegetable substances has been long distinguished by physicians and apothecaries. It contains many active substances much employed in medicine; and they certainly possess a sufficient number of peculiar properties to entitle them to be ranked apart. Unfortunately these substances have not yet attracted much of the attention of chemists. Their properties and constituents of course are but imperfectly ascertained. Of late, however, they have engaged the attention of M. Braconnot, who has recently published a very detailed examination of several of them, and has promised to examine in the same manner the whole of the class.† They may be distinguished by the following characters.

Properties.

They are usually opaque, or at least their transparency is inferior to that of resins. They are always solid, and most commonly brittle, and have sometimes a fatty appearance.

When heated they do not melt as the resins do; neither are they so combustible. Heat, however, commonly softens them, and causes them to swell. They burn with a flame.

They have almost always a strong smell, which in several instances is alliaceous. Their taste also is often acrid, and always much stronger than that of the resins.

They are partially soluble in water; but the solution is always opaque, and usually milky.

Alcohol dissolves only a portion of them. The solution is transparent; but when diluted with water it becomes milky; yet no precipitate falls, nor is any thing obtained by filtering the solution.

Vinegar and wine likewise dissolve them partially; and the solution, like the aqueous, is opaque or milky.

According to Hermbstadt, they are insoluble in sulphuric ether.

* Hatchett, Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

† Ann. de Chim. lxxviii. 19.

The action of alkalis on them has been examined only by Mr. Hatchett. All of them tried by that celebrated chemist dissolved readily in alkaline solutions when assisted by heat. We may therefore consider them as soluble in alkalis like resins.

From the experiments of Mr. Hatchett, we learn that the acids act on these bodies nearly as they do on the resins. Sulphuric acid dissolves them, and gradually converts them into artificial tannin and charcoal. The following are the quantities of charcoal obtained by this chemist from 100 grains of different gum resins digested in alcohol :

Ammoniac	58	Myrrh	40
Asafœtida	51	Gamboge	31
Olibanum	44		

Nitric acid acts upon them with energy; converting them first into a brittle mass, and then, with the assistance of heat, dissolving them. By evaporating this solution, Mr. Hatchett obtained, from ammoniac and asafœtida, a portion of artificial tannin; but he did not succeed in procuring it by the same means from olibanum, myrrh, and gamboge.*

Their specific gravity is usually greater than that of the resins.

Their other properties still continue unknown. They all either exude spontaneously from plants, or are obtained by incisions. At first they seem to be in a liquid state; but they gradually harden when exposed to the air and weather.

They have been usually considered by chemists as composed of gum and resin; but their properties are not consistent with that supposition. They all contain a volatile oil, or a substance intermediate between an oil and resin. To this substance we are to ascribe the milky solution which they form with water. The other constituent, in most cases, bears a much closer resemblance to extractive than to gum. Perhaps, then, we shall not err very far, if we consider the gum resins as composed of a gum or an extractive substance, and a body intermediate between oil and resin; to which last they owe their most peculiar properties.

The gum resins, which have been hitherto applied to any useful purpose, are the following :

* Phil. Trans. 1806.

Book IV.
Galbanum.

1. *Galbanum*. It is obtained from the *bubon galbanum*, a perennial plant, and a native of Africa. When this plant is cut across a little above the root, a milky juice flows out, which soon hardens and constitutes galbanum. It comes to this country from the Levant, in small pieces composed of tears, agglutinated together, of a yellowish or white colour. Its taste is acrid and bitter, and its smell peculiar. Water, vinegar, and wine, dissolve part of it, but the solution is milky. Alcohol dissolves about $\frac{3}{5}$ ths. When distilled it yields about half its weight of volatile oil, which has at first a blue colour. Its specific gravity is 1.212.*

According to the analysis of M. Meisner, it is composed as follows:

Resin	65.8
Gum	22.6
Cerasin	1.8
Malic acid	0.2
Volatile oil	3.4
Vegetable debris	2.8
Loss	3.4
	100.0 †

Ammoniac.

2. *Ammoniac*.—This substance is brought from the East Indies. Nothing certain is known concerning the plant which yields it; though from analogy it has been suspected to be a species of *ferula*. † It is in small pieces agglutinated together, and has a yellowish-white colour. Its smell is somewhat like that of galbanum, but more pleasant. Its taste is a nauseous sweet mixed with bitter. It does not melt. Water dissolves a portion of it; the solution is milky, but gradually lets fall a resinous portion. More than one-half is soluble in alcohol. This portion is a resin. Colour white, soft, and ductile. Melts when heated, and burns like a resin. When heated becomes harder, but not brittle. Nearly tasteless. Soluble in ether and in nitric acid. Precipitated from it in the form of an orange substance, partly resin, partly bitter. A portion

* Brisson.

† Annals of Philosophy, xiv. 385.

‡ Wildenow made the seeds vegetate which are observed in aminoniac. Of the plant produced he has formed a new genus under the name of *heracleum gummifera*. But as the roots of this plant were not observed to yield *ammoniac*, it is still doubtful how far it is the plant from which ammoniac is procured. See Ann. de Chim. lxi. 267.

remains, and gives the liquid a yellow colour. Taste of the solution slightly acid and bitter. Not precipitated by carbonate of soda, ammonia, nitrate of silver, nor acetate of lead. The orange substance has a bitter taste. When heated, readily swells and blackens, but does not flame. Burns without leaving any residuum. Lighter than water. When agitated in water, tinges it yellow, but does not all dissolve. The specific gravity of ammoniac is 1.207. Mr. Hatchett found it soluble in alkalies. Neither alcohol nor water, when distilled off it, bring over any thing.

According to the analysis of Braconnot, ammoniac is composed of the following ingredients:

Resin	70.0
Gum	18.4
Glutinous matter	4.4
Water	6.0
Loss	1.2
	100.0

The resin he found brittle and yellow. In these respects it differs from the resin which I extracted from ammoniac, which was soft, and continued so after being exposed to the air for two months. This difference is probably owing to the state of the ammoniac, sometimes it is brittle and yellow, at other times soft and white. It was in this last state that I examined it. Braconnot found the yellow matter into which this resin is converted by nitric acid soluble in hot alcohol and water. It had the property of dyeing silk a fine yellow colour, not altered by chlorine. The gum which he extracted from ammoniac possessed the properties of common gum as far as he examined them. It is transparent, yellowish, brittle, soluble in water, and precipitated by subacetate of lead, but not by the acetate, nor the nitrate of lead. The mercurial salts render the solution milky. By nitric acid it is converted into saccharic and oxalic acids, and furnishes also a little malic acid. The glutinous matter was insoluble in water and alcohol; it became black when dried, and yielded a yellow matter and some oxalic acid when treated with nitric acid.*

3. *Aloes*.—This is the inspissated juice of the leaves of a Aloes.

* Ann. de Chim. lxxviii. 69.

Book IV.

variety of the *aloe perfoliata*, a plant which grows in Socotora, an island in the East Indies. Aloes has a resinous appearance, a reddish-yellow colour, a bitter taste, and an aromatic smell. It has been recently examined by Braconnot, who found it composed chiefly of a peculiar matter similar to that detected by Vauquelin in many febrifuge barks, and to which Braconnot proposes to give the name of *resinous bitter principle*.* Trommsdorf, on the other hand, and Bouillon La Grange and Vogel, consider it as composed of resin and a peculiar extractive matter.† Fabroni has discovered that the recent juice of the leaves of the aloe has the property of absorbing oxygen, and of assuming a fine reddish-purple, and of yielding a pigment which he strongly recommends to the painter.‡

Olibanum.

4. *Olibanum*.—This substance is the frankincence of the ancients. They obtained it, as Dioscorides informs us, from Arabia and India. From Pliny we learn that the tree which yielded it was neither known to the Greeks nor Romans. At present olibanum is imported to London in chests, containing each about a hundred weight. It comes from different places, among others from the East Indies, but the Indian olibanum is least esteemed. The tree which yields this gum resin continues still doubtful. The probability is, that different species of trees furnish it in different countries. Lamark is of opinion that the tree which yields the Arabian olibanum is the *amyris gileadensis*. It is called by the natives *soukiou*. According to Foerskal, it is the *amyris kataf*. Mr. Colebrook has shown that the Indian olibanum comes from the *boswellia serrata* of Roxbourgh, a large tree, which grows on the mountains of India. §

Olibanum is a semitransparent, brittle, whitish-yellow substance, usually covered externally by a whitish farinaceous substance, produced by the pieces rubbing against each other. Its taste is acrid and aromatic, and when burnt, it diffuses an agreeable odour. When heated, it melts with difficulty, but burns brilliantly, and leaves a whitish ash composed of phosphate of lime, carbonate of lime and sulphate, muriate and carbonate of potash. According to the analysis of Braconnot, it is composed of the following constituents:

* Ann. de Chim. lxxviii. 18.

† Ibid. p. 11 and 155.

‡ Ibid. xxv. 301.

§ Asiatic Researches, ix. 377.

Volatile oil	8
Resin	56
Gum	30
Matter like gum, insoluble in water and alcohol	5·2
Loss	0·8

Chap. I.

100·0 *

The oil had a lemon colour, and a smell similar to the oil of lemons.

The resin is reddish-yellow, brittle, tasteless, and very similar in appearance to rosin. Boiling water softens it; but a higher temperature is necessary to melt it. When burnt, it emits rather an agreeable odour. When heated to dryness with potash ley, it leaves a matter which is capable of forming a kind of emulsion with boiling water.

The gum possesses few peculiarities. The infusion of nut-galls occasions a precipitate in its aqueous solution. Nitric acid converts it partly into saccharic acid.†

5. *Sagapenum*.—The plant which yields this gum resin is *Sagapenum*. not well known; but it is suspected to be the *ferula persica*. The substance itself is brought to Europe from Alexandria. It is commonly in tears agglutinated together. Colour yellow. Taste hot and bitter. Smell alliaceous. Softens between the fingers, but does not melt when heated. Sparingly soluble in water, but almost completely soluble in alcohol. When distilled with water it yields a little volatile oil. The water is strongly impregnated with the flavour of the sagapenum. †

6. *Asafetida*.—This substance is obtained from the *ferula Asafetida*, a perennial plant which is a native of Persia. When the plant is about four years old, its roots are dug up and cleaned. Their extremity being then cut off, a milky juice exudes, which is collected. Then another portion is cut off, and more juice exudes. This is continued till the roots are exhausted. The juice thus collected soon hardens and constitutes *asafetida*. It comes to Europe in small grains of different colours, whitish, reddish, violet, brown. Pretty hard, but brittle. Its taste is acrid and bitter; its smell strongly alliaceous and fetid. Alcohol, according to Neumann, dissolves about $\frac{1}{4}$ ths of this substance; and water takes up nearly

* Ann. de Chim. lxxviii. 60.

† Braconnot, Ann. de Chim. lxxviii. 60.

‡ Neumann's Chem. p. 316.

Book IV. $\frac{1}{4}$ th if applied before the spirit. A considerable portion of earthy matter remains undissolved. It yields a volatile oil, both when distilled with water and alcohol. This oil possesses the active properties of the asafœtida itself.* The specific gravity of the gum resin is 1.327.

Scammony. 7. *Scammony*.—This substance is obtained from the *convolvulus scammonia*, a climbing plant which grows in Syria, and was first correctly described by Dr. Russel.† The roots when cut yield a milky juice. This when collected and allowed to harden constitutes scammony. Colour dark-grey or black. Smell peculiar and nauseous; taste bitter and acrid. With water it forms a greenish coloured opaque liquid. Alcohol dissolves the greatest part of it. It is usually mixed with the expressed juice of the root, and frequently also with other impurities, which alter its appearance. In medicine it operates as a strong cathartic. Its specific gravity is 1.235.‡

Vogel and Bouillon La Grange have analyzed the two varieties of scammony that come from Aleppo and from Smyrna. The scammony of Aleppo was composed of

Resin	60
Gum	3
Extractive	2
Vegetable debris, earth, &c.	35
	<hr/>
	100

The scammony of Smyrna was composed of

Resin	29
Gum	8
Extractive	5
Vegetable debris, &c.	58
	<hr/>
	100 §

Opoponax. 8. *Opoponax*.—This substance is obtained from the *pastinaca opoponax*, a plant which is a native of the countries round the Levant. The gum resin, like most others, is obtained by wounding the roots of the plant. The milky juice, when dried in the sun, constitutes the opoponax. It is in lumps of

* Neumann's Chemistry, p. 312.

† See an abridgment of his account by Dr. Lewis, Neumann's Chem. p. 303.

‡ Brisson.

§ Ann. de Chim. lxxii. 69.

a reddish-yellow colour, and white within. Smell peculiar. Taste bitter and acrid. With water it forms a milky solution, and about one-half of it dissolves. Alcohol acts but feebly. When distilled with water or alcohol, these liquids acquire the flavour of opoponax, but no oil separates.* Its specific gravity is 1.622.†

When distilled it yields a brown oil, acetic acid in which a bituminous oil swims, the residual charcoal weighs $\frac{2.8}{100}$ of the opoponax distilled. When incinerated it left $\frac{2.7}{100}$ of its weight of ashes composed of

Carbonate of lime	18	
Silica	2	
Carbonate	} of potash	
Sulphate		15
Muriate		—
	35	

According to the analysis of Pelletier, to whom we are indebted for the preceding distillation, opoponax is composed of the following constituents :

Resin	42.0
Gum	33.4
Wood	9.8
Starch	4.2
Malic acid	2.8
Extractive	1.6
Caoutchouc	Trace
Wax	0.3
Volatile oil and loss	5.9
	—
	100.0 †

9. *Gamboge* or *Gumgutt*.—This substance is obtained from *Gamboge*, the *stalagmitis gambogioides*, a tree which grows wild in the East Indies. In Siam it is obtained in drops by wounding the shoots; in Ceylon it exudes from wounds in the bark. It is brought to Europe in large cakes. Its colour is yellow; it is opaque, brittle, and breaks vitreous. It has no smell, and very little taste. With water it forms a yellow turbid liquid. Alcohol dissolves it almost completely; and when mixed with water becomes turbid, unless the solution contain ammonia.

* Neumann's Chem. p. 316.

† Brisson.

‡ Ann. de Chim. lxx. 20.

Book IV.

In that case acids throw down an insoluble yellow precipitate. It operates, when taken internally, as a most violent cathartic. Its specific gravity is 1.221.* It appears that it was brought to Europe by the Dutch about the middle of the 17th century.† It forms a fine yellow paint, and stains hot marble of a fine lemon-yellow.‡

Braconnot analyzed it, and found it composed of one part of a gum which possessed the properties of cherry tree gum, and four parts of a reddish brittle resin which possessed the characteristic properties of the resins. It dissolved in alcohol and alkalis, and by nitric acid was converted into a yellowish bitter matter. Chlorine deprived it of its dark colour, and a combination took place between it and muriatic acid, in which it neutralized that acid. §

Myrrh.

10. *Myrrh.*—The plant from which this substance is obtained is unknown. If we believe Bruce it belongs to the genus of *mimosa*. It grows in Abyssinia and Arabia. It is in the form of tears. Colour reddish-yellow; when pure somewhat transparent, but it is often opaque. Odour peculiar. Taste bitter and aromatic. Does not melt when heated, and burns with difficulty. With water it forms a yellow opaque solution. The solution in alcohol becomes opaque when mixed with water, but no precipitate appears. By distillation with water it yields an oil heavier than water; but nothing comes over with alcohol.|| Its specific gravity is 1.360.** It is employed in medicine. Mr. Hatchett found it soluble in alkalis.

From the analysis of Braconnot it appears that myrrh is composed of about

23 resin
77 gum
<hr style="width: 10%; margin: 0 auto;"/>
100

The resin is reddish, has a bitter taste and the peculiar odour of myrrh. The gum, according to Braconnot, differs in its properties from every other gummy substance hitherto examined. It has a dark-brown colour; is at first soluble in water, but by boiling the liquid, or by exposing the gum to

* Brisson.

† See the *Anatomia Essentiarum Vegetabilium* of Angelus Sala, p. 31.‡ Lewis, *Neumann's Chem.* p. 360.§ *Ann. de Chim.* lxxiii. 33.|| Lewis, *Neumann's Chem.* p. 317.

** Brisson.

heat, it acquires cohesive properties, and becomes insoluble in water. When distilled it yields ammonia, and when dissolved in nitric acid azotic gas is disengaged. It is precipitated by the salts of lead, mercury, and tin, and it falls in combination with the oxides of these metals.*

But the result of the examination of myrrh by Pelletier differs materially from that of Braconnot. According to him it is composed of

Resin containing some volatile oil	33
Gum	67
	100

The resin owed its taste and solubility in water to the volatile oil, when deprived of which it became tasteless and insoluble in water and infusible at the temperature of 212°. The gum possessed the properties of common gum. It was soluble in water after the aqueous solution had been evaporated to dryness. When treated with nitric acid it yielded oxalic acid but no saccharic acid.†

11. *Euphorbium*.—This substance is obtained from the *euphorbia officinalis*. The milky juice which exudes from that plant, when dried in the sun, constitutes euphorbium. It is brought from Africa in small yellow tears. It has no smell, and is mostly soluble in alcohol. Its specific gravity is 1.124.‡ It is considered as poisonous.

According to the analysis of Braconnot it is composed of

Resin	37.0
Wax	19.0
Malate of lime	20.5
Malate of potash	2.0
Water	5.0
Woody matter	13.5
Loss	3.0
	100.0

The resin has peculiar properties. It is reddish and transparent, excessively acrid, and possesses poisonous qualities. It is insoluble in alkalis, but soluble in sulphuric and nitric acid. These properties show it to be a peculiar vegetable

* Ann. de Chim. lxxviii. 52.

† Ibid. lxxx. 45.

‡ Brisson.

Book IV. principle. The wax possesses the properties of bees wax. The malate of lime had been mistaken for gum.*

Bdellium. 12. *Bdellium*.—This gum resin comes from Arabia, and is supposed to be the produce of a species of *amyris*: though nothing certain has been published on that subject. It is in yellowish transparent tears. When triturated between the teeth it becomes soft. Its taste is somewhat acrid. When thrown upon burning coals it becomes soft, and burns like a resin. Its specific gravity is 1.371.† Alcohol dissolves about three-fifths of its weight of bdellium, the undissolved portion is a mixture of gum and cerasin. According to the analysis of Pelletier, its constituents are:

Resin	59
Gum	9.2
Cerasin	30.6
Volatile oil and loss	1.2
	———
	100.0 ‡

Caranna. 13. Little is known concerning the substance called *caranna*, reckoned among the gum resins. Its specific gravity is 1.124.§ The substance extracted from ivy, and known by the name of *gummi hederæ*, is considered at present as a gum resin; but I do not know that it has been chemically examined. Its specific gravity is 1.294.

Ipecacuan. 14. From the experiments made upon ipecacuan, the root of the *cephaelis ipecacuanha*, especially by Dr. Irvine, we learn that it also contains a gum resin. The same remark applies to several other vegetable substances employed in medicine.

It deserves attention, that the gum resins, when subjected to destructive distillation, yield all of them a portion of ammonia; a proof that they all contain azote. In this respect they agree with extractive.

SECT. XXX.

OF CAOUTCHOUC.

History. ABOUT the beginning of the 18th century, a substance called *caoutchouc* was brought as a curiosity from America. It was soft, wonderfully elastic, and very combustible. The pieces of

* Ann. de Chim. lxxiii. 44.

† Ann. de Chim. lxxx. 39.

‡ Brisson.

§ Brisson.

it that came to Europe were usually in the shape of bottles, birds, &c. This substance is very much used in rubbing out the marks made upon paper by a black lead pencil; and therefore in this country it is often called *Indian-rubber*. Nothing was known of its production, except that it was obtained from a tree, till the French academicians went to South America in 1735 to measure a degree of the meridian. M. de la Condamine sent an account of it to the French Academy in the year 1736. He told them, that there grew in the province of Esmeraldas, in Brazil, a tree, called by the natives *Hheré*; that from this tree there flowed a milky juice, which, when inspissated, was *caoutchouc*. Don Pedro Maldonado, who accompanied the French academicians, found the same tree on the banks of the Maragnon; but he died soon after, and his papers were never published. Mr. Fresnau, after a very laborious search, discovered the same tree in Cayenne. His account of it was read to the French Academy in 1751.*

It is now known that there are at least two trees in South America from which caoutchouc may be obtained; the *hævea caoutchouc*, and the *jatropha elastica*; and it is exceedingly probable that it is extracted also from other species of *hævea* and *jatropha*. Several trees likewise which grow in the East Indies yield caoutchouc; the principal of these are, the *ficus indica*, the *artocarpus integrifolia*, and the *urceola elastica*; a plant discovered by Mr. Howison, and first described and named by Dr. Roxburgh.† Dr. Benjamin Smith Barton is said to have obtained it from the juice of the *smilax caduca*, which grows abundantly in the neighbourhood of Philadelphia.‡ Mr. Woodcock found it in the milky juice of the *asclepias vincetoxicum*.§

When any of these plants is punctured, there exudes from it a milky juice, which, when exposed to the air, gradually lets fall a concrete substance, which is caoutchouc.

If chlorine be poured into the milky juice, the caoutchouc precipitates immediately, and at the same time the acid loses its peculiar odour. This renders it probable that the formation of the caoutchouc is owing to its basis absorbing oxygen.|| If the milky juice be confined in a glass vessel containing

* Mem. Par. 1751, p. 319.

† Asiatic Researches, v. 167. London Edition.

‡ Phil. Mag. xl. 66.

§ Ibid. 125.

|| Fourcroy, Ann. de Chim. xi. 259.

Book IV.

common air, it gradually absorbs oxygen, and a pellicle of caoutchouc appears on its surface.*

Caoutchouc was no sooner known than it drew the attention of philosophers. Its singular properties promised that it would be exceedingly useful in the arts, provided any method could be fallen upon to mould it into the various instruments for which it seemed particularly adapted. Messrs. de la Condamine and Fresnau had mentioned some of its properties; but Macquer was the first person who undertook to examine it with attention. His experiments were published in the Memoirs of the French Academy for the year 1768. They threw a good deal of light on the subject; but Macquer fell into some mistakes, which were pointed out by Mr. Berniard, who published an admirable paper on caoutchouc in the 17th volume of the Journal de Physique. To this paper we are indebted for the greater number of facts at present known respecting caoutchouc. Mr. Grossart and Mr. Fourcroy having likewise added considerably to our knowledge of this singular substance; both of their treatises have been published in the 11th volume of the Annales de Chimie.

Properties.

Caoutchouc, when pure, is of a white colour,† and without either taste or smell. The blackish colour of the caoutchouc of commerce is owing to the method employed in drying it after it has been spread upon moulds. The usual way is to spread a thin coat of the milky juice upon the mould, and then to dry it by exposing it to smoke; afterwards another coat is spread on, which is dried in the same way. Thus the caoutchouc of commerce consists of numerous layers of pure caoutchouc alternating with as many layers of soot.

Caoutchouc is soft and pliable like leather. It is exceedingly elastic and adhesive; so that it may be forcibly stretched out much beyond its usual length, and instantly recover its former bulk when the force is withdrawn. It cannot be broken without very considerable force. Its specific gravity is 0.9335.‡

Elasticity
owing to
latent heat.

To that acute philosopher Mr. Gough of Kendal, I am indebted for some very important experiments on the connexion between the temperature of caoutchouc and its elasticity. They have been since published in the second volume

* Fourcroy, Ann. de Chim. xi. 259.

† I have some pieces of it from the East-Indies, which had been allowed to inspissate in the open air. They are white, with a slight cast of yellow, and have very much the appearance and feel of white soap.

‡ Brisson.

of the Manchester Memoirs, second series. It is necessary to premise, that Mr. Gough has been blind from an infant, and that therefore his sense of touch is peculiarly delicate; so much so, that he is an excellent botanist, and can distinguish plants with the utmost certainty by the feel: a power so extraordinary, that we who enjoy the advantage of sight can scarcely conceive how it can be acquired. Mr. Gough's experiments are as follows:

Take a thong of this substance two or three inches long, and a few lines in breadth and thickness; put it in warm water till it becomes quite pliant: then, holding it merely extended between the two hands, bring the edge of it in contact with the lips, and observe the temperature (of the variations of which that part of the face is a very nice judge); then remove the thong a few lines from the lips, and stretch it forcibly, and bring it again in contact with the lips, and a very sensible increase of temperature will be perceived. Allow it to relax to its former state, and the temperature will be perceived immediately to sink. If we stretch the thong again, and then plunge it immediately into cold water, keeping it extended for a minute or more in the liquid, on letting go one end it will be found to have lost much of its contractile power; for it will not return to its former dimensions. But if we plunge it into warm water, or warm it by holding it for some time in the shut hand, it will begin to contract again, and soon return to its former figure and size. These experiments are of great importance, as they furnish a very palpable and convincing proof that ductility is owing to latent heat as well as fluidity. They afford a fine illustration of Dr. Black's theory of latent heat. We see clearly that the elasticity of caoutchouc and the ductility of metals are different cases of one and the same thing.

Caoutchouc is not altered by exposure to the air; it is perfectly insoluble in water: but if boiled for some time its edges become somewhat transparent, owing undoubtedly to the water carrying off the soot; and so soft, that when two of them are pressed and kept together for some time, they adhere as closely as if they formed one piece. By this contrivance pieces of caoutchouc may be soldered together, and thus made to assume whatever shape we please.*

* Grossart, Ann. de Chim. xi. 153. See a method of making caoutchouc tubes by means of this property, Phil. Mag. xxii. 340.

Book IV.

Action of
alcohol.

Caoutchouc is insoluble in alcohol. This property was discovered very early, and fully confirmed by the experiments of Mr. Macquer. The alcohol, however, renders it colourless.

Ether.

Caoutchouc is soluble in ether. This property was first pointed out by Macquer. Berniard, on the contrary, found that caoutchouc was scarcely soluble at all in sulphuric ether, which was the ether used by Macquer, and that even nitric ether was but an imperfect solvent. The difference in the results of these two chemists was very singular; both were remarkable for their accuracy, and both were too well acquainted with the subject to be easily misled. The matter was first cleared up by Mr. Cavallo. He found that ether, when newly prepared, seldom or never dissolved caoutchouc completely; but if the precaution was taken to wash the ether previously in water, it afterwards dissolved caoutchouc with facility. Mr. Grossart tried this experiment, and found it accurate.* It is evident from this that these chemists had employed ether in different states. The washing of ether has two effects. It deprives it of a little alcohol with which it is often impregnated, and it adds to it about one-tenth of water, which remains combined with it. Alcohol precipitates the caoutchouc from this solution.

When the ether is evaporated, the caoutchouc is obtained unaltered. Caoutchouc, therefore, dissolved in ether, may be employed to make instruments of different kinds, just as the milky juice of the *hævea*; but this method would be a great deal too expensive for common use.

Oils.

Caoutchouc is soluble in volatile oils;† but, in general, when these oils are evaporated, it remains somewhat glutinous, and therefore is scarcely proper for those uses to which, before its solution, it was so admirably adapted.

Alkalies.

It is said by Berniard to be insoluble in alkalies; but I find upon trial that this is a mistake. I was led to make the experiment by an accident. I employed a caoutchouc bottle fitted with a stop-cock in the usual way for holding ammoniacal gas. The gas very soon disappeared, though the bottle was perfectly air-tight, as I learned by plunging it in water. This induced me to fill it repeatedly with gas. In a short time it became evident that the gas had been absorbed by the bottle itself. It became soft and then glutinous, and never recovered its elasticity. I then tried the alkalies in general, and found

* Grossart, Ann. de Chim. xi. 147.

† Berniard.

that they were all capable of producing the same changes on caoutchouc, and even of dissolving it, though in a very minute proportion.

The acids act but feebly upon caoutchouc. Sulphuric acid, Acids. even after a very long digestion, only chars it superficially. The proportion of charcoal obtained in Mr. Hatchett's experiments was only 12 per cent. and he could observe no traces of artificial tannin.* But when heat is applied the caoutchouc is completely decomposed. When treated with nitric acid, there came over azotic gas, carbonic acid gas, prussic acid gas; and oxalic acid is said to be formed.† Muriatic acid does not affect it.‡ The other acids have not been tried.

Fabroni has discovered that rectified petroleum dissolves it, Heat. and leaves it unaltered when evaporated.§

When exposed to heat it readily melts; but it never afterwards recovers its properties, but continues always of the consistence of tar. It burns very readily with a bright flame, and diffuses a fetid odour. In those countries where it is produced, it is often used by way of candle.

Fourcroy assures us|| that caoutchouc when distilled gives out ammonia. I have not tried this experiment. But I have strong reasons for considering it as inaccurate. I passed a grain of caoutchouc through a copper tube filled with red hot peroxide of copper. I obtained 4.69 cubic inches of carbonic acid gas, and 0.8 grains of water, and nothing else. Hence there seems no reason to consider azote as a constituent of caoutchouc. The carbon in 4.69 cubic inches of carbonic acid gas weighs

	0.593 gr.
The hydrogen in 0.8 gr. water =	0.089
	<hr style="width: 10%; margin: 0 auto;"/>
	0.682
Deficiency	0.318
	<hr style="width: 10%; margin: 0 auto;"/>
	1.000

This deficiency must be owing to the presence of oxygen in the caoutchouc. Now if these three ingredients be conceived to exist in the gaseous state their volumes will be as follows:

* Third Series of Experiments on Artificial Tannin, Phil. Trans. 1806.

† Ann. de Chim. xi. 232.

‡ Berniard.

§ Ann. de Chim. xi. 195; and xii. 156.

|| Ann. de Chim. xi. 232.

Book IV.

	Cubic inches.	Volumes.
Carbon	4.69	or 5
Hydrogen	4.20 4½
Oxygen	0.938 1

This is equivalent to

Composi-
tion.

10 atoms carbon	= 7.5
9 atoms hydrogen	= 1.125
4 atoms oxygen	= 4.0
	12.625

Plants con-
taining it.

Caoutchouc seems to exist in a great variety of plants; but it is usually confounded with the other ingredients. It may be separated from resins by means of alcohol. It may be extracted from the different species of misletoe by water, with which, in the fluid state in which it exists in these plants, it readily combines. When mixed with gum or extractive, it may be separated by the following process: Digest a part of the plant containing it first in water and then in alcohol, till all the substances soluble in these liquids are extracted. Dry the residuum, and digest it in five times its weight of rectified petroleum. Express the liquid part by squeezing the substances in a linen cloth. Let this liquid remain several days to settle, then decant off the clear liquid part, mix it with a third part of water and distil; the caoutchouc remains behind.*

According to Bucholz, opium contains a considerable portion of caoutchouc.† That part of mastich which is insoluble in alcohol possesses the properties of caoutchouc, as has been formerly observed.

SECT. XXXI.

OF COTTON.

Origin.

COTTON is a soft down which envelopes the seeds of various plants, especially the different species of *gossypium*, from which the cotton of commerce is procured. These plants are natives of warm climates; grow wild in Asia, Africa, and America, within the tropics; and are cultivated in the East and West Indies. The finest cotton, according to Mr. Edwards, is distinguished by the name of *green seed cotton*, from the colour

* Hermbstadt, Med. and Phys. Jour. iii. 372.

† Ann. de Chim. xxxiv. 133.

of its seeds,* and is perhaps the produce of the *gossypium hirsutum*. There are two species of it; in one of which the cotton does not easily part from the seeds. But the cotton plant commonly cultivated is a shrub, of which Mr. Edwards enumerates five kinds; namely, the *common Jamaica*, the *brown bearded*, the *nankeen*, the *French* (*gossypium arboreum*), and the *kidney cotton*.† When the seeds are ripe, the pods open and display the cotton, which is collected and separated from the seeds by means of rollers.

Cotton, when spun and woven into cloth, furnishes garments to a very considerable portion of the civilized world. The quantity annually brought into this country, and spun by machinery, is not less than 20 millions of pounds; and the number of individuals employed in manufacturing it cannot be less than 700,000. It constitutes therefore one of the most important of our manufactures.

Though no correct chemical investigation of the properties of cotton has hitherto been made, yet as its obvious qualities distinguish it sufficiently from every other vegetable substance, we must consider it as a peculiar vegetable principle; and I have introduced it here, in hopes that some person or other will be induced to examine its nature in detail. The following are the particulars at present known.

This substance is in threads differing in length and in fineness. No asperities can be discovered on the surface of these threads; but if Lewenhoeck's microscopical observations are to be trusted, they are all triangular, and have three sharp edges. Cotton differs considerably in colour; but when bleached it becomes of a fine white. Properties.

Cotton is tasteless and destitute of smell. It is completely insoluble in water, alcohol, ether, and oils, and in all the vegetable acids.

The diluted alkaline leys have no perceptible action on cotton; but when very strong they dissolve it if assisted by a sufficient degree of heat. The new products obtained by this solution have not been examined.

Cotton has a strong affinity for some of the earths, especially for alumina. Hence this substance is used to fix co-

* History of the West Indies, ii. 264.

† Perhaps the first species are only varieties of the *gossypium herbaceum*.

Book IV. lours on cotton. The cloth is steeped in a solution of alum or acetate of alumina, and afterwards dyed.

Several of the metallic oxides also combine with it readily, and remain united with much obstinacy. Oxide of iron is one of the most remarkable. When cotton is dipped into a solution of iron in an acid, it comes out yellow, and the iron is neither separated by alkalis nor soap, nor even by acids, unless when the combination is quite recent. The colour gradually deepens by exposure to the air, owing no doubt to the oxidizement of the iron, unless the cloth be steeped in an aluminous solution, which prevents the colour from becoming disagreeable, probably by diluting it.* Oxide of tin also combines with cotton, and is frequently used as a mordant.

Cotton combines readily with tannin, and forms a yellow or brown compound. Hence the infusion of galls, and of other astringent substances, is often used as a mordant for cotton.

Nitric acid decomposes cotton when assisted with heat, and oxalic acid is formed; the other products have not been examined. Sulphuric acid likewise chars it. Chlorine gas bleaches it, and probably alters and dissolves it when applied in a concentrated state.

Cotton is extremely combustible, and burns with a clear lively flame. The ashes left behind, according to Neumann, contain some potash. When distilled it yields a great portion of acidulous water, and a small quantity of oil, but no ammonia.†

SECT. XXXII.

OF SUBER.

THIS name has been introduced into chemistry by Fourcroy, to denote the outer bark of the *quercus suber*, or the common cork; a substance which possesses properties different from all other vegetable bodies.

It is exceedingly light, soft, and elastic; very combustible, burning with a bright white flame, and leaving a light black bulky charcoal; and when distilled it yields a little ammonia.

Chevreul has subjected cork to an elaborate analysis, which

* See Chaptal, Ann. de Chim. xxvi. 266.

† Neumann's Chem. p. 430.

however, has not led to any consequences of much importance. He treated it with water in his silver digester, till that liquid was capable of taking up nothing more. He obtained an aromatic principle, and a little acetic acid, which passed over into the receiver; the extract formed by the water contained two colouring matters, the one yellow, the other red; an acid, the nature of which was not determined, gallic acid, an astringent substance; a substance containing azote; a substance soluble in water and insoluble in alcohol; gallate of iron; lime and traces of magnesia. Twenty parts of cork thus treated left 17·15 parts of insoluble matter.

The undissolved residue being treated a sufficient number of times with alcohol in the same apparatus, yielded *resin*, oil, and a substance which Chevreul distinguished by the name of *cerin*; but as this name has been already applied by John, to another substance, perhaps it would be requisite to give to this new principle of Chevreul, the name of *suberin*.*

Twenty parts of cork, when thus treated, are reduced to 14 parts. This residue consists of suber, but probably not in a state of absolute purity.†

The properties of suber have not yet been accurately determined, owing to the difficulty of obtaining it in a state of purity. Sulphuric acid readily chars it. Nitric acid gives it a yellow colour, corrodes, dissolves, and decomposes it; converting it partly into suberic acid, partly into a substance resembling wax, partly into artificial tannin, and partly into a kind of starchy matter.‡

Fourcroy supposes, from some experiments which he does not relate, that the epidermis of all trees is a substance possessed of the same properties with cork; and perhaps we may conclude from the experiments of Link, that all vegetable membrane possesses nearly the same chemical characters. At least it agrees with cork in furnishing suberic acid when treated with nitric acid.§

* The characters of this substance have been given in the 22d section of this chapter, while treating of wax.

† Ann. de Chim. ii. 141.

‡ Bouillon La Grange, Ann. de Chim. xxiii. 50.

§ Nicholson's Jour. xxiii. 155.

SECT. XXXIII.

OF MEDULLIN.

Book IV.

THIS is a name given by Dr. John to the pith of the sunflower (*helianthus annuus*,) the *syringa vulgaris*, &c.

According to him it is distinguished by the following properties:

Properties.

1. It is insoluble in water, ether, alcohol, and oils.
2. It is destitute of taste and smell.
3. Its structure is peculiar, being full of pores.
4. It is soluble in nitric acid; but instead of forming suberic acid, as is the case with suber, it furnishes a quantity of oxalic acid.
5. When distilled it furnishes a considerable quantity of ammonia, and leaves a charcoal having a metallic appearance and a colour similar to that of bronze.*

SECT. XXXIV.

OF WOOD, OR LIGNIN.

ALL trees, and most other plants, contain a particular substance, well known by the name of *wood*. If a piece of wood be well dried, and digested, first in a sufficient quantity of water and then of alcohol, to extract from it all the substances soluble in these liquids, there remains only behind the *woody fibre*.

This substance, which constitutes the basis of wood, is composed of longitudinal fibres, easily subdivided into a number of smaller fibres. It is somewhat transparent; is perfectly tasteless; has no smell; and is not altered by exposure to the atmosphere.

Properties.

It is insoluble in water and in alcohol. The fixed alkalies, when assisted by heat, give it a deep brown colour, render it soft, and decompose it. A weak alkaline solution dissolves it without alteration; and it may be thrown down again by means of an acid. By this property we are enabled to separate wood from most of the other vegetable principles, as few of them are soluble in weak alkaline leys.

Action of heat.

When heated, it blackens without melting or frothing up, and exhales a disagreeable acrid fume, and leaves a charcoal which retains exactly the form of the original mass. When

* John's Chemische Tabellen der Pflanzen Analysen, p. 9.

distilled in a retort, it yields an acid liquor of a peculiar taste and smell, distinguished by the name of pyrolignous, and formerly considered as a distinct acid; but Fourcroy and Vauquelin ascertained that it is merely the acetic acid combined with an empyreumatic oil.*

By nitric acid Fourcroy converted the residuum of quinquina, which does not seem to differ from the woody fibre, into oxalic acid; at the same time there was a little citric acid formed, and a very small quantity of malic and acetic acids. Some azotic gas also was disengaged.

By this process he obtained from 100 parts of woody fibre

Oxalic acid	56·250
Citric acid	3·905
Malic acid	0·388
Acetic acid	0·486
Azotic gas	0·867
Carbonate of lime	8·330
	<hr/>
	70·226
Residuum	32·031
	<hr/>
	102·257

There was likewise a quantity of carbonic acid gas disengaged, the weight of which was unknown. This increase of weight in the product was evidently owing to the oxygen derived from the nitric acid.†

When this residuum was distilled in a retort, 100 parts yielded the following products:

26·620 of a yellow liquid, containing alcohol, and an acid which had the smell of pyromucous.

6·977 of concrete oil, mostly soluble in alcohol.

22·995 charcoal

3·567 carbonate of lime } in the retort.

60·159

39·841 gas, half-carbonic acid, half-carburetted hydrogen.

100·000‡

* Mollerat has succeeded in France in making acetic acid from wood as pure as radical vinegar. It answers very well for aromatic vinegar, but possesses a little acrimony, which makes it less fit for the table. See Nicholson's Jour. xxiv. 70. Very good vinegar from wood is now made both in London and Glasgow. Animal charcoal is probably employed to remove the empyreumatic oil.

† Ann. de Chim. viii. 153.

‡ Ibid. 151.

Book IV.
Constitu-
ents.

Several analyses of pure woody fibre were made by Gay-Lussac and Thenard, by burning it along with chlorate of potash and ascertaining the products. The result of their experiments gives its constituents as follows:

	Oak.		Beech.		Mean,
Oxygen	41.78	42.73	42.25
Carbon	52.53	51.45	52.00
Hydrogen	5.69	5.82	5.75
	<hr/>		<hr/>		<hr/>
	100.00		100.00		100.00*

As we are unacquainted with the weight of an ingredient particle of woody fibre, it is impossible from this analysis to deduce the number of atoms of which it is composed. But the smallest number that corresponds with the analysis is the following:

4 atoms oxygen	=	4.00	41.02
7 atoms carbon	=	5.25	53.86
4 atoms hydrogen	=	0.5	5.12
		<hr/>		<hr/>
		9.75		100.00

Were these numbers the representation of the true constitution of wood, it would follow that it might yield more than half its weight of acetic acid. For if we suppose it deprived of 1 atom water and 3 atoms carbon, the residue would be acetic acid.

Charcoal
yielded by.

When wood is burnt with a smothered flame, it leaves, as is well known, a quantity of charcoal behind it, which exhibits the exact form, and even the different layers of the original wood. As it is the wood alone which undergoes this change, while other component parts of the plant are dissipated, we may form some notion of the relative proportion of wood which different plants contain, by the proportion of charcoal which they yield. Now, the quantity of charcoal yielded by 1.00 of different trees is, according to the experiments of Proust, as follows:

Black ash	0.25	Heart of oak ..	0.19
Guaiacum	0.24	Wild ash	0.17
Pine	0.20	White ash	0.17
Green oak	0.20		

Count Rumford, by continuing a very moderate fire for 96 hours, converted different kinds of wood into charcoal, and

* Recherches Physico-chimiques, ii. 294.

obtained a much greater proportion of charcoal than was obtained in the preceding experiments of Proust. The following table exhibits the weight of charcoal left in these experiments by 1·00 of the different woods tried.*

Poplar	0·4357	Maple	0·4223
Lime	0·4359	Elm	0·4327
Fir	0·4418	Oak	0·4300

Braconnot has made some experiments on lignin, which serve still farther to extend our knowledge of it.

When the dried sawings of hornbeam are mixed and agitated with sulphuric acid, considerable heat is evolved, sulphurous acid is disengaged, and the whole becomes black. When mixed with water a black powder falls, which burns with flame and possesses nearly the characters of sawings of hornbeam exposed to the air and moisture for several years. The sulphurous liquor is nearly colourless, and being saturated with carbonate of lime and filtered, it yields by evaporation a gum possessing many characters in common with gum arabic.—Hemp cloth dissolves in sulphuric acid in the same way, but without the evolution of sulphurous acid, and is wholly converted into the same kind of gum.

This gum is transparent, yellowish, without smell or taste, and breaks with a vitreous fracture. It reddens vegetable blues, adheres strongly to bodies, and forms a very shining varnish. Its solution in water has less tenacity than that of gum arabic. When heated sufficiently it burns, giving out a strong odour of sulphurous acid. Its solution in water is not precipitated by nitrate of barytes, nor by acetate of lead. But subtriacetate of lead throws down a copious precipitate which is totally dissolved by acetic acid.

If this gum be boiled for some time in dilute sulphuric acid it is converted into two new principles; namely, *sugar*, and an acid which Braconnot has called *vegeto-sulphuric acid*. The sugar possesses the properties of starch sugar. The vegeto-sulphuric acid forms soluble salts with all bases, and none of them seems capable of crystallizing.

When equal weights of caustic potash and sawings of wood are heated in a silver or iron crucible, stirring the mixture constantly at a certain temperature, it softens and dissolves completely, swelling up greatly at the same time. If it be

* Gilbert's *Annalen der Physik*, xlv. 75.

Book IV. now taken from the fire and water poured upon it, the whole dissolves with the greatest facility, and we obtain a dark-brown liquid, consisting of a combination of potash and ulmin dissolved in water. If an acid be dropped into this liquid, the ulmin falls down in great abundance, and requires only to be washed and dried.

Artificial ulmin thus prepared has a brilliant black colour like jet. It is very brittle and easily broken into angular fragments. It has no smell and little taste. In this state it is insoluble in water; but when just precipitated and still moist, water dissolves $\frac{1}{2500}$ th of its weight of it. Nitrate of silver, persulphate of iron, nitrate of barytes, acetate of alumina, muriate of lime, muriate of soda, throw down brown precipitates from it; but they do not appear till some time after the reagents have been added. Lime and litharge deprive the water entirely of its colour. It is soluble in concentrated sulphuric acid, and precipitated by water. Alcohol dissolves it with facility and forms a dark-brown liquid which is precipitated by water.*

SECT. XXXV.

OF FUNGIN.

THIS is a name given by Braconnot to the fleshy part of mushrooms, which he considers as a peculiar vegetable principle. It approaches in its chemical characters so closely to woody fibre, that perhaps it would be better to consider it merely as a variety of that substance. Fungin is the substance which remains after the mushroom has been deprived of every thing soluble, either in water or alcohol. It is distinguished by the following properties:

Properties. It is white, soft, and insipid, possesses but little elasticity, and divides easily between the teeth. In this state it constitutes a nourishing article of food.

It is insoluble in water, alcohol, ether, and oils.

Action of
alkalies,

Alkalies have but little action on fungin; a property which distinguishes it from woody fibre, which is very readily dissolved even by a weak alkaline ley. Yet when boiled in a concentrated alkaline ley fungin is partly dissolved, and a saponaceous liquid is obtained from which the acids throw down a flocky matter.

* Annals of Philosophy, xvi. 92.

Ammonia dissolves a little of it when digested over fungin, and on exposure to the air allows it to precipitate again in the state of white flocks. Chap. I.
Of acids.

Diluted sulphuric acid has no action on it; but concentrated sulphuric acid chars it, while acetic and sulphurous acid are evolved.

Muriatic acid, at first, seems to have no action on it, even though assisted by heat; but it gradually dissolves it and converts it into a gelatinous matter soluble in water. Potash dropped into the solution throws down the fungin, probably altered in its properties.

Chlorine gas converts it into a yellow matter which has an acrid taste while wet, but which it loses when dried. This yellow substance is a compound of fungin, muriatic acid, and a kind of adiporesinous matter.

Weak nitric acid disengages azote from it. When distilled with six times its weight of nitric acid it becomes yellow, swells considerably, and effervesces at first very much; but the violent action soon subsides. Prussic acid is formed by this process, together with a considerable quantity of oxalic acid, two fatty bodies resembling tallow and wax and a small quantity of yellow-bitter principle, and a yellow resinous looking substance.

When put into infusion of nutgalls it absorbs the greatest part of the tannin contained in that liquid, and acquires a fawn-colour.

When mixed with water and left to spontaneous putrefaction it emits, at first, the smell of putrid cheese; but this smell soon goes off. The liquid covering it, after an interval of three months, was neither acid nor alkaline; but held in solution a mucilaginous matter. The fungin still retained its original shape. When washed it is easily reduced to a pulp, which may be kneaded between the fingers; but has not the elastic consistence of gluten.

Fungin when dry burns with vivacity, and leaves a white ash consisting chiefly of phosphate of lime.

Thirty-eight parts being distilled yielded 8 parts of a brown empyreumatic oil, and $11\frac{1}{2}$ parts of a liquid containing an excess of ammonia and holding in solution acetate of ammonia mixed with oil. The charcoal in the retort weighed 10 parts.* When fungin was distilled by Vauquelin, the

* Braconnot, Ann. de Chim. lxxix. 267.

Book IV. liquid which he obtained contained ammonia; but it reddened vegetable blues, and therefore contained an excess of acetic acid, as is the case with the liquid obtained from the distillation of wood.*

SECT. XXXVI.

FOREIGN SUBSTANCES FOUND IN PLANTS.

THESE are substances which seem rather to belong to the mineral than the vegetable kingdom: but which require to be enumerated because they constantly make their appearance, though but in small quantity, chiefly in the vegetable juices. These substances may be arranged under four heads; namely, *acids, alkalies, earths, and metals.*

I. ACIDS.

The mineral acids found in plants are usually combined with a base constituting different salts. They are the phosphoric, silica, sulphuric, nitric, muriatic, and carbonic acids.

Phosphoric, 1. Phosphoric acid has been found in different plants, but only in very small quantities: it is almost constantly combined with lime or potash, though it exists free in the onion. Meyer found it in the leaves of many trees; † Thuren found phosphate of lime in the *aconitus napellus*; ‡ and Bergman found it in all kinds of grain. § Phosphate of potash exists in barley and other species of corn, and has, I presume, been often confounded with phosphate of lime.

Silica, 2. Silica exists in many plants, particularly in grasses and equisetums. Sir H. Davy has ascertained that it forms a part of the epidermis, or outermost bark of these plants; and that in some of them almost the whole epidermis is silica.

	Parts Silica.
100 parts of the epid. of bonnet cane yielded ..	90
————— bamboo	71·4
—— (arundo phragm.) common reed	48·1
————— stalks of corn	6·5

The concretions which are sometimes found in the bamboo

* Ann. de Chim. lxxxv. 12.

† Encyc. Meth. Phisiol. Veget. i. 100.

‡ Ann. de Chim. ii. 308.

§ Bergman, v. 96.

cane called tabasheer, have been ascertained by Mr. Macie to be composed of pure silica.*

Chap. I.

Any attempt to enumerate the plants in which traces of sulphuric, muriatic, nitric, and carbonic acids have been found would be superfluous. It may be sufficient to say that common salt is an ingredient in almost all plants that grow in the sea, and in many that vegetate on the sea-shore; that nitre exists in sun-flower, nitrate of soda in barley, and sulphate of lime in clover.

II. ALKALIES.

The only alkalies found in plants are potash and soda. Ammonia may indeed be obtained by distilling many vegetable substances, but it is produced during the operation. One or other of these alkalies is found in every plant which has hitherto been examined. The quantity indeed is usually very small. From the experiments of Vauquelin, it is probable that the alkalies are combined in plants with acetic and carbonic acids.

1. Potash is found in almost all plants which grow at a distance from the sea. It may be extracted by burning the vegetable, washing the ashes in water, filtrating the water, and evaporating it to dryness. It is in this manner that all the potash of commerce is procured.

Dr. Peschier of Geneva has pointed out an ingenious method of detecting this alkali in the juices of plants without subjecting them to combustion. He agitates the liquid for some time with a quantity of magnesia. The acid previously combined with the potash (usually carbonic, oxalic or tartaric) forms an insoluble compound with the magnesia, while the potash thus set at liberty remains in solution in the liquid, and may be detected by its properties.†

The following table exhibits the quantity of ashes and potash which may be extracted from 100 parts of various plants:

	Ashes.	Potash.
Sallow	2·8 0·285 †
Elm	2·36727 0·39 ‡

* Fourcroy and Vauquelin examined a tabasheer from the Andes, and found it composed of 70 silica and 30 potash. Gehlen's Jour. Second Series, ii. 112.

† Annals of Philosophy, xii. 336.

‡ Those marked ‡ are from Kirwan, Irish Trans. v. 164. The rest from Pertuis, Ann. de Chim. xix. 178. The following table by Redi, of the

Book IV.

Oak	1.35185	0.15343
Poplar	1.23476	0.07481
Hornbeam	1.1283	0.1254
Beech	0.58432	0.14572
Fir	0.34133	—
Vine branches	3.379	0.55 §
Common nettle	10.67186	2.5033
Common thistle	4.04265	0.53734
Fern	4.00781	0.6259
Cow thistle	10.5	1.96603
Great river rush	3.85395	0.72234
Feathered rush	4.33593	0.50811
Stalks of turkey wheat	8.86	1.75 §
Wormwood	9.744	7.3
Fumitory	21.9	7.9 §
Trifolium pratense ..	—	0.078 §
Vetches	—	2.75 §
Beans with their stalks	—	2.0 §

ashes and salts yielded by different plants, is worth inserting. It is printed in the Phil. Trans. for 1698 (vol. xx. p. 281.)

Pounds.	Vegetables.	Ashes.			Salts.	
		lbs.	oz.	dr.	oz.	dr.
100	Of dried flowers of oranges	4	6	0	0 5
800	Of gourds new gathered, which dried in the oven were 36 lbs.	4	0	0	10 0
400	Red onions (being 720) roasted, the coals turned to 16 lbs. to the coals new added 4 oz. of sulphur	1	6	0	2 2
150	Eyebright fresh, and afterwards stilled and burnt	5	0	0	4 0
120	Distilled roses	4	0	1	0 0
100	Of maidenhair	9	0	0	0 4
150	Roots of black hellebore, which dried came to 50 lbs.	6	0	0	1 0
150	Roots of white hellebore, fresh, which dried came to 50 lbs.	2	0	0	4 0
96	Roots dried and burnt of fresh esula	3	0	0	2 0
30	Roots of liquorice	2	0	0	1 4
20	Pellitory	1	0	0	0 6
100	Green endive	2	0	0	2 0
90	Green bindweed	1	0	0	2 0
2000	Leaves of laurel	33	0	4	0 0
500	Leaves of laurel	6	0	0	10 0
1000	Water melons well ripe, the seeds being taken	25	0	1	0 0
2400	Cucumbers	18	0	0	0 0

In general, three times as much ashes are obtained from shrubs, and five times as much from herbs, as from trees. Equal weights of the branches of trees produce more ashes than the trunk, and the leaves more than the branches. Herbs arrived at maturity produce more ashes than at any other time. Green vegetables produce more ashes than dry.*

The salt which is obtained from plants does not consist wholly of potash, there are other salts mixed with it; these usually are sulphate of potash, muriate of potash, sulphate of lime, phosphate of lime, &c.; but these bear, in general, but a small proportion to the potash. The ashes consist of potash, mixed with earths.

Some judgment may be formed of the quantity of potash which a plant contains, from the quantity of ashes which it yields: but the foregoing table is sufficient to show us, that were we to trust to that we should often be misled.

2. Soda is found in almost all the plants which grow in the sea, and in many of those which grow on the shore. In general, the quantity of soda which plants contain bears a much greater proportion to their weight than the potash does which is found in inland vegetables. One hundred parts of the *salsola soda*, for instance, yield 19.921 of ashes; and these contain 1.992 parts of soda; some of which, however, is combined with muriatic acid.† The plants from which the greater part of the soda, or *barilha* as it is called, which is imported from Spain, is extracted, are the *salsola sativa* and *vermiculata*.

Pounds.	Vegetables.	Ashes.			Salts.	
		lbs.	oz.	dr.	oz.	dr.
300	Wood of ivy	9	0	0	0	0
50	Scorzonera dried	8	0	0	0	0
300	Pine apples, the nuts taken out	3	0	0	0	0
150	Mugwort dried	8	0	0	0	0
130	Leaves of cyprus	6	0	0	0	0
10	Peel of pomegranates dried	0	8	0	0	0
2	Sassafras.....	0	0	0	0	0
12	Lignum sanctum	2	6	0	0	0
4	Yellow sanders	0	1	4	0	0
4	Black pepper.....	0	2	4	0	0
30	Ginger.....	1	7	0	0	0
12	Turbith	1	0	0	0	0
	Wood of fir	3	0	0	3	0
	Scopæ	16	0	1	4	0
	Scopæ	16	0	1	0	0

* Ann. de Chim. xix. 174.

† Vauquelin, *ibid.* xviii. 77.

III. EARTHS.

Book IV.

The only earths hitherto found in plants are the three following; *lime*, *magnesia*, *alumina*.

Lime,

1. Lime is usually the most abundant of the earths of plants, and the most generally diffused over the vegetable kingdom. Indeed it is a very uncommon thing to find a plant entirely destitute of lime: *salsola soda* is almost the only one in which we know for certain that this earth does not exist.*

Magnesia,

2. Magnesia does not exist so generally in the vegetable kingdom as the preceding earth. It has been found, however, in considerable quantities in several sea plants, especially *fuci*;† but the *salsola soda* contains a greater proportion of magnesia than any plant hitherto examined. Mr. Vauquelin found that 100 parts of it contained 17·929 of that earth.‡

Alumina,

3. Alumina has only been found in very small quantities in plants.

The following table will show the quantity of these four earths which exist in several vegetables.

100 parts of Oak contain of earths	1·03 §
Beech	0·453
Fir	0·003
Turkey wheat	7·11
Sunflower	3·72
Vine branches	2·85
Box	2·674
Willow	2·515
Elm	1·96
Aspin	1·146
Fern	3·221 **
Wormwood	2·444 ††
Fumitory	14·000 ††

This table shows us that the quantity of earth is greater in herbs than in trees.

Bergmann found all the four earths in every kind of grain which he analysed.‡‡

* Vauquelin, Ann. de Chim. xviii. 76.

† Ibid. 86, and ix. 94.

§ Watson.

** Home.

‡‡ Opusc. v. 94.

‡ Ibid. p. 78.

|| Kirwan, Irish Trans. iii. 35.

†† Wiegleb.

Vauquelin found that 100 parts of oat grain left 3·1591 of residuum. This residuum is composed of

Silica	60·7
Phosphate of lime	39·3
	100·0*

When the whole of the *avena sativa*, however, stalk and seed together, is burnt, it leaves a residuum composed of

Silica	55
Phosphate of lime ..	15
Potash	20
Carbonate of lime ...	5

95 and a little oxide of iron.†

This shows us that the stalk contains several substances not to be found in the grain.

The following table exhibits the quantity of earths and metallic oxides in grains, obtained by Schräder from 32 ounces of the seeds of the following kinds of corn; wheat (*triticum hybernum*), rye (*secale cereale*), barley (*hordeum vulgare*), oats (*avena sativa*), and likewise from the same quantity of rye straw.‡

	Wheat.	Rye.	Barley.	Oats.	Rye Straw.	Earths in corn.
Silica	13·2	15·6	66·7	144·2	152	
Carbonate of lime	12·6	13·4	24·8	33·75	46·2	
Carbonate of magnesia	13·4	14·2	25·3	33·9	28·2	
Alumina	0·6	1·4	4·2	4·5	3·2	
Oxide of manganese ..	5·0	3·2	6·7	6·95	6·8	
Oxide of iron	2·5	0·9	3·8	4·5	2·4	
	47·3	48·7	131·5	227·8	238·8	

Saussure junior, in his *Chemical Researches on Vegetation*, a work published in 1804, has given us the most copious table of the earthy and saline constituents of vegetables which has yet appeared. This table is the more valuable, as it was constructed entirely from his own experiments. For that reason I shall insert it in this place:

* Ann. de Chim. xxix. 17.

† Ibid. 19.

‡ Gehlen's Jour. iii. 525.

Book IV.
Saussure's
table.

TABLE OF INCINERATIONS.					Constit. of 100 parts of ashes.					
	Names of Plants.	Ashes from 1000 parts of the plant, green.	Ditto, dry.	Water from 1000 parts of the plant, green.	Soluble salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
1	Leaves of oak (<i>quercus robur</i>), May 10.	13	53	745	47	24	0.12	3	0.64	25.24
2	Ditto, Sept. 27.	24	55	549	17	18.25	23	14.5	1.75	25.5
3	Wood of a young oak, May 10.		4		26	28.5	12.25	0.12	1	.58
4	Bark of ditto		60		7	4.5	63.25	0.25	1.75	22.75
5	Perfect wood of oak.		2		38.6	4.5	32	2	2.25	20.65
6	Alburnum of ditto.		4		32	24	11	7.5	2	23.5
7	Bark of ditto.		60		7	3	66	1.5	2	21.5
8	Cortical layers of ditto.		73		7	3.75	65	0.5	1	12.75
9	Extract of wood of ditto.		61		51					
10	Soil from wood of ditto.		41		24	10.5	10	32	14	8.5
11	Extract from ditto.		111		66					
12	Leaves of poplar (<i>populus nigra</i>), May 26.	23	66	652	36	13	29	5	1.25	15.75
13	Ditto, Sept. 12.	41	93	565	26	7	36	11.5	1.5	18
14	Wood of ditto, Sept. 12.		8	26		16.75	27	3.3	1.5	24.5
15	Bark of ditto.		72		6	5.3	60	4	1.5	23.2
16	Leaves of hazel (<i>corylus avellana</i>), May 1.		61		26	23.3	22	2.5	1.5	24.7

Table continued.

Chap. IV.

TABLE OF INCINERATIONS.				Constit. of 100 parts of the ashes.						
	Names of Plants.	Ashes from 1000 parts of the plant, green.	Ditto, dry.	Water from 1000 parts of the plant, green.	Soluble Salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
17	Ditto, washed in cold water.		57		8.2	19.5	44.1	4	2	22.2
18	Leaves of ditto, June 22.	28	62	655	22.7	14	29	11.3	1.5	21.5
19	Ditto, Sept. 20.	31	70	557	11	12	36	22	2	17
20	Wood of ditto, May 1.		5		24.5	35	8	0.25	0.12	32.2
21	Bark of ditto.		62		12.5	5.5	54	0.25	1.75	26
22	Perfect wood of mulberry (<i>morus nigra</i>), November.		7		21	2.25	56	0.12	0.25	20.38
23	Alburnum of ditto.		13		26	27.25	21	1	0.25	21.5
24	Bark of ditto.		89		7	8.5	45	15.25	1.12	23.13
25	Cortical layers of ditto.		88		10	16.5	48	0.12	1	24.38
26	Perfect wood of hornbeam (<i>carpinus betulus</i>), Nov.	4	6	346	22	23	26	0.12	2.25	26.63
27	Alburnum of ditto.	4	7	390	18	36	15	1	1	29
28	Bark of ditto.	88	134	346	4.5	4.5	59	1.5	0.12	30.38
29	Wood of horse-chesnut (<i>asculus hippocastanum</i>), May 10.		35		9.5					
30	Leaves of ditto, May 10.	16	72	782	50					

Table continued.

Book IV.

TABLE OF INCINERATIONS.				Constit. of 100 parts of the ashes.						
	Names of Plants.	Ashes from 1000 parts of the plant, green.	Ditto, dry.	Water from 1000 parts of the plant, green.	Soluble Salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
31	Ditto, July 23.	29	84	652	24					
32	Ditto, Sept. 27.	31	86	636	13.5					
33	Flowers of ditto, May 10.	9	71	873	50					
34	Fruit of ditto, October 5.	12	34	647	82	12	0	0.5	0.25	5.25
35	Plants of pease (<i>pisum sativum</i>), in flower.		95		49.8	17.25	6	2.3	1	24.65
36	Ditto, ripe.		81		34.25	22	14	11	2.5	17.25
37	Plants of vetches (<i>vicia faba</i>), before flowering, May 23.	16	150	895	55.5	14.5	3.5	1.5	0.5	24.50
38	Ditto, in flower, June 23.	20	122	876	55.5	13.5	4.12	1.5	0.5	24.38
39	Ditto, ripe, July 23.		66		50	17.75	4	1.75	0.5	26
40	Ditto, seeds separated.		115		42	5.75	36	1.75	1	12.9
41	Seeds of ditto.		33		69.28	27.92	0	0	0.5	2.3
42	Ditto, in flower, raised in distilled water.		39		60.1	30	0	0	0.5	9.4
43	<i>Solydago vulgaris</i> , before flowering, May 1.		92		67.5	10.75	1.25	1.5	0.75	18.25
44	Ditto, just in flower, July 15.		57		59	8.5	9.25	1.5	0.75	21

Table continued.

Chap. I

TABLE OF INCINERATIONS.				Constit. of 100 parts of the ashes.						
	Names of Plants.	Ashes from 1000 parts of the plant, green.	Ditto, dry.	Water from 1000 parts of the plant, green.	Soluble Salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
15	Ditto, Seeds ripe Sept. 20.		50		48	11	17.25	3.5	1.5	18.75
16	Plants of turusol (<i>helianthus annuus</i>), a month before flowering, June 23.		147		63	6.7	11.56	1.5	0.12	16.67
17	Ditto in flower, July, 23.	13	137	877	61	6	12.5	1.5	0.12	18.78
18	Ditto, bearing ripe seeds, Sept. 20.	23	93	753	51.5	22.5	4	3.75	0.5	17.75
19	Wheat (<i>triticum sativum</i>), in flower.				43.25	12.75	0.25	32	0.5	12.25
20	Ditto, seeds ripe.				11	15	0.25	54	1	18.75
21	Ditto, a month before flowering.		79		60	11.5	0.25	12.5	0.25	15.5
22	Ditto, in flower, June 14.	16	54	699	41	10.75	0.25	26	0.5	21.5
23	Ditto, seeds ripe.		33		10	11.75	0.25	51	0.75	23
24	Straw of wheat.		43		22.5	6.2	1	61.5	1	7.8
25	Seeds of ditto.		13		47.16	44.5	0	0.5	0.25	7.6
26	Bran.		52		44.16	46.5	0	0.5	0.25	8.6
27	Plants of maize (<i>zea mays</i>), a month before flowering, June 23.		122		69	5.74	0.25	7.5	0.25	17.25

Table continued.

Book IV.

TABLE OF INCINERATIONS.				Constit. of 100 parts of the ashes.						
	Names of Plants.	Ashes from 1000 parts of the plant, green.	Ditto, dry.	Water from 1000 parts of the plant, green.	Soluble Salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
58	Ditto, in flower, July 23.		81		69	6	0.25	7.5	0.25	17
59	Ditto, seeds ripe.		46							
60	Stalks of ditto,		84		72.45	5	1	18	0.5	3.05
61	Spikes of ditto.		16							
62	Seeds of ditto.		10		62	36	0	1	0.12	0.88
63	Chaff of barley (<i>hordeum vulgare</i>).		42		20	7.75	12.5	57	0.5	2.25
64	Seeds of ditto.		18		29	32.5	0	35.5	0.25	2.8
65	Ditto.				22	22	0	21	0.12	29.88
66	Oats.		31		1	24	0	60	0.25	14.75
67	Leaves of <i>rhododendron ferrugineum</i> , raised on Jura, a limestone hill, June 20.		30		23	14	43.25	0.75	3.25	15.63
68	Ditto, raised on Breven, a granite hill, June 27.		25		21.1	16.75	16.75	2	5.57	31.53
69	Branches of ditto, June 20.		8		22.5	10	39	0.5	5.4	22.48
70	Spikes of ditto, June 27.		8		24	11.5	29	1	11	24.5

Table continued.

Chap. I.

TABLE OF INCINERATIONS.				Constit. of 100 parts of the ashes.					
	Names of Plants.	Ashes from 1000 parts of the plant, green.		Soluble Salts.	Earthy Phosphates.	Earthy Carbonates.	Silica.	Metallic Oxides.	Loss.
		Ditto, dry.	Water from 1000 parts of the plant, green.						
71	Leaves of fir (<i>pinus abies</i>), raised on Jura, June 20.	29		16	12.27	13.5	2.5	1.6	24.13
72	Ditto, raised on Breven, June 27.	29		15	12	29	19	5.5	19.5
73	Branches of pine, June 20.	15		15					
74	Blackberry (<i>vaccinium myrtillus</i>), raised on Jura, Aug. 29.	26		17	18	12	0.5	3.12	19.38
75	Ditto, raised on Breven, Aug. 20.	22		21	22	22	5	9.5	17.5
76	Soil from <i>rhododendron ferrugineum</i> of Nos. 67 and 69.	65		0.5	6	29	28	18	15.5
77	Extract from soil of ditto.	140		53	17.25	21.5	3.25	3	21.88
78	Soil of <i>rhododendron ferrugineum</i> , raised on a siliceous earth of Nos. 68 and 70.	620							
79	Extract from the preceding soil.	112		21	13	17	14	10	21.88

For some other tables of a similar nature, I refer the reader to Dr. John's Tabellen der Pflanzen Analysen, p. 63.

IV. METALS.

Several metallic substances have also been found in the ashes of vegetables, but their quantity is exceedingly small; so small, indeed, that without very delicate experiments their presence cannot even be detected.

The metals hitherto discovered are iron, which is by far the most common, manganese, and, if we believe some chemists, gold.

1. Iron has been found in many plants; the ashes of salsola contain a considerable quantity of it. From the experiments of Mr. Grimshaw, it appears that six ounces of cotton cloth, unbleached, contain about four grains of iron; and six ounces of linen in the same state about three grains; but it is likely, as Mr. Grimshaw observes, that part of this iron is communicated to the cloth by the weaver's dressing. When it gets acid, and is allowed to remain in iron pans, it must take up a part of the metal.*

2. Scheele first detected manganese in vegetables.† Proust found it in the ashes of the pine, calendula, vine, green oak, and fig-tree.‡

3. With respect to the minute portion of gold extracted from the ashes of plants by Kunkel, Sage, &c. it is probable that it proceeded rather from the lead which they employed in their processes than from the ashes.

SECT. XXXVII.

OF VEGETABLE PRINCIPLES IN GENERAL.

IN the preceding sections a pretty full account has been given of the different vegetable principles hitherto examined by chemists. They amount to no fewer than 36 genera. They consist of substances which owe their formation to the process of investigation. It is of them, of course, that vegetables are formed; they are the substances which come into view when vegetables are analyzed. It is necessary therefore to be well acquainted with their essential characters, that we may know the marks by which they are to be recognized.

* Phil. Mag. xvi. 33

† Opusc. i. 106.

‡ Phil. Mag. v. 99.

These unfortunately are sometimes ambiguous; so that a good deal of skill and experience are necessary before we can distinguish them readily. Some of them indeed are so imperfectly known, that we are in possession of no good characteristic marks. The following table contains a general view of the most striking properties of the different genera.

Chap. I.

1. *Sugar*.—Crystallizes. Soluble in water and alcohol. Taste sweet. Soluble in nitric acid, and yields oxalic acid.

Table of essential characters.

2. *Sarcocoll*.—Does not crystallize. Soluble in water and alcohol. Taste bitter sweet. Soluble in nitric acid, and yields oxalic acid.

3. *Gum*.—Does not crystallize. Taste insipid. Soluble in water, and forms mucilage. Insoluble in alcohol. Precipitated by silicated potash. Soluble in nitric acid, and forms sacelactic and oxalic acids.

4. *Mucus*.—Does not crystallize. Taste insipid. Soluble in water, and forms mucilage. Insoluble in alcohol. Not precipitated by silicated potash. Precipitated by alcohol from water in a fibrous form without rendering the liquid opaque.

5. *Ulm*.—Does not crystallize. Taste insipid. Sparingly soluble in water and alcohol, and does not form mucilage. Very soluble in liquid carbonate of potash. Precipitated by acids and sulphate of iron.

6. *Nicotin*.—Volatile, and when drawn into the nose excites sneezing. Precipitated by tincture of nutgalls.

7. *Extractive*.—Soluble in water and alcohol, insoluble in ether. Precipitated by chlorine, muriate of tin, and muriate of alumina; but not by gelatine. Dyes fawn.

8. *Morphia*.—Soluble in water and alcohol. Crystallizes. Converts vegetable blues to green. Forms neutral salts with acids.

9. *Asparagin*.—Crystallizes. Taste cooling and nauseous. Soluble in hot water. Insoluble in alcohol, soluble in nitric acid, and converted into bitter principle or artificial tamin.

10. *Cerasin*.—Insoluble in water and alcohol. When put into water it swells up into a transparent jelly, which may be employed to paste paper. Soluble in water acidulated with nitric acid. Without taste or smell.

11. *Inulin*.—A white powder. Insoluble in cold water. Soluble in boiling water; but precipitates unaltered after the solution cools. Insoluble in alcohol. Soluble in nitric acid, and yields oxalic acid.

Book IV.

12. *Starch*.—A white powder. Taste insipid. Insoluble in cold water. Soluble in hot. Solution opaque and glutinous. Precipitated by infusion of nutgalls; precipitate re-dissolved by a heat of 120°. Insoluble in alcohol. Soluble in dilute nitric acid, and precipitated by alcohol. Yields with nitric acid oxalic acid and a waxy matter.

13. *Indigo*.—A blue powder. Taste insipid. Insoluble in water, alcohol, ether. Soluble in sulphuric acid. Soluble in nitric acid, and converted into bitter principle and artificial tannin. Volatilized by heat in a purple smoke.

14. *Gluten*.—Forms a ductile elastic mass with water. Partially soluble in water; precipitated by infusion of nutgalls and chlorine. Soluble in acetic acid and muriatic acid. Insoluble in alcohol. By fermentation becomes viscid and adhesive, and then assumes the properties of cheese. Soluble in nitric acid, and yields oxalic acid.

15. *Pollenin*.—Insoluble in water, alcohol, ether, oils, and petroleum. Forms a brown coloured solution in caustic alkalies. Burns rapidly when thrown into the flame of a candle. Colour yellow. Not altered by exposure to the air. Destitute of taste and smell.

16. *Fibrin*.—Tasteless. Insoluble in water and alcohol. Soluble in diluted alkalies and in nitric acid. Soon putrifies.

17. *Fixed oils*.—No smell. Insoluble in water, alcohol. Form soaps with alkalies. Coagulated by earthy and metallic salts.

18. *Wax*.—Insoluble in water. Soluble in alcohol, ether, oils. Forms soaps with alkalies. Fusible.

19. *Volatile oil*.—Strong smell. Insoluble in water; soluble in alcohol. Liquid. Volatile. Oily. By nitric acid inflamed, and converted into resinous substances.

20. *Camphor*.—Strong odour. Crystallizes. Insoluble in water; soluble in alcohol, oils, acids; insoluble in alkalies. Burns with a clear flame, and volatilizes before melting.

21. *Bird-lime*.—Viscid. Taste insipid. Insoluble in water. Partially soluble in alcohol. Very soluble in ether; solution green.

22. *Olivile*.—White flattened needles. No smell. A sweet and bitter taste. Melts at 158°. Insoluble in cold water, soluble in 32 parts of hot water. Very soluble in alcohol. Insoluble in sulphuric ether, and in the fixed and volatile oils. Soluble in nitric acid, and converted into oxalic acid, and bitter principle. Soluble in concentrated acetic acid. Solution colourless, and not precipitated by water.

23. *Resins*.—Solid. Melt when heated. Insoluble in water. Soluble in alcohol, ether, and alkalies. Soluble in acetic acid. By nitric acid converted into artificial tannin.

24. *Guaiacum*.—Possesses the characters of resins, but dissolves in nitric acid, and yields oxalic acid and no tannin.

25. *Balsams*.—Possess the characters of the resins, but have a strong smell; when heated, benzoic acid sublimes. It sublimes also when they are dissolved in sulphuric acid. By nitric acid converted into artificial tannin.

26. *Caoutchouc*.—Very elastic. Insoluble in water and alcohol. When steeped in ether, reduced to a pulp which adheres to every thing. Fusible and remains liquid. Very combustible.

27. *Gum resins*.—Form milky solutions with water, transparent with alcohol. Soluble in alkalies: With nitric acid converted into tannin. Strong smell, brittle, opaque, infusible.

28. *Cotton*.—Composed of fibres. Tasteless. Very combustible. Insoluble in water, alcohol, and ether. Soluble in alkalies. Yields oxalic acid to nitric acid.

29. *Suber*.—Burns bright and swells. Converted by nitric acid into suberic acid and wax. Partially soluble in water and alcohol.

30. *Wood*.—Composed of fibres. Tasteless. Insoluble in water and alcohol. Soluble in weak alkaline ley. Precipitated by acids. Leaves much charcoal when distilled in a red heat. Soluble in nitric acid, and yields oxalic acid.

31. *Medullin*.—Insoluble in water, alcohol, ether, and oils. Destitute of taste and smell. Structure porous. Soluble in nitric acid, and forms oxalic acid, but no suberic acid. When distilled it yields ammonia, and leaves a charcoal having a metallic lustre.

32. *Fungin*.—Insoluble in water, alcohol, and ether. Partially soluble in a strong boiling alkaline ley. Soluble in hot muriatic acid, and forming a gelatinous matter. When distilled it yields ammonia.

But these are not the only substances which occur in the vegetable kingdom. It cannot be doubted that there are several others easily recognized by their peculiar properties in certain vegetable bodies; though chemists have not yet succeeded in obtaining them in a separate state, and in ascertaining their characters with precision. It was this deficiency of precise information that induced me to omit them; for nothing

Book IV.

can be more hazardous than to swell the list of vegetable constituents, by introducing supposed substances from vague analogies and imperfect experiments.

Acrid principle.

There are many vegetable substances remarkable for a peculiar acrid property. Most of the *ranunculi*, the *polygonum hydropiper*, *mustard*, &c. are well known instances. In some of these, the acidity disappears on drying, while in others, as *mustard*, it remains. In some, as the *ranunculi*, the acrid portion is taken up by water; while in *mustard*, water dissolves only a very small part, and acquires the taste of garlic. This acrid property certainly depends upon the presence of some principle or principles with which we are unacquainted.

We are equally ignorant of the nature of the peculiar liquid to which the leaves of the common nettle owe their property of irritating the skin. We do not know the substance which gives the alliaceous tribe their peculiar taste and smell; a substance amazingly penetrating, but which is destroyed or dissipated by the heat of boiling water. We know little of the poisonous juices with which many vegetables are filled. Indeed the only vegetable poison with which we can say we are acquainted is *prussic acid*. These are but a small number of instances, but abundance of others will occur to every reader. Indeed whole regions of the vegetable kingdom remain still unexplored. Much curious information may be expected from the analysis of the *fungi*, and other kindred plants; from the analyses of the *algæ*, and of the *lichens*.

Vegetable analysis imperfect.

Indeed the state of the chemical analysis of vegetables is still very imperfect. No general rules have yet been laid down; no successful method has yet been ascertained. In every particular case the analyst must follow his own judgment, and be guided by his own experience. Thus every one is obliged, in some measure, to invent a method for himself, and to create the whole as it were anew. Thus much time is wasted before the chemist is upon a level with his predecessors: and every one striking, as it were, from the same point, the progress is much slower than it otherwise would be.

The older chemists confined their analysis entirely to destructive distillation. By this process they obtained nearly the same products from every vegetable. For every plant when distilled yields water, oil, acid, and carburetted hydrogen and carbonic acid gas; while a residuum of charcoal remains in the retort. For the first introduction of solvents we

are indebted to the apothecaries. The experiments of Boerhaave and Neumann were doubtless of value; but it was Rouelle who first made the great step in vegetable analysis, by pointing out the uses of various solvents to separate the different constituents of vegetables from each other. He was the first that attempted a precise description of the vegetable principles, and that gave marks by which their presence might be ascertained. The next great step was made by Scheele, who detected the different acids of vegetables, ascertained their properties, and pointed out the method of separating them from those vegetable substances into which they enter as constituents. The experiments of some of the German chemists, especially of Hermbstadt, added considerably to the discoveries of Rouelle and Scheele, and made us better acquainted with some of the constituents of vegetables. But of all the modern chemists none is entitled to greater praise than Vauquelin. His analyses have been numerous, and his discoveries important. He has introduced into the subject that precision which distinguishes all his researches, and his method has been followed by all the other French chemists. Fourcroy has been often associated with Vauquelin, and his analysis of the quinquina, which he published alone, is entitled to considerable praise. Proust has turned his attention to the same subject, and has favoured the world with very instructive papers. Till lately the analysis of vegetable substances was almost entirely overlooked by British chemists; but the fineness of the field has now begun to attract their attention. Experiments of great importance have been published by Davy, Chenevix, &c. and above all by Hatchett, to whom vegetable chemistry lies under deep obligations; and we may expect still more important information from the zeal and skill of those enlightened philosophers. A few years will probably change the appearance of this branch of chemistry almost entirely, and bring it nearer to the state which the chemistry of minerals has already acquired.

Thus we have noticed all the substances which have been hitherto examined from vegetables. By analysing each of them, or by exposing them repeatedly to destructive distillation, combustion, &c. we come at last to those bodies which we are at present obliged to consider as simple, because they have not yet been decomposed, and of which accordingly we must suppose that vegetables are ultimately composed. These amount to 15, namely,

Chap. I.
Improvements of
Rouelle.

Book IV.
 Substances
 found in
 plants.

- | | |
|-------------------|----------------|
| 1. Oxygen. | 9. Iron. |
| 2. Sulphur. | 10. Manganesc. |
| 3. Phosphorus. | 11. Potash. |
| 4. Carbon. | 12. Soda. |
| 5. Hydrogen. | 13. Lime. |
| 6. Azote. | 14. Magnesia. |
| 7. Muriatic acid. | 15. Alumina. |
| 8. Silica. | |

But of these substances there are twelve which compose but a very small proportion indeed of vegetables. Almost the whole of vegetable substances are composed of four ingredients, namely,

Carbon.	Oxygen.
Hydrogen.	Azote.

Of these the last, namely azote, forms but a small proportion even of those vegetable substances of which it is a constituent part, while into many it does not enter at all: so that, upon the whole, by far the greater part of vegetable substances is composed of carbon, hydrogen, and oxygen. I do not mention caloric and light; concerning the nature of which too little is known to enable us to determine with certainty into what substances they enter.

The substances at present known to chemists, which they have not been hitherto able to decompose, amount (omitting caloric and light) to 49. Fifteen of these exist in plants; the rest belong exclusively to the mineral kingdom: for it is a fact, that no substance (I mean simple substance) has been hitherto found in the animal kingdom which does not exist also in vegetables.

On the contrary, all the simple substances at present known may be found in minerals. This indeed ought not to surprise us, if we recollect that the spoils of animals and vegetables, after they have undergone decomposition, are ultimately confounded with minerals, and consequently arranged under the mineral kingdom. Besides, if vegetables draw their food from the mineral kingdom, it would be absurd to suppose that they contain substances which they could not have procured from minerals.

Only ten of these principles (omitting the acids) have been analysed with sufficient precision to give us an idea of the number of atoms of which they are composed. The following table exhibits the number of atoms of oxygen, carbon,

and hydrogen, which enter into the composition of these ten bodies: Chap. II.

	Oxygen.	Atoms of Carbon.	Hydrogen.		
Wax	1	20	18	=	39 atoms
Rosin	3	10	11	=	24
Sandarach	1	14	13	=	28
Copal	2	19	18	=	39
Caoutchouc	4	10	9	=	23
Woody fibre	4	7	4	=	15
Starch sugar	5	5	5	=	15
Common sugar	5	6	5	=	16
Gum arabic	6	6	6	=	18
Starch	9	10	10	=	29

From this table it appears that the composition of all of them is very complicated. We have no conception at present in what way these atoms are united together. Most probably they are first grouped together in binary or ternary compounds, and a certain number of these primary groups go to the composition of the vegetable principle.

CHAP. II.

OF THE PARTS OF PLANTS.

We have, in the preceding Chapter, enumerated and described the different substances hitherto detected in plants. But a more difficult task remains for the present Chapter; namely, to explain the composition of each vegetable organ in all the numerous families of plants which constitute the vegetable kingdom. This task, indeed, in the present state of vegetable chemistry, cannot be fulfilled. By far the greater number of plants have never been examined at all; and even of those which, from their medicinal virtues or nutritive qualities, have attracted the attention of chemists, only particular organs have been analysed, while the rest have been neglected as unworthy of notice. Nothing, therefore, either resembling a complete view, or an exact arrangement, is to be looked for in this Chapter. I shall satisfy myself with stating the most important facts hitherto discovered respecting the composition of plants, as far as I am acquainted with them, under the fifteen following heads:

Book IV.

- | | |
|--------------------|----------------|
| 1. Sap. | 9. Seeds. |
| 2. Juices. | 10. Fruits. |
| 3. Air. | 11. Bulbs. |
| 4. Wood and roots. | 12. Lichens. |
| 5. Bark. | 13. Mushrooms. |
| 6. Leaves. | 14. Fuci. |
| 7. Flowers. | 15. Diseases. |
| 8. Pollen. | |

These form the subjects of the following Sections.

SECT. I.

OF THE SAP OF PLANTS.

Definition.

It is the general opinion of physiologists, that plants receive a considerable part of their nourishment by the root; that it enters into them in a liquid state, and passes up in proper vessels towards the leaves. This liquid is distinguished by the name of *sap*. In the spring, when the buds begin to expand themselves into leaves, if we break off the extremity of a branch, or cut into the wood of a tree, this sap flows out, and may be obtained in considerable quantities. It was first examined by Dr. Hales; but chemical analysis had not made sufficient progress in his time to enable him to ascertain its constituents. Deyeux and Vauquelin have more recently analyzed the sap of different trees. To them we are indebted for most of the facts known respecting this liquid. A few additional facts have been ascertained by John.

Constituents.

The sap in all the vegetables hitherto examined is nearly as liquid as water. It always contains an acid, sometimes free, but more commonly combined with lime and potash. Various vegetable principles are also present: of these sugar is the most remarkable, and mucilage. Sometimes albumen and gluten, and sometimes tannin, can be detected. When left to itself, the sap soon effervesces and becomes sour; or even vinous, when the proportion of sugar is considerable.

Hitherto the sap of a few species of trees only has been examined. We are not in possession of any means of collecting the sap of the inferior orders of plants. The expressed juices of a considerable number of vegetables, indeed, have been prepared for medicinal purposes; but these are not sap, but a collection of all the liquid substances which the plant

contained. At present, then, it is not possible to present a general view of the properties of sap. The following are the particular species which have been examined: Chap. II.

1. *Sap of the Elm, ulmus campestris.*

Vauquelin collected three different specimens of the sap of this tree; the first portion towards the end of April, the second in the beginning of May, and the third about the end of May. It had a reddish-brown colour; its taste was sweet and mucilaginous; and it scarcely altered the colour of the infusion of litmus. Ammonia, barytes, and lime water throw down a copious yellow-coloured precipitate, which dissolves with effervescence in acids. Oxalic acid and the nitrate of silver throw down a white precipitate. Diluted sulphuric acid occasions a brisk effervescence, and disengages the odour of acetic acid. Chlorine destroys the colour of the sap, and throws down a brown precipitate. Alcohol produces a flaky precipitate. When evaporated by a gentle heat, a pellicle forms on the surface; brown flakes precipitate, and an earthy matter is deposited on the sides of the vessel. The earthy matter was a mixture of carbonate of lime and vegetable substance. The liquid, after depositing these bodies, and being evaporated to $\frac{2}{10}$ ths of its original bulk, contained a considerable portion of acetate of potash. Properties.

1039 parts of this sap were composed, according to Vauquelin's analysis, of

Water and volatile matter	1027·904	Constituents.
Acetate of potash	9·240	
Vegetable matter	1·060	
Carbonate of lime	0·796	

The vegetable matter was partly extractive and partly mucilaginous.*

On analyzing the same sap somewhat later in the season, Mr. Vauquelin found the quantity of vegetable matter a little increased, and that of the carbonate of lime and acetate of potash diminished. Still later in the season the vegetable matter was farther increased, and the other two ingredients farther diminished. The carbonate of lime was held in solution by carbonic acid, of which there existed a considerable Changes of the sap by vegetation.

* Ann. de Chim. xxxi. 20.

Book IV. excess in the sap. It is to this acid gas that the air bubbles, which so often accompany the sap as it issues from the tree, is owing.*

2. *Sap of the Beech, fagus sylvatica.*

Vauquelin collected two different specimens of this sap; the first in the end of March, the second about the end of April. **Properties.** It had a reddish-brown colour, and a taste similar to the infusion of tan. It slightly reddened vegetable blues. Barytes, ammonia, carbonate of potash, and oxalate of ammonia, occasion precipitates in it; chlorine throws down yellow flakes; sulphuric acid blackens it, and disengages the odour of acetic acid; sulphate of iron strikes a black, and glue throws down a copious whitish precipitate. When gently evaporated to dryness, it leaves a brown extract amounting to about $\frac{1}{4\frac{1}{5}8}$ of its weight, ductile while hot, but brittle when cold, and having the smell and somewhat of the taste of new baked bread. It absorbs moisture from the atmosphere, and increases in weight about $\frac{1}{7}$ th. Lime disengages ammonia, and sulphuric acid acetic acid, from this extract. Alcohol dissolves only a small part of it. This sap contained the following ingredients:

**Constitu-
ents.**

Water.
Acetate of lime with excess of acid.
Acetate of potash.
Gallic acid.
Tannin.
A mucous and extractive matter.
Acetate of alumina.

It contained, besides, a colouring matter, which may be fixed on cotton and linen by means of alum, and dyes them of a fine solid reddish-brown colour.†

3. *Sap of the Hornbeam, carpinus sylvestris.*‡

Three specimens of this sap were collected by Vauquelin **Properties.** during the months of March and April. It was limpid, and its colour was whitish; its taste slightly sweet, and its smell analogous to that of whey. Barytes throws down from it a copious white precipitate, soluble in muriatic acid. Carbonate of potash likewise throws down a precipitate soluble in acids with

* See Coulomb, Jour. de Phys. xlix. 392.

† Ann. de Chim. xxxi. 26.

‡ I presume the carpinus betulus is meant.

effervescence. Sulphuric acid deepens the colour, and evolves the odour of vinegar. Oxalic acid throws down a copious precipitate, and nitrate of silver gives the solution a fine red colour. 3918 parts, when distilled, left an extract of a reddish yellow colour, amounting to 8.279 parts. It had a sharp taste, and attracted humidity from the air.

When the extract is digested in alcohol, about the half of it dissolves. This portion consists of extractive, a saccharine matter, and acetate of potash. The residue, which is soluble in water, consists of mucilaginous matter, acetate of lime, and a colouring substance.

When this sap was left exposed to the air in an open glass vessel, it became milky, disengaged carbonic acid, acquired a spirituous smell and taste, and its acidity increased. After some weeks this odour was dissipated, and carbonic acid was no longer extricated. Its acidity continued still to increase, white flakes fell to the bottom, and the liquid became transparent. After 50 days the acidity was found diminished, a mucous pellicle formed on the surface, which became at last blackish brown, and the liquid had only a mouldy taste. In a close bottle the sap never became transparent; and when the bottle was opened after three months, the air which it contained was found converted into azote and carbonic acid. The liquid had a very strong taste of vinegar.*

Action of
the air.

4. Sap of the Birch, *betulus alba*.

The sap of this tree is colourless; it has a sweet taste; reddens vegetable blues. Neither ammonia, alcohol, nor chlorine, produce any change upon it. Barytes and lime throw down a precipitate which dissolves in muriatic acid. Hydrosulphurets, sulphate of iron, and glue, produce no effect. Oxalic acid throws down a white precipitate. Sulphuric acid disengages the odour of vinegar. Nitrate of silver strikes a red colour. When evaporated to about $\frac{1}{4}$ th, it lets fall a reddish-brown powder insoluble in water. When 3918 parts of the sap were evaporated to dryness, they left 34 parts of brown extract. This had an agreeable taste, attracted moisture from the atmosphere, and was almost completely soluble in alcohol. When this extract is dissolved in water and mixed with yeast, it ferments, and the fermented liquor yields a considerable proportion of alcohol; it yields also a considerable portion of vinegar. Vauquelin did not succeed in his attempts to obtain crystallized sugar from this sap; but he ascertained that it con-

Properties.

Fermented.

* Ann. de Chim. xxxi. 51.

Book IV.

tained a portion of extractive which dyes wool of a brownish yellow colour. Acetate of lime and of alumina were also present, and probably also acetate of potash.*

More lately Dr. John has made a few experiments on the sap of the birch. He found it transparent, colourless, destitute of smell, and scarcely sweetish. Its specific gravity was 1.004. One portion of birch sap which had been frozen did not alter vegetable blues; but another portion which had not been frozen reddened them. From 1730 grains of sap he obtained only 4 grains of sugar. He found in it likewise a little carbonate of lime, which precipitated of its own accord when the sap was left exposed to the air. He thought likewise that it contained traces of a tartrate, some gummy matter, and some albumen.†

5. Sap of the Yoke Elm.

Properties.

This sap has a bitter taste. When evaporated to dryness it leaves a brown extract, in which crystals of nitrate of potash are gradually formed. This extract was scarcely soluble in alcohol. It gave traces of the presence of acetate of potash and acetate of lime.‡

6. Sap of the Vine, *vitis vinifera*.

This sap was examined by Dr. Prout. It had the whitish appearance of common river water. Its taste was sweetish but not rough. It did not alter litmus paper, nor did its specific gravity differ from that of pure water. Potash and ammonia gave it a fine red colour, and threw down red flocks readily dissolved by acetic acid. It was slightly precipitated by oxalate of ammonia, ferrocyanate of potash, nitrate of silver, and subacetate of lead. Two thousand three hundred parts of it when evaporated left only one part of residuum, half of which was carbonate of lime and the remainder a peculiar vegetable matter, which was insoluble in alcohol. The sap had contained both carbonic and acetic acids, and likewise an alkali.§

7. Sap of the common maple, *acer compestre*.

The sap of this tree was examined by Professor Scherer, of Vienna. It has a milky aspect, a sweetish taste, and its specific gravity varies. It does not alter litmus or turmeric paper. It is precipitated by oxalate of potash, nitrate of silver, and

* Ann. de Chim. xxxi. 36.

+ Chemische Untersuchungen, ii. 4.

‡ Ann. de Chim. xxxi. 38.

§ Annals of Philosophy, v. 109.

barytes water; but not by muriate of barytes. When boiled it lets fall gluten in flocks. It yields when evaporated a salt with basis of lime possessing peculiar properties. The acid is destroyed by heat, and according to Scherer differs from every other vegetable acid. He therefore distinguishes it by the name of *aceric acid*. The acerate of lime is white, slightly translucent; has a weak acidulous taste, and is not altered by exposure to the atmosphere. One thousand parts of cold water dissolve 9 parts, and 1000 parts of boiling water 17 parts of this salt.*

Chap. II.

Aceric acid.

SECT. II.

OF THE PECULIAR JUICES.

THE sap passes from the roots in peculiar vessels to the leaves, where it is altered by a process similar to that of digestion in animals, and formed into all the liquid substances requisite for the purposes of the plant. These liquids flow from the leaves towards the root in appropriate vessels, and have received the name of the *peculiar juices* of vegetables. They differ very considerably from each other in different plants. They have all a certain degree of consistency, and always contain much more vegetable matter than the sap. In the present state of vegetable chemistry, an accurate detail of their properties cannot be attempted. Indeed it is often difficult to procure them from any plants unmixed with the sap. They sometimes exude spontaneously, and may always be procured in smaller or greater quantity by incisions through the bark of the plants containing them. The following are the species of peculiar juices which have been hitherto attended to.

1. *Milky Juices*.—Many plants, when wounded, emit a considerable quantity of a milky liquid, which may, in most cases at least, be considered as one of the peculiar juices of the vegetable from which it flows. The nature of this juice is extremely various.

The root of the *campanula rotundifolia* yields a milky juice, of a peculiar, and not unpleasant smell and taste. Children in some parts of Scotland collect the plant for the sake of this juice, which they suck with avidity. Its chemical properties have not been examined.

* Schweigger's Jour. iv. 362.

Book IV.

Humboldt mentions a tree which is a native of South America, which yields a milky juice, employed by the natives as a nourishing article of food. It is called in the country the cow tree, and the milk of it on standing forms a film on the surface, resembling in its properties the curd of milk.*

Euphorbia.

The different species of *euphorbia* yield a milky juice, which has a hot taste analogous to that of pepper, but more acrid, and which continues for a long time in the mouth. When chlorine was poured into this juice, a very copious white precipitate fell down. This powder, when washed and dried, had the appearance of fine starch, and was not altered by keeping. It was neither affected by water nor alkalies. Alcohol, assisted by heat, dissolved two-thirds of it; which were precipitated by water, and had all the properties of resin. The remaining third part possessed the properties of the *woody fibre*. Mr. Chaptal tried the same experiment on the juices of a great number of other plants, and he constantly found that chlorine precipitated from them *woody fibre*.†

Dr. John has made a set of experiments on the milky juice of the *euphorbia cyparissias*. Its colour is milk-white. It is opaque and has a slippery feel. Its taste is at first sweetish with an impression of bitterness, and it leaves a feeling of burning heat in the mouth and on the tongue, which continues for some time. He found its constituents as follows :

Water	77
Tartaric acid	—
Resin	13·80
Gum	2·75
Extractive	2·75
Albumen	1·37
Caoutchouc	2·00
Fixed oil	—

 99·67 †

Poppy.

The different species of the poppy (*papaver*) and lettuce (*lactuca*) yield a milky juice, which possesses narcotic qualities, and is distinguished by a peculiar taste and smell. These juices have been more examined than the preceding. They are of a very complex nature, containing a great variety of constituents. Gum and resin, and extractive, are the most

* Ann. de Chim. et de Phys. vii. 132.

† Ann. de Chim. xxi. 235.

‡ Chemische Untersuchungen, i. 3.

prominent of these; and morphia and meconic acid constitute essential ingredients in these juices.

Caoutchouc.

The milky juice which exudes from the *jatropha elastica*, the *hevea caoutchouc*, the *artocarpus integrifolia*, the *urceola elastica*, and from several other plants, gradually becomes inspissated when exposed to the air, and constitutes caoutchouc. This juice has been hitherto examined only by Fourcroy. Though kept in close vessels, the caoutchouc for the most part gradually separates, and forms a white solid elastic mass. The juice, however, still continues milky. When evaporated in the open air, a pellicle of caoutchouc gradually forms on the surface, and when this is removed another succeeds. Fourcroy ascribes this formation to the absorption of oxygen from the atmosphere. Besides the caoutchouc, Fourcroy obtained from the juice transparent prismatic crystals, which had a sweet taste, and which he considered as saccharine matter approaching to the nature of an acid.*

The juice of the papaw tree possesses properties which distinguish it from most others. According to the analyses of Vauquelin, detailed in the preceding Chapter, its constituents resemble very closely the constituents of blood.

Papaw.

Dr. John has made some experiments to ascertain the nature of the milky sap of the *Asclepias Syriaca*, which he was enabled to examine from the plant being cultivated in the Botanic garden at Berlin. The sap of this plant is milky, reddens vegetable blues, has a peculiar smell similar to that of narcotic substances. Its taste is sourish, peculiar but not hot. Its constituents were as follows:

Resin	26.5
Caoutchouc	12.5
Gluten	4
Extractive	4
Tartaric acid } Held in solution	
Albumen . . } by water	73
120.0†	

The milky juice of the *Chondrilla Juncea*, a plant which vegetates abundantly in the neighbourhood of Frankfort, consists, according to the same chemist, of water, holding in solu-

* Ann. de Chim. xi. 225.

† Chemische Untersuchungen, i. 20.

Book IV. tion mucilage, resin, and caoutchouc, in the following proportions:

Mucilage	$\frac{1}{3}$
Resin	$\frac{1}{4}$
Caoutchouc	$\frac{1}{2}$ *

Besides the milky juices enumerated above, there are many others. Almost all the gum resins indeed are originally in that state; but as no chemical analysis of these juices has yet been made, it is needless to dwell upon them.†

Mucilaginous juices.

2. *Mucilaginous Juices*.—The peculiar juices of many plants are not milky, and some not distinguished by any strong taste or smell. In these mucilage seems to be the predominating matter. Under this head may be reckoned the juices of most of the mucilaginous plants enumerated in the fifth Section of the preceding Chapter.

Cambium.

The substance called *cambium*, too, if we may be allowed to consider it as a peculiar juice, since it is obviously different from the sap, is entirely mucilaginous. It makes its appearance, according to Mirbel, in all those parts of vegetables where new matter is to be formed, and seems necessary for all such formations, either as the matter employed in their formation, or as furnishing a proper bed for them to be formed in. It does not appear to be confined in vessels like the other juices.‡

Turpentine.

3. Some juices are intermediate between volatile oils and resins. These may be distinguished by the name of *turpentine*. Common turpentine, balsam of copaiva, and probably also opobalsamum, belong to this class. These juices are originally limpid and transparent, have a strong smell and taste, and gradually acquire consistency by exposure to the air, probably by absorbing oxygen.

Resins.

4. Other juices obtained by excision possess at first the properties of resins, or at least acquire them before they are brought into this country. Such, for example, are tacamahac and mastich, and most of the resinous bodies described in Section XXVI. of the preceding Chapter.

Balsams.

5. Others are entitled to the name of *balsam*, according to

* *Chemische Untersuchungen*, ii. 1.

† The reader may consult a dissertation on the milky juices of plants by Carradori. See Geblen's *Jour.* vi. 630.

‡ Mirbel, *Ann. de Mus. d'Hist. Nat.* No. xl. p. 294.

the strict meaning of the word; as balsam of tolu, of Peru, styrax, and benzoin. Chap. II.

6. Some peculiar juices are composed almost entirely of tannin, or at least are characterised by containing a superabundance of that substance. Such probably are the juices of oak, sumach, and of most vegetables that yield abundance of the tanning principle. In some cases, it would seem that these juices exude spontaneously; though in general they are obtained by artificial means. Tannin.

7. Some vegetables possess juices characterized by the great quantity of sugar which they contain. Such, for example, are the sugar-cane, the carrot, and the various species of beet. For it is surely more reasonable to consider the saccharine matter in these plants as belonging to the peculiar juices, than as confined to the sap. Sugar.

8. Finally, the peculiar juices of some plants are characterized by containing a considerable portion of saline matter. Thus the various species of sorrel contain a notable quantity of binoxalate of potash, and several of the sedums malate of lime. In short, the peculiar juices of plants are nearly as numerous as the vegetable principles themselves; and when the fungi, algæ, lichens, and several other of the numerous inferior tribes of vegetables have been once examined, it cannot be doubted that the number will greatly increase. Saline matter.

Braconnot has lately examined the juices of a number of plants, in order to ascertain the peculiar acids which they contained. The following is a short abstract of the results which he obtained.*

The expressed juice of the *aconitum lycortonum* evaporated to dryness and incinerated leaves about 0.01 of carbonate of potash. This juice contains a considerable proportion of citric acid, partly combined with potash and partly with lime. Perhaps also malic and acetic acids are present in the juice of this plant.

The juice of the *delphinium elatum*, the *ranunculus aconitifolius*, the *thalictrum flavum*, the *clematis recta* and *viticella*, likewise contains a quantity of citric acid like the preceding juice.

The juice of the *salvia sclarea* contains benzoic acid probably combined with potash.

* Ann. de Chim. lxxv. 277.

Book IV.

The juice of the *ruta graveolens* contains malic acid combined with potash and with lime.

The juice of the *eupatorium cannabinum* contains an acid which appears to be the malic mixed with some phosphoric.

The juice of the *nicotiana rustica* and *tabacum* contains malic acid combined with potash and lime.

The juice of the *mirabilis jalapa* contains nitric acid, muriatic acid, malic acid, and a little sulphuric acid chiefly combined with potash.

The *spinacia oleracea* contains oxalates of lime, and of potash, malate and phosphate of potash.

The *tropæolum majus* contains phosphoric acid, nitric acid, and malic acid united to lime and to potash.

The *ricinus communis* contains malic acid doubtless combined with potash.

The *phytolacca decandra* contains an uncommon proportion of potash, and an acid which possesses the properties of the oxalic.

The juice of the *vaccinium myrtillus* or *bilberry*, contains a colouring matter, citric and malic acids, and a considerable quantity of uncrystallizable sugar. Charcoal removes the colouring matter completely from this juice. When diluted with water, and mixed with yeast, it ferments readily, and yields a considerable quantity of sugar.*

SECT. III.

OF THE GASES IN PLANTS.

IN many plants the stem is hollow and filled with air. In others, as the onion, the leaves are filled with air. Air is lodged in the pod of the pea, in the leaves of some species of fuci. In short, there is hardly a plant that does not possess some part more or less hollow, and of course filled with air. Now it is a question of some curiosity to determine what species of gas thus fills up the hollow parts of plants. Is it common air? or is it a secretion by the plant itself? On the latter supposition it may be hydrogen gas, azotic gas, carbonic acid gas, or any other gas whatever.

A few experiments were made by Dr. Priestley on the air

* Vogel, *Annals of Philosophy*, xii. 232.

in the sea-weed. He found it sometimes the same as common air; in other cases it contained a greater proportion of oxygen, in others a greater proportion of azote. The air within the leaves of onions and in the bladder of senna he found the same as common air. Air pressed from the stalks of the common *flag* contained a greater proportion of azote than common air.*

The air in two or three plants examined by Dr. Darwin was the same as common air. Hubert found the air from the *arundo bambos* worse than common air.

M. Bidault de Villiers examined the air in a considerable number of plants. That in the leaves of the onion was the same as common air. The same observation applies to the air in the petals of the *melo peppo* in the capsules of the *colutea arborescens*, the pods of the *pisum sativum*, the capsules and vesicles of the *staphylea pinnata*, and *nigella damascena*, the stems of the *borago officinalis*, *conium maculatum*, *sonchus oleraceus*. In some cases he found the air in the leaves of onions extinguish a candle; it must of course have contained a considerable excess of azote.†

From these few observations it seems to follow, that the air contained in plants is common air, frequently unaltered; but sometimes deprived of a portion of its oxygen. In Dr. Priestley's experiment it would appear that a redundancy of oxygen was present. But this experiment would require repetition before it could be considered as authenticated.

SECT. IV.

OF THE WOOD AND ROOTS.

I. THE roots of a great variety of plants are employed in medicine and the arts; but few of them have been hitherto subjected to chemical examination. The substances found in them are various; and indeed, as the peculiar juices of the roots are always included in such examinations, it is clear that almost all the vegetable principles will be found in them.

The process usually followed in examining these substances, is first to separate every thing soluble in cold water, then to boil the root in water. Alcohol is applied next; and afterwards different acids or alkalis, according to the nature of the

* Priestley, iii. 279.

† Ann. de Chim. lxxxviii. 89.

Book IV.

substances suspected to be present. The roots of the following plants have been subjected to chemical examination :

List.	Helleborus hyemalis.	Rubia tinctorum.
	Brionia alba.	Curcuma longa.
	Calaguala.	Valeriana officinalis.
	Ho-ang-lien.	Cochlearea armorica.
	Cephælis, or calicocca.	Glycyrrhiza glabra.
	ipecacuanha.	Inula helenium.
	Convolvulus jalappa.	Acorus calamus.
	Gentiana lutea.	Andropogon schænanthus.
	Rheum palmatum.	

Helleborus
hyemalis.

1. *Helleborus hyemalis*.—The root of this plant is tuberous, has a yellowish-white colour, and is covered with a black skin. At first it has no taste: but in a short time a strong acridity becomes sensible in the mouth and throat. Vauquelin has subjected it to a chemical examination, chiefly in order to ascertain the nature of the bitter and acrid principle which it contains. This principle he found to be an oil of a peculiar nature, possessing properties intermediate between the fixed and volatile oils. He obtained it by digesting the root in alcohol, and then distilling off the alcohol in an alembic. The oil gradually separated and concreted on cooling. Its taste was extremely acrid, and it had a yellowish-brown colour. When dissolved in weak alcohol, it precipitates the sulphates of iron of a fine purplish-red colour, which becomes green by means of alkalis. This oil is extremely poisonous, and, according to Vauquelin, exists in many plants, and is the cause of their poisonous qualities.

Contains a
poisonous
oil.

When the root is digested in water, and the liquid passed through a cloth, it is obtained opaque and milky, and gradually deposits a white powder which possesses the properties of starch. The liquid, by evaporation, deposits a brownish matter, which forms successive pellicles on the surface. This substance possesses the properties of extractive. Besides these substances, Vauquelin detected a substance analogous to gluten, some sugar, and a portion of woody fibre.*

Brionia
alba.

2. *Brionia alba*.—This root has been used in medicine, and was known to contain a considerable portion of starch, and a bitter principle soluble in water and alcohol. Vauquelin lately examined it. By maceration in water, and subsequent

* Ann. de Mus. d'Hist. Nat. No. xliii. 32.

pressure in a linen-cloth, the starch was separated and obtained in a state of purity. The bitter substance was soluble both in alcohol and water, and appeared to possess the properties of pure bitter principle. The root was found likewise to contain a considerable portion of gum; a substance which is precipitated by infusion of galls, and which Vauquelin denominates *vegeto-animal matter*, some woody fibre, a small portion of sugar, and a quantity of supermalate of lime and phosphate of lime.*

Chap. II.

Constitu-
ents.

3. *Calaguala root*.—This is a root brought from America for medicinal purposes, which has acquired considerable celebrity on the continent. The plant from which it is obtained is not known; but from the observations of Vauquelin is supposed to be a species of *polypodium*. It has a brown colour, is partly covered with scales like the roots of fern, and is very hard and difficult to reduce to powder. Vauquelin, who has subjected it to analysis, found it a very complex substance, and separated from it no fewer than the following substances:

Calaguala.

Woody fibre.
Gum.
Resin.
Sugar.
Starch.

Colouring matter.
Malic acid?
Muriate of potash.
Lime.
Silica.

Constitu-
ents.

Alcohol dissolves the resin and sugar. By evaporating the solution to dryness, and treating the residue with water, the sugar is separated and the resin left. This resin has a reddish-brown colour, and a bitter and acrid taste. It dissolves in alkalies, communicating a brown colour and a bitter taste, and is again separated by acids. Vauquelin suspects that this principle is the constituent of the root, both of this plant and of the other filices which possess vermifuge properties. Water dissolved the gum and the muriate of potash, which were obtained by evaporation. Diluted nitric acid dissolved the starch and the colouring matter, and let fall the former when mixed with four times its bulk of alcohol. The woody fibre remained, which when incinerated left carbonate of lime, muriate of potash, and a little silica. As the decoction reddened vegetable blues, it is possible that the lime was in combination with malic acid.†

* Ann. de Mus. d'Hist. Nat. No. xliii. 88.

† Vauquelin, Ann. de Chim. lv. 22.

Book IV. 4. *Ho-ang-lien*.—This is a bitter tasted root, brought from
 Ho-ang- China, and used on the continent as a medicine. It has been
 lieu. slightly examined by Bouillon La Grange. From his experi-
 ments, it seems to owe its properties chiefly to the bitter prin-
 ciple, of which it contains a considerable portion. It contains
 also a resinous matter, taken up by alcohol and thrown down
 by water. This substance when dry has a brown colour,
 melts when heated, and burns, emitting an aromatic odour.*

Ipecacuan. 5. *Ipecacuan*.—This is the root of a plant which grows
 spontaneously in Brazil, and probably in other parts of South
 America. It was first correctly described and figured by Dr.
 Brotero, under the name of *callicocca ipecacuanha*.† The
 root is about the thickness of a quill, unequal and knotty, and
 varies considerably in its colour. When pounded it forms the
 mildest and safest emetic in the whole *Materia Medica*.
 Though probably employed in America from time immemo-
 rial, it was not introduced into Europe till the time of Lewis
 XIV, when one Grenier, a French merchant, brought 150lbs.
 of it from Spain; with which trials were made at the Hotel
 Dieu. Helvetius first made known its use in dysentery, for
 which he was rewarded by Lewis XIV. with 1000*l.* sterling.‡

Properties. This substance has been recently analysed by M. M.
 Majendie and Pelletier. The following are the constituents
 which they obtained :

Oil	2
Emetin	16
Wax	6
Gum	10
Starch	42
Wood	20
Loss	4

100§

Jalap. 6. *Jalap*.—This very active cathartic is the root of the *con-
 volvulus jalappa*, a plant which is a native of Xalapa, a pro-
 vince of New Spain.|| It is brought to Europe in thin slices,
 which are hard, of a brown colour, exhibiting resinous streaks
 and circles, easily inflammable, of scarcely any smell, and a
 weak acrid taste. From the few experiments detailed by

* Vauquelin, *Ann. de Chim.* iv. 40. † *Linnean Trans.* vi. 137.

‡ *Neumann's Chem.* p. 357.

§ *Ann. de Chim. et Phys.* iv. 180.

|| Hence the term *Jalap*. *Neumann's Chem.* p. 352.

Neumann, it appears to contain a resin, to which its active properties are ascribed, and likewise a portion of extractive and of mucilaginous matter. It has been analysed by Mr. Henry, who found its constituents resin, extractive, starch, and woody fibre.* According to him, the proportions of these different substances, contained in 500 parts of the three varieties of jalap which occur in commerce, are the following:

	Resin.	Extract.	Starch.	Woody Fibre.
Jalap leger	60	75	95	270
— sain	48	140	102	210
— piqué	72	125	103	200

7. *Rhubarb*.—This is the root of the *rheum palmatum*, and perhaps also of some other species of *rheum*, brought chiefly from the northern parts of China by the way of Russia, though of late it has been cultivated also in Britain. The root is large, of an oblong or roundish shape; of a dark-brown colour externally, with black and reddish streaks; internally it is reddish-yellow, and when fresh contains a juice of the same colour. No accurate chemical analysis of rhubarb has yet been made; but from the experiments of Neumann, it appears that nearly one half of it is soluble in water, and that alcohol scarcely takes up any thing from the residue. From the properties of the watery extract, enumerated by that laborious chemist, we may infer, with some probability, that it consists chiefly of the extractive and bitter principle, and that it contains some tannin. A small quantity of greenish-yellow resinous matter seems also to be present.† Scheele separated from the root about $\frac{1}{4}$ th of its weight of oxalate of lime.‡ But this salt is not taken up by the water.

8. *Gentian*.—This is the root of the *gentiana lutea*, a plant which grows spontaneously in the mountainous parts of France, Switzerland, Hungary, &c. It is said to have been named after Gentius, a king of Illyria, who is supposed to have first discovered its virtues. The root is externally brown; internally yellow, with a spongy pith in the middle. Its taste is intensely bitter. From the experiments of Neumann, we may conclude that the soluble parts of it consist chiefly of bitter principle, mucilaginous matter, resin, and extractive. The first seems to be the most copious and most active ingredient, and the one to which it owes its medicinal properties.

* Ann. de Chim. lxxii. 275.

† Neumann's Chem. p. 360.

‡ Croll's Annals, i. 34. Eng. Trans.

Book IV.

It has been subjected to a chemical examination by Mr. Henry, who extracted from it the following substances :

(1) A substance possessing very nearly the properties of birdlime.

(2) A resin united to a little oil which gives its odour to gentian.

(3) An extract, similar to that of cinchona, which constitutes the greatest part of gentian, and which possesses the bitter taste and the other characters which distinguish this plant. It does not possess alkaline characters but agrees with those already recognized in the bitter principle.

(4) Gum united to a colouring matter.

(5) A salt with a calcareous base, possessing the characters of a phosphate.

(6) It contains neither starch nor inulin, nor any alkaline substance.*

Madder.

9. *Madder*.—This is the root of the *rubia tinctorum*, a plant cultivated in Zealand, &c. for the dyers. The root is the useful part, forming one of the most valuable of the red dyes. It is about the thickness of a goose quill, somewhat transparent, of a reddish colour, and a strong smell. The colouring matter of madder is the only part of this root that is valued. Mr. Watt has published a dissertation on it; to which we owe almost every thing that we know respecting its chemical nature.

Properties.

It appears that there are two distinct colouring matters in madder: a red, which is extracted readily by cold water; and a brown, which requires a boiling heat to separate it in any quantity. Hence the reason that dyers are cautious not to raise the infusion of madder to a boiling temperature till towards the end of their process. Both of these colouring matters appear to possess the properties of extractive. They are soluble in water and alcohol; and when the solution is evaporated, pellicles successively form on the surface. These gradually subside, and form a reddish-brown matter, only imperfectly soluble in water. When alum liquor is poured into the infusion of madder, a deep brownish-red flaky precipitate falls, and the liquid assumes a brownish-yellow colour; and alkaline carbonates and lime water throw down a blood-red lake, differing in the shade of its colour according to the proportion of alum employed. Acetate of lead thrown down a brown-coloured precipitate.

* Annals of Philosophy, xvi. 89.

When carbonate of magnesia is mixed with infusion of madder, the liquid assumes a bright blood-red colour; and by evaporation an extract is obtained, which readily dissolves in water, and when spread upon paper becomes yellow in a few days, if exposed to the sun.

Water, holding a little alkali in solution, forms with madder a dark brownish-red solution; but if the water be mixed with a little acid, the infusion is yellowish.*

From the late observations of Merimé, it appears probable that a third colouring matter is present in madder; that it is more soluble than the red; and that it may be separated by steeping the madder a certain time in water, and then precipitating the solution with an alkali. After this a much finer red lake may be obtained from madder.†

10. *Turmeric*.—This is the root of the *curcuma longa*, a Turmeric. plant which grows in the East Indies. It has a fine yellow colour, a slight smell, and an aromatic bitter taste. To water it gives its own yellow colour, and the infusion is rendered brown by alkalies. To alcohol it gives a red colour, which stains hot marble of a durable colour.‡ The constituents of this root are but imperfectly known. From the experiments of Neumann, we learn that it contains a little volatile oil; that water dissolves nearly one-third of the root, but that a very small portion only is soluble in alcohol.§ It forms the chief ingredient in *curry powder*.

11. *Valerian*.—Trommsdorf has examined the root of the Valerian. *valeriana officinalis*. It loses three-fourths of its weight by drying. Distilled with water it yields a volatile oil, very liquid, and of a greenish-white colour. Its odour is strong and camphoric; its specific gravity, at the temperature of 77°, is 0.9340; its taste is aromatic and camphoric without being acrid. When exposed to the light it becomes yellowish. Nitric acid converts it into a resinous substance, or, if it be used in a sufficient dose, into oxalic acid. The expressed juice of the roots of this plant is muddy, and has a strong odour. It lets fall a portion of starch. It contains a peculiar substance approaching the nature of extractive, soluble in water, but insoluble in ether, and in pure alcohol. It is precipitated from water by the salts of lead, silver, mercury, and antimony. This juice contains also a portion of gum. The roots deprived

* See Watt, Ann. de Chim. iv. 104.

† Berthollet, Elemens de l'Art de la Teinture, ii. 120.

‡ Lewis, Neumann's Chem. p. 384.

§ Ibid. p. 388.

Book IV.

of this juice yield a portion of black-coloured resin, but consist chiefly of woody fibre.*

Horse-radish.

12. *Horse radish*.—Einhoff has examined the root of this vegetable, the *cochlearea armorica*, and found that its acidity is owing to the presence of a small quantity of volatile oil, which he obtained by distilling the mashed roots in a water bath. A liquid came over which was at first milky, but gradually deposited a little of the volatile oil in question. This oil had a pale yellow colour, and the consistence of the oil of canella. Its odour was that of horse-radish excessively strong. Its taste was at first sweetish; but it left an acrid impression, and excited inflammation in those parts of the tongue and lips to which it was applied. It is heavier than water, with which it forms a milky liquid by agitation. It dissolves readily in alcohol. It is volatilized at the temperature of 60°. The liquid obtained from the roots of the horse radish by distillation yielded traces of sulphur.*

13. *Liquorice*.—The *glycyrrhiza glabra*, or liquorice, is a native of the south of Europe, but is cultivated in considerable quantities in England for the sake of the root, from which is extracted the black substance well known under the name of *liquorice*, or *black sugar*. This root was subjected to analysis by M. Robiquet, who obtained the following substances from it:

(1.) Starch, first observed in it by M. Lautour.

(2.) Gluten, which is separated by boiling.

(3.) Liquorice sugar, a substance which approaches sarcocoll in its properties. It is soluble in alcohol and water; but insoluble in cold water. It has a sweet taste, does not yield alcohol by fermentation, nor is it converted into oxalic acid by the action of nitric acid.

(4.) Phosphates and malates of lime, and magnesia.

(5.) A brown thick resinous oil, which gives an acrid character to the decoction of liquorice.

(6.) A substance similar in its properties to asparagin; but crystallizing in octahedrons.

(7.) Woody fibre.†

Elicampane.

14. *Elicampane*.—The roots of the *inula helenium* or *elicampane* were examined by Rose, who extracted from them the peculiar vegetable principle called *inulin*. M. Funke afterwards subjected them to a new analysis, and obtained the following substances:

* Ann. de Chim. lxx. 95.

† Ibid. 185.

‡ Ibid. lxxii. 145.

- A crystallizable volatile oil.
- Inulin.
- Extractive.
- Acetic acid.
- A crystallizable resin.
- Gluten.
- A fibrous matter (woody fibre?)*

15. *Sweet flag*.—The roots of the *acorus calamus*, or *sweet flag*, have been subjected to analysis by Trommsdorf. According to him 64 ounces of this root are composed of the following ingredients :

	oz.	dr.	gr.
Volatile oil	0	.. 0	.. 13·33
Inulin	1	.. 0	.. 1
Extractive with some muriate of potash	2	.. 1	.. 10
Gum with some phosphate of potash ..	3	.. 4	.. 0
A viscid resin	1	.. 4	.. 0
Woody fibre	13	.. 6	.. 0
Water	42	.. 0	.. 35·61
	64	0	0 †

16. *Andropogon schænantus*.—The root of this plant from the isle of France was examined by M. Vauquelin. It was found to contain, 1. A resin similar to that of myrrh. 2. A bitter colouring matter soluble in water. 3. An acid. 4. Oxalate of lime. 5. Abundance of oxide of iron. 6. Much woody fibre. †

Many other roots have been employed either in medicine or the arts; but as they have not yet been subjected to chemical analysis, I think it unnecessary to enumerate them here.

II. The *wood* of different trees differs materially in hardness, strength, durability, and beauty. But from the experiments of Count Rumford, there is reason to believe that the mere woody fibres of all plants are nearly the same, and that the differences are owing almost entirely to the various proportions of liquids and empty spaces with which the woody fibres are intermixed. He found the specific gravity of the wood of different trees as follows :

* Ann. de Chim. lxxvi. 98.

† Ibid. lxxxii. 332.

‡ Ibid. lxxii. 302.

Book IV.
 Sp. gravity
 of woods.

Poplar	1·4854
Lime	1·4846
Birch	1·4848
Fir	1·4621
Maple	1·4599
Beech	1·5284
Elm	1·5186
Oak	1·5344

A cubic inch of oak and poplar consisted respectively of the following proportions of wood, sap, and air :

	Wood.	Sap.	Air.
Oak	0·39353 0·36122 0·24525
Poplar	0·24289 0·21880 0·53831

He found likewise that the same tree in winter contains more sap than in summer; and that in summer it contains more air than in winter.*

Out of the great number of woods known and employed in the arts, I shall mention only the following.

Logwood.

1. *Logwood*.—This is the wood of the *hæmatoxylon campechianum*, a low prickly tree, which grows abundantly about Campeachy, in the bay of Honduras. It comes to Europe in large logs, cleared from the bark, and is very hard, compact, heavy, and of a red colour. It has a sweetish astringent taste, but no perceptible smell. It is very much used by dyers, especially in dyeing black, to which indeed it is in some degree essential. For the facts known respecting its properties, we are chiefly indebted to Dr. Lewis.† Several of its constituents have lately been ascertained by Chevreul.‡

Properties.

Logwood gives out its colouring matter both to water and alcohol; but to neither with much facility. The liquors assume a fine red colour with a shade of purple. The infusion becomes gradually deeper, and at last almost black. To cloth, previously boiled in alum and tartar, it gives a very beautiful, but perishable violet colour. Alkalies deepen the colour of the infusion; acids change it to yellow. With sulphuric, nitric, muriatic, and acetic acids, slight reddish-brown precipitates fall. Alum throws down a copious violet precipitate. Sulphate of iron strikes a dark bluish-black like ink, and throws down slowly a precipitate of the same colour.

* Nicholson's Jour. xxxiv. 319.

† Neuman's Chem. p. 385.

‡ Ann. de Chim. lxvi. 254; and lxxxii. 128.

Sulphate of copper throws down a copious brownish-black precipitate, and acetate of lead an equally copious reddish-black precipitate; and muriate of tin throws down instantly a copious purple lake, while the liquid becomes transparent and colourless.* Chevreul has shown that this wood contains a volatile oil, tannin, a yellow colouring matter, and acetates of lime and of potash.† But the most important of its constituents is the substance to which Chevreul has given the name of hematin, and which has been described in the preceding chapter.

2. *Brazil wood*.—This is the wood of the *cæsalpinia crista*, Brazil wood. a tree which grows in Brazil, and likewise in other countries, where it is cultivated for the dyer. The tree is very large, and the wood very hard and heavy. Its colour is at first whitish, but it becomes red by exposure to the air. Water in which it is boiled acquires a fine red colour; while the wood itself, by repeated decoctions, becomes black. After water has ceased to act, alkaline leys still deprive it of a considerable quantity of colouring matter. Alcohol likewise takes up the colouring matter of Brazil wood, and acquires a deeper colour than water.

From the decoction in water, sulphuric and muriatic acids throw down a scanty red precipitate, while the liquid becomes yellow. An additional dose of acid restores the red colour again. Nitric acid changes it to yellow, and then to orange. Oxalic acid throws down an orange-red precipitate. The alkalis change the colour to purple, and throw down a scanty precipitate. Alum precipitates slowly a copious reddish lake, and muriate of tin an abundant rose-coloured powder, while the liquid becomes nearly colourless. Sulphate of iron strikes a black with a tint of violet, and throws down a precipitate of the same colour. Acetate of lead throws down a fine dark red precipitate.‡ An elaborate set of experiments on Brazil wood has lately been published by Chevreul. The following are the results which he obtained:

The decoction of yellow Brazil wood is yellow. When Properties. distilled it yields water containing a portion of acetic acid, and a volatile oil having the odour and taste of pepper. As the liquid is concentrated by evaporation, a viscid matter is

* Berthollet, *Elemens de l'Art de la Teinture*, ii. 344.

† *Ann. de Chim.* lxxvi. 254.

‡ Berthollet, *Elemens de l'Art de la Teinture*, ii. 229.

Book IV.

perceptible, of a brown colour, which is a compound of tannin and the colouring matter. The liquid retains a portion of the same substances. It contains also traces of ammonia, and of lime probably combined with acetic acid. Chevreul suspects also the presence of sugar and of gallic acid; both of which he detected in the extract of Brazil wood of commerce. All his attempts to separate the colouring matter from the tannin, and obtain it in a separate state, failed. The acids change it from yellow to red in the infusions containing it. Alkalies and alkaline earths make it pass to violet. Alumina strikes a carmine, protoxide of tin a violet, and peroxide of tin a red colour. Several neutral salts, as sulphate of soda, &c. when added in considerable quantity, change the colour to red. Sulphuretted hydrogen gas combines with it, and destroys the colour; but it appears again, and the sulphuretted hydrogen is separated. It has a stronger affinity for alkalies than litmus: and accordingly it renders litmus paper red, by depriving it of the alkali with which it is in combination.*

Red saun-
ders.

3. *Red saunders*.—This is the wood of the *pterocarpus santalinus*, a large tree which grows on the Coromandel coast, and in other parts of India. The wood is at first light-red, but becomes very dark by exposure to the air. It is very compact and ponderous, and has no smell, and very little taste. This wood is chiefly used for colouring drugs, and has not yet been subjected to chemical examination. The colouring matter is of a resinous nature, and probably similar to that of dragon's blood. It is taken up by alcohol, but not by water. The alcoholic tincture is a fine red, but becomes yellow when diluted with a large quantity of spirits.† By this property of yielding no colouring matter to water, we can easily distinguish red saunders from Brazil wood; a distinction first pointed out by Dr. Lewis. ‡

Fustic.

4. *Fustic*.—This is the wood of the *morus tinctoria*, a large tree which grows in the West India islands. It is used in dyeing yellow, and is very rich in colouring matter. It is neither very hard nor heavy; its colour is yellow with orange veins. It gives a deep orange-colour to water. The decoction gives a precipitate with glue, but still retains its colour. Acids throw down a slight greenish-yellow precipitate, which is re-dissolved by alkalies. These give the decoction a dark-

* Ann. de Chim. lxxvi. 225.

† Neumann's Chem. p. 337.

‡ Ibid. p. 336.

red colour, and gradually separate a yellowish matter. Alum throws down a scanty yellow precipitate; sulphate of iron, a precipitate at first yellow, but which becomes brown; sulphate of copper, a yellowish-brown precipitate; acetate of lead, an orange-yellow precipitate; and muriate of tin, a very copious fine yellow precipitate.*

Chap. II.

5. *Sumach*.—This is the shoots of the *rhus coriaria*, a shrub which grows in the countries bordering on the Levant. They are cut down yearly, dried, and ground in a mill. Like the preceding substances, it is used in dyeing, and chiefly as a mordant, on account of the tannin which it contains. It gives to water a greenish-yellow colour, which soon becomes brown by exposure to the air. It has not yet been subjected to a regular chemical analysis. The tannin appears to be the principal ingredient; and, according to Proust, it differs in its properties from the tannin contained in other plants. Bartholdi published a set of experiments on the decoction of this plant. According to him, it contains a great variety of saline matters; nitrate of potash, muriate of soda, sulphate of lime, gallate of magnesia, carbonate of lime, &c. But these experiments stand in need of repetition.†

Sumach.

SECT. V.

OF BARKS.

THE bark is the outermost part of vegetables. It covers the whole plant from the extremity of the roots to the extremity of the branches. It is usually of a green colour: if a branch of a tree be cut across, the bark is easily distinguished from the rest of the branch by this colour. If we inspect such a horizontal section with attention, we shall perceive that the bark itself is composed of three distinct bodies, which, with a little care, may be separated from each other. The outermost of these bodies is called the *epidermis*, the middlemost is called the *parenchyma*, and the innermost, or that next the wood, is called the *cortical layers*.

Bark.
Consists of
three sub-
stances.

The *epidermis* is a thin transparent membrane, which covers all the outside of the bark. It is pretty tough. When inspected with a microscope, it appears to be composed of a number of slender fibres crossing each other, and forming a

1. Epider-
mis.

* Berthollet, *Elements de l'Art de la Teinture*, ii. 259.

† *Ann. de Chim.* xii. 305.

Book IV.

kind of net-work. It seems even to consist of different thin retiform membranes, adhering closely together. This, at least, is the case with the epidermis of the birch, which Mr. Duhamel separated into six layers. The epidermis, when rubbed off, is reproduced. In old trees it cracks and decays, and new epidermes are successively formed. This is the reason that the trunks of many old trees have a rough surface.

2. The parenchyma.

The *parenchyma* lies immediately below the epidermis; it is of a deep green colour, very tender, and succulent. When viewed with a microscope, it seems to be composed of fibres which cross each other in every direction, like the fibres which compose a hat. Both in it and the epidermis there are numberless interstices, which have been compared to so many small bladders.

3. The cortical layers.

The *cortical layers* form the innermost part of the bark, or that which is next to the wood. They consist of several thin membranes, lying the one above the other; and their number appears to increase with the age of the plant. Each of these layers is composed of longitudinal fibres, which separate and approach each other alternately, so as to form a kind of network. The meshes of this network correspond in each of the layers; and they become smaller and smaller in every layer as it approaches the wood. These meshes are filled with a green-coloured cellular substance, which has been compared by anatomists to a number of bladders adhering together, and communicating with each other.

Nature of barks.]

Fourcroy supposes that the epidermis is the same in its nature in all trees, and that it possesses constantly the properties of suber; but this opinion is not likely to be verified. The cortical layers seem, at least in many cases, to have a similar fibrous basis; a basis possessing essentially the properties of flax, which is itself merely the cortical layers of the *linum usitatissimum*. The matter of the parenchyma, and the juices which exist in barks, vary extremely, and probably occasion most of the differences between them. Some, as oak bark, are characterized by their astringency, and contain tannin; others, as cinnamon, are aromatic, and contain an essential oil; others are bitter, as Jesuit's bark; some are chiefly mucilaginous, others resinous, &c. But in the present state of the subject, an enumeration of the different kinds of barks is not to be expected. I shall therefore, as in the preceding sections, satisfy myself with detailing the properties of those

barks that have been subjected to examination. The following are the most remarkable of these: Chap. II.

1. *Bark of the cinchona floribunda, or quinquina of St. Domingo.*—This bark has been analyzed with great address by Fourcroy. His analysis still constitutes one of the completest which has hitherto been made on any of the substances belonging to the vegetable kingdom. Quinquina
of St. Do-
mingo.

This bark is in rolled pieces, six or seven inches long, and three or four lines in thickness. Its colour is greyish-green externally, but within it exhibits different shades of green, purple, white, brown, &c. Its taste is bitter and disagreeable; its odour strong and unpleasant.

It gives out nearly half its weight to water, provided it be boiled in a sufficient quantity of that liquid. The residue possesses the properties of woody fibre. When burnt it leaves a coal, which contains the following saline constituents. The quantities are those obtained from a pound, or 9216 grains of the bark.

	Grains.	
Carbonate of potash	10	Earthy re- sidue.
Sulphate of potash	12	
Muriate of potash	38	
Phosphate of lime	20	
Carbonate of lime	420	
Silica	4	

The decoction of the bark has a reddish-brown colour, and an extremely bitter taste. It deposits, on cooling, a blackish substance, soft and tenacious, which does not dissolve in cold water, though it is soluble in hot water and in alcohol. More of this substance precipitates as the liquor is evaporated. When the inspissated juice, freed from this precipitate, is mixed with alcohol, a quantity of gummy matter separates. When the black matter which precipitates as the decoction cools is treated with hot alcohol, the greatest part of it is dissolved; but a fine red powder remains mixed with some mucilage, which is easily separated by water. When the alcoholic solution is exposed to the air, it deposits light-yellowish crystals of a saline nature. When mixed with water, white flakes are thrown down, which possess the properties of gluten; but the greatest part remains in solution. Thus the soluble part of the bark may be separated into five distinct substances; namely, *gum, gluten, a red powder, a*

Book IV. *saline matter*, and a *brownish bitter substance*, retained in solution by the diluted alcohol. The last is by far the most abundant. To it the peculiar qualities of the decoction of this bark are to be ascribed. The following are the quantities obtained by Fourcroy from 9216 grains of the bark :

	Grains.
Gum	648
Gluten	84
Red powder	144
Saline crystals	72
Brownish matter	4076
Total	5024

Properties. The gummy matter and gluten possessed very nearly the properties of these substances, as described in the last chapter of this work. The red powder is insoluble in water and alcohol, but dissolves readily in alkalies. When distilled it yields much oil and ammonia, and little water. According to Fourcroy, it is *extractive* saturated with oxygen. The nature of the saline matter was not determined in a satisfactory manner. The brownish matter which remained dissolved in the diluted alcohol, being by far the most abundant and important ingredient, was examined with particular care. From the experiments of Fourcroy, it appears to approach nearest to *extractive* in its properties.

Its colour is reddish-brown, and its taste very bitter. When dry it appears black, is very brittle, and breaks with a vitreous fracture. It is insoluble in cold water, but dissolves very readily in that liquid when heated, precipitating, however, in part, as the water cools. But if one part of it be dissolved in 24 parts of hot water, no portion of it precipitates. If lime water be dropped into this solution, it throws down a reddish powder, redissolved by the addition of a sufficient quantity of water. Muriatic acid precipitates nothing from this solution. The brown matter dissolves readily in alcohol, if it be previously reduced to powder. When distilled it gives out carbonic acid and inflammable air, water impregnated with pyromucous acid, and a quantity of oil. The watery portion contains likewise ammonia. The charry residue is very bulky, ten times as much so as the original matter.

When chlorine gas is passed through this extractive substance dissolved in water, it becomes immediately reddish, and then a precipitate of the same colour gradually falls. This

precipitate possesses exactly the properties of the *red powder* obtained from the decoction of the bark; a proof that this red powder is merely the brown extractive altered in its properties by the absorption of oxygen. If the current of chlorine gas be continued, the whole liquor gradually loses its colour, and the precipitate also becomes a pale yellow. If it be separated in this state by filtration, it assumes the appearance of gamboge. Alcohol dissolves it, but not a mixture of alcohol and water.*

2. *Bark of cinchona officinalis*.—This tree grows in Quito; it is confined to the high grounds, and when stripped of its bark soon dies. There are three different kinds of bark to be found in commerce, but whether they be all obtained from the same trees is not known; the contrary is probable. The following are the most remarkable of these varieties.

Red Peruvian bark.—This bark is usually in large pieces, and is reducible to powder with more ease than the preceding. Its powder is reddish-brown, and has a slightly bitter taste, with a good deal of astringency. We are indebted to Fourcroy likewise for the analysis of this bark.

Water in which the powdered bark is macerated, acquires the property of reddening vegetable blues, and contains a portion of citric acid, some muriate of ammonia, and muriate of lime. When 576 parts of the bark are boiled repeatedly in water till that liquid ceases to have any action on it, the water acquires a reddish orange colour, and when evaporated to dryness leaves a residue which amounts to 38 parts. This contains a small portion of the acid and salts just mentioned, but consists chiefly of a matter which possesses nearly the properties of extractive. When the residual bark is treated with a sufficient quantity of alcohol, that liquid takes up 24 parts of a substance which has a red colour, and appears intermediate between resin and extractive. The residue of the bark possesses the properties of woody fibre. When burnt it leaves twelve parts of a grey ash, consisting of

Carbonate of potash	1·9
Muriate of potash	0·6
Sulphate of potash	0·5
Carbonate of lime	9·0
	<hr/>
	12·0

* Fourcroy, Ann. de Chim. viii. 120.

Book IV.

Such is the result of Fourcroy's analysis.* Subsequent experiments have detected a portion of resin and of bitter principle. The presence of tannin in small quantity was indicated by the property which the decoction showed of striking a black with sulphate of iron. The bitter principle seems to belong to the second species, or that which was first observed in coffee.

2. Yellow.

Yellow Peruvian bark.—This species of bark, first brought into use in this country about the year 1790, has not yet been subjected to a rigorous analysis; but its constituents do not appear, from the trials which have been made, to differ much from those of the red species. M. Dechamps has published a method of extracting from it a considerable quantity of an undescribed salt with base of lime, which Vauquelin has shown to be kinate of lime. It is white, and crystallizes in plates. The process consists in macerating the bark in a sufficient quantity of cold water, in concentrating the liquid, and setting it aside to spontaneous evaporation. The crystals gradually separate, and are purified by repeated crystallizations. The quantity procured amounts to about seven parts for every hundred of bark employed.†

3. Pale,

Pale Peruvian bark.—This is the common variety of the bark. It has not yet been subjected to a correct chemical analysis. Its taste is astringent and bitter, and very disagreeable. It is supposed to contain bitter principle, tannin, extractive, and resin. Besides these, it contains a principle first pointed out by Seguin, and upon which Dr. Duncan junior published some experiments.‡ It is distinguished by the property of precipitating infusion of galls; but as this property is common to a considerable number of substances, it is not sufficient alone to characterize it.

3. *Bark of cinchona Carribæa.*—This bark was first made known by Dr. Wright, who published a botanical description of the tree, with a figure, in the Philosophical Transactions; § and an account of the medicinal properties of the bark in the London Medical Journal for 1787. || A description of a tree to which the same name is given, together with a chemical analysis of the bark, was published in the Journal de Physique for 1790** by M. Vavasseur; but it is not quite certain that the plants are the same.

* Ann. de Chim. viii. 174.

† Nicholson's Jour. vii. 226.

|| Vol. viii. 239.

† Ibid. xlvi. 65.

§ Vol. lxvii. p. 504.

** Vol. xxxvii. p. 241.

4. Vauquelin has examined all the different species of cinchona, or Peruvian bark, which he could procure, in order to determine, if possible, how far they differ from each other, and what the constituents are that act as febrifuge.*

They may be divided into three sets:

Species of bark.

I. The infusions of the first set precipitate the infusion of nutgalls, but not that of gluc.

II. The infusions of the second set precipitate gluc, but not the infusions of nutgalls.

III. The infusions of the third set precipitate at once glue, nutgalls, and tartar emetic.

The following table exhibits the effects of the different reagents upon all the barks tried. As Vauquelin has not given the botanical name in all cases, it is difficult, notwithstanding his descriptions, to ascertain the real names of the different species tried. Indeed, there is reason to believe that the same species occur in different parts of the table under different names:

Barks.	Precipitate by glue.	Precipitate by tannin.	Ditto by tartar emetic.
Yellow bark	White.		Copious.
Quinquina of S ^{ta} Fé . .	Reddish.	Copious.	0
Grey quinquina	White.	Ditto.	White.
Quinquina gris canelle .	Brown.	0	0
Red quinquina	Red.	0	Yellowish white.
Grey quinquina	White.	Yellow.	Ditto.
Quinquina gris plate . .	0		0
Cinchona pubescens . .	0	Yellow.	Yellowish white.
Cinchona officinalis . .	0	Ditto.	0
Cinchona magnifolia . .	Copious.	0	0
Quinquina pitton vrai . .	0	Copious.	Copious.

Barks brought from Peru by Humboldt.

Quinquina of Loxa . .	Copious.	Copious.	Copious.
Quin. white of S ^{ta} Fé . .	0	0	0
Quin. yellow of S ^{ta} Fé . .	0	Copious.	Copious.
Quin. red of S ^{ta} Fé . .	Copious.	0	0
Q. yellow of Cuença . .	0	0	0
Quinquina, ordinary . .	Copious.	Copious.	Copious.
Infusions of nutgalls . .	Copious.		Yellow white.
Oak bark	Ditto.		0
Cherry-tree bark	0	0	0

All the barks strike a green with iron, and the most of them precipitate it of that colour.

* Ann. de Chim. lix. 113.

Book IV.

The substance in these barks which precipitates tannin has a brown colour and a bitter taste; it is less soluble in water than in alcohol. It precipitates tartar emetic, but not glue. It has some resemblance in its properties to the resins, though it gives out ammonia when distilled.

Croton.

5. *Bark of croton eleutheria*.—This bark constitutes the cleutheria and cascarilla of the shops. It is usually in rolled pieces, and has some resemblance to Peruvian bark. It has an aromatic smell, and an agreeable bitter taste. When burnt it emits an aromatic odour resembling that of musk. For the only chemical analysis of this bark hitherto published we are indebted to Professor Trommsdorf; 4696 parts of it yielded him the following products.*

Mucilage and bitter principle ..	864 parts
Resin	688
Volatile oil	72
Water	48
Woody fibres	3024
	4696

White willow.

6. *Bark of the white willow (salix alba)*.—The bark of this tree, which is common enough in Scotland, is remarkable for its astringent taste, and has been often used in intermittents by the common people. It has lately been proposed by Bouillon La Grange as an excellent substitute for Peruvian bark; being composed, according to him, of the very same constituents to which that bark owes its medicinal virtues. A very superficial examination, however, may satisfy any one, that the properties of the two are very far from similar.

The decoction of this bark has a dark-reddish colour; and, according to Bouillon La Grange, when repeated decoctions are made with the same portion of bark, the last is always deepest coloured. It is precipitated abundantly by glue, carbonate of potash, and carbonate of ammonia. Lime-water throws down a precipitate, at first blue, and afterwards buff-coloured. Sulphate of iron throws down a very dark-green precipitate. Alcohol separates white flakes. When evaporated to dryness, a reddish brittle substance remains, which has a very bitter taste, and does not attract moisture from the atmosphere.

When alcohol is digested over this bark, it acquires a green-

* Ann. de Chim. xxii. 219.

ish-yellow colour. The tincture is rendered muddy by water. When evaporated, it leaves a brilliant yellow substance of a very bitter taste, which melts at a moderate heat, and emits an aromatic odour.*

These experiments indicate the presence of tannin, bitter principle, extractive, and gluten, in this bark.

7. *Bark of quercus nigra*.—This tree, to which the name of *Quercitron*. *quercitron* has been given, grows spontaneously in North America. Dr. Bancroft discovered, about the year 1784, that the inner bark of this tree contains a great quantity of colouring matter; and since that time, it has been very generally used by the dyers. To prepare it for them, the epidermis (which contains a brown colouring matter) is shaved off, and then the bark is ground in a mill. It separates partly into stringy filaments, and partly into a fine light powder.

Quercitron bark readily gives out its colouring matter to water heated to the temperature of 100°. At that heat water dissolves about $\frac{1}{2}$ th of the bark employed. The infusion has a yellowish-brown colour, and is rendered lighter by acids, and darker by alkalies. Alum occasions a scanty precipitate of a deep-yellow colour; muriate of tin a copious bright-yellow precipitate; sulphate of iron a dark-olive precipitate; and sulphate of copper a precipitate of a yellow colour, inclining to olive.† From these experiments it is obvious, that the colouring matter of this bark is a species of extractive. The other ingredients have not been examined.

8. *Bark of the æsculus hippocastanum*.—The bark of this *Horse-chestnut*. tree, which forms so fine an ornament in parks and avenues, has lately been proposed in Paris as a remedy in intermittents, and a set of observations on it have been published by Mr. Henry.‡ Water or moderately strong spirits are the best solvents of it. Strong alcohol dissolves but little. The watery infusion has a fawn colour and a bitter taste, without astringency. It is precipitated abundantly by gelatine, and sparingly by acids. It precipitates iron green. With nitrate of mercury it forms a copious precipitate. The infusion of nutgalls and tartar emetic produce no effect upon it. From these properties it is obvious that it differs entirely in its constituents from the different species of cinchonæ examined by Vauquelin.

* Ann. de Chim. liv. 290.

† Berthollet, Elements de l'Art de la Teinture, ii. 262.

‡ Ann. de Chim. lxvii. 205.

Book IV.
Cinnamon.

9. *Cinnamon*.—This is the inner bark of the *laurus cinnamomum*, a tree which grows chiefly in Ceylon, where it is cultivated for the bark.

This has a reddish-brown colour, and consists of long rolled pieces which break into splinters. It has a very pleasant aromatic smell, and a pungent but agreeable taste. Its properties are owing entirely to the volatile oil which it contains. This oil may be separated by infusing the bark in alcohol, and then separating the alcohol from the oil by distillation. It has a whitish-yellow colour, and an extremely pungent taste and smell. The quantity yielded by cinnamon is very small. From 16 ounces of bark, Neumann obtained only two scruples and a half of oil. When water is distilled off this bark, it comes over milky, from the accompanying oil, which it retains with great obstinacy; very little separating till the mixture has stood a very considerable time.*

10. *Anthemis pyrethrum*, or *Pellitory of Spain*. This plant when chewed excites a burning sensation in the mouth, and occasions a copious flow of saliva. It was examined by M. Gautier, with a view to ascertain the nature of the substance which occasions this peculiar sensation. He found it to be a peculiar fixed oil residing in the bark, and which on a close inspection may be seen lodged in it in minute vesicles. This oil has a reddish colour, a strong smell, is insoluble in water, solid while cold, but melting when heated. The following are the substances which Gautier extracted from the pyrethrum:

Volatile oil	Trace
Fixed oil	5
Yellow colouring matter	14
Gum	11
Inulin	33
Muriate of lime	Trace
Lignin	35

98†

* Neumann's Chemistry, p. 402.

† Annals of Philosophy, xiii. 69.

SECT. VI.

OF LEAVES.

THE leaves of plants bear a greater resemblance to each other than any of the parts already treated of. They are covered only with an epidermis; the cortical layers of the bark being wanting. Below the epidermis is a pulpy matter, usually green, which contains a resinous substance and a portion of gluten. Below this is the fibrous matter, or rather the vessels which form the outline of the leaf. Alcohol extracts the green colouring matter of leaves, and so do oils but not water. It is remarkable, that even those leaves which are not green still give a green tincture to alcohol. The green colour soon fades when exposed to air, either in solution, or separated by evaporation.* Many leaves give a yellow colour to cloth, which may be fixed by means of alum. Hence the much greater number of yellow dyes than of any other colour. The leaves of woad (*reseda luteola*) are those which give the finest yellow, and they are much employed by the dyers on that account.

Chap. II.

Properties.

No extensive series of chemical experiments having hitherto been made upon leaves in general, I must satisfy myself with noticing the few specimens which have been examined by chemists. The following are the most remarkable of these.

1. *Senna*.—This drug, much used in medicine as a mild purgative, is composed of the leaves of *cassia senna*, a small shrub which grows chiefly in Egypt, from which country it is imported into Europe. It has been chemically examined by Bouillon La Grange.

The leaves are lanceolate, of a yellowish-green colour; an aromatic and somewhat disagreeable odour, and a bitter taste. Steeped in water, senna communicates to that liquid its peculiar smell and taste, and a reddish-brown colour. The strong mineral acids, oxalic acid, and the alkaline carbonates, throw down a precipitate; so do barytes water, and the nitrates of mercury, silver, and lead. The infusion of galls strikes a yellowish-buff colour, and throws down a slight precipitate. Alcohol throws down a precipitate, which, if we believe Bouillon La Grange, is composed of about 5 parts of carbonate of lime and 1 part of silica.† The residual liquor

* Lewis, Neumann's Chem. p. 434.

† I do not clearly understand this part of the analysis. The precipitate

Book IV. contains extractive, with some salts, whose bases are potash and magnesia.

When the infusion is exposed to the air, a substance precipitates from it. A precipitate is obtained also by chlorine. This precipitate is of a lemon-yellow colour, insoluble in water, has a bitter taste, dissolves in alcohol and alkalies, and, in short, possesses the characters of resin.

The chief matter taken up from senna by water is a peculiar extractive, which is converted into a species of resin by the absorption of oxygen. To this substance the qualities of senna are owing. If we believe Bouillon La Grange, its activity depends upon its absorbing oxygen.

A hundred parts of water boiled on senna are sufficient to deprive it of all the soluble matter which water is capable of extracting. Alkali separates from the residue a portion of matter similar to the extractive, altered by exposure to the air. The leaves are now green, and give out a green matter to alcohol. This substance is insoluble in water and alkalies, and is whitened by chlorine, and rendered less soluble in alcohol. These properties indicate a peculiar substance. 244·6 parts of senna yield to these different menstrua 79·6 parts, or rather more than $\frac{1}{3}$ d of soluble matter. When the watery infusion of senna is evaporated to dryness, and the residue burnt, it yields potash, sulphate of potash, carbonate of lime, magnesia, and silica.*

Worm-
wood.

2. *Wormwood*.—This is the leaves, and likewise the stalks and flower-tops of the *artemisia absinthium*, a plant which grows wild in many parts of Britain. Wormwood has a strong smell, rather disagreeable, and an intensely bitter taste. Its smell is owing to a volatile oil which it contains. It contains, besides, a considerable portion of bitter principle. No late analysis of wormwood has been made. The last is that of Kunsmuller,† who obtained, by water, from 12 parts of wormwood, the following products:

Resin	0·48
Muriate of potash	0·12
Vegetable acid	0·50
Ditto combined with potash	2·14
	<hr/>
	3·24

is stated to be partly soluble in water and alcohol, and yet to be obtained by alcohol. These characters surely do not apply to carbonate of lime and silica.

* Ann. de Chim. xxvi. 3.

† Ibid. vi. 35.

The residue, after the action of water, being burnt, left an ash, which, from 12 ounces of wormwood, contained the following substances: Chap. II.

Muriate of potash	3 grains
Sulphate of potash . . .	1
Carbonate of lime	59
Alumina	5
Sulphate of lime	5
Silica	4
Oxide of iron	3

3. *Tobacco*.—This substance is formed from the leaves of Tobacco. the *nicotiana tabacum*, a plant which is a native of Virginia, where it is cultivated very extensively for the purposes of commerce. Its name was given it from Tabaco, a province of Mexico, where it was first observed, and whence it was originally brought to Europe.*

Tobacco, when properly cured, has a yellowish-green colour, a strong and rather unpleasant smell, and a very acrid taste. When burnt it emits a strong odour, which to many is peculiarly grateful. When swallowed, it acts very violently upon the system; and when taken in sufficient quantity, proves fatal. Vauquelin subjected the leaves of different species of *nicotiana*, particularly the *latifolia*, to analysis. The expressed juice contained the following substances:

A considerable quantity of vegetable albumen or gluten.

Supermalate of lime.

Acetic acid.

A notable quantity of nitrate and muriate of potash.

A red matter soluble in alcohol and water which swells considerably when heated.

Muriate of ammonia.

Nicotin.

Green fecula composed chiefly of gluten, green resin, and woody fibre.†

4. *Asparagus*.—This plant has been analysed by Robiquet. Asparagus. I place the result here rather than any where else, because the whole plant was employed, and the leaves constitute the principal part of it.

M. Robiquet reduced the asparagus to the state of a pulp,

* For a detailed account of the culture and management of tobacco, see Tatham's Essay on its culture and commerce, published in 1800

† Ann. de Chim. lxxi. 139.

Book IV.

and separated its juice by expression. A feculous matter was separated from this juice by filtration. Boiling alcohol dissolved a portion of this fecula, and left matter which possessed the properties of gluten. As the alcohol cooled, a portion of wax precipitated. The alcohol, by evaporation, left a green acrid matter, intermediate in its properties between resin and volatile oil.

The filtered juice had the appearance of whey, and reddened the infusion of litmus. When heated it deposited flakes, which were considered by Robiquet as albumen. When left for a long time to evaporate in the open air, a quantity of asparagin and of saccharine matter, having the appearance of manna, separated in crystals. The juice was inspissated, and treated with alcohol. A portion dissolved; the residue was dissolved in water.

The alcoholic solution, when treated with sulphuric acid, and distilled, yielded vinegar. When this solution was evaporated to dryness and burnt, the ashes consisted of potash, carbonate of lime, and phosphate of potash and of lime. Hence it was probable that the acetic acid was combined with potash and lime.

The watery solution reddened vegetable blues. The infusion of nutgalls throws down a precipitate from this solution. The residue contains a peculiar extractive matter and a colouring matter; and doubtless, also, asparagin and sugar.

Such is the result of Robiquet's very ingenious and interesting analysis of this substance.*

Deadly
nightshade.

5. *Atropa belladonna*, *deadly nightshade*.—The leaves of this plant have been examined by Melandri. He found in them oxalate of magnesia with excess of acid, oxalate of lime, muriate of potash, a soft green-coloured resin, mucus, extractive, and a substance which Melandri called animal extractive, no doubt analogous to gluten or albumen.† The leaves of this plant have been examined likewise by Vauquelin. He obtained from them the following constituents:

(1.) Vegetable albumen, or gluten.

(2.) A bitter nauseous substance, which constitutes the narcotic part of the plant.

(3.) Nitrate, muriate, sulphate, superoxalate, and acetate of potash.‡

* Ann. de Chim. lv. 152.

† Ibid. lxx. 222.

‡ Ibid. lxxii. 53.

6. *Rheum palmatum*, *rhubarb*.—The leaves of this plant have been examined by Bouillon La Grange, who found them to contain a considerable quantity of superoxalate of potash and of malic acid.* Chap. II.
Rhubarb.

7. *Isatis tinctoria*, *woad*.—The leaves of this plant have been subjected to analysis by Chevreul. The expressed juice, separated by filtration from the *green fecula*, with which it was mixed, deposited on standing a blue powder, which possessed the properties of *indigo*. When heated, white flocks separated, consisting no doubt of *gluten*. When distilled, the liquid that came over yielded traces of a *volatile oil*, of *ammonia*, and of *sulphur*. When distilled with sulphuric acid, the liquid that came over contained *acetic acid* and *prussic acid*. Evaporated to dryness, and digested in alcohol, it yielded to that liquid a quantity of matter described in the last Chapter of this Work, under the Section entitled *Of Indigo* by the name of *green matter*. It yielded also a yellow-coloured *extractive*, and the following salts; *nitrate of potash*, *muriate of ammonia*, *acetate of potash*, *muriate of potash*, and pure *acetic acid*. The residue thus deprived of the substances soluble in alcohol consisted of *gum*, *yellow extractive*, *supermalate of lime*, *sulphate of lime*, and a salt which Chevreul considered as *citrate of lime-and-magnesia*. The green fecula being examined, was found to consist of *green coloured resin*, *wax*, *indigo*, *gluten*, and *woody-fibre*. The dry matter of the leaves, from which the juice had been expressed, yielded *woody fibre*, *green resin*, *wax*, *indigo*, *nitre*, a *red-coloured matter*, *malate of lime*, and traces of some other salts.† Woad.

8. The leaves of the *indigofera anil*, examined by the same chemist, yielded the same products with the *isatis*, but the quantity of *indigo* obtained was much greater.‡ Indigofera
anil.

9. The leaves of the *agathophyllum ravenara* were found by Vauquelin to contain an oil similar to the oil of cloves, gum, and carbonate of lime.§ Agathophyl-
lum raven-
sara.

10. *Gratiola officinalis*.—This plant was found by Vauquelin to contain, 1. A brown gummy matter, 2. A resinous matter, differing from other resins in being soluble in hot water; but much more soluble in alcohol than in water, and having a very bitter taste. It is in this substance that the purgative quality of the plant resides. 3. A small quantity of Gratiola of-
ficinalis.

* Ann. de Chim. lxxvii. 91.

† Ibid. p. 305.

‡ Ibid. lxxviii. 284.

§ Ibid. lxxviii. 306.

Book IV. animal matter. 4. Common salt in considerable quantity, and another salt which is probably malate of potash. It contained likewise oxalate and phosphate of lime and oxide of iron probably united to phosphoric acid.*

Menianthes trifoliata. 11. *Menianthes trifoliata*.—This plant was analysed by Trommsdorf. The fresh plant consists of 75 parts of moisture and 25 of solid matter. The expressed juice, when boiled, lets fall green flocks composed of 75 parts of vegetable albumen, and 25 of green resin. The liquid contains disengaged malic acid, acetate of potash, and a peculiar matter which is precipitated by tannin, but not coagulated by heat, and which is soluble in alcohol. It contains likewise a bitter extractive matter, a brown gum, and a quantity of inulin.†

Conium oleracea. 12. *Conium oleracea*.—An elaborate analysis of this plant has been published by Schrader. One thousand grains of it yielded the following constituents:

Extractive	27·3
Gum	35·2
Resin	1·5
Vegetable albumen	3·1
Green fecula	8·0
	75·1‡

Brassica oleracea. 13. *Brassica oleracea*.—Schrader obtained from 1000 grains of this plant the following substances:

Extractive	23·4
Gum	28·9
Resin	0·5
Vegetable albumen	2·9
Green fecula	6·3
	62·0§

Chenopodium olidum. 14. *Chenopodium olidum*.—This plant contains uncombined ammonia, to which, in the opinion of M. M. Chevalier and Lasseigne, it owes its peculiar smell. The substances extracted from the plant by these chemists, were the following:

- Carbonate of ammonia.
- Albumen.
- Ozmazome.

* Ann. de Chim. lxxii. 191.

† Ibid. lxxx. 330.

‡ Schweigger's Jour. v. 19.

§ Ibid.

An aromatic resin.

A bitter matter.

Nitrate of potash in considerable quantity.

Acetate and phosphate of potash.

Tartrate of potash.

100 parts of the plant yielded $5\frac{1}{2}$ of potash.*

15. *Mesembryanthemum crystallinum*.—The juice of this plant is said to be always colder than the surrounding air. Dr. John assures us that on the 13th of September, when he was engaged in examining it in Berlin, the thermometer in the shade stood at $54\frac{1}{2}^{\circ}$, but when plunged into the sap of this plant it sunk to 41° . He found this sap composed of the following substances:

Mesembryanthemum, &c.

$\frac{1}{2}$ of its weight of water.

Lignin.

Extractive.

Gum.

Resin in minute quantity.

Albumen.

Green powder.

Nitrate of potash in considerable quantity.

Common salt.

Phosphate of lime.

Carbonate of lime.†

SECT. VII.

OF FLOWERS.

The flowers of vegetables, on account of the great variety and beauty of their colour, attracted at an early period the attention of dyers and painters; but, upon trial, few have been found capable of yielding serviceable materials for either of these artists. The blue-coloured flowers, in general, readily give out their colour to water; though there are some from which water gains only a reddish or purplish hue. Scarcely any give a blue colour to alcohol: some give a red colour to that liquid, and some no colour whatever.‡ The expressed

Blue flowers.

* Annals of Philosophy, xii. 231.

† Chemische Untersuchungen, ii. 7.

‡ Lewis, Neumann's Chem. p. 430.

Book IV.

juice of most of these flowers is blue. This colour is changed into red by all acids, and into green by all alkalies and alkaline earths.

The nature of the colouring matter of these flowers is still unknown. But whatever it be, it is of a very perishable nature; for the flowers cannot be dried without losing their colour almost entirely. The more rapidly they are dried, and the more completely the air is excluded, the more of their colour they retain: but even after that, they mostly lose it altogether, and become white or yellow. The expressed juice gradually becomes purple, and soon fades altogether. Even when reddened by acids, these juices soon become purple, and at last colourless. The green colour induced by alkalies gradually changes to a yellow, which has some permanence.*

Red.

The colour of the red flowers, in general, is scarcely more permanent than the blue. They all give out a red colour to water. The pale-red and purple flowers scarcely tinge alcohol; resembling in this respect the blue flowers, to which indeed, in point of colouring matter, they seem to approach very nearly. The flowers, which have a full red colour, give a deep-red tincture to alcohol; the pale-red are tinged green by alkalies, and their red heightened by acids, like the blue flowers; but this is not the case with the full red flowers.†

Yellow.

The colour of the yellow flowers, whether pale or deep, is much more durable. They, in general, dry very well, and give out their colouring matter both to water and alcohol. When wool or silk is previously impregnated with alum and tartar, it receives a durable yellow colour when boiled with yellow flowers. Some of them, indeed, appear to be used by dyers for this purpose. Neither acids nor alkalies destroy the yellow colour of flowers, though both alter its shade; the acids rendering it paler, and the alkalies deeper. Alum likewise heightens the colour, though not so much as alkalies. ‡

White.

Even white flowers are by no means destitute of colouring matter. Many of them strike a green with alkalies, though acids do not turn them red. Some of them seem to contain the same colouring matter as yellow flowers. Thus the petals of the *convolvulus sepium*, though white, give a deep-yellow or orange colour to water, which is rendered paler by acids, and heightened by alkalies and alum. The vapours of sul-

* Lewis Neumann's Chem. p. 431.

† Ibid.

‡ Ibid.

phurous acid do not alter white flowers, though they destroy the colour of all others.*

For these curious observations on the colours of flowers, we are indebted to Dr. Lewis. Most of them contain a peculiar volatile oil, to which they owe their smell. Their other constituents remain still unknown, as few of them have been subjected to chemical analysis. The following are the species which have hitherto attracted the attention of chemists or artists:

1. *Flowers of the carthamus tinctorius.*—This is an annual plant, cultivated chiefly in Egypt for the sake of the flowers. The petals are originally of a fine red, but lose much of their colour on drying. It is from this substance that the *rouge*, or red pigment, employed by the ladies, is extracted. When the flowers have come to maturity, they are pulled, their juice expressed, the residue washed repeatedly with salt water, and exposed upon mats to dry. Care is taken to cover them from the action of the sun during the day, and to expose them to the action of the evening dews, in order to prevent them from drying too fast. In this state they are brought to Europe under the name of *safflower*, and the colouring matter is extracted from them by different artists. The process has been kept secret.

Carthamus
tinctorius.

Beckmann published a very elaborate set of experiments on the flowers of the carthamus, in 1774.† Little addition was made to our knowledge of this substance after that period, till 1804, when Dufour published an excellent analysis of it; ‡ and his observations have been since confirmed by those of Marchais. §

The flowers of the carthamus contain two colouring mat-
ters: a yellow, which is soluble in water, and has hitherto been applied to no use; and a red, which is employed by the dyers, &c. and which constitutes the pigment called *rouge*. The yellow-colouring matter dissolves readily in water; but it is very difficult by that means to separate the whole of it. The method followed by Dufour was to put the carthamus in a linen cloth, and knead it continually between the fingers, while a streamlet of water dropped upon it. By continuing this process, and then steeping it in water, it assumed a fine

Analysis.

* Lewis, Neumann's Chem. p. 462.

† In the Gottingen Transactions for that year.

‡ Ann. de Chim. xlvi. 283.

§ Ibid. i. 73.

Book IV. red colour. The liquid which passed, after being filtered to separate some portion of the carthamus, and some impurities with which it was mixed, had a yellow colour. When heated to about 145° , a number of flakes formed in it, consisting of albumen, or rather, of gluten. These were separated by filtration, and the liquid evaporated to dryness. The extract obtained had a yellow colour and a strong taste. Water dissolved the whole of this extract, except a portion of brown matter, which possesses the properties of resin. The watery solution gives a red colour to vegetable blues, and is precipitated abundantly by infusion of galls, but not by glue. Chlorine destroys the yellow colour, and renders it white. When evaporated to dryness, and treated with alcohol, a portion dissolves, and a portion remains undissolved. The soluble portion is chiefly extractive: the insoluble yellow-colouring matter.

When the residual carthamus is digested for a sufficient time in alcohol, that liquid assumes a brick-red colour. When this alcoholic solution is sufficiently concentrated by evaporation, a matter separates in grains, which has the appearance of honey, and possesses properties analogous to those of wax. Scarcely any colouring matter could be procured from the alcoholic solution.

Dufour, after various unsuccessful attempts to obtain the red-colouring matter of carthamus in a separate state, at last succeeded by the following process, founded on the strong attraction between the red matter and cotton. Having deprived the carthamus as much as possible of its yellow-colouring matter, by means of water, he macerated it for an hour in a weak solution of carbonate of soda. He then decanted off this solution, put into it a quantity of cotton, and then dropped into it lemon-juice, till the liquid acquired a fine cherry-red colour. After standing 24 hours, the liquid lost its red colour, the whole colouring matter having united with the cotton and dyed it red. The cotton was taken out, and washed repeatedly in water, to separate any of the yellow-colouring matter which might have adhered. It was then put into a very diluted solution of carbonate of soda. This alkali separated the colouring matter from the cotton, dissolved it, and assumed a yellow colour. The cotton being removed, and lemon-juice dropped into the solution, a fine rose-coloured powder gradually separated, and at last precipitated. This was the red-colouring matter.

From this process, we see that the red-colouring matter has a stronger attraction for cotton than the yellow, and that by means of cotton the two pigments may be separated. We see, too, that it is soluble in alkaline carbonates, but precipitated by acids. The alkalies dissolve it also but alter its nature. The solution of it in the alkaline carbonates is yellow. It is insoluble in water. Alcohol dissolves it readily, and assumes a fine rose-red colour. When this tincture is heated, it assumes a shade of orange. It dissolves likewise in ether, but not well. The oils, both fixed and volatile, have no action on it. When distilled it yields very little water, scarcely any gases, a little oil, and a portion of charcoal equal to $\frac{1}{3}$ d of its original weight. When this charcoal is burnt, no sensible portion of ashes remains behind. From these properties, for which, as well as for almost all the other facts respecting carthamus, we are indebted to Dufour, it appears that the red-colouring matter of the carthamus differs from every other known vegetable substance.

When the carthamus has been deprived of its colouring matter, the residue consisted chiefly of woody fibre and some earthy matter. The following table exhibits the constituents of 1000 parts of carthamus, according to the experiments of Dufour.

Moisture	62	Constituents.
Sand and small particles of the plant	34	
Gluten	55	
Yellow colouring matter	268	
Extractive	42	
Resin	3	
Wax	9	
Red colouring matter	5	
Woody fibre	196	
Alumina and magnesia	5	
Red oxide of iron	2	
Sand	12	
Loss	7	
	—————	
	1000	

If we believe Marchais, a considerable portion of what was considered by Dufour as woody fibre is in reality albumen.

2. *Flowers of the arnica montana.*—These flowers have been employed in medicine, especially in Germany, where their virtues have been much celebrated. They have a yellow co-

Arnica montana.

Book IV.

lour, a bitter and sharp taste, and a very weak aromatic smell. When macerated in water, the liquid acquires a brown colour and a bitter taste. It reddens the infusion of litmus, does not form a precipitate with glue, and does not alter the solution of tartar emetic. With sulphate of iron it strikes a deep green, which passes into black on drying. The mineral acids occasion brown precipitates, but the alkaline carbonates produce no change in it.*

When the flowers are macerated in alcohol, the liquid acquires a yellow colour. It reddens vegetable blues, strikes a green with sulphate of iron, and becomes milky when mixed with water. When this tincture is distilled, the alcohol which comes over has a bitter taste, but does not alter vegetable blues; while the liquid which remains in the retort is very acid, precipitates with water, and strikes a green with sulphate of iron. When evaporated to dryness, a bitter acid powder remains, seemingly of a peculiar nature.†

When distilled, the flowers of the arnica yield a yellowish acid liquor, which strikes a black with sulphate of iron. This liquid is covered by a quantity of oil. When these flowers are burnt, the residual ashes contain potash, carbonate of potash, and muriate and sulphate of potash.‡

These flowers have been lately subjected to a chemical examination by MM. Chevallier and Lassaigne, who extracted from them the following substances:

A resin having the odour of arnica.

A bitter nauseous substance resembling the emetic substance from the *cytisus laburnum*.

Gallic acid.

A yellow colouring matter.

Albumen.

Gum.¶

Muriate and phosphate of potash.

Traces of sulphates.

Carbonate of lime.

Silica.§

Saffron.

3. *Saffron*.—This yellow pigment is composed of the pistils of the *crocus sativus*. The flowers of this plant are of a purplish-blue colour, but the upper part of the pistil is of a deep

* Bouillon La Grange, Ann. de Chim. lv. 38.

† Ibid. lv. 39.

‡ Ibid.

§ Annals of Philosophy, xvi. 90.

orange, and divided into three fleshy filaments. These are carefully picked out, dried, and pressed into cakes. Chap. II.

Saffron is in long broad filaments of a deep-red colour without any yellow; moderately dry, yet flexible and soft, and difficultly reduced to powder. It has a strong agreeable smell, affecting the eyes, and drawing tears from them, and a pungent, and somewhat bitter taste.*

Hitherto saffron has not been subjected to a correct chemical analysis. From the experiments of Neumann, it does not appear that any volatile oil can be procured from it by distillation. It is probable, however, that it owes its strong smell to such a principle, though in too small a quantity to be easily obtained separate. The colouring matter of saffron is equally soluble in alcohol and water; and is the substance called polychroite by Vogel.

SECT. VIII.

OF POLLEN.

THE antheræ of most vegetables, at the time that the flower is in perfection, is covered with a considerable quantity of a powdery matter, usually of a yellow colour, which, falling upon the pistil, is supposed to impregnate it. This matter is known by the name of *pollen*. In some plants, especially those in which the male and female flowers are separate, the quantity of it is so great, that it may easily be collected.

The first person who examined this powder was Dr. Lewis; Experiments of Lewis. but in his time chemistry had not made sufficient progress to enable him to make a complete analysis. He found that alcohol, in which it is macerated, acquires a bright yellow colour, and that water takes a duller yellow, while the undissolved portion is of a yellowish-white colour. The colour of these infusions is heightened by alkalies, and turned *red* by acids. When alkalies are dropped into the reddened solution, a deep yellow colour is restored.†

Fourcroy and Vauquelin have published a detailed analysis of the pollen of the *phœnix dactilifera*, or date tree, which may be considered as furnishing a pretty correct view of the properties of this class of vegetable substances. At least it is

* Neumann, p. 401.

† Neumann's Chem. p. 231.

Book IV. } probable, that the pollen of different vegetables does not differ nearly so much from each other as the other parts of plants.

Pollen of the
date tree.

The pollen of this tree is so abundant, that at the season of fructification, it surrounds the plants like thick dust. The specimen examined was brought from Egypt by Delisle. Its colour was sulphur-yellow, its taste was sourish and unpleasant, it reddened the infusion of litmus, and communicated to water a yellow colour, and a sensible degree of acidity. The watery infusion was precipitated by lime water and ammonia, by alcohol, and by some metallic salts.

Action of
water and

When pollen was washed in cold water, that liquid acquired a reddish-brown colour, a sourish taste and smell, not unlike that of beer. When evaporated, the liquid left a reddish-brown matter, which had the consistence and odour of melasses; its taste was acid, but disagreeable. Cold alcohol produced no effect upon this matter, but by the assistance of heat it dissolved a portion of it, and assumed a dark colour. The residue had more consistence, and was less coloured. It dissolved with facility in water, allowing at the same time a greyish bulky matter to separate. Thus the matter taken up by the watery infusion from pollen has been divided into three portions.

Alcohol.

The alcoholic solution being concentrated by evaporation to the consistence of an extract, had a red colour, the odour of baked apples, and a taste strongly acid, but disagreeable. It presented all the characters of malic acid. The grey matter, which refused to dissolve in water, possessed the properties of phosphate of lime.

The aqueous solution was ascertained to consist chiefly of phosphate of magnesia, malic acid, and an animal matter, which formed a copious precipitate with the infusion of nut-galls.

Insoluble
parts.

The pollen, thus deprived of its soluble matter, being exposed on filtering paper to the open air for a week, instead of drying assumed the form of a paste, and ran rapidly to putrefaction, emitting an odour similar to that of old cheese. When dried, this matter became semitransparent, and of a consistence resembling that of glue. Before being dried it mixed readily with water, and exhibited the characters of soap. The fixed alkalis showed the presence of ammonia in it. When distilled there came over a watery liquid, which gradually became coloured as the distillation advanced: some time after there came over a red coloured oil and carbonate of

ammonia, partly in crystals, and partly in solution. The liquid also contained a portion of the oil in the state of a soap. The charcoal left was voluminous, and difficult of incineration. When burnt it left a considerable portion of phosphate of lime.

Chap. II.

The pollen is partly soluble in muriatic acid. The nitric acid reduces it to a paste, azotic gas is disengaged, and afterwards, on the application of heat, nitrous gas and carbonic acid. A yellowish oil separated and swam upon the surface, and alkalis separated earthy phosphates and oxalates. When the liquid was evaporated to dryness, it left a yellowish-red matter, extremely bitter, tenacious and adhesive, and perfectly soluble in water. This matter, from the description of it, must have been a species of artificial tannin. Thus it appears, that by the action of nitric acid, the portion of the pollen which is insoluble in water is converted into oil and artificial tannin. The oily matter became solid on cooling, and possessed almost exactly the properties of tallow treated with nitric acid.

Action of acids.

Strong alkaline leys dissolve the pollen, and assume with it the properties of soap, while ammonia is disengaged.

When the pollen, mixed with water, is set aside to ferment, it disengages the odour of new cheese, and assumes the form of a tenacious mass. Its taste is extremely sharp; it contains abundance of ammonia, partly united to oil, and partly to malic acid. From all these phenomena, there is reason to believe that the insoluble part of the pollen approaches vegetable gluten in its properties; or rather, that it is intermediate between gluten and albumen.

Such are the properties and constituents of the pollen of the date tree, as far as they have been ascertained by Fourcroy and Vanquelin.* We may conclude from them, that it approaches very much to the nature of animal matter.

Professor Link has examined the pollen of the hazel nut. It differs a good deal from that of the date tree just described. He found in it tannin in great abundance, a resin, a good deal of gluten, and a little fibrin.† The pollen of the tulip, which has been carefully analyzed by Grotthus, is also very different from that of the date tree. He could neither detect in it phosphorus nor pollenin. From 26 grains of it he extracted

Pollen of the hazel.

* Ann. de Mus. d'Hist. Nat. i. 417.

† Ann. de Chim. lxxi. 292.

Book IV.
Pollen of
the tulip.

Vegetable albumen	20·25
Malate of lime with trace of malate of magnesia .	3·50
Malic acid	1·00
Malate of ammonia }	1·25
Colouring matter ..	
Saltpetre?	

26·00 *

Pollen of
the pinus
sylvestris.

The pollen of the pinus sylvestris or *Scotch fir*, according to the analysis of John, is composed of

Cerin	2·00
Oil	Trace
Brownish-yellow resin	3·75
Malates of potash, lime, and magnesia, with some } gum	5·00
Sugar, with a trace of extractive	
Albumen	4·00
Pollenin	77·25
Sulphate } Muriate } of potash } Phosphate }	3·00
Phosphate of lime	
Oxide of iron	
Malate of ammonia	Trace
Sulphur?	Trace

100·00

Pollen of
the pinus
abies.

The constituents of the pollen of the pinus abies he found as follows :

Cerin	2·25
Oil	Trace
Brownish-yellow resin	4·00
Malates of potash, lime, and magnesia, with some } gum	6·00
Sugar	
Extractive	Trace
Albumen	4·00
Pollenin	75·25
Sulphur?	Trace
Malate of ammonia } Sulphate } Muriate } of potash } Phosphate }	3·00
Phosphate of lime	
Oxide of iron	

100·00

* Schweigger's Jour. xi. 281.

The pollen of *juglans regia*, *cannabis sativa*, *zea mays*, and of tulips, was likewise examined by Dr. John, and found to contain nearly the same ingredients, though not in quite the same proportions.* Chap. II.

SECT. IX.

OF SEEDS.

THE seeds of vegetables, constituting one of the most important articles of food, and being the materials from which a variety of useful substances are procured, have been more subjected to chemical experiments than any of the other parts of plants.

They all contain a considerable portion of starch, though this matter differs somewhat in its properties in different seeds. Constitu-
ents. Gluten is likewise a common ingredient, but extremely various both in its quantity and in its appearance. Some seeds contain a saccharine matter, as those of the *lycopodium*; but the quantity is always small, and its properties are hardly ever the same as those of true sugar. Fixed oil is also a common ingredient. In the seeds of the grasses, the proportion of it, if any can be detected, is always small: the same remark applies to the seeds of the papilionaceous plants; but the seeds of the cruciform plants abound in a limpid and nearly insipid oil. In some seeds a volatile oil is present, but this is not common. Some possess a peculiar colouring matter; some possess a portion of bitter principle, as the seeds of the common broom and of the coffee plant.

Saussure junior has shown that wheat, and probably every other kind of corn, yields a quantity of phosphorus, if it be charred in a moderate heat, and the charcoal be afterwards violently heated in a porcelain retort. He has shown that the phosphorus, in this case, is owing to the decomposition of phosphate of potash, which exists in these seeds, and which is decomposed by the charcoal at a very high temperature. He has shown that phosphate of soda, and even phosphate of lime, may be decomposed in the same way.† Thus the observations of Margraff, and some of the earlier chemists, respecting the phosphorus which may be obtained from the seeds of plants have been verified.

* Chemische Untersuchungen, iv. 27.

† Ann. de Chim. lxx. 189.

Book IV.

I shall in this section enumerate, in the first place, the properties of those seeds which are known by the name of *corn*, and are used as articles of food; then the seeds of the papilionaceous plants; namely, peas, beans, vetches, &c. which are likewise employed as food. The seeds of the remaining tribes of plants have been hitherto but imperfectly examined; we can only therefore detail the experiments hitherto made, without any attempt at arrangement.

The most important genera of corn are, wheat, rye, oats, barley, rice.

Wheat.

1. Wheat, the *triticum hybernum* (and other species) of botanists, has been cultivated from time immemorial in Europe and the northern parts of Africa, and the seeds of it employed as one of the most important articles of food. Indeed wheat flour is the only substance known from which good loaf bread can be made. The seeds, when ripe, are ground to a fine powder, and by passing this powder through cloth sieves, of various degrees of fineness, it is separated into distinct portions. The fine flour constitutes the greatest portion; and the *bran*, which consists of the outer coat of the seed, and which is the coarsest of all, constitutes the next greatest portion.

Constituents.

Hitherto the flour of wheat only has been subjected to a chemical examination; while the bran, which may vary from one-third to one-fourth, has been neglected. The two most remarkable substances in wheat flour are the *starch*, which constitutes by far the greatest part of it, and the *gluten*, which is the next ingredient in point of quantity. It appears also, from my trials, to contain a small portion of *bitter principle*. Saccharine matter exists in it also, though in a very small proportion. Mucilage is mentioned as a constituent by Fourcroy and Vauquelin; and from their experiments, also, it appears that phosphate of lime is present, but that they did not detect any phosphate of potash.*

Water in which wheat flour has been macerated for six hours, has an opal colour, and becomes transparent very slowly. Its taste is sweetish, it has the odour of unripe grain, and does not alter vegetable blues. The infusion of nutgalls, chlorine, and other acids, occasion a precipitate in it. It becomes very rapidly sour. When heated, yellowish flakes are separated. When concentrated by evaporation, its taste be-

* Ann. de Mus. d'Hist. Nat. No. xxxvii. p. 3.

comes more perceptibly sweet. When inspissated sufficiently, it assumes a golden yellow colour; has a sweet taste, but is, at the same time, acid and acrid; yellowish flakes separate, and a crust of phosphate of lime coats the capsule in which the evaporation is performed. Alcohol coagulates it, and throws down a substance which possesses the properties of gluten. Hence it is evident, that the water owes its properties to the gluten which it holds in solution,* and probably also to a little bitter principle. It contained also some sugar and mucilage, and phosphate of lime.

2. Rye is the seed of the *secale cereale*, a plant cultivated in the northern parts of Europe, in considerable quantity, as an article of food. Bread made of it is much denser than wheaten bread, and has a brownish colour, and a peculiar sweetish taste, which to most persons is rather agreeable. We are indebted to Einhof for an elaborate analysis of rye-meal. He published the result of his experiments in the summer of 1805.†

A determinate portion of rye-meal was formed into a paste, and washed in pure water till that liquid ceased to acquire any colour, or take up any thing. The liquid, after filtration, had a yellowish colour, an insipid taste, and a smell like that of new whey. It slightly reddened litmus paper, and was rendered muddy by carbonate of potash, sulphate of silver, and infusion of galls. When raised to the boiling temperature, it became muddy, and a number of white flakes separated. These, when collected andedulcorated, had the appearance of curd. They dissolved in alkaline leys, were insoluble in alcohol, and possessed the properties of *vegetable albumen*. When the liquid was concentrated by evaporation, it deposited a small additional portion of albumen. This being separated, the liquid was evaporated to the consistence of honey, and digested in alcohol repeatedly till nothing more was taken up. The residue was greyish-white and insipid: water dissolved it slowly, and the solution, being evaporated, deposited more flakes of albumen. The residue had the appearance of a solution of gumi arabic, and when evaporated to dryness, left behind it a portion of gummy matter.

The alcoholic solution became muddy when mixed with water. The alcohol being drawn off by distillation, there remained in the retort an aqueous solution, of a wine-yellow

* Fourcroy and Vauquelin, *ibid.*

† Gehlen's *Jour.* v. 131.

Book IV. colour with large brownish flakes swimming in it. These flakes, when collected, were found to possess the properties of gluten. They formed a glutinous mass with cold water, the bulk of which contracted when the water was made to boil. Boiling alcohol dissolved it; but ether took up only the colouring matter. Alkalies dissolved it, and acids threw it down again from its solution.

The watery solution, thus freed from the gluten, being evaporated, left an extract of a wine-yellow colour. By repeated digestions in alcohol, dilutions with water, and evaporations, it was freed from a portion of gluten, which still adhered to it. Its taste was then sweet, but harsh; its colour wine-yellow; and it was soluble in water, alcohol, and ether. It was considered as the saccharine matter of rye.

Insoluble. The rye-meal, thus freed from the matters soluble in water, was mixed with a great quantity of water, and by repeatedly agitating the liquid, and decanting it off after standing some time, it was separated into two portions; namely, a greyish-coloured substance, which, being lightest, did not so soon subside, and a white powder which possessed the properties of starch. The grey-coloured substance, by repeated digestions in alcohol and water, was separated into three portions; namely, gluten, starch, and the coats of the rye-seeds. Such are the constituents of rye-meal, according to Einhof. The following are the proportions of these different substances, detected by this chemist, in the best rye-seeds and rye-meal. 3840 parts of good rye-seeds were composed of

Constitu-	Husk	930
ents.	Moisture	390
	Pure meal	2520
		3840

A hundred parts of good rye-meal contained

Albumen	3·27
Gluten, not dried	9·48
Mucilage	11·09
Starch	61·09
Saccharine matter	3·27
Husk	6·38
Loss	5·42
	100·00

But the proportion of these substances must vary extremely according to the soil, the climate, and the age of the rye. The gluten of rye differs in several particulars from that of wheat. It is less tenacious and more soluble. When it was allowed to ferment, Einhof perceived a strong smell of nitric acid, which is peculiar to this species of gluten. The starch of rye bears a striking resemblance to that of wheat. Like this last, it does not form a colourless solution with boiling water, and always precipitates at last, when the solution is left a sufficient time at rest.

Einhof did not examine the ashes of rye; but from the experiments of Schrader, we know that the same quantity, analyzed by Einhof, namely 3840, yielded the following fixed substances :*

Silica	3·90
Carbonate of lime	3·35
Carbonate of magnesia .	3·55
Alumina	0·35
Oxide of manganese ...	0·80
Oxide of iron	0·22
	12·17

3. Oats are the seeds of the *avena sativa*, a plant cultivated in considerable quantities as an article of food. No correct chemical analysis of this seed has yet been made. The husk is uncommonly thick, and constitutes a considerable portion of the corn. The proportion of starch is likewise considerable, though it is difficult to free it from another substance with which it is united.

4. Barley is the seed of the *hordeum vulgare*, a plant too well known to require any description. Great crops of it are reared annually, partly as an article of food, and partly as a material from which malt liquors and ardent spirits are drawn. This species of corn has been examined of late with considerable attention by chemists, partly in order to form correct conceptions, if possible, of the nature of the process of fermentation, and partly to ascertain the constituents of barley. Fourcroy and Vauquelin published several ingenious remarks and experiments on it in 1806,† and Einhof published a still

* Gehlen's Jour. iii. 525.

† Ann. de Mus. d'Hist. Nat. No. xxxvii. p. 5.

Book IV. more elaborate analysis about the commencement of the same year; having examined this grain in different stages of its growth, and after it was fully ripe.*

Unripe barley. When unripe barley-corns are triturated with water, the liquid acquires a milky colour. If this process be continued, adding fresh portions of water as long as the liquid passes off muddy, there remains only a green husky matter. When this matter is macerated a sufficient time in cold water, it acquires a greenish-grey colour, and when dry has the appearance of vegetable fibre. The water in which it was macerated, when boiled, deposits a few flakes of albumen, and when evaporated to dryness leaves a small portion of extractive.

The water with which the barley was triturated is at first milky, and gradually deposits a white powder; yet it does not become transparent, though allowed to stand a considerable time. When filtered, it passes through transparent, while a slimy substance of a greenish-grey colour remains upon the filter. This substance possesses the properties of gluten. When the solution, now transparent and of a yellowish colour, is boiled, it deposits flakes of albumen. It reddens litmus paper, and is strongly precipitated by lime-water, nitrate of lead, and sulphate of iron, indicating the presence of phosphoric salts.

The liquid being evaporated to the consistence of a syrup, and the residue treated with alcohol, the solution diluted with water, and the alcohol distilled off, to separate some gluten which still remained, a syrupy matter was obtained, having a sweet taste, which was considered as a saccharine matter of the barley. A portion refused to dissolve in alcohol. This portion was considered as extractive.

The white powder which precipitated from the water in which the barley had been originally triturated possessed the properties of starch.

Such are the constituents of unripe barley, according to the experiments of Einhof. The following are the proportions of each which he obtained from 2880 parts of unripe barley.†

* Gehlen's Jour. vi. 62.

† Ibid. 83.

Green husk, &c	478
Albumen, with phosphate of lime	13
Gluten	51
Saccharine matter	160
Extractive	76
Starch	420
Volatile matter	1500
Loss	182
	<hr/>
	2880

Chap. II.
 Constitu-
 ents.

When ripe barley is steeped in water a sufficient time, and then cautiously kneaded between the fingers in a cloth, every part of the grain is washed away except the husk, which by this process may be dried and weighed. Ripe barley.

When barley meal, previously made into a paste, is treated in the same way, a brownish residuum remains, consisting chiefly of the husk, though it contains also portions of starch and gluten which cannot well be separated. The water in which the meal has been washed gradually deposits a white powder, but does not become transparent though left at rest. It runs very soon into acidity. Indeed, if we believe Fourcroy and Vauquelin, barley often contains an acid; the water in which it has been macerated reddening the infusion of litmus: this acid is the acetic. The colour of this water is reddish-brown. It holds in solution a considerable portion of matter, which, according to Fourcroy and Vauquelin, consists chiefly of gluten, but which Einhof found to be of a more complicated nature, consisting of albumen, or rather gluten, mucilage, and saccharine matter. It contains in solution likewise a notable portion of phosphate of lime.

When barley meal is macerated a sufficient time in alcohol, that liquid acquires a yellow colour, and becomes muddy and more odorous, and by evaporation leaves an oily matter of a yellow colour and an acrid taste, having the consistence of butter. This oil burns like a fat oil, and forms soap with alkalies. It is but imperfectly soluble in alcohol.* This oil escaped the observation of Einhof. I obtained it by a process similar to that afterwards described by Fourcroy and Vauquelin; but its colour was asparagus green, and it did not burn with the same readiness as an oil. It has very much the appearance of olive oil coagulated, but its consistence is less, and

* Fourcroy and Vauquelin, Ann. de Mus. d'Hist. Nat. No. xxxvii. p. 8.

Book IV.

its colour is darker. To this oil the peculiar flavour of spirits from raw grain is ascribed at present. If this opinion be well founded, the oil must be dissipated or destroyed by the process of malting.

Constituents.

The following are the proportions of the constituents obtained by Einhof from 3840 parts of barley-corns:

Volatile matter	430
Husk	720
Meal	2690
		<hr/>
		3840

From the same quantity of barley-meal he obtained

Volatile matter	360
Albumen	44
Saccharine matter	200
Mucilage	176
Phosphate of lime, with some albumen		9
Gluten	135
Husk, with some gluten and starch	..	260
Starch, not quite free from gluten	..	2580
Loss	76
		<hr/>
		3840

Besides these substances, Fourcroy and Vauquelin ascertained the presence of phosphates of lime and magnesia, and of silica and iron; and I found in it, besides these substances, phosphate of potash and nitrate of soda. The presence of phosphate of potash was ascertained likewise by Saussure junior.

Rice.

5. Rice, the seeds of the *oryza sativa*, have been analysed by Braconnot. He found in this grain the following constituents:

	Carolina Rice.	Piedmont Rice.
Water	5.00	7.00
Starch	85.07	83.80
Parenchyma	4.80	4.80
Vegeto-animal matter	3.60	3.60
Uncrystallizable sugar	0.29	0.05
Gummy matter approaching starch	0.71	0.10
Oil	0.13	0.25
Phosphate of lime	0.13	0.40
	<hr/>	<hr/>
	99.73	100.00

Besides traces of muriate of potash, phosphate of potash, acetic acid, sulphur and lime and potash united to a vegetable alkali.*

Vauquelin was unable to detect any saccharine matter in rice. He found in it a little phosphate of lime and gum.†

6. Maize. This is the seed of the *zea mais*, or Indian corn. Maize. Some experiments on it have been made by Proust. It is converted into charcoal in the usual way, and leaves about $\frac{1}{4}$ th its weight of charcoal. This charcoal is extremely difficult of incineration. It contains some phosphate of potash.‡

Having now given an account of the constituents of the various species of corn hitherto subjected to chemical examination, let us proceed to the seeds of the papilionaceous plants. The most remarkable of these are peas, beans, kidney-beans, lupines, lentiles, &c.

7. Peas. The seeds of the *pisum sativum* constitute a very Peas. common and nutritive article of food. They have been examined in different states by Einhof, § who has devoted his chief attention to this peculiar branch of chemical investigation.

By treating the green plant nearly in a similar manner with Green plant barley, he obtained from 3840 parts the following constituents:

Volatile matter	3000
Starch	53
Vegetable fibre	400
Gluten 	70
Albumen	35
Phosphate of lime ..	4
Saccharine matter. ..	176
Extractive	25
Loss	77
	3840

The green pod or husk of the pea, by a similar treatment, Pod. yielded, from 3840 parts,**

* Ann. de Chim. et Phys. iv. 370.

† Annals of Philosophy, xii. 151.

‡ Nicholson's Jour. xviii. 239.

§ Gehlen's Jour. vi. 115.

|| In the same state as that from the leaves of plants, and mixed with the green colouring matter.

** Gehlen's Jour. vi. 119.

Book IV.

Volatile matter	3120
Vegetable fibre	344
Starch	90
Green gluten	22
Albumen	17½
Phosphate of lime	3½
Saccharine syrup	182
Loss	61.
	<hr/>
	3840

Juice.

The peas themselves, when very young, are often filled with a sweet juice, which may be obtained by a slight pressure. Einhof examined this liquid. It has a greenish-yellow colour and a very sweet taste. When exposed to the air, thin cuticles formed on its surface, and white flakes precipitated. It gradually underwent a kind of fermentation and became sour. From 1440 parts of this juice, Einhof obtained, by analysis,*

Albumen	10
Extractive	18
Saccharine syrup	155

This syrup had much the taste of raw sugar, but could not be made to crystallize.

Ripe peas.

From ripe peas, by macerating them in water, and employing a mode of analysis similar to that used for ascertaining the constituents of barley, Einhof obtained the following products. The quantity employed was 3840 parts.†

Volatile matter	540
Starchy fibrous matter, with the } coats of the peas	840
Starch	1265
Animo-vegetable matter	559
Albumen	66
Saccharine matter	81
Mucilage	249
Earthy phosphates	11
Loss	229
	<hr/>
	3840

Starchy
fibrous mat-
ter.

The second of these constituents, entitled *starchy fibrous matter*, was what remained after the peas, reduced to a pulp with water in a mortar, had been washed with water till they

* Gehlen's Jour. vi. 120.

† Ibid. 13.

ceased to discolour it, or to give out any soluble matter to it. This residue was in part composed of the coats of the peas, and partly of a white fibrous matter without taste or smell. When dried, it became yellowish, and was easily reduced to powder; and the powder formed a paste with water, and dissolved like starch in hot water. It then bore a certain resemblance to starch, and agreed in its properties with the fibrous matter of potatoes.*

The fourth of the preceding constituents, entitled *animo-vegetable matter*, was obtained in this manner: the milky water in which the peas had been macerated, after depositing the starch, still continued muddy, but neither deposited any thing, nor would pass through the filter; but when diluted with its own bulk of water, it gradually deposited a powder, the whole of which, being collected on a filter, constituted the substance to which Einhof gave the name of *animo-vegetable matter*. It approaches most nearly to gluten; but as it differs in several particulars both from gluten and from all other vegetable constituents, we must consider it as a peculiar principle.

Animo-vegetable matter.

Its colour was at first white; it had no taste nor smell, but reddened vegetable blues, even after having been repeatedly washed in cold water. It was glutinous and adhesive, and could be kneaded into a paste. It was insoluble in water, both cold and hot; but when mixed with that liquid it soon putrefied; and on being treated with lime, gave out the smell of ammonia. In the pure alkaline leys it dissolves readily, and forms a kind of soap; but the carbonates require heat to produce the solution. Sulphuric acid dissolves it; the solution is light-brown, and when diluted with water, lets fall a white thread-like substance. It dissolves also in muriatic acid and chlorine, and in acetic acid. Nitric acid gives it a lemon-yellow colour. It dissolves also in alcohol, and the solution becomes milky when mixed with water. The tincture of galls throws down a copious white precipitate. Ether and volatile oils have no action on it. When dried it assumes a light-brown colour and the semitransparency of glue, and is easily reduced to powder.† Such are the properties of this substance, as far as they have been ascertained by M. Einhof.‡

* Gehlen's Jour. vi. p. 123.

† Einhof remarks, that he has seen the gluten of wheat assume this appearance. I have observed the same thing twice. In both cases the wheat was very inferior in quality, and had been the growth of a very rainy season.

‡ Gehlen's Jour. vi. 124.

Book IV.

When 3840 parts of ripe peas were reduced to ashes, the residue weighed 112 parts. From these ashes Einhof extracted phosphoric acid, sulphuric acid, muriatic acid, alumina, silica, carbonate of lime, phosphate of lime, oxide of iron, and phosphate of ammonia and magnesia.*

Marsh bean. 8. The seeds of the *vicia faba*, a small bean, becoming blackish when ripe, and used as an article of food, have likewise been examined by Einhof. The analysis was conducted in the same way as his other experiments already described. From 3840 parts of the ripe beans he obtained the following substances :

Volatile matter	600
Skins	386
Starchy fibrous matter	610
Starch	1312
Vegeto-animal matter	417
Albumen	31
Extractive, soluble in alcohol. ..	136
Gummy matter	177
Earthy phosphate	37½
Loss	133½
	3840†

Fourcroy and Vauquelin, who made experiments upon this substance likewise, found the ashes which it leaves, when incinerated, to consist of the phosphates of lime, magnesia, potash, and iron, and of uncombined potash. They could detect no sugar in it.‡

Kidney beans.

9. Kidney beans, which are the seeds of *phaseolus vulgaris*, have been likewise analysed by Einhof. They are characterized by the great proportion of animo-vegetable matter which they contain. From 3840 parts of these beans Einhof obtained the following substances :

* Gehlen's Jour. vi. 132.

† Ibid. 136.

‡ Ann. de Mus. d'Hist. Nat. No. xxxvii. 9.

Skins	288
Starchy fibrous matter	425
Starch	1380
Animo-vegetable matter, not quite free from starch	799
Extractive	131
Albumen, with some vegeto-animal matter	52
Mucilage	744
Loss	21

 3840*

10. Lentiles. To the same indefatigable chemist we are indebted for the analysis of the seeds of the *eruum lens*. From 3840 parts he obtained the following substances: Lentiles.

Fibrous matter	720
Albumen	44
Earthy phosphates, with a little albumen ..	22
Extractive, soluble in alcohol	120
Gummy matter	230
Starch	1260
Vegeto-animal matter	1433
Loss	11

 3840†

Here the proportion of animo-vegetable matter is still greater than in the kidney bean.

Lentiles were examined by Fourcroy and Vauquelin, but whether the same species as the preceding they do not put it in our power to determine, as they omit the botanical name. The liquid in which the flour of these seeds was macerated was not acid: it had a slight and rather disagreeable taste, and was precipitated copiously by infusion of galls, chlorine, and sulphate of iron. In short, it exhibited nearly the same phenomena as are described by Einhof. When the flour of lentiles is digested in alcohol, the liquid assumes a greenish-yellow colour, and a bitter acrid taste. When distilled it has the odour of vanilla very strongly, but acquires a disagreeable smell when mixed with water. The residue is greenish-yellow, has the appearance of a thick solution of soap, and a green oil swims upon the surface. The pod of the lentiles contains a portion of tannin.†

* Gehlen's Jour. vi. 545.

† Ibid. 542.

‡ Ann. de Mus. d'Hist. Nat. No. xxxvii. 10.

Book IV.
White lupine.

11. White lupine. To Fourcroy and Vanquelin we are indebted for the analysis of the seeds of the *lupinus albus*; the farina of which, according to their experiments, differs from that of all the other leguminous plants hitherto tried, as it contains neither starch nor saccharine matter. They found in it a yellowish-green oil of an acrid nature, which approaches the fixed oils in its properties; a vegeto-animal matter, which appears to bear a strong resemblance to gluten: a considerable portion of phosphate of lime and magnesia, and traces of phosphates of potash and of iron.*

Hemp-seeds.

12. Hemp-seeds. These seeds have been subjected to a chemical examination by Bucholz. The following are the results which he obtained.

Sixteen parts of hemp-seed yielded by expression rather more than three parts of a yellow-coloured oil. Its taste was mild, and it possessed all the characters of the fixed oils. From the residue he procured, by digestion in water and coagulation by heat, about $3\frac{1}{2}$ parts of albumen, and not quite half a part of fibrous matter. The insoluble coats and husks of the seeds weighed $6\frac{1}{8}$ parts. About one-fourth of a part of a brown coloured resin was obtained by means of alcohol, and about the same quantity of a substance to which Bucholz gives the name of *mucilaginous sugar*, and *soapy extract*, and about $1\frac{1}{2}$ part of *gummy extractive*.†

Coffee.

13. Coffee. This is the seeds of the *coffea arabica*, an evergreen-tree, usually of a small size. It was originally brought from Arabia, and was unknown in Europe till 1650. ‡ The entire fruit has some resemblance to a cherry. Under a pulpy skin it contains an oval kernel, which separates longitudinally into two parts, each covered with a thin coat, and marked longitudinally with a furrow on the flat side, where they are joined. The ripe fruit is dried in the sun, and heavy rollers passed over it, by which the seeds are parted and the coats broken. They are afterwards separated by winnowing.

Many experiments have been made by chemists to ascertain the constituents and qualities of coffee; Neumann, Geoffroy, Dufour, Kruger, Westfield, &c. published successively the result of their trials on it. The latest analyses are those of Hermann§ and Cadet.|| They neither agree well with each

* Ann. de Mus. d'Hist. Nat. No. xxxvii. 12.

† Gehlen's Jour. vi. 615.

‡ Neumann's Chemistry, p. 378.

§ Crell's Annals, 1800, ii. 108.

|| Ann. de Chim. lviii. 266.

other, nor with those of preceding writers. It is probable that these differences ought to be ascribed partly to the different mode of analyses followed, and partly to variations in the berries examined.

The infusion of coffee in boiling water is of a yellowish-green colour; but the decoction, by continuing the boiling, becomes brown. It becomes turbid on cooling. The alkalis render it more brown. It strikes a black with sulphate of iron, but does not precipitate with glue. Chlorine nearly destroys the colour; but if an alkali be added the liquid becomes red. Cadet obtained a decoction with coffee which produced no effect upon vegetable blues; but other chemists describe it as changing them to red, and Hermann obtained with alcohol a solution which produced the same effect. Infusion.

When water was distilled from coffee, what came over had an aromatic odour, and a few drops of a substance similar to myrtle wax swam on the surface of it; the residual liquid became milky when mixed with alcohol, and let fall a substance possessing the properties of gum.

When alcohol is digested on coffee, it acquires some colour. The solution when mixed with water becomes milky, and lets fall a whitish resinous matter. The residue being evaporated to dryness leaves a substance composed partly of extractive, and partly of the bitter principle discovered by Chenevix. From these experiments, chiefly made by Cadet, we see that coffee contains an aromatic principle, a little oil, gallic acid, mucilage, extractive, and bitter principle. It contains also, according to Cadet, sulphate of alumina, lime, muriate of potash, and a little albumen, which separates from the decoction in the form of scum. The result of Cadet's experiments was nearly as follows. Sixty-four parts of coffee yielded.

Gum	8.0	
Resin	1.0	Constitu-
Extract and bitter principle	1.0	ents.
Gallic acid	3.5	
Albumen	0.14	
Fibrous insoluble matter	43.5	
Loss	6.86	
	64.00	

Hermann obtained from 1920 parts of Levant and Martinique coffee the following proportions respectively:

Book IV.

	Levant.	Martinique.
Resin	74	68
Extractive	320	310
Gum	130	144
Fibrous matter	1335	1386
Loss	61	12
	<hr/>	<hr/>
	1920	1920

Roasted
coffee.

When coffee is roasted, a portion of tannin is formed in it by the action of the heat. It appears also that a new substance, having a peculiar agreeable smell, is developed; but the nature of this new principle has not yet been ascertained. It is developed also by roasting barley, beans, and a great variety of other vegetables, which are on that account occasionally employed as substitutes for coffee.

An elaborate analysis of coffee has been published by Payssé. He examined with particular care the bitter principle of coffee, first pointed out by Chenevix, and has endeavoured to show that it is in reality a peculiar acid, to which he has given the name of *caffic acid*. It reddens vegetable blues, but in other respects does not seem better entitled to the name of acid than *tannin*. He detected in coffee, besides this substance, a quantity of albumen, resin, and extractive. It contains a great proportion of matter insoluble in water. Its ashes yield muriate of potash and carbonate of lime. Payssé could not succeed in ascertaining the presence of tannin in roasted coffee. When coffee is distilled it yields, according to him, an acidulous water, a thick brown oil, and carbonate of ammonia. Very little gaseous matter was obtained.*

Coco.

14. Coco. This is the fruit of the *cocos nucifera*, one of the most valuable vegetable productions of India. A detailed account of the management and products of this tree has been published by Le Goux de Flaix;† but we are still unacquainted both with the constituents and chemical properties of the nuts. The outer coat is fibrous, and formed by the natives into excellent cordage. The kernel contains a considerable proportion of fixed oil, used by the Indians for lamps.‡

Almonds.

15. Almonds. This is the fruit of the *amygdalus com-*

* Ann. de Chim. lix. 196.

† Phil. Mag. xx. 316; and xxi. 77 and 110.

‡ See Tennant's Indian Recreations, ii. 282.

munis, used both as an article of food and in medicine. There are two kinds, the *bitter* and *sweet*. The latter consists chiefly of fixed oil and starch, or albumen; chemists have not yet ascertained correctly which. Bitter almonds, besides these two substances, probably contain also a portion of bitter principle. Prussic acid is likewise a constituent of them; but hitherto almonds have not been subjected to a regular chemical analysis. Chap. II.

16. Nutmeg. This is the seed of the *myristica moschata*, a tree which is cultivated in the Asiatic islands. The covering of the nut is known by the name of *mace*. The nutmeg varies in size and figure; it is furrowed on the outside, and greyish brown internally. Those that want white streaks are the best. From the experiments of Neumann, we learn that this substance contains two species of oil; a volatile oil, to which it owes its peculiar smell and taste, and which in his trials amounted to about $\frac{1}{3}$ d part of the nut; and a solid fixed oil resembling wax, and amounting to about $\frac{1}{3}$ d of the nutmeg. He detected also a quantity of gum;* and it is probable, from the appearance of the kernel, that it contains likewise starch. By expression, the solid oil is separated, and mixed with the volatile oil. In that state it is sold under the name of *oil of mace*. Nutmeg.

17. Pepper. This is the seed of the *piper nigrum*, a shrub cultivated in India for the sake of its berries. Its outer coat is of a dark brown colour; when deprived of it, the berry is known by the name of *white pepper*. Pepper.

Pepper, when macerated in cold water, does not lose its shrivelled appearance. The infusion acquires a brown colour, reddens vegetable blues, and has both the taste and odour of pepper. If we repeat the maceration a number of times, the liquid still continues to acquire a colour, but it ceases to have the smell and flavour of pepper. White pepper does not communicate any colour to water. Hence the colouring matter must reside in the outer coat. It possesses most of the properties of extractive. When pepper is macerated in alcohol, the liquid acquires a light yellowish green colour. When distilled, it leaves a green coloured matter, partly resinous, and partly oily. This oil is the source of the odour and taste of pepper. Its taste is extremely hot, and its smell, when dissolved in alcohol, and diluted with water, extremely

* Neumann's Chem. p. 104.

Book IV. pleasant. It possesses nearly the properties of volatile oil. When water is boiled for some time with ground pepper, it acquires the property of forming, with the infusion of nut-galls, a precipitate, which dissolves again when the liquid is heated to 120°. Hence it contains a portion of starch. These three substances, starch, oil, and extractive, are the most remarkable of the constituents of pepper.*

Mustard. 18. Mustard. The seeds of the *senapis nigra* are well known for the strong biting taste which they possess, and for their property of acting as a vesicatory. This acrid quality resides in a yellow oil, which may be distilled over with water. It sinks in that liquid, and according to the experiments of M. Thibierge is not soluble in alcohol or ether, but soluble in water. Besides this acrid oil mustard contains a fixed tasteless oil, which may be obtained by expression. Sulphur, albumen, gum, phosphoric acid, several phosphates with traces of sulphates and muriates, may be likewise detected in mustard.†

Anatto. 19. Anatto. This pigment is obtained from the seeds of the *bixa orellana*, a tree cultivated in Guiana and other parts of the West Indies. When ripe, the capsules are gathered, and the seeds being separated, are pounded, steeped in water for weeks or months, then subjected to the press; and the colouring matter thus obtained is suffered to subside, collected, and dried. A shorter process has been lately proposed by Leblond. The colouring matter, which is confined to the surface of the seed, is separated by maceration and washing, and then thrown down from the water by an acid.‡ Anatto is usually in hard cakes, brown without, and red within. It dissolves much more easily in alcohol than in water. Weak alkaline leys dissolve it likewise with facility. The decoction of anatto has a peculiar smell and a disagreeable taste. Its colour is reddish yellow; alkalies render it orange yellow; and from this mixture the acids throw down an orange-coloured precipitate.§ An elaborate chemical examination of this pigment has been made by Dr. John. He found its constituents to be as follows :

* See Nicholson's Jour. ii. 7.

† See Annals of Philosophy, xvi. 90: and John's Chemische Untersuchungen, iii. 153.

‡ Ann. de Chim. xlvii. 113.

§ Berthollet, Elemens de l'Art de la Teinture, ii. 268.

Aroma, trace.
An acid, trace.

Chap. II.

Resin, with colouring matter	28
Mucilage	26.5
Albumen	20
Extractive	20
A peculiar substance, approaching mucilage	4
	98.5*

20. Berries of the *Phytolacca decandra*. These berries give a beautiful purple colour to water, of a very fugitive nature. A few drops of lime water change it to yellow; and this yellow liquid is the most delicate test of acids hitherto observed. The smallest quantity of acid restores its purple colour. Braconnot, to whom we are indebted for these observations, has shown, that it is at least four times as delicate as the infusion of litmus. Unfortunately it alters its nature in a few hours, and then loses its delicacy as a reactive. It can only be used when recently prepared. †

21. The seeds of the *Lycopodium clavatum*, or common club moss, have been long remarkable for their combustibility. They are extremely small; and when blown into the flame of a candle they take fire, with a small explosion, and are used in theatres to imitate the appearance of lightning. They have been subjected to chemical analysis by Bucholz. From 1000 parts of the seeds he obtained 60 parts of a fixed oil, soluble in alcohol like castor oil; 30 parts of sugar; 15 of a mucilaginous extract. There remained a matter insoluble in water, alcohol, ether, oil of turpentine, alkaline leys. But when boiled with the last of these agents, ammonia was disengaged, and a kind of extract formed. When 2000 grains of this substance were distilled, they yielded 290 ounce measures of gas, which was a mixture of carhuretted hydrogen and carbonic acid; 1080 grains of a brownish oil, containing ammonia; 330 grains of a watery liquid, holding acetate of ammonia in solution. The residue in the retort, amounting to 310 parts, had very much the appearance of anthracolite. When the insoluble part of the seeds was boiled in nitric acid, a portion of oil soluble in alcohol was likewise formed. ‡

* Chemische Untersuchungen, i. 56.

† Ann. de Chim. lxii. 31.

‡ Gehlen's Jour. vi. 593.

Book IV.
Husk of
nuts.

22. To Braconnot we are indebted for an elaborate analysis of the outer husks of the common nuts which are frequently employed in France as a dye stuff, probably as a substitute for nutgalls. He obtained from the fresh husks of nuts the following substances :*

Starch.	Citric acid.
An acrid and bitter substance.	Phosphate of lime.
Malic acid.	Oxalate of lime.
Tannin.	Potash.

SECT. X.

OF FRUITS.

Constitu-
ents.

THE fruits of vegetables are equally various with the seeds. But few of them have been hitherto subjected to chemical analysis. They almost all contain an acid; and this acid is usually the tartaric, the oxalic, the citric, or the malic, or a mixture of two or more of them. Hardly any other, except perhaps the acetic, has hitherto been found in fruits. They usually contain likewise a portion of gummy matter, sometimes starch; and the fleshy fruits contain also a fibrous matter, not yet accurately examined. The colouring matters of fruits, especially the red, dissolve readily both in water and alcohol; but very speedily decay when exposed dry to the action of the sun and weather. Hence they cannot be used as dyes.

A list of the acids contained in a considerable number of fruits has been given already in a former part of this work. As very few vegetable fruits have hitherto been analysed, little can be added to what was given there. The only analyses, indeed, hitherto published, are those of tamarinds, grapes, and cucumber.

Tamarinds.

1. *Tamarinds*.—This substance consists chiefly of a pulpy matter which fills the pods of the *tamarindus indica*, and covers the seeds. It is brought to Europe preserved in sugar. We are indebted to Vauquelin for an analysis of this substance, published however at a very early period of his chemical career. By treating the pulp of tamarinds, such as they are sold by the apothecaries, first with cold water, and afterwards with hot, he separated the following substances:†

* Ann. de Chim. lxxiv. 303.

† Ibid. v. 92.

Supertartrate of potash	300
Gum	432
Sugar	1152
Jelly	576
Citric acid	864
Tartaric acid	144
Malic acid	40
Feculent matter	2880
Water	3364
		<hr/>
		9752

2. *Grapes*.—Though grapes have been repeatedly examined Grapes. by chemists, we are not yet in possession of an accurate chemical analysis of their constituents. It has been ascertained, however, that they contain supertartrate of potash, tartaric acid, citric and malic acids.* They contain also abundance of sugar, a portion of mucilage and jelly, some albumen and colouring matter, and, according to Proust, a portion of gluten.

3. *Cucumber*.—This is the fruit of the *cucurbita citrullus*, Cucumber. and is well known as an article of food. It was subjected to analysis by Dr. John, who found 600 parts of it composed of the following ingredients:

Water	582·8
Substance similar to fungin	3·2
Soluble vegetable albumen	0·8
Resin	0·25
Extractive with sugar	9·95
Mucus	} 3·0
Phosphate of lime	
Phosphate of potash	..	
Phosphoric acid	
Ammoniacal salt	
A malate	
Sulphate of potash	...	
Muriate of potash	
Phosphate of iron	
		<hr/>
		600·0†

Pears, apples, lemons, oranges, &c. have not yet been sub-

* Bouillon, Jour. de Phys. xxix. 3. Parmentier, Ann. de Chim. liii. 119.

† Schweigger's Jour. ix. 34.

Book IV. } jected to chemical examination. The few facts which have been ascertained respecting them are detailed in a different part of this work.

SECT. XI.

OF BULBS.

By *bulbs* are understood tubercles connected with the roots of vegetables, very analogous to the *buds*, and containing the embryo of a future plant. The *potatoe* is a well known example of a bulb.

Bulbs vary considerably in their constituents and properties, as well as the other parts of plants; but they bear, upon the whole, a considerable resemblance to the seeds of those grasses which are cultivated under the name of *corn*; at least starch constitutes a notable portion of most of the bulbs hitherto examined. Hence several of them are employed as articles of food. Few of them hitherto have been subjected to a chemical analysis. I can only therefore specify the following:

Potatoes.

1. *Potatoes* are the bulbs of the *solanum tuberosum*, an American plant, first brought to this country by Sir Walter Raleigh, and now cultivated in every part of Europe, and in Britain constituting one of the most important articles of food. The plant and the bulb are too well known to require description.

Potatoes have been repeatedly subjected to chemical examination. Parmentier published an elaborate dissertation on their culture, uses, and properties, about the year 1776, which contributed considerably to promote the cultivation of them on the continent. Dr. Pearson contributed to the Board of Agriculture a valuable essay on potatoes, containing a set of chemical experiments on them, performed with his usual skill and ingenuity; and Einhof has published a very elaborate analysis of the root in the fourth volume of Gehlen's Journal.

The variety of potatoe which chiefly occupied the attention of Einhof, was that which has a red skin and flesh-coloured juice. When dried by a moderate heat, till they ceased to lose any weight, potatoes were reduced to $\frac{1}{4}$ th of their original weight.*

The analysis of this root was conducted by Einhof pretty much in the same manner as his analysis of barley and rye.

* Gehlen's Jour. iv. 457.

A determinate quantity of potatoes was reduced with water to a pulp, and then washed on a scarce till the liquid ceased to come off milky, or to hold any thing in solution. What remained on the cloth was the fibrous matter of the potatoe; but it differed essentially from the fibrous matter of most plants. With boiling water it formed a paste similar to that made by means of starch, and when dry it assumed a semi-transparent appearance. This matter, when triturated in a mortar, and again washed with water, yielded a considerable portion of starch. The residue, which was of a light-grey colour, being triturated a second time, formed a powder which bore a considerable resemblance to starch, both in its appearance and properties.

The liquid with which the potatoe was washed was at first milky, but deposited, on standing, a heavy white powder, which was *starch*. When filtered it had a carmine-red colour, and reddened vegetable blues. When boiled a flaky precipitate separated, partly white and partly red. This precipitate possessed the properties of albumen. The residue, evaporated to the consistence of an extract, had a brownish colour, was insoluble in alcohol and ether, soluble in water, and, according to Einhof, was a mucilaginous matter. The following were the proportions of these various substances obtained from 100 parts of potatoes :

Starch	15	Constitu- ents
Fibrous starchy matter	7	
Albumen	1·4	
Mucilage, in the state of a thick syrup	4	
	27·4	

To ascertain the nature of the acid which exists in potatoes, Einhof separated the juice of potatoes by a gentle pressure. He had first frozen and then thawed them, to facilitate the separation. Lime-water was added in excess to this liquid, and the precipitate was digested in diluted sulphuric acid, to separate the lime from the acid. Thus obtained, it was found to be a mixture of tartaric and phosphoric acids.* The sap, thus deprived of its acid, contained an excess of lime in com-

Contains
tartaric and
phosphoric
acids.

* Though Einhoff obtained phosphoric acid by the method described in the text, it does not follow that it existed in the potatoes in an uncombined state. It might have been in combination with lime, and held in solution in the potatoe juice by means of the tartaric acid.

Book IV. bination with the mucilage. Einhof found, that this lime became gradually saturated with carbonic acid, even though the sap was kept in close vessels; and that in process of time, if a sufficient quantity of lime were present, the mucilage acquired a sweet taste; and when treated with alcohol, a portion dissolved, which yielded crystals of sugar. This he considered as a conversion of the mucilage into sugar.

Ashes. From 1820 parts of dried potatoes, Einhof obtained 96 parts of a greyish-white ash. Of these, 64 parts were soluble in water. They consisted chiefly of carbonate of potash; but contained likewise $10\frac{1}{4}$ parts of phosphoric acid, $3\frac{1}{2}$ of sulphuric acid, and two of muriatic. The insoluble 35 parts consisted of earths and oxides. From 20 parts of it he obtained

2.5 silica.

6.0 lime.

4.0 alumina.

7.0 magnesia, with some manganese or oxide of iron.

19.5

Einhof examined different kinds of potatoes. He found the same ingredients in all, but the proportions varied considerably. It is not necessary to state the results of his experiments, because the same variations doubtless occur even in the same kind of potatoe.

But as it may be useful to know the quantity of starch furnished by different varieties of potatoe, the following table, drawn up from the experiments of Mr. William Skrimshire, junior, is subjoined. Five pounds avoirdupois of fresh potatoes were used, and the starch was separated by grating the potatoe, and pouring water upon it placed upon a searce.*

* Nicholson's Jour. xxi. 71.

Substances.	Varieties of Potatoes used.							
	Captain Hart.	Rough Red.	White Kidney.	Moulton White.	Yorkshire Kidney.	Hundred Eyes.	Purple Red.	Ox Noble.
Fine starch . . .	0 9	0 7 $\frac{1}{4}$	0 0	0 9	0 8 $\frac{1}{4}$	0 8 $\frac{1}{4}$	0 8	0 6 $\frac{1}{2}$
Ditto, slightly } coloured ..	0 3	0 3 $\frac{1}{4}$	0 9 $\frac{1}{2}$	0 2 $\frac{1}{4}$	0 2 $\frac{1}{4}$	0 0 $\frac{1}{4}$	0 0 $\frac{1}{2}$	0 1 $\frac{1}{4}$
Pulp dried . . .	0 6	0 6 $\frac{1}{2}$	0 3 $\frac{1}{4}$	0 5 $\frac{1}{2}$	0 6 $\frac{1}{2}$	0 6 $\frac{1}{2}$	0 5	0 8
Water, mu- } cus, and ex- } tractive . . .	3 14	3 15	4 2 $\frac{1}{4}$	3 14 $\frac{1}{2}$	3 14 $\frac{1}{4}$	4 0 $\frac{1}{4}$	4 2 $\frac{1}{2}$	3 15 $\frac{1}{4}$
Total . . .	5 0	5 0	5 0	5 0	5 0	5 0	5 0	5 0

When potatoes are exposed to the action of frost, it is well known that they become soft, and acquire a sweet taste. This taste is succeeded by a sour taste, owing to the rapid evolution of acetic acid, and the root soon passes to putrefaction. From the experiments of Einhof, we learn that the sugar is formed at the expense of the mucilage; for the other ingredients were found, in potatoes sweetened by frost, in the usual proportions. He considers this sweetening process as connected with the vegetative powers of the root.

Action of
frost on po-
tatoes.

When potatoes are boiled, they lose from 1 to 1 $\frac{1}{4}$ per cent. of their weight. The juice, which may be separated from them, is sweet-tasted. The meal is insoluble even in boiling water, though potatoe starch forms a transparent solution with hot water. Thus it appears, that by boiling, the albumen, fibrous matter, and starch combine together, and form an insoluble compound.*

Boiling.

From these experiments, it appears that potatoes differ essentially from wheat and barley by containing no gluten. They approach, in some measure, to the nature of rye.

Dr. Peschier, of Geneva, has detected the presence of mucous sugar and of gum in the potatoe. This explains why they are capable of undergoing the vinous fermentation.†

2. *Garlic*.—This is the bulbous part of the root of the *allium sativum*, and is well known, and remarkable for its strong

Garlic.

* Gehlen's Jour. iv. 485.

† Annals of Philosophy, xii. 337.

Book IV. smell and peculiar taste. It was much celebrated by the ancients, both as an article of food and as a medicine. It has been repeatedly examined by chemists. The analysis of Neumann, considering the state of the art of examining vegetables at that time, must be considered as very exact.* Cadet has subjected it to a chemical examination.†

Contains an
acid oil.

When dried, it loses nearly two-thirds of its weight; but this proportion is doubtless subject to considerable variation. The expressed juice of garlic is of a thick consistence like mucilage, and slightly reddens vegetable blues. When diluted with water, and filtered, it yields flakes of albumen when boiled. The residue consists chiefly of mucilage, of which garlic yields a very great proportion, and of extractive. This last is somewhat acrid in its nature. When garlic is distilled with water, it yields a portion of yellow-coloured volatile oil, at first lighter than water, but gradually becoming heavier as the distillation advances. To this oil garlic owes its most remarkable properties. Its taste is very acrid, and its smell strong. When applied to the skin, it produces an irritation not inferior to cantharides, and, like that drug, might be employed to blister the skin. When triturated with oxide of iron, it immediately strikes with it a black colour; but it has no effect upon any other metallic oxide.

Other con-
stituents.

When garlic is treated with alcohol, the liquid assumes a reddish-yellow colour, and leaves, when evaporated, a brown extract, very acrid, which attracts moisture from the air.

When garlic is distilled, it yields first a liquid slightly coloured, and having a very acrid taste; then a thick brown oil, and abundance of inflammable air and carbonic acid. The liquid in the receiver emits the smell of ammonia when mixed with lime. When 40,320 parts of garlic were incinerated, they left 4896 parts of ashes, or about $\frac{1}{8}$ th of the original weight. From 172 parts of these ashes Cadet obtained the following substances:

* Neumann's Chem. p. 431.

† Ann. de Chim. lix. 106.

Potash	33·0
Sulphate of potash with some muriate	58·0
Alumina	2·0
Phosphate of lime	15·6
Oxide of iron	1·5
Magnesia	9·0
Lime	14·0
Silica	8·0

 141·1

From 1406 parts of fresh garlic he obtained

Mucilage	520
Albumen	37
Fibrous matter	48
Water, by estimate ...	801

 1406

Bonillon-Lagrange has detected in garlic, besides the acrid oil, a quantity of sulphur, starch, and saccharine matter.*

3. *Onion*.—This is the bulbous root of the *allium cepa*. A Onion. few experiments on it had been made by Neumann and Cadet. But Fourcroy and Vauquelin have published a very curious and accurate analysis of it. When reduced to a pulp, and subjected to the press, it yields a viscid juice, somewhat opaque, at first colourless, but becoming gradually red in consequence of the oil which it contains. It has a strong smell, and reddens vegetable blues. It is precipitated by acetate of lead, lime water, oxalic acid, nitrate of silver, and potash. When distilled it yields a milky liquid, slightly acid, on the surface of which swim some drops of oil. It contains a little sulphur dissolved in the oil; for chlorine gives it the property of precipitating nitrate of barytes, and when distilled in a copper vessel some sulphuret of copper is formed. The portion of juice not distilled deposits a fawn-coloured sediment, having a strong oniony odour. Alcohol deprives this sediment of oil and of sulphur; the residue seems to possess properties analogous to those of gluten. The liquid, from which this precipitate has separated, contains phosphoric acid, sulphur, and gluten.

When onion juice is kept at a temperature between 60° and 70° it emits no gas, but it changes its colour successively

* *Annal. of Philosophy*, vol. 50.

Book IV. to red and yellow, and lets fall a fawn-coloured sediment. It is now converted into vinegar still retaining the oniony odour; a proof that the volatile matter remains undecomposed. This vinegar contains in it a good deal of manna, which readily crystallizes in needles. On examining recent onion juice no manna could be found in it, merely a considerable quantity of uncrystallizable sugar. This sugar, in a heat between 60° and 70°, is gradually decomposed, and converted into acetic acid and manna. The fawn-coloured matter, which precipitated during this acetification, consisted of gluten, oil, and sulphur.

Onion juice is capable of undergoing the vinous fermentation. When eight ounces troy were mixed with 122 cubic inches of water, and somewhat less than one ounce of barm, they underwent a brisk fermentation, and yielded by distillation about 2·35 ounces troy of strong alcohol. By Lavoisier's experiments 3·7 ounces of sugar would have yielded the same quantity of alcohol.

Besides the preceding constituents, Fourcroy and Vauquelin found in onion juice mucilage, phosphate of lime, and citrate of lime.*

Squills.

4. *Squills*.—This is the bulbous root of the *scilla maritima*, and is employed in medicine as a diuretic, and to excite nausea and vomiting. It has been subjected to a chemical examination by Vogel, and was found by him to owe its peculiar properties to a species of bitter principle which he distinguished by the name of *scillitin*, which has been described in the preceding chapter. From dried squills he extracted the following substances: †

Gum	6
Scillitin	35
Tannin	24
Citrate of lime	
Sugar	
Woody fibre . .	30

Lathyrus tuberosus.

5. Tubercles of *lathyrus tuberosus*.—This plant is cultivated in Holland, where the tubercles are used as an article of food. They are called by the French *macjon*. The constituents of these tubercles, according to the analysis of Braconnot, are as follows:

* Ann. de Chim. lxx. 161.

† Ibid. lxxxiv. 147.

Water	327·98
Starch	84·00
Common sugar	30·00
Lignin	25·20
Animalized matter ..	15·00
Albumen	14·00
Oxalate of lime	1·80
Rancid oil }	0·90
Adipocire }	
Phosphate of lime...	0·50
Sulphate of potash ..	0·22
Malate of potash ...	0·20
Phosphate of potash .	0·10
Muriate of potash ...	0·10
Odorous principle ...	—
	500·00*

SECT. XII.

OF LICHENS.

THE lichens are a class of plants which differ almost in every respect from other vegetables. Many of them have not the smallest appearance of plants, but form hard crusts, which cover rocks, wood, trees, &c.; others have the form of leaves or of branches, but nothing resembling flowers is visible. It was to Tournefort and Micheli that botany is indebted for first fixing the word *lichen*, formerly vague and ill defined, to a peculiar set of plants. Afterwards Linnæus placed them among the algæ, and described 81 species. Since that time a great variety of botanical writers have devoted considerable attention to them, particularly in ascertaining and describing their parts of fructification, and the various changes which they undergo in the different periods of their vegetation; but few only have attempted to analyze them, or to point out the useful purposes to which they may be applied. Wellemet has given us a historical account of 41 species of lichens, and detailed their medical and economical uses with considerable accuracy. Amoreux, in a dissertation on the subject has given us still more copious details, and has published likewise a brief

* *Annals of Philosophy*, xiii. 70.

Book IV.

chemical analysis of some of the most remarkable lichens. Hoffmann, who had previously distinguished himself by his botanical arrangement of the lichens, published an account of their chemical and economical properties in 1787, and has given us the chemical analysis of several, made by Georgi with considerable care. Westring turned his attention particularly to the uses of the lichens in dyeing; and, in seven dissertations, published successively in the Stockholm Transactions from 1792 to 1797, has examined almost all the lichens of the north, and described the colours which each of them is capable of yielding, and the manner of obtaining it. It is to these writers, and to Georgi, that we are indebted for the few facts known respecting the composition and chemical properties of the lichens.

The lichens are found in all countries and climates, and are very numerous; considerably more than two hundred species have been described by botanists as natives of Britain.

Some lichens contain gum.

From the experiments of Georgi, professor of chemistry in Petersburg, we learn that the lichen *farinaceus*, *glaucus*, and *physodes*, form with water a mucilage which yields, when evaporated, a gum as transparent and tasteless as gum arabic. Lichen *pulmonarius* yields likewise a gum, but its taste is somewhat bitter. The gum yielded by these lichens amounted to $\frac{1}{8}$ th of their weight. When treated with alcohol, the liquid acquires a green colour and a bitter taste.* Amoreux, who repeated these experiments, obtained from the lichen *pulmonarius* a reddish gum, much less transparent than gum arabic. This lichen gave a yellow colour to alcohol. Probably he had examined a different species from that tried by Georgi, or if not, the lichens must have been of very different ages.

Amoreux found, that when the lichen *prunastri* was steeped in water, its branches became transparent like animal membrane, and adhered strongly to paper. In this state it is insipid, but as friable as celery. He obtained abundance of gum from the lichen *islandicus*, and from all the broad-leaved lichens tried. He succeeded in extracting gum from the lichen *fraxineus*, *caninus*, and *caperatus*. This last gave a lemon-yellow colour to ammonia.†

* See the experiments of Georgi, as quoted by Amoreux, in his *Recherches et Experiences sur les divers Lichens*, p. 94.

† *Ibid.* p. 95.

Chap. II.
Some may
be used as
food.

Georgi found, that when the lichen physodes, hirtus, farinaceus, and pulmonarius, were boiled in water, they yielded a yellowish mucilage nearly insipid, and that the lichens thus treated might be eaten with salt. They all yielded a portion of resin to alcohol, but it did not give a taste to the water in which they were boiled. When incinerated, these lichens yielded a little potash, lime, and silica, but no sulphuric or muriatic salt. When distilled, they yielded an acidulous water, and a yellow or blackish oil which sunk in water.

Such are the imperfect experiments hitherto made on the constituents of a few of the lichens. One of them, however, the *lichen islandicus*, has been subjected to a rigorous and curious analysis by Berzelius. He obtained from 100 parts of this lichen the following constituents :

Syrup	3·6
Bitartrate of potash with some tartrate of lime and phosphate of lime ... }	1·9
Bitter principle	3·0
Green wax	1·6
Gum	3·7
Extractive colouring matter	7·0
Starch	44·6
Starchy insoluble matter	36·6
	—————
	102·0 *

Berzelius afterwards examined the lichen plicatus, lichen barbatus, lichen fastigiatus, and the lichen fraximens. He found them all characterized by the presence of a species of starch which possesses several peculiar properties.†

I shall now mention such of the lichens as are most remarkable for the colouring matters which they yield. Archil.

1. *Lichen roccella*.—This lichen, which grows abundantly in the Canary islands, but which is found also on the south coast of England and France, yields the dye stuff called *archil*. If we believe Tournefort, this dye stuff was known to the ancients, and was employed to produce the colour known among them by the name of *purple of amorgos*. Be that as it may, the colouring property of the lichen roccella was accidentally discovered by a Florentine merchant in 1300, who observed that

* Ann. de Chim. xc. 277.

† Afhandlingar, iii. 381.

Book IV.

urine gave it a fine violet colour. Since that time it has been an article of commerce, and was long prepared in Florence. The lichen, reduced to powder, is put into a vat with urine and quicklime, and regularly stirred, fresh urine and lime being added till it has acquired the destined colour. Other substances were likewise added; but it has been shown that these only are essential. Hellot ascertained, that by treating lichens with lime and ammonia, their colouring matter, if they have any, will be developed; and Westring has simplified the process by substituting sal-ammoniac for ammonia.

Archil thus prepared is used occasionally to brighten other colours, but the colour which it yields is far from fixed. It is employed to tinge the alcohol used in thermometers. Nollet observed that this tincture became colourless when kept excluded from the air, but that it recovered its colour again when air was admitted. The nature of the colouring matter of archil has not yet been ascertained by chemists.

2. *Lichen parellus*.—From this lichen, which grows abundantly in the mountains of Auvergne, and other parts of France, and which is common also in Britain, the pigment called *archil of Auvergne* is prepared. The process is pretty much the same as that by which the lichen roccella is prepared, and the pigment is distinguished by the same name, and applied to the same use, but is not considered as so valuable. It is obvious that the colouring matter of each is analogous.

Other lichens yielding colouring matters.

3. *Lichen pertusus*.—Treated with lime and sal-ammoniac, it yielded a brown colouring matter to Westring.

4. *Lichen ventosus*.—This lichen dyed wool of a brown colour, which resisted the action of alkalies.

5. *Lichen hæmatoma* yielded a wax-yellow colour.

6. *Lichen corralinus*.—This lichen was found by Westring to abound in colouring matter. By simple infusion in water with a little common salt, it dyed wool yellow. Without addition, it gives a deep-brown of considerable permanence. It yielded the same colour when treated with sal-ammoniac and lime.

7. *Lichen pseudo-corrallinus* yielded a fine orange, which was brightened by muriate of cobalt.

8. *Lichen tartareus* yielded a fine brown.

9. *Lichen centrifugus*, with fixed alkalies, yielded a fine wax-yellow; with water, a brown; with common salt and nitre, an orange.

10. *Lichen saxatilis*.—This lichen, with soda, yields a yellow; with lime and sal-ammoniac, a nankeen; and with muriate of soda and nitre, an orange.

11. *Lichen physodes*, by the same reagents, yields various shades of yellow and brown; *lichen juniperinus*, yellow and brown; *lichen tenellus*, yellow, olive, and reddish-brown; *lichen furfuraceus*, yellow and brown. The same colours were obtained from a considerable number of leafy lichens.

Lichen croceus, with lime and sal-ammoniac, gave out a red colour. Westring obtained several colours from other lichens; and by mixing several of them together, he varied the shade, and produced a new set of colours, differing both in their intensity and fixity. But for the particulars of his numerous experiments, the reader is referred to his dissertations on the subject.*

SECT. XIII.

OF MUSHROOMS.

THE mushrooms are a class of plants proverbial for the rapidity of their growth and their speedy decay. When they putrefy, they give out an extremely unpleasant odour, and appear to approach animal matter much more closely than other vegetable substances. They have lately attracted the attention of M. Vauquelin, and M. Braconnot, to whom we are indebted for the analysis of no fewer than 17 species of this hitherto neglected tribe of vegetables. Braconnot has distinguished the insoluble spongy portion which characterizes the mushrooms by the name of *fungin*, and under that name it has been described in the preceding chapter. It approaches woody fibre in its properties; but seems to be sufficiently distinguished from it by various characters, particularly by constituting a nourishing article of food, and by being much less soluble in alkaline leys. Braconnot likewise ascertained the existence of two new acids in mushrooms. One of these has been described in a former part of this work under the name of *boletic acid*. The other, which appears to constitute a very general ingredient in

* The first has been translated into French, and printed in vols. xv. and xvii. of the *Annales de Chimie*. The others are inserted in *Crell's Annals* for 1796, 1797, and 1799.

Book IV. mushrooms, he has on that account distinguished by the name of *fungic acid*.*

Fungic acid. Fungic acid was extracted from the *boletus juglandis* in sufficient quantity to examine its properties. He found it also in the *merulius cantharellus*, *boletus pseudo-ignarius*, *peziza nigra*, and the *phallus impudicus*. Braconnot obtained it from the *boletus juglandis* by the following process. The juice of that mushroom was pressed out, boiled, and filtered, in order to separate the vegetable albumen. The filtered liquid was evaporated to the consistence of an extract, and digested repeatedly with alcohol. The undissolved portion was dissolved in water, and the solution mixed with acetate of lead. A copious precipitate fell, consisting chiefly of *fungate of lead*. This precipitate was decomposed in a gentle heat by means of dilute sulphuric acid. The fungic acid thus disengaged was saturated with ammonia, and the fungate of ammonia purified by repeated solutions and crystallizations. It was then dissolved in water, the fungic acid was precipitated by means of acetate of lead, and the fungate of lead, after being carefully washed, was decomposed by means of diluted sulphuric acid. Fungic acid thus obtained possesses the following properties :

It is colourless, has a very sour taste, does not crystallize, and deliquesces if it be evaporated to dryness and exposed to the air.

Fungates. With potash and soda it forms uncrystallizable salts, very soluble in water, but insoluble in alcohol.

With ammonia it forms a salt which crystallizes in regular six-sided prisms.

With lime a salt not altered by exposure to the air; and soluble at the temperature of 73° in 18 times its weight of water.

With barytes, a salt difficultly crystallizable, and soluble in 15 times its weight of water, at the usual temperature of the atmosphere.

With magnesia, a salt soluble in water and in small granular crystals.

With alumina, an uncrystallizable salt having the appearance of gum.

* This name is unhappy, because according to the received principle of naming acids it indicates that the acid so named is a compound of *fungin* and *oxygen*; an opinion so far from being established that it has not even been advanced.

With oxide of manganese, a salt similar to the fungate of alumina. Chap. II.

With oxide of zinc, a salt moderately soluble in water, which crystallizes readily in parallelopipedons. When exposed to the heat of a candle it burns without frothing and leaves the oxide of zinc pure.

When dropped into acetate of lead it occasions a white precipitate soluble in acetic acid. When dropped into nitrate of silver no precipitate appears; but a precipitate falls when an alkaline fungate is dropped into that salt.*

The following are the different species of mushroom hitherto subjected to analysis.

1. *Agaricus campestris*.—This agaric, which is a common article of food, was analysed by M. Vauquelin, who found in it the following substances: *Agaricus campestris.*

(1.) Adipocire. This substance was obtained by boiling in alcohol the matter that remained after the juice of the agaric was pressed out. The alcohol on cooling deposited the adipocire in flocks. It has a brownish-white colour, a fatty feel like that of spermaceti, melts when heated, and gives off a white vapour, having the odour of fat.

(2.) An oily or fatty matter.

(3.) Vegetable albumen.

(4.) The sugar of mushrooms.

(5.) An animal matter soluble in water and alcohol. When heated it gives out the smell of roasting meat. Similar to the substance called *osmazome*.

(6.) An animal substance insoluble in alcohol.

(7.) Fungin.

(8.) Acetate of potash.†

2. *Agaricus volvaceus*.—This agaric, according to the analysis of Braconnot, contains the following constituents: *Agaricus volvaceus.*

(1.) Much water.

(2.) Fungin.

(3.) Gelatin.

(4.) Vegetable albumen.

(5.) A great quantity of phosphate of potash.

(6.) Acetate of potash.

(7.) Sugar of mushrooms.

(8.) A fluid brown oil.

(9.) Adipocire.

* Ann. de Chim. lxxxvii. 244.

† Ibid. lxxxv. 7.

Book IV.

- (10.) Wax.
- (11.) A very fugaceous deleterious matter.
- (12.) An uncombined acid, suspected to be the acetic.
- (13.) Benzoic acid.
- (14.) Muriate of potash.*

Agaricus acris.

3. *Agaricus acris* or *piperatus*.—This agaricus was examined by Dr. Lister in 1672.† He obtained from it a milky juice with taste hotter than pepper, not discoloured by exposure to the air, nor by the blade of a knife. It speedily coagulated when kept in a glass vessel; but did not lose its hot taste. Trommsdorf made some experiments on it, and extracted from it a peculiar acrid matter and vegetable albumen. When distilled it yielded a considerable quantity of carbonate of ammonia.‡ Braconnot subjected it to an elaborate analysis, and extracted from it the following substances :

- Water.
- Fungin.
- Vegetable albumen.
- Gelatin.
- Much adipocire.
- Acetate of potash.
- Sugar of mushrooms.
- Phosphate of potash.
- A peculiar vegetable acid united to potash.
- An oily matter.
- A very acrid and fugaceous matter.
- Muriate of potash.§

Stypticus,

4. *Agaricus stypticus*.—Twenty parts of this agaric analysed by Braconnot, yielded

Fungin	16·7
Resin	} 1·8
Adipocirc.	
Unknown gelatinous substance. . .	} 1·5
Combustible acid united to potash	
A fugaceous acrid principle	
20·0	

Bulbosus,

5. *Agaricus bulbosus*.—Vauquelin subjected this agaric to

* Ann. de Chim. lxxix. 276.

† Phil. Trans. vii. 5116.

‡ Ann. de Chim. xxii. 220.

§ Ibid. lxxix. 285.

|| Ibid. lxxxvii. 260.

some experiments, and extracted from it the following substances. Obap. II.

An animal matter insoluble in alcohol.

Osmazome.

A fatty soft matter, of a yellow colour and acrid taste.

An acid salt, which is not a phosphate.

The skeleton of the agaric when distilled yielded an acid.*

6. *Agaricus theogalus*.—Vauquelin extracted from this agaric, Theogalus,

Sugar of mushrooms.

A fatty matter of a bitter and acrid taste.

An animal matter insoluble in alcohol.

Osmazome.

A salt with a vegetable acid.†

7. *Agaricus muscarius*.—This agaric was examined by Vauquelin, who extracted from it Muscarius.

The two animal matters contained in the last agaric.

A fatty matter.

Muriate, phosphate, and sulphate of potash.

The skeleton both of this and the preceding species yielded an acid when distilled. †

8. *Boletus juglandis*.—Braconnot subjected this boletus to an elaborate and ingenious analysis. He found 1260 parts of Boletus juglandis, it composed of the following constituents:

Water	1118·3
Fungin	95·68
Animal matter insoluble in alcohol	18
Osmazome	12
Vegetable albumen	7·2
Fungate of potash	6
Adipocire	1·2
An oily matter	1·12
Sugar of mushrooms	0·5
Phosphate of potash	Trace

1260·0 §

9. *Boletus viscidus*.—Braconnot found this boletus composed almost entirely of an animal mucus, which acquires cohesion when heated, and becomes in part insoluble in water. viscidus,

* Ann. de Chim. lxxxv. 23.

† Ibid. p. 24.

‡ Ibid.

§ Ibid. lxxxvii. 237.

|| Ibid. lxxix. 302.

Book IV.
Pseudo-ignarius,

10. *Boletus pseudo-ignarius*.—A large boletus of this species was subjected to an elaborate analysis by Braconnot, who extracted from it the following substances:

Water.

Fungin.

A sweet-tasted mucilaginous matter.

Boletate of potash.

A yellow fatty matter.

Vegetable albumen.

Phosphate of potash in small quantity.

Acetate of potash.

Fungic acid in a state of combination.*

Laricis,

11. *Boletus laricis*.—This plant, in a dry state, is used on the continent as a medicine, and sold under the name of *agaric*. It has been examined by Bouillon La Grange.†

It is in pieces, which are white, light, and friable. The outer skin is leathery and dark-coloured. Its taste is at first sweetish, but it leaves a bitter and acrid impression in the mouth. When steeped in water, it communicates a yellowish colour and a sweetish taste to the liquid. The infusion reddens vegetable blues; and holds in solution sulphate of potash, sulphate of lime, and muriate of potash.

When this substance is boiled in water, the liquid acquires a gelatinous form as it cools. Evaporated to dryness, and treated with lime, the odour of ammonia becomes perceptible. Alcohol, boiled upon the boletus, acquires a red colour; and when mixed with water, lets fall a copious precipitate, which exhibits the properties of a resin. This resin has a yellow colour, is brittle, semitransparent, and has a sour and bitter taste. When treated with lime, and the solution afterwards decomposed with muriatic acid, a quantity of benzoic acid is obtained from it. From these experiments it is evident that this substance contains resin, benzoic acid, different salts, some extractive, and some animal matter, to which the gelatinous form of the decoction must be ascribed.

Sulphuric acid dissolves and rapidly chars the boletus. Nitric acid acts with energy; nitrous gas is disengaged, and the boletus becomes brown. By continuing the action of the acid, crystals of oxalic acid are obtained; malic acid is likewise formed, together with some resin, and a substance which approaches the nature of wax in its properties. The fixed alk-

* Ann. de Chim. lxxx. 272.

† Ibid. li. 76.

lies give it a red colour, render it gelatinous, and a great quantity of ammonia is disengaged.*

Chap. II.

12. *Boletus igniarius*.—This boletus is not uncommon in this country on the trunks of trees. A variety of it is used in Germany and some other countries as *tinder*. It is prepared for this purpose by boiling it in a solution of saltpetre, beating it till it becomes sufficiently soft, and then boiling it a second time in the same solution. In 1750 it was recommended by Brossard, a French surgeon, as an excellent styptic; and various trials were made with it, both in France and Britain. It has been subjected to a chemical examination by Bouillon La Grange.†

Boletus igniarius.

When boiled in water the liquid acquires a deep brown colour and a slightly astringent taste. It holds in solution sulphate of lime and muriate of potash. When evaporated to dryness, it leaves a brown-coloured extract, which attracts moisture from the atmosphere. This substance, when incinerated, left a white ash, containing a considerable portion of potash; and when dissolved in water, exhibited, by reagents, the presence of lime, and of muriatic and sulphuric acids. The residual portion of the boletus, being calcined, was found to contain phosphates of lime and magnesia, and some iron.

Alcohol has scarcely any action on this boletus; but when assisted by heat, it dissolves a small portion of resin. Nitric acid dissolves it readily; malic and oxalic acids are formed, and probably also a portion of bitter principle, while carbonic acid and nitrous gas are disengaged. Alkaline leys dissolve it with difficulty; forming, however, a soapy liquid, and separating a small portion of ammonia. From these experiments, we learn that this boletus differs in many respects from the preceding. It contains much less resin, and a much smaller proportion of animal matter, and yields no traces of benzoic acid.‡

13. *Hydnum repandum*.—This hydnum is not uncommon in the woods. It has a yellow colour. Braconnot subjected it to analysis, and extracted from it the following substances:

Hydnum repandum.

Water.

Fungin.

Gelatin in small quantity.

* Bouillon la Grange, Ann. de Chim. li. 76.

† Ann. de Chim. liv. 92.

‡ Bouillon la Grange, Ann. de Chim. liv. 92.

Book IV.

Sugar of mushrooms in considerable quantity.
 Much acetate of potash.
 Fungate of potash.
 A little phosphate of potash.
 Another vegetable alkali united to potash.
 An oil.
 Adipocire.
 A very fugacious acrid matter.
 Muriate of potash.*

Hibridum. 14. *Hydnum hibridum*.—This hydnum has a blackish-brown colour, has not an acrid taste, and when distilled yields no ammonia, but leaves a very alkaline ash. Its constituents according to Braconnot are

Water.	Phosphate of potash.
Fungin.	Fungate of potash.
Gelatin.	Adipocire.
Sugar of mushrooms.	A brown oil.
Vegetable albumen.	A colouring matter.
Acetate of potash.	Muriate of potash.†

Merulius cantbarellus.

15. *Merulius* or *agaricus cantharellus*.—This substance has a yellow colour and an agreeable odour. When taken into the mouth its taste is at first acrid; but soon becomes exceedingly agreeable. It is very much used as an article of food. According to the analysis of Braconnot its constituents are

Water.
 Fungin.
 Gelatin.
 A considerable proportion of sugar of mushrooms.
 Acetate of potash.
 Phosphate of potash.
 Fungate of potash.
 Oil.
 Adipocire.
 An acrid matter destroyed by heat.
 An uncombined acid similar to the acetic.‡

Phallus impudicus.

16. *Phallus impudicus*.—This fungus, well known in consequence of the intolerable odour which it exhales, was found by Braconnot to contain the following constituents:

* Ann. de Chim. lxxix. 291.

† Ibid. p. 297.

‡ Ibid. p. 300.

Water.
 Fungin, much animalized.
 Albumen.
 Animal matter.
 Mucus.
 Superacetate of potash.
 A little superacetate of ammonia.
 Adipocire.
 An oily matter.
 Sugar of mushrooms.
 A little phosphate of potash.
 A combustible acid united to potash.*

17. *Peziza nigra* or *Lycoperdon truncatum*.—This fungus is Peziza ni-
gra. common upon the bark of oaks soon after they have been cut down. Its surface has a black colour, while the fleshy part is gelatinous, semitransparent, and elastic. Four hundred parts of it analysed by Braconnot yielded the following constituents :

Water	376·0
Cerasin	18·4
Gum	3·6
Fungic acid chiefly free	8·0
Sugar of mushrooms	0·4
Matter soluble in alcohol very little animalized ..	0·4
Fatty matter, rendered purple by potash	4
Potash and lime	Trace
	407·2†

18. *Lycoperdon tuber*, or *tuber cibarium*; *truffles*.—This is a Truffles. globular fungus found four or five inches under ground, in woody places, and is considered as one of the best of the eatable mushrooms. Bouillon La Grange has subjected it to a chemical analysis.† It was attempted nearly in the manner formerly described for ascertaining the constituents of the different species of corn.

It was grated down small, and then washed with water upon a sarsee, till the liquid ceased to carry off any thing. A blackish fibrous matter remained upon the sarsee. The liquid let fall a brown coloured matter when left at rest. The liquid produced no effect upon vegetable blues. The brown

* Ann. de Chim. lxxx. 291.

† Ibid. lxxxvii. 253.

‡ Ibid. xlvi. 197.

Book IV.

deposite does not resemble starch in its properties. Water produces but little effect upon truffles. Warm water, however, dissolves a portion, which possesses the characters of albumen.

When the plant was treated with nitric acid, a solution was accomplished; nitrous gas, carbonic acid, and azotic gas were disengaged; and the solution, by distillation, yielded a liquor containing prussic acid. The residual liquid yielded bitter principle, an oily matter, and small crystals, which Bouillon La Grange considered as a combination of oxalic acid and bitter principle. He suspected also the presence of malic acid.

Alcohol dissolves a small portion of brownish bitter matter, which acquires the properties of resin by exposure to the atmosphere.

When truffles are left in the state of a paste with water, they acquire the smell of cheese. When mixed with sugar and water they undergo fermentation, carbonic acid is disengaged, and alcohol formed.

When distilled, truffles yield an acid liquid, a black oil, carbonate of ammonia, carbonic acid, and carburetted hydrogen. The charcoal contains magnesia, phosphate of lime, iron, and silica.*

Several experiments have been likewise made upon this plant by Robert and Antoine, which confirm those of Bouillon La Grange.

Mucor septicus.

19. *Reticularia portensis*, or *mucor septicus*.—This singular vegetable substance appears upon the leaves of ivy and beech, but it is most common upon tan in hot houses. It has no certain size or figure, but has a fine yellow colour and at first resembles cream beat up to froth. In 24 hours it acquires a thin filmy coat, becomes dry and full of a sooty powder adhering to downy threads. According to the analysis of Braconnot its constituents are,

Fungin, very much divided.

A notable quantity of carbonate of lime.

Albumen.

An animal matter.

A yellow fatty matter.

Acetate of potash.†

* Bouillon La Grange, Ann. de Chim. xlvi. 197.

† Ann. de Chim. lxxx. 283.

20. *Tremella nostoc*.—I place this singular vegetable substance here on account of its resemblance to the preceding species. It has more similarity to an animal than a vegetable body. Some naturalists indeed are of opinion that it is the spawn of the frog or of some similar animal. Braconnot subjected it to analysis and extracted from 200 parts of it the following ingredients:

Chap. II.
Tremella
nostoc.

Water	185·0	
Cerasin	13·8	
Mucus, having a spermatic smell	1·2	
Fat		} in small quantity.
Phosphate of lime ..		
Carbonate of lime ..		
Muriate of potash ..		
Sulphate of potash ..		
Potash		

200·0*

SECT. XIV.

OF FUCI.

FUCI are well known plants that vegetate in the sea, and by the combustion of which the impure alkaline body called Kelp is obtained. Great quantities of fuci are burnt for this manufacture on the coasts of Scotland and Ireland. The same substance is made on the coast of Normandy, and is known in France by the name of *varec*. Some of these *fuci* were examined in 1777 by Macquer and Poulletier de la Salle, chiefly in order to determine the different saline substances which they contain, and the result of their experiments was published by Macquer in the second edition of his dictionary.† But chemical analysis had at that time made too little progress to enable these gentlemen to determine with accuracy the nature of the different salts which they had to examine and separate. Besides, they did not distinguish between the different species of fuci; but subjected them indiscriminately to the destructive agency of fire.

After the discovery of iodine in kelp the attention of chemists was naturally turned towards the fuci, by means of which

* Ann. de Chim. lxxvii. 265.

† Under the article *Varec*.

Book IV.

kelp is procured. Several of them accordingly were examined in 1814, by Sir H. Davy, who detected traces of iodine in their ashes.* In 1815, an elaborate analysis of six different species of fucus was published by M. Gaultier de Claubry.† During the same year Dr. John, actuated by the same motive, a desire to determine whether iodine could be detected in fuci, made a set of experiments on the fucus vesiculosus, which he likewise published.‡ Vauquelin had previously made some observations on the saccharine matter contained in different fuci, and had shown that it possessed the characters of manna, and this was fully confirmed by the subsequent experiments of M. Gaultier de Claubry. I shall in this section give the constituents of the different species of fucus hitherto subjected to analysis.

Fucus saccharinus,

1. *Fucus saccharinus*.—This which is very common on our coasts consists of a single linear elliptic leaf, without any midrib. It has a tawny-green colour, and when full grown is about 5 feet long and 3 inches wide. But it varies prodigiously in size. Gaultier de Claubry analysed this fucus by treating it first with water, then with alcohol. He examined it also by means of sulphuric acid, nitric acid, and potash. Finally he subjected it to combustion and analyzed the ashes which remained behind. The constitution of it was found very complicated, as no fewer than 21 ingredients were extracted from it. The following were the substances:

A saccharine matter.	Manna.	Muriate of magnesia.
Mucilage in considerable quantity.		Hyposulphite of soda. §
Vegetable albumen.		Carbonate of potash.
Green colouring matter.		Carbonate of soda.
Oxalate of potash.		Hydriodate of potash.
Malate of potash.		Silica.
Sulphate of potash.		Subphosphate of lime.
Sulphate of soda.		Subphosphate of magnesia.
Sulphate of magnesia.		Oxide of iron probably united with phosphoric acid.
Muriate of potash.		Oxalate of lime.
Muriate of soda.		

* Phil. Trans. 1814, p. 505.

† Ann. de Chim. xciii. 75 and 113.

‡ Schweigger's Jour. xiii. 464.

§ I am not acquainted with this salt, nor do I know any method by which it could be formed. Probably the salt obtained by M. Gaultier de Claubry was sulphite of soda.

|| Ann. de Chim. xciii. 83.

Chap. II.
Digitatus,

2. *Fucus digitatus*.—This fucus is olive-coloured: it consists of a cylindrical stem sometimes as thick as a walking stick and about two feet long. The summit of the stalk suddenly dilates into a plain broad leaf, sometimes about a foot in breadth and a foot and a half in length, and divided almost to the base into sword-shaped, ribless segments, from four to about twelve in number. It is much employed in Scotland as a manure for the land. It was subjected to analysis by Gaultier de Claubry, who found in it the same constituents, as in the preceding species, but in different proportions. The quantity of hydriodate of potash which it contained was much smaller.*

3. *Fucus vesiculosus*.—This very common fucus consists of a dichotomous leaf, the edges of which are entire, and in the disc of which near the edges are immersed a number of air bladders about the size of a hazel nut, the use of which seems to be to cause the leaf to float in the water. From the analysis of Gaultier de Claubry, it appears to contain nearly the same constituents as the preceding species. He did not obtain the mucilaginous matter which the fucus saccharinus yields in such abundance. It contains less malic acid and less iodine than the fucus saccharinus, but its other constituents are the same.† John from 100 parts of this fucus in a dry state obtained the following constituents:

A brownish-red slimy matter	}	4
Flesh-red extract with some sulphate and muriate of soda		
A peculiar acid	—	
Resinous fat	2	
Sulphate of soda with some common salt	3·13	
Sulphate of lime with much sulphate of magnesia and some phosphate of lime	12·87	
Trace of oxides of iron and manganese	—	
Membranous matter or albumen of fuci	78	
Silica?		
		100·00 †

4. *Fucus serratus*.—This fucus, like the preceding, is a dichotomous leaf; but its edges are serrated, and it wants the air bladders which characterize the vesiculosus. When analyzed by Gaultier de Claubry, it yielded abundance of

* Ann. de Chim. xciii. 113.

† Ibid. 116.

‡ Schweigger's Jour. xiii. 464.

Book IV. mucilage and albumen. Alcohol separated from it a bitter substance of a brownish-red colour; likewise a green matter which precipitated when the alcohol was cooling, and which had a bitter and disagreeable taste. The salts are the same as in the preceding species; but it contains more iodine than *fucus digitatus* or *vesiculosus*; it contains likewise a greater quantity of carbonate of soda.*

Siliquosus, 5. *Fucus siliquosus*.—This fucus consists of a waved coriaceous stalk, about four feet long, greatly branched, dark-olive when fresh but quite black when dry. The fructifications have the appearance of little flat pods. When examined by Gaultier de Claubry he obtained abundance of albumen, mucilage having a brownish-red colour, a very bitter substance soluble in alcohol, a substance of a greenish-brown colour, soluble in hot but insoluble in cold alcohol; the same salts as in the preceding species; but very little iodine. When dried it becomes covered with a notable quantity of manna sugar.†

Filum. 6. *Fucus filum*.—This consists of a cylindrical thread from the size of a fiddle string to that of a whip cord, and often twelve feet long. Its colour is olive-green. When dry it becomes nearly black; but by exposure to the air it acquires a yellowish or nearly white colour. When examined by Gaultier de Claubry, he found in it scarcely any sensible quantity of albumen, a good deal of mucilage, a very little green matter soluble in hot alcohol, and precipitating as the liquid cools: The same salts as in the preceding species; but very little iodine.‡

SECT. XV.

OF THE DISEASES OF PLANTS.

HAVING now given such an account of the parts of plants as the present imperfect state of vegetable chemistry permitted, it remains for us to point out the new substances generated in plants by disease, and to describe the properties of these substances. But though the diseases of plants are numerous and well marked, and though the effects which they produce on those vegetables which are employed as articles of food are often severely felt, yet little progress has been

* Ann. de Chim. xciii. 119.

† Ibid. p. 121.

‡ Ibid. p. 123.

made, either in ascertaining the nature of these diseases, or the method of preventing or removing them. The following are the only observations, strictly chemical, which can be at present offered :

1. Old trees are frequently affected with a kind of ulcer, Ulcers. the chief seat of which is under the bark. There the juices undergo a change, and acquire a certain acidity, which enables them to corrode and destroy the solid parts of the plant. We are indebted to Vauquelin for an examination of the sanies or morbid liquid which flows from the ulcer in these cases. He collected the matter which he examined from the elm, which is particularly subject to this disease.*

This liquor is sometimes nearly transparent like water, and having an acrid and saline taste; sometimes it is slightly coloured, sometimes blackish or brown. It deposits upon the sides of the ulcer a soft matter, which is insoluble in water. The bark over which the transparent sanies passes becomes white like chalk, acquires an alkaline taste, and effervesces powerfully with acids. It becomes friable, loses its fibrous texture, and exhibits the appearance of crystals. By means of a glass small rhomboids and four-sided prisms can be distinguished in it. When the liquid is dark-coloured, the bark assumes a black appearance, and looks as if covered with a coat of varnish. This black matter is sometimes in such quantity that it assumes the appearance of stalactites. It is soluble in water, has an alkaline taste, effervesces with acids.

The white matter which was deposited on the bark of the elm round the ulcer, was found composed of

Vegetable matter	60·5
Carbonate of potash	34·2
Carbonate of lime	5·0
Carbonate of magnesia	0·3
	100·0

The black shining matter consisted of carbonate of potash, and a peculiar substance which was obviously ulmin.

The destructive tendency of such ulcers in trees is sufficiently obvious. Vauquelin has shown, that to produce the quantity of white matter which surrounded the ulcer of the

* Ann. de Chim. xxi. 39.

Book IV. elm, and on which he made his experiments, 500 lbs. weight of the wood of the tree must have been destroyed.*

Blight in wheat,

2. The blight is a disease which attacks the different species of corn, especially wheat, and in some seasons nearly destroys the crop. It begins first upon the leaves and stem, and at last, attacking the seed, destroys a considerable part, or even the whole of it, leaving a black matter in its place. Botanists have shown, that this black matter is a species of small fungus, which draws its nourishment from the wheat.† The wheat thus loaded with the black fungus has been lately examined by Fourcroy and Vauquelin, in order to ascertain how far the experiments of former chemists were precise. The result of their examination has been published in the sixth volume of the *Annales de Museum d'Histoire Naturelle*. ‡

When the wheat was macerated with alcohol, it gave out a portion of acrid oil, of the consistence of butter, of a deep green colour. Water in which it was steeped acquired acidity, and when saturated with potash let fall an animal matter, mixed with phosphate of ammonia-and-magnesia, and held phosphate of potash in solution: of course, the acid which it contained was the phosphoric. It still retained in solution a matter similar to gluten of wheat allowed to run into putrefaction. The residue, distilled, yielded water, holding in solution acetic acid and acetate of ammonia, a brown oil, and a quantity of charcoal, amounting to nearly $\frac{1}{4}$ th of the original weight. The result of the experiments was, that blighted wheat contained an acrid oil, putrid gluten, chareoal, phosphoric acid, phosphate of magnesia and ammonia, and phosphate of lime; but no traces of starch could be detected.

In barley.

3. Barley is subject to a similar disease with wheat, and, doubtless, from the same cause. Einhof has subjected this grain, damaged by the rubigo, or blight, to a chemical examination. The result of his trials coincides pretty nearly with the experiments of Fourcroy and Vauquelin on blighted wheat. He could detect no starch in it. The infusion of it in water reddened vegetable blues, and contained an acid, which appeared to be the phosphoric. There was present a

* *Ann. de Chim.* xxi. 39.

† See an account of this fungus, with figures of it, and ingenious conjectures respecting its propagation, by Sir Joseph Banks. *Nicholson's Jour.* x. 225.

‡ No. xxxv. p. 332.

peculiar animal matter, differing in its properties from any of the principles of barley, and a quantity of charcoal.* Chap. III.

CHAP. III.

OF VEGETATION.

WE have now seen the different substances which are contained in plants, and of which they are composed; but we have still to examine the manner in which these substances are produced, and to endeavour to trace the different processes which constitute vegetation. But I must warn the reader not to expect complete information in this chapter. The wonders of the vegetable kingdom are still but very imperfectly explored; many of the organs of plants are too minute for our senses; and scarcely a single process can be completely traced.

The multiplicity of operations continually going on in vegetables at the same time, and the variety of different, and even opposite substances, formed out of the same ingredients, and almost at the same place, astonish and confound us. The order, too, and the skill with which every thing is conducted, are no less surprising. No two operations clash; there is no discord, no irregularity, no disturbance; every object is gained, and every thing is ready for its intended purpose. This is too wonderful to escape our observation, and of too much importance not to claim our attention. Many philosophers, accordingly, distinguished equally by their industry and sagacity, have dedicated a great part of their lives to the study of *vegetation*. But hitherto their success has not been equal to their exertions. No person has been able to detect this *Agent*, always so busy, and performing such wonders, or to discover him at his work; nor have philosophers been much more fortunate in their attempts to ascertain the instruments which he employs in his operations. A great variety, however, of curious and interesting facts have been discovered. These I shall attempt in the following sections to collect and arrange, to point out their dependence on each other, and perhaps to deduce such consequences as obviously result from this mutual dependence.

* Gehlen's Jour. vi. 91.

SECT. I.

OF GERMINATION.

Book IV.
Plants produced from seeds.

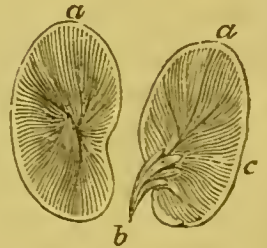
1. NATURAL historians have proved, by a very complete induction of facts, that all plants arise from seeds. The pretended exceptions have disappeared, one after another, as our knowledge of vegetables increased: and now there remains scarcely a single objection entitled to the smallest regard. The late attempt of Girtanner* to revive the doctrine of equivocal generation, deserves no attention whatever; because his conclusions are absolutely incompatible with the experiments of Mr. Sennebier upon the very substance on which his theory is founded.

Seeds composed of three parts;

A seed consists of three parts; namely, the *cotyledons*, the *radicle*, and the *plumula*, which are usually inclosed in a cover.

viz. Cotyledon,

If we take a garden bean, we may perceive each of these three parts with great ease; for this seed is of so large a size, that all its organs are exceedingly distinct. When we strip off the external coats of the bean, which are two, and of different degrees of thickness in different parts, we find that it easily divides into two lobes, pretty nearly of the same size and figure. Each of these lobes is called a *cotyledon* (fig. *a*.) The cotyledons of the bean, then, are two in number.



Radicle,

Near that part of the lobes which is contiguous to what is called the *eye* of the bean, there is a small round white body (*b*), which comes out between the two lobes. This body is called the *radicle*.

Plumula.

Attached to the radicle there is another small round body (*c*), which lies between the cotyledons, and wholly within them, so that it cannot be seen till they are separated from each other. This body is called the *plumula*.

The appearance and shape of these three parts vary much in different seeds, but there is no seed which wants them. The figure and size of the seed depend chiefly upon the cotyledons. This is evidently the case with the bean, and it is so with all other seeds. The number of cotyledons is different in different seeds. Some seeds have only one coty-

* Ann. de Chim. xxxiv. 35.

ledon, as the seeds of wheat, oats, barley, and the whole tribe of grasses; some have three; others six, as the seeds of the garden cress; but most seeds, like the bean, have two cotyledons.

Chap. III.

2. When a seed is placed in a situation favourable to vegetation, it very soon changes its appearance. The radicle is converted into a root, and sinks into the earth, the plumula on the other hand, rises above the earth, and becomes the trunk or stem. When these changes take place, the seed is said to *germinate*: the process itself has been called *germination*. Seeds do not germinate equally and indifferently in all places and seasons. Germination, therefore, is a process which does not depend upon the seed alone: something external must also affect it.

Germination

3. It is a well known fact, that seeds will not germinate unless *moisture* have access to them; for seeds, if they are kept perfectly dry, never vegetate at all, and yet their power of vegetation is not destroyed. *Water*, then, is essential to germination. Too much water, however, is no less prejudicial to most seeds than none at all. The seeds of water plants, indeed, germinate and vegetate extremely well in water; but most other seeds, if they are kept in water beyond a certain time, are rotted and destroyed altogether.

Requires moisture

4. It is well known also, that seeds will not germinate, even though supplied with water, provided the temperature be below a certain degree. No seed, for instance, on which the experiment has been tried, can be made to vegetate at or below the freezing point: yet this degree of cold does not injure the vegetating power of seeds; for many seeds will vegetate as well as ever after having been frozen, or after having been kept in frozen water. We may conclude, then, that a certain degree of heat is necessary for the germination of seeds: and every species of plant seems to have a degree peculiar to itself, at which its seeds begin to germinate; for every seed has a peculiar season at which it begins to germinate, and this season varies with the temperature of the air. Mr. Adanson found that seeds, when sown at the same time in France and in Senegal, always appeared sooner above ground in the latter country, where the climate is hotter than in France.*

Heat,

5. Seeds, although supplied with moisture and placed in a proper temperature, will not germinate, provided atmospheric

And oxygen gas.

* Encyc. Method. Physiol. Veget. 124.

Book IV.

cal air be completely excluded from them. Mr. Ray found that grains of lettuce did not germinate in the vacuum of an air-pump, but they began to grow as soon as air was admitted to them.* Homberg made a number of experiments on the same subject, which were published in the Memoirs of the French Academy for the year 1693. He found, that the greater number of seeds which he tried refused to vegetate in the vacuum of an air-pump. Some, however, did germinate: but Boyle, Muschenbroek, and Boerhaave, who made experiments on the same subject in succession, proved beyond a doubt that no plant vegetates in the vacuum of an air-pump; and that in those cases in which Homberg's seeds germinated, the vacuum was far from perfect, a quantity of air still remaining in the receiver. It follows, therefore, that no seed will germinate unless atmospherical air, or some air having the same properties, have access to it. It is for this reason that seeds will not germinate at a certain depth below the surface of the earth.

Mr. Scheele found that beans would not germinate except oxygen gas were present: Mr. Achard afterwards proved, that oxygen gas is absolutely necessary for the germination of all seeds, and that no seed will germinate in azotic gas, or hydrogen gas, or carbonic acid gas, unless these gases contain a mixture of oxygen gas. These experiments have been confirmed by Mr. Gough, Mr. Cruickshank, and many other philosophers.† It follows, therefore, that it is not the whole atmospheric air, but merely the oxygen gas which it contains, that is necessary for the germination of seeds.

Nay, Mr. Humboldt has ascertained that seeds vegetate more rapidly when steeped in chlorine, or when watered with it; and this substance is well known for the facility with which it decomposes water and sets at liberty oxygen. It seems even to augment the vegetative power of seeds. At Vienna several seeds which had been long kept, and which had constantly refused to germinate, grew readily when treated with it.‡

* Phil. Trans. No liii.

† Carradori, indeed, has lately attempted to show that seeds will begin to germinate without oxygen. (Ann. de Chim. xlvi. 188.) But we learn from the experiments of Saussure, that the water in which the seeds in these cases are steeped, contains in solution as much oxygen as is sufficient to enable the seeds just to commence germination, but not to make any progress. (Recherches Chimiques sur la Vegetation, p. 3.)

‡ Jour. de Phys. xlvii. 63.

6. Light also has considerable influence on the germination of seeds. Ingenhousz found that seeds always germinate faster in the dark than when exposed to the light.* His experiments were repeated by Mr. Sennebier with equal success.† But the Abbé Bertholin, who distinguished himself so much by his labours to demonstrate the effect of electricity on vegetation, objected to the conclusions of these philosophers, and affirmed that the difference in the germination of seeds in the shade and in the light was owing, not to the light itself, but to the difference in the moisture in the two situations; the moisture evaporating much faster from the seeds in the light than from those in the shade; and he affirmed, that when precautions were taken to keep the seeds equally moist, those in the sun germinated sooner than those in the shade.‡ But when Mr. Sennebier repeated his former experiments, and employed every possible precaution to ensure equality of moisture in both situations, he constantly found the seeds in the shade germinate sooner than those in the light.§ We may conclude, therefore, that light is injurious to germination; and hence one reason for covering seeds with the soil in which they are to grow. But from the recent experiments of Saussure, there is reason to believe that light is only injurious in consequence of the heat which it produces; for when the direct rays of the sun were intercepted, though light was admitted, the germination of the seeds was not sensibly retarded.||

Chap. III.
Impeded by
light.

7. Thus we have seen that seeds will not germinate unless *moisture, heat, and oxygen* be present. Now, in what manner do these substances affect the seed? What are the changes which they produce?

It was observed before, that all seeds have one or more cotyledons. These cotyledons contain a quantity of farinaceous matter, laid up on purpose to supply the embryo plant with food as soon as it begins to require it. This food, however, must undergo some previous preparation, before it can be applied by the plant to the formation or completion of its organs. It is probable that all the phenomena of germination, which we can perceive, consist in the chemical changes which

Phenomena
of germination.

* Experiences sur les Vegetaux, ii.

† Mem. Physico-Chimiques, iii. 41.

‡ Jour. de Phys. 1789, December.

§ Encyc. Method. Physiol. Veget. 126.

|| Recherches Chimiques sur la Vegetation, p. 23.

Book IV. are produced in that food, and the consequent developement of the organs of the plant.

Carbonic acid emitted.

Oxygen of the air diminished.

When a seed is placed in favourable circumstances, it gradually imbibes moisture, and very soon after emits a quantity of carbonic acid gas, even though no oxygen gas be present.* If no oxygen gas be present, the process stops here, and no germination takes place; but if oxygen gas be present, a portion of it is converted into carbonic acid gas. From the experiments of Saussure, it appears, that if seeds be left to germinate in a determinate portion of oxygen gas or common air, the bulk of that gas is not altered; the carbonic acid formed being equal to the oxygen which has disappeared. Hence it follows, that the carbonic acid contains in it exactly the whole oxygen consumed.† No oxygen, then, is absorbed by the seed; or at least, if it be absorbed, none of it is retained, the whole being thrown out in combination with carbon. The quantity of oxygen thus changed into carbonic acid by the germination of seeds, is in some measure proportional to the weight of the seed; but some seeds require more than others. In the experiments of Saussure, wheat and barley, weight for weight, consumed less oxygen than peas; while peas consumed less than beans and kidney-beans. The oxygen consumed by wheat and barley amounts to between $\frac{1}{1000}$ th and $\frac{1}{2000}$ th of their weight; while that consumed by beans, and kidney-beans may amount to $\frac{1}{100}$ th part of their weight.‡ Similar experiments were made by Dr. Woodhouse.§

It does not appear that any water is decomposed during the process of germination, at least we have no evidence that it is so. Neither hydrogen nor oxygen gas are emitted. It would not be surprising if a portion of water, so far from being decomposed, were actually formed by the union of its constituents previously existing in the grain. When Saussure dried seeds at a certain temperature before germination, and afterwards brought them to the same degree of dryness after germination, he always found that the loss of weight was greater than it ought to have been. Thus 73 peas, which together

* Gough, Manch. Mem. iv. 315. Cruickshank, Rollo on Diabetes, p. 452.

† Jour. de Phys. xlix. 92.

‡ Recherches, p. 13.

§ He tried the seeds of the zea mays, apium petroselinum, lactuca sativa, cucurbita citrulla, phaseolus sativus, sisymbrium sativum, raphanus sativa. They changed oxygen into carbonic acid. Nicholson's Jour. ii. 161.

(when dried at 77°) weighed 200 grains, when left with five times their weight of water in a vessel full of air, and standing over mercury for two days, germinated, and produced $4\frac{1}{2}$ cubic inches of carbonic acid. They were taken out, and dried slowly at the same temperature; during their drying they changed $4\frac{1}{2}$ cubic inches more of oxygen into carbonic acid. Now these nine inches of carbonic acid contain 1·7 grains of carbon. The water in which peas had been placed while germinating, when evaporated to dryness, left 0·75 grains of mucilaginous matter. Thus the peas, by germinating and drying, ought to have lost only 2·45 grains, leaving a residue of 197·55 grains. But their weight was only 189 grains. So that they had lost eight grains more than can be accounted for by the mucilage taken up by the water and carbonic acid formed. We are obliged to suppose this loss owing to water. From Saussure's experiments, this water seems to be formed or set at liberty during the drying of the seeds, for the quantity of it always increased with the slowness of the process of drying.*

Several seeds, by germinating, acquire a sweetish taste. Hence it has been supposed, that the mucilage which they contain is converted into saccharine matter. We see that the carbon is diminished by germination; and unless water be formed, the proportion of hydrogen and oxygen must be proportionably augmented. It is obvious from the process of malting, that heat is generated during germination. I have seen the radicles of barley, when kept without turning on the malt-floor, shoot out half an inch in a single night, and the heat rise as high as 100°.

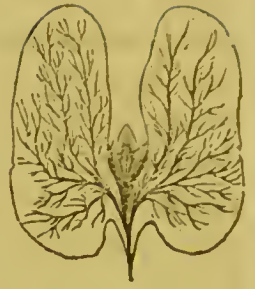
Food prepared in the cotyledons,

So far seems to be the work of chemistry alone; at least we have no right to conclude that any other agent interferes, since hay, when it happens to imbibe moisture, exhibits nearly the same processes. Carbonic acid gas is evolved, oxygen gas is absorbed, heat is produced so abundantly, that the hay often takes fire: at the same time a quantity of sugar is formed. It is owing to a partial change of the same kind that old hay generally tastes much sweeter than new hay. Now we have no reason to suppose that any agents peculiar to the vegetable kingdom reside in hay; as all vegetation, and all power of vegetating, are evidently destroyed.

* Saussure, Recherches Chimiques sur la Vegetation, p. 17.

Book IV.
 And sent
 into the ra-
 dicle.

But when the farina in the seeds of veget-
 ables is converted into sugar, a number of
 vessels make their appearance in the coty-
 ledon. The reader will have a pretty dis-
 tinct notion of their distribution by inspect-
 ing the figure. These vessels may indeed be
 detected in many seeds before germination
 commences, but they become much more



distinct after it has made some progress. Branches from
 them have been demonstrated by Grew, Malpighi, and Hed-
 wig, passing into the radicle, and distributed through every
 part of it. These evidently carry the nourishment prepared
 in the cotyledons to the radicle; for if the cotyledons be cut
 off, even after the processes above described are completed,
 germination, as Bonnet and Sennebier ascertained by experi-
 ment, immediately stops. The food therefore is conveyed
 from the cotyledons into the radicle; the radicle increases in
 size, assumes the form of a *root*, sinks down into the earth,
 and soon becomes capable of extracting the nourishment ne-
 cessary for the future growth of the plant. Even at this period,
 after the radicle has become a perfect root, the plant, as Sen-
 nebier ascertained by experiment, ceases to vegetate if the
cotyledons be cut off. They are still, then, absolutely neces-
 sary for the vegetation of the plant.

Which be-
 comes a
 root.

Seminal
 leaves.

The cotyledons now assume the appearance of leaves, and
 appear above the ground, forming what are called the *seminal*
leaves of the plant. After this the *plumula* gradually increases
 in size, rises out of the earth, and expands itself into branches
 and leaves. The seminal leaves, soon after this, decay and
 drop off, and the plant carries on all the processes of veget-
 ation without their assistance.

As it does not appear that there is any communication be-
 tween the cotyledons and the plumula, it must follow that the
 nourishment passes into the plumula from the radicle; and
 accordingly we see that the plumula does not begin to vege-
 tate till the radicle has made some progress. Since the plant
 ceases to vegetate, even after the radicle has been converted
 into a root, if the cotyledons be removed before the plumula
 is developed, it follows that the radicle is insufficient of itself
 to carry on the processes of vegetation, and that the cotyle-
 dons still continue to perform a part. Now we have seen
 already what that part is; they prepare *food* for the nourish-
 ment of the plant. The root, then, is of itself insufficient for

Digest the
 food sent by
 the root.

this purpose. When the cotyledons assume the form of seminal leaves, it is evident that the nourishment which was originally laid up in them for the support of the embryo plant is exhausted, yet they still continue as necessary as ever. They must therefore receive the nourishment which is imbibed by the root; they must produce some changes on it, render it suitable for the purposes of vegetation, and then send it back again to be transmitted to the plumula. Chap. III.

After the plumula has acquired a certain size, which must be at least a *line*, if the cotyledons be cut off, the plant, as Mr. Bonnet ascertained by a number of experiments, afterwards repeated with equal success by Mr. Sennebier, does not cease to vegetate, but it continues always a mere pigmy: its size, when compared with that of a plant whose cotyledons are allowed to remain, being only as two to seven.*

When the plumula has expanded completely into leaves, the cotyledons may be removed without injuring the plant, and they very soon decay of themselves. It appears, then, that this new office of the cotyledons is afterwards performed by that part of the plant which is above ground.

Thus we have traced the phenomena of germination as far as they have been detected. The facts are obvious; but the *manner* in which they are produced is a profound secret. We can neither explain how the food enters into the vessels, how it is conveyed to the different parts of the plant, how it is deposited in every organ, nor how it is employed to increase the size of the old parts, or to form new parts. These phenomena are analogous to nothing in mechanics or chemistry, but resemble exactly the organization and nourishment of animals. They belong therefore to that difficult branch of science known by the name of Physiology.

SECT. II

OF THE FOOD OF PLANTS.

PLANTS, after they have germinated, do not remain stationary, but are continually increasing in size. A tree, for instance, every season adds considerably to its former bulk. The root sends forth new shoots, and the old ones become larger and thicker. The same increment takes place in the

* Encyc. Method. Physiol. Veget. 12.

Book IV.

branches and the trunk. When we examine this increase more minutely, we find that a new layer of wood, or rather of alburnum,* has been added to the tree in every part, and this addition has been made just under the bark. We find, too, that a layer of alburnum has assumed the appearance of perfect wood. Besides this addition of vegetable fibre, a great number of leaves have been produced, and the tree puts forth flowers, and forms seeds.

Plants require food.

It is evident from all this, that a great deal of new matter is continually making its appearance in plants. Hence, since it would be absurd to suppose that they *create* new matter, it must follow that they receive it by some channel or other. Plants, then, require food as well as animals. Now, what is this food, and whence do they derive it? These questions can only be answered by an attentive survey of the substances which are contained in vegetables, and an examination of those substances which are necessary for their vegetation. If we could succeed completely, it would throw a great deal of light upon the nature of soils and of manures, and on some of the most important questions in agriculture. But we are far indeed at present from being able to examine the subject to the bottom.

Water necessary.

1. In the first place, it is certain that plants will not vegetate without water; for whenever they are deprived of it, they wither and die. Hence the well-known use of rains and dews, and the artificial watering of ground. Water, then, is at least an essential part of the food of plants. But many plants grow in pure water; and therefore it may be questioned whether water is not the only food of plants. This opinion was adopted very long ago, and numerous experiments have been made in order to demonstrate it. Indeed it was the general opinion of the 17th century; and some of the most successful improvers of the physiology of plants, in the 18th century, have embraced it. The most zealous advocates for it were, Van Helmont, Boyle, Bonnet, Duhamel, and Tillet.

* The *wood*, when we inspect it with attention, is not, through its whole extent, the same; the part of it next the bark is much softer and whiter, and more juicy than the rest, and has for that reason obtained a particular name; it has been called the *alburnum* or *aubier*. The *perfect wood* is browner, and harder, and denser than the alburnum, and the layers increase in density the nearer they are to the centre.

Van Helmont planted a willow which weighed five pounds, in an earthen vessel filled with 200 lbs. of soil previously dried in an oven, and moistened with rain water. This vessel he sunk into the earth, and he watered his willow, sometimes with rain, and sometimes with distilled water. After five years it weighed 169½ lbs. and the earth in which it was planted, when again dried, was found to have lost only two ounces of its original weight.* Here, it has been said, was an increase of 16½ lbs. and yet the only food of the willow was pure water; therefore it follows that pure water is sufficient to afford nourishment to plants. The insufficiency of this experiment to decide the question was first pointed out by Bergman in 1773.† He showed, from the experiments of Margraff, that the rain water employed by Van Helmont contained in it as much earth as could exist in the willow at the end of five years. For, according to the experiments of Margraff, one pound of rain water contains one grain of earth.‡ The growth of the willow, therefore, by no means proves that the earth which plants contain has been formed out of water. Besides, as Mr. Kirwan has remarked,§ the earthen vessel must have often absorbed moisture from the surrounding earth, impregnated with whatever substance the earth contained; for unglazed earthen vessels, as Hales|| and Tillet** have shown, readily transmit moisture. Hence it is evident, that no conclusion whatever can be drawn from this experiment; for all the substances which the willow contained, except water, may have been derived from the rain water, the earth in the pot, and the moisture imbibed from the surrounding soil.

The experiments of Duhamel and Tillet are equally inconclusive; so that it is impossible from them to decide the question, Whether water be the sole nourishment of plants or not? But all the attempts hitherto made to raise plants from pure water have failed: the plants vegetating only for a certain time, and never perfecting their seeds. These experiments were made by Hassenfratz, Saussure, and others, with the same unfavourable result. Duhamel found, that an oak, which he had raised by water from an acorn, made less and

Chap. III.
 Supposed
 the sole food
 of plants;

But without
 reason.

* Opera Van Helmont, p. 105. Complexionum atque Mistionum Elementalium Figmentum, Sect. 30.

† Opusc. v. 92.

‡ Ibid. ii. 15 and 19.

§ Irish Trans. v. 160.

|| Veget. Stat. i.

** Mem. Par. 1772, 29.

Book IV.

less progress every year. We see, too, that those bulbous roots, such as hyacinths, tulips, &c. which are made to grow in water, unless they be planted in the earth every other year, refuse at last to flower, and even to vegetate; especially if they produce new bulbous roots annually, and the old ones decay. From all these facts and experiments, it is reasonable to conclude that water is not the sole food of plants.*

A certain proportion only requisite.

So far, indeed, is water from being the sole food of plants, that in general only a certain proportion of it is serviceable, too much being equally prejudicial to them as too little. Some plants, it is true, grow constantly in water, and will not vegetate in any other situation; but the rest are entirely destroyed when kept immersed in that fluid beyond a certain time. Most plants require a certain degree of moisture in order to vegetate well. This is one reason why different soils are required for different plants. Rice, for instance, requires a very wet soil: were we to sow it in the ground on which wheat grows luxuriously, it would not succeed: and wheat, on the contrary, would rot in the rice ground.

We should, therefore, in choosing a soil proper for the plants which we mean to raise, consider the quantity of moisture which is best adapted for them, and choose our soil accordingly. Now the dryness or moisture of a soil depends upon two things; the nature and proportions of the earths which compose it, and the quantity of rain which falls upon it. Every soil contains at least three earths, silica, lime, and alumina, and sometimes also magnesia. The silica is always in the state of sand. Now soils retain moisture longer or shorter according to the proportions of these earths. Those which contain the greatest quantity of sand retain it the shortest, and those which contain the greatest quantity of alumina retain it longest. The first is a dry, the second a wet soil. Lime and magnesia are intermediate between these two extremes: they render a sandy soil more retentive of moisture, and diminish the wetness of a clayey soil. It is evident, there-

* The experiments of Braconnot, who has endeavoured to prove that water is the sole food of plants, are by no means decisive. See *Ann. de Chim.* lxi. 187. He raised plants on sand and metallic oxides by means of water, and found that the plants had all the usual vegetable, earthy, and alkaline constituents. But in experiments of that nature it is impossible to guard against every channel, by means of which these substances may have access to plants.

fore, that by mixing together proper proportions of these four earths, we may form a soil of any degree of dryness and moisture that we please.

But whatever be the nature of the soil, its moisture must depend in general upon the quantity of rain which falls. If no rain at all fall, a soil, however retentive of moisture it be, must remain dry; and if rain were very frequently falling, the soil must be open indeed, if it be not constantly wet. The proportion of the different earths in a soil, therefore, must depend upon the quantity of rain which falls. In a rainy country, the soil ought to be open; in a dry country, it ought to be retentive of moisture. In the first, there ought to be a greater proportion of sand; in the second of clay.

2. That air is necessary for the vegetation of plants has been completely proved; and that a considerable portion at least of the carbonaceous matter which they contain is absorbed from the air, has been rendered probable by the late experiments of chemical philosophers. Three articles which furnish nourishment to plants are supplied by the atmosphere; namely, *carbonic acid*, *oxygen*, and *moisture*; but it has been disputed how far plants are capable of absorbing carbonic acid, without the assistance of the soil on which they usually vegetate.

Plants absorb food from the atmosphere:

Mr. Hassenfratz analysed the bulbous roots of hyacinths, in order to discover the quantity of water, carbon, and hydrogen, which they contained; and by repeating the analysis on a number of bulbs, he discovered how much of these ingredients was contained in a given weight of the bulb. He analysed also kidney beans and cress seeds in the same manner. Then he made a number of each of these vegetate in pure water, taking the precaution to weigh them before hand, in order to ascertain the precise quantity of carbon which they contained. The plants being then placed, some within doors, and others in the open air, grew and flowered, but produced no seed. He afterwards dried them, collecting with care all their leaves and every other part which had dropped off during the course of the vegetation. On submitting each plant to a chemical analysis, he found that the quantity of carbon which it contained, was somewhat less than the quantity which existed in the bulb or the seed from which the plant had sprung.*

These experiments have been repeated by Saussure with a very different result. Sprigs of *mentha piperita*, allowed to

* Ann. de Chim. xiii. 188.

Book IV. } vegetate for some time in distilled water, nearly doubled the portion of carbon which they originally contained;* but when the same experiment was repeated in a place where but little light had access, the carbon, instead of being increased, was somewhat diminished, as Hassenfratz had found. Hence it is probable, that the result obtained by Hassenfratz was owing to the want of light. Thus it cannot be doubted, that plants, even when they vegetate in pure water, are capable of absorbing nourishment from the air; but as plants in this situation cannot produce perfect seeds, and as they gradually decay and cease to vegetate, it is obvious that water and air alone are not sufficient.

But this not sufficient.

3. The third, and only remaining source from which plants can draw their food, is the soil on which they grow. Now this soil consists of two parts; namely, *pure earths*, which constitute its basis, and the *remains* of animals and vegetables applied as manure.

One use of the earthy part of the soil is to furnish a support for the plants, and to administer the proper quantity of water to them; but as all plants contain earthy and saline matters, and as these substances are constantly present, we cannot avoid considering them as in some measure constituting an essential part of plants. Now there is reason to believe that the proportion of earthy and saline matters in plants is considerably influenced by the soil on which they vegetate. Saussure made beans grow in three different situations. The first set was supported by distilled water; the second was planted in sand and supported by rain water; the third set was planted in a pot filled with mould, and placed in a garden. The ashes yielded by these different plants were to each other in the following proportions.†

Proportion of earth in plants varies with the soil.

1. Those fed by distilled water 3·9
2. Those fed by rain water 7·5
3. Those growing in soil 12·0

Here the quantity of fixed matter yielded by the beans which had vegetated in soil was more than three times greater than what was obtained from those which had been fed solely by distilled water.

The same philosopher examined the ashes of various plants

* Recherches Chimiques sur la Vegetation, p. 51.

† Ibid. p. 231.

growing on a granitic and on a calcareous soil.* The plants which grew on the granitic soil contained a considerable proportion of silica and metallic oxides; those which grew in the calcareous, little or none of these bodies, but a greater proportion of calcareous earth than the granitic plants. Thus the pinns abies, the product of each of these soils, yielded the following proportions of fixed bodies:

	Granitic.	Calcar.
Potash	3·60	} 15
Alkaline sulphates and muriates ..	4·24	
Carbonate of lime	46·34	63
Carbonate of magnesia	6·76	0
Silica	13·49	0
Alumina	14·86	16
Metallic oxides	10·52	0
	<hr/>	<hr/>
	99·81	94†

Thus it cannot be doubted, that the proportion of earthy matter contained in plants is considerably influenced by the nature of the soil on which they grow; but whether plants derive the whole of these fixed principles from the soil, or whether they are capable of forming them to a certain extent by the unknown powers of vegetation, are questions not yet finally decided. The experiments of Saussure would lead us to believe, that all the earths found in plants are absorbed from the soil; while those of Schrader seem to prove, that a portion of them is formed by vegetation, even when plants are so situated that they can derive no fixed principle from the soil on which they grow.

Whether plants generate earth.

The Berlin Academy proposed as a prize question, *To determine the earthy constituents of the different kinds of corn, and*

* These soils were composed of the following ingredients:

<i>Granitic.</i>	<i>Calcareous.</i>
Silica	Carbonate of lime..
75·25	98·000
Alumina	Alumina
13·25	0·625
Lime	Oxide of iron
1·74	0·625
Iron and manganese	Petroleum
9·00	00·25
<hr/>	<hr/>
99·24	99·275

† Phil. Mag. viii. 185. Jour. de Phys. lii. 27. See the table of incinerations printed in p. 192 of this volume, where the result of all the experiments of this philosopher is given.

Book IV. *to ascertain whether these earthy parts are formed by the process of vegetation.* The prize was gained by Schrader, an apothecary in Berlin, and the result of his experiments was published by the Academy in 1800. He analysed the seeds of *wheat*, *rye*, *barley*, and *oats*, and ascertained the portion of earth which each contains. He analysed in the same manner rye-straw.* After having in this manner determined the proportion of earth which these seeds contained, he endeavoured to make them grow in some medium which could not furnish any earthy ingredient whatever. For a long time his attempts were baffled, every substance tried containing less or more of earth, and being therefore improper. At last he found that sulphur, in the state of flowers, might be used with success, as it contained no earthy matter whatever, and as the seeds grew in it, and sent out their roots perfectly well, when it was properly moistened with water. The oxides of antimony and zinc were the substances which answered best after sulphur. The seeds, then, were planted in sulphur, placed in a garden at a distance from all dust, put into a box to which the light and the air had free access, but from which all dust and rain were carefully excluded, and they were watered with distilled water. The corn raised in this manner was found by Schrader to contain *more* earthy matter than had existed in the seeds from which it had grown.† Here, then, it would appear was the formation of earthy matter, unless we conceive that the air might have contained a sufficient quantity floating in it to furnish all that was found. Since the publication of his prize dissertation, Schrader has given to the public additional experiments on the same subject. In these he notices the trials of Saussure, and the results which were obtained from vegetables growing in calcareous and granitic soils; and particularly draws the attention of chemists to the fact ascertained by Saussure, that plants vegetating in a calcareous soil, which contained little or no silica, were yet found to yield a considerable portion of that earth.‡ Einhof likewise found in the ashes of a *pinus sylvestris*, which had grown in a soil that yielded no traces of lime, no less than 65 per cent. of that earth. He informs us also, that he has frequently observed

Experiments of Schrader.

* The result of these analyses have been given in page 191 of this Volume.

† Braconnot's experiments agree with those of Schrader, but they were not made with the same precautions. See Ann. de Chim. lxi. 187.

‡ Gehlen's Jour. iii. 538.

the *lichen prunastri* and *ciliaris* encrusted in the months of August and September with a coat of carbonate of lime, when no lime was to be found in the neighbourhood, and when other plants were not encrusted in that manner.* Chap. III.

Saussure has observed very justly, that the absorption of earthy matter does not depend so much upon the earths which constitute the basis of the soil on which they grow, as upon the portion of earth held in solution in the liquid part of the soil. This observation will explain several of the facts above stated, but is by no means sufficient to set aside the experiments of Schrader, which go directly to prove that earths are formed by the processes of vegetation.

4. Besides earths, plants always contain a portion of saline matter. Those that grow at a distance from the sea contain potash, while sea-plants contain soda and common salt. Phosphate of lime is a constant ingredient, and phosphate of potash a very common one. Some plants contain peculiar salts. Thus nitrate of soda is usually present in barley, nitrate of potash in nettles and the sun-flower, &c. It appears from the experiments that have been made, that peculiar salts promote the vegetation of peculiar plants. Sea-plants require common salt, and languish in soils where it is not to be found. Borage, nettles, and pellitory, thrive only in soils which contain nitrate of lime or potash; gypsum promotes the vegetation of lucern and clover.† Soil contains salts,
Which act on plants.

Salts then are not inactive; and when properly applied, they promote the growth of vegetables. In these cases they appear to be absorbed by the plants which they invigorate. Duhamel found that sea-plants made little progress in soils which contain no common salt. Bullion made seeds of sun-flower to grow in a sandy soil containing no trace of nitre. On examining the plants, no nitrate of potash could be detected in them, but the salt made its appearance as usual when the plants were watered with a weak nitrous solution.

From the experiments of Saussure, we learn that plants absorb saline solutions in very different proportions, and that in general those are absorbed in greatest quantity which are most injurious to vegetation. He dissolved the following substances in water, in such proportions, that each solution contained $\frac{1}{100}$ th part of its weight of the substance dissolved, except the last, which contained $\frac{1}{5}$ th part. Plants absorb different proportions of salts.

* Gehlen's Jour. iii. 563.

† Saussure, Recherches, p. 264.

1. Muriate of potash.
2. Muriate of soda.
3. Nitrate of lime.
4. Sulphate of soda effloresced.
5. Muriate of ammonia.
6. Acetate of lime.
7. Sulphate of copper.
8. Crystals of sugar.
9. Gum arabic.
10. Extract of soil.

Into each of these solutions he put plants of *polygonum persicaria*, or of *bidens cannabina*, furnished with their roots.

The polygonum grew for five weeks in the solutions of muriate of potash, nitrate of lime, muriate of soda, sulphate of soda, and extract of soil; and the roots increased in them as usual. It languished in the solution of sal-ammoniac, and the roots made no progress. It died in eight or ten days in the solutions of gum and acetate of lime, and in less than three days in the solution of sulphate of copper.

When such a number of plants of polygonum were put into the solutions as to absorb one half of each in two days, the remaining half was found to have lost very different proportions of the salt which it had originally contained. Suppose the portion of salt originally in solution to be 100, the following table exhibits the quantity of each which had disappeared when one half of the liquid was absorbed :

Muriate of potash	14·7
Muriate of soda	13·0
Nitrate of lime	4·0
Sulphate of soda	14·4
Muriate of ammonia	12·0
Acetate of lime	8·0
Sulphate of copper	47·0
Sugar	29·0
Gum	9·0
Extract of soil	5·0

The *bidens* absorbed pretty much the same proportions, but in general did not vegetate so long as the polygonum. In these trials, it was the sulphate of copper and the sugar that were absorbed in greatest abundance, and these were the substances which proved most injurious to the plant. Saussure

explains this apparent anomaly by supposing that a portion of the roots were soon destroyed in these liquids, and that then they absorbed the solution indiscriminately.*

When various salts were dissolved at once in the same solutions, and plants made to vegetate in them, it was found that different proportions of the salts were absorbed. The following table exhibits the result of these trials, supposing, as before, the original weight of each salt to have been 100. Each solution contains $\frac{1}{10}$ th part of its weight of each salt.

1	{ Sulphate of soda effloresced ..	11·7
	{ Muriate of soda	22·0
2	{ Sulphate of soda effloresced ..	12·0
	{ Muriate of potash	17·0
3	{ Acetate of lime	8·0
	{ Muriate of potash	33·0
4	{ Nitrate of lime	4·5
	{ Muriate of ammonia	16·5
5	{ Acetate of lime	31·0
	{ Sulphate of copper	34·0
6	{ Nitrate of lime	17·0
	{ Sulphate of copper	34·0
7	{ Sulphate of soda	6·0
	{ Muriate of soda	10·0
	{ Acetate of lime	0·0
8	{ Gum	26·0
	{ Sugar	34·0

These experiments succeeded nearly equally with other plants, as *the mentha piperita* and the Scotch fir. When the roots were cut or removed, the plants absorbed all solutions indiscriminately. On examining the plants, the salts absorbed were found in them unaltered.

Thus it appears that plants do not absorb saline bodies indiscriminately. Saussure supposes that the difference depends rather upon the degree of liquidity which the solution possesses, than upon any discriminating power in the root. But if this were the case, it would be difficult to explain how so much greater a proportion of water should be absorbed than of the salt which it holds in solution.

* Saussure, Recherches, p. 247, 261.

Book IV.
 Manure necessary for the growth of plants.

5. Water, then, carbonic acid, and oxygen, and perhaps also earths and salts, constitute a part of the food of plants; but it is very clear that the whole food is not furnished by these substances. It is well known, that if vegetables be successively raised on the same ground, they at last exhaust it, or render it sterile; and to prevent this, farmers are obliged to supply their grounds annually with a quantity of manure. Without this manure, or some equivalent, plants cannot be made to thrive, or to perfect their seeds. Neither water, air, nor earths, nor salts, will prevent them from perishing. Giobert mixed together the four earths, silica, alumina, lime, and magnesia, in the proper proportions to constitute a fertile soil; and after moistening them with water, planted several vegetables in them: but none of them grew well till he moistened his soil with water from a dunghill. Lampadius planted different vegetables in compartments of his garden, filled each with one of the pure earths, and watered them with the liquor which exuded from a dunghill. They all grew, notwithstanding the diversity of the soil; and each contained the usual earthy constituents of plants, notwithstanding the absence of these constituents from the soil.

It is not the earths which constitute a fertile soil, but the remains of animal and vegetable substances, and the proportion of these capable of being held in solution by water. It appears from the experiments of Mr. Hassenfratz, that substances employed as manures produce effects in times proportioned to their degree of putrefaction; those substances which are most putrid producing the most speedy effects, and of course soonest losing their efficacy. Having manured two pieces of the same kind of soil, the one with a mixture of dung and straw highly putrefied, the other with the same mixture newly made, and the straw almost fresh, he observed, that during the first year, the plants which grew on the land manured with the putrefied dung produced a much better crop than the other: but the second year (no new dung being added), the ground which had been manured with the unputrefied dung produced the best crop; the same thing took place the third year, after which, both seemed to be equally exhausted.* Here it is evident that the putrefied dung acted soonest, and was soonest exhausted. It follows from this, that carbon only acts as a manure when in a parti-

* Ann. de Chim. xiv. 57.

cular state of combination; and this state, whatever it may be, is evidently produced by putrefaction. Another experiment of the same chemist renders this truth still more evident. He allowed shavings of wood to remain for about ten months in a moist place till they began to putrefy, and then spread them over a piece of ground by way of manure. The first two years this piece of ground produced nothing more than others which had not been manured at all; the third year it was better, the fourth year it was still better, the fifth year it reached its maximum of fertility; after which it declined constantly till the ninth, when it was quite exhausted.* Here the effect of the manure evidently depended upon its progress in putrefaction.

When vegetables are allowed to putrefy in the open air, they are converted into a loose black substance, well known under the name of *vegetable mould*. On this mould plants grow with great vigour. It is the substance which renders newly cultivated lands in America, &c. so fertile. When exposed to the air, in the course of cultivation, it is gradually wasted and destroyed, and the lands are thus impoverished. This vegetable mould, therefore, is obviously one of the grand sources of the food of plants. It deserves, therefore, an accurate examination.

Vegetable manure;

To Saussure and Einhof we are indebted for a chemical examination of its properties and constituents. Saussure employed in his experiments pure vegetable soil, which he procured either from the trunks of trees or from elevated rocks, where it was unmixed with any animal matter; and by passing it through a sieve, he removed from it the remains of undecayed vegetables with which it was mixed.† By distilling 200 grains (French) of mould from the oak, he obtained the following products; while the same quantity of undecayed oak yielded him the following proportions of the same constituents:

Its properties,

	Mould.	Oak.
Carburetted hydrogen gas . . .	124 inches (French)	116
Carbonic acid	34	29
Water, containing pyrolig- } nate of ammonia }	53 grains (French)	80
Empyreumatic oil	10	13
Charcoal	51	41½
Ashes	8	0½

* Ann. de Chim. xiv. 58.

† Recherches sur la Vegetation, p. 162.

Book IV.

Nearly the same results were obtained by making similar experiments on other vegetables, and the mould which they yield when decayed. It appears from them, that the mould contains more charcoal, weight for weight, than the vegetables from which it proceeded. It yields also more ammonia, and therefore contains more azote.

Acids do not act powerfully on mould. Alcohol takes up a portion of extractive and resin. The fixed alkalies dissolve it almost completely, and ammonia is disengaged during the solution. Acids throw down a small portion of a brown combustible powder. Water dissolves a portion of extractive; but the quantity is small, especially from the mould of fertile soils. Ten thousand parts of water, left for five days on the mould of turf, were found to have dissolved only 26 parts of extract; and the same quantity, left on the soil of a field which bore a fine crop of wheat, was found to have dissolved but four parts. It appears from the experiments of Saussure, that a mould which yielded to boiling water, by repeated decoctions, $\frac{1}{11}$ th of its weight of extract, did not produce so good an effect upon beans as a mould which contained only half that quantity of soluble matter. But when mould is deprived of its soluble part by boiling water, though its appearance is not altered, yet it does not support plants nearly so well as mould which has not been thus exhausted.*

The extract thus obtained from mould by Saussure did not deliquesce when exposed to the air. When distilled it yielded carbonate of ammonia. It produced no change on vegetable blues. When reduced to the consistence of a syrup it had a sweet taste; it precipitated by exposure to the air, and became turbid when mixed with lime water, carbonate of potash, and most metallic solutions. Alcohol dissolves only a portion of it, and this part is very deliquescent.†

And constituents.

From the experiments of Einhof, it appears that the extract obtained from mould possesses very nearly the properties of the *extractive principle*. The mould which he employed in his experiments was from the soil of a wood, and had been formed by the leaves of the trees and the putrefied herbs. It was black, firm, produced no change on vegetable blues, and contained no undecayed plants.‡ Water in which this mould was boiled was at first colourless; but by exposure to the air

* Recherches, p. 170.

† Ibid. p. 174.

‡ Geblen's Jour. vi. 373.

it acquired a brownish tint. The substance which is held in solution possessed exactly the characters of extractive.

Experiments upon vegetable mould have been made also by Braconnot. He found no portion of it soluble in water. His other results resembled those of Saussure. When an alkaline ley was boiled with mould a portion was dissolved. The residue had the exact appearance of pit coal.*

Besides this fertile vegetable mould, Einhof has examined another of a different nature, to which he has given the name of *acid vegetable mould*. It occurs in low-lying meadows and marshes, and the plants which grow upon it in these situations are the different species of *carex*, *juncus*, and *eriphorum*. It constitutes also the principal part of the mould in high-lying situations, and moors where the soil is covered with heath (*erica vulgaris*.) This mould is distinguished from the preceding by containing a notable portion of phosphoric and acetic acids, which give it the property of reddening vegetable blues. The extractive which it contains is chiefly insoluble in water.† This sour vegetable mould bears a considerable resemblance to peat, into which indeed it probably passes. Like it, peat contains a portion of phosphoric acid, and probably also of acetic acid, and an extractive readily soluble in alkalies, though but sparingly in water. ‡

Acid vegetable mould.

Einhof has observed, that acid vegetable mould never occurs in those soils which abound in lime, and that it is counteracted and brought to the state of good mould by the action of lime and marl. It is clear that these manures will neutralize the acids, and thus enable the extractive, and other vegetable substances, to be acted upon by the atmosphere, to yield carbonic acid, and to assume those states which are proper for the nourishment of vegetables. It is probable that they act also directly upon the vegetable matter, and occasion decompositions favourable for vegetation. Hence the efficacy of lime when applied to peat moss, and to sour lands in general.

Effect of lime on mould.

Upon the whole, then, it appears, that plants are fed chiefly by that portion of vegetable matter which becomes soluble in water, and assumes the properties of extractive; that the quantity of it in soil must neither be too great nor too small; that the insoluble part of vegetable mould gradually assumes

* Ann. de Chim. lxi. 191.

† Gehlen, vi. 379.

‡ Einhof, Gehlen's Jour. iii. 400

Book IV. this taste, either by the action of the atmosphere, or of earths or salts; that the presence of an acid, by counteracting this change, injures the nourishing quality of vegetable mould; and that lime is serviceable, partly by neutralizing the acid, and partly by accelerating the decomposition of vegetable mould.

Animal manures.

Animal manures probably supply similar materials with vegetable mould. They probably prevent the formation of acids, or neutralize them when formed: and likewise promote the decomposition and solubility of vegetable matter. The striking effects which animal manure produces on the growth of vegetables are well known; though it is not possible at present to point out the way in which they act. Two elaborate and ingenious sets of experiments on this subject have been published. The first, by Einhof and Thaer, consists of an examination of the excrements of black cattle;* and the second, by Berzelius, is upon the human excrementitious matter.† These add considerably to our knowledge of animal substances, and shall therefore be noticed in a subsequent part of this Work; but as they throw little light on the use of these matters as manures, it is unnecessary to give a detail in this place of the constituents of which they found them composed.

Food absorbed by the ends of the roots.

Such is the present state of our knowledge respecting the food of plants, as far as it is supplied by the soil in which they vegetate. It is probable that it is imbibed by the extremities of the roots only; for Duhamel observed, that the portion of the soil which is soonest exhausted is precisely that part in which the greatest number of the extremities of roots lies.‡ This shows us the reason why the roots of plants are continually increasing in length. By this means they are enabled, in some measure, to go in quest of nourishment. The extremities of the roots seem to have a peculiar structure adapted for the imbibing of moisture. If we cut off the extremity of a root, it never increases any more in length: therefore its use as a root has been in a great measure destroyed. But its sides send out fibres which act the part of roots, and imbibe food by their extremity. Nay, in some cases, when the extremity of a root is cut off, the whole decays, and a new one is formed in its place. This, as Dr. Bell informs us, is the case with the hyacinth.§

* Gehlen's Jour. iii. 276.

† Ibid. vi. 509.

‡ Physique des Arbres, iii. 239.

§ Manch. Mem. ii. 412.

The extremities of the roots contain no visible opening. Hence we may conclude that the food which they imbibe, whatever it may be, must be in a state of solution; while the absolute necessity of water renders it probable that water is the solvent. And, in fact, the carbonaceous matter in all active manures is in such a state of combination that it is soluble in water. All the salts which we can suppose to make a part of the food of plants, are more or less soluble in water. This is the case also with lime, whether it be pure or in the state of a salt; magnesia and alumina may be rendered so by means of carbonic acid gas; and Bergman, Macie, and Klaproth, have shown that even silica may be dissolved in water. We can see, therefore, in general, though we have no precise notions of the very combinations which are immediately imbibed by plants, that all the substances which form essential parts of that food *may* be dissolved in water.

Chap. II I.
It must be
in a state of
solution.

SECT. III.

OF THE MOTION OF THE SAP.

SINCE the food of plants is imbibed by their roots in a fluid state, it must exist in plants in a fluid state; and unless it undergoes alterations in its composition just when imbibed, we may expect to find it in the plant unaltered. If there were any method of obtaining this fluid food from plants before it has been altered by them, we might analyse it, and obtain by that means a much more accurate knowledge of the food of plants than we can by any other method. This plan indeed must fail, provided the food undergoes alteration just when it is absorbed by the roots: but if we consider, that when one species of tree is grafted upon another, each bears its own peculiar fruit, and produces its own peculiar substances, we can scarcely avoid thinking that the *great* changes at least which the food undergoes after absorption, are produced, not in the roots, but in other parts of the plant.

If this conclusion be just, the food of plants, after being imbibed by the roots, must go directly to those organs where it is to receive new modifications, and to be rendered fit for being assimilated to the different parts of the plant. There ought therefore to be certain juices continually ascending from the roots of plants; and these juices, if we could get them pure and unmixed with the other juices or fluids which the plant must contain, and which have been secreted and formed from these primary juices, would be, very nearly at least, the

Sap of
Plants.

Book IV. food as it was imbibed by the plant. Now, during the vegetation of plants, there actually is a *juice* continually ascending from their roots. This juice has been called the *sap*, the *succus communis*, or the *lymph* of plants.

Flows out in
the spring.

The sap is most abundant during the spring. At that season, if a cut be made through the bark and part of the wood of some trees, the sap flows out very profusely. The trees are then said to *bleed*. By this contrivance any quantity of sap we think proper may be collected. It is not probable, indeed, that by this method we obtain the ascending sap in all its purity; it is no doubt mixed with the peculiar juices of the plant: but the less progress vegetation has made, the purer we may expect to find it; both because the peculiar juices must be in much smaller quantity, and because its quantity may be supposed to be greater. We should therefore examine the sap as early in the season as possible, and at all events before the leaves have expanded.

Though a number of experiments, as we have seen in the last chapter, have been made upon the sap, they are not of such a nature as to throw much light upon the food absorbed by plants. Chemistry has not made such progress as to enable even the most expert analysis to separate and distinguish small quantities of the vegetable matter. It is even possible that the food, after it has been imbibed, may be, to a certain extent, modified and altered by the roots. In what manner this is done we cannot say, as we know very little about the vascular structure of the roots. We may conclude, however, that this modification is nearly the same in most plants: for one plant may be grafted on another, and each continue to produce its own peculiar products; which could not be, unless the proper substances were conveyed to the digestive organs of all. There are several circumstances, however, which render the modifying power of the roots somewhat probable. It is even possible that the roots may, by some means or other, throw out again some part of the food which they have imbibed as excrementitious. This has been suspected by several physiologists; and there are several circumstances which render it probable. It is well known that some plants will not vegetate well after others; and that some again vegetate unusully well when planted in ground where certain plants had been growing. These facts, without doubt, may be accounted for on other principles. If there be any excrementitious matter emitted by the roots, it is much more pro-

bable that this happens in the last stage of vegetation; that is to say, when the food, after digestion, is applied to the purposes which the root requires. But the fact ought to be supported by experiments, otherwise it cannot be admitted.

Chap. III.

Mr. Knight, to whose ingenious and important experiments and observations vegetable physiology lies under so many obligations, has rendered it extremely probable that the sap, as it ascends, is mixed with a quantity of matter, previously deposited in the alburnum for that purpose, and ready prepared to be assimilated to the different vegetable organs.* According to him, plants, after they have attained their full growth, are employed during the latter part of summer in preparing food for the expanding of the buds and blossoms in the succeeding spring. This food, when prepared, is deposited in the alburnum. Here it continues during the winter: and in next spring, mixing with the ascending sap, it affords nourishment to the buds and leaves. This ingenious opinion Mr. Knight has supported by experiments and observations, sufficient, I think, to establish its truth. It constitutes a very important step in vegetable physiology, as it enables us to explain many circumstances in a satisfactory manner that appeared formerly altogether anomalous.

Mixed with true sap in its ascent.

He ascertained, by experiments, that the sap increases in density as it ascends towards the leaves. Sap extracted from the sycamore, close to the ground, was of the specific gravity of 1.004; while that which flowed out at the height of seven feet was 1.008, and at the height of 12 feet 1.012. The sap of the birch was somewhat lighter, but its comparative increase of density, according to its height, was the same. When extracted near the ground, the sap, both of the sycamore and birch, was nearly tasteless; but it became sensibly sweet at some height, and the sweetness increased with the distance from the ground. Thus it appears, that the quantity of vegetable matter in sap increases as it flows towards the leaves; a direct proof that it imbibes and mixes with something during its passage. That this matter was lodged in the alburnum was rendered probable by comparing with each other the alburnum in winter and in summer. For if nutriment be laid up in the alburnum in winter, and employed in summer for the purposes of vegetation, it is obvious that the alburnum

* On the State in which the True Sap of Trees is deposited during Winter. Phil. Trans. 1655.

Book IV.

during winter ought to be denser, and ought to yield more extract to water, than the same substance in summer: both of which Mr. Knight found to be the case. Oak poles, of the same age, and growing from the same stool, were felled, partly in December and partly in May. They were placed in the same situation, and dried for seven weeks by a fire. The specific gravity of the winter felled wood was 0·679, of the summer 0·609. When the alburnum alone was weighed, the specific gravity of the winter felled was 0·583; of the summer felled 0·533. One thousand grains of each being mixed with six ounces of boiling water, and left to macerate for 24 hours, the winter felled infusion was much deeper coloured than the other. Its specific gravity was 1·002; while that of the summer felled infusion was 1·001. This deposition of nutritious matter explains why the alburnum of trees felled in winter is much more solid and valuable than the alburnum of trees felled in summer.

The sap, as Dr. Hales has shown us, ascends with a very considerable force. It issued during the bleeding season with such impetuosity from the cut end of a vine branch, that it supported a column of mercury $32\frac{1}{2}$ inches high.*

Now, what is the particular channel through which the sap ascends, and what is the cause of the force with which it moves? These are questions which have excited a great deal of the attention of those philosophers who have made the physiology of vegetables their particular study; but the examination of them is attended with so many difficulties that they are very far from being decided.

Sap ascends, It is certain that the sap flows from the roots towards the summit of the tree. For if in the bleeding season a number of openings be made in the tree, the sap begins first to flow from the lowest opening, then from the lowest but one, and so on successively, till at last it makes its appearance at the highest of all; and when Duhamel and Bonnet made plants vegetate in coloured liquors, the colouring matter, which was deposited in the wood, appeared first in the lowest part of the tree, and gradually ascended higher and higher, till at last it reached the top of the tree, and tinged the very leaves.

Through the wood,

It seems certain too that the sap ascends through the wood, and not through the bark of the tree: for a plant continues to grow even when stripped of a great part of its bark; which

* Veg. Stat. i. 105.

could not happen if the sap ascended through the bark. When an incision, deep enough to penetrate the bark, and even part of the wood, is carried quite round a branch, provided the wound be covered up from the external air, the branch continues to vegetate as if nothing had happened; which could not be the case if the sap ascended between the bark and the wood. It is well known, too, that in the bleeding season little or no sap can be got from a tree unless our incision penetrate deeper than the bark. Cchap. III.

These conclusions have been confirmed by the experiments made by Coulomb and Knight. Coulomb observed that no sap ever flows from the poplar till the tree be cut nearly to the centre.* Mr. Knight observed that coloured infusions always pass through the alburnum, and that whenever the alburnum is cut through the plant dies.†

As the sap is never found in the parenchyma, it must of necessity be confined in particular vessels; for if it were not, it would undoubtedly make its appearance there. Now, what are the vessels through which the sap ascends? In peculiar vessels.

Grew and Malpighi, the first philosophers who examined the structure of plants, took it for granted that the woody fibres were tubes, and that the sap ascended through them. For this reason they gave these fibres the name of *lymphatic vessels*. But they were unable, even when assisted by the best microscopes, to detect any thing in these fibres which had the appearance of a tube; and succeeding observers have been equally unsuccessful. The conjecture therefore of Malpighi and Grew, about the nature and use of these fibres, remains totally unsupported by any proof. Duhamel has even gone far to overturn it altogether: for he found that those woody fibres are divisible into smaller fibres, and these again into still smaller; and even, by the assistance of the best microscopes, he could find no end of this subdivision.‡ Now, granting these fibres to be vessels, it is scarcely possible, after this, to suppose that the sap really moves through tubes, whose diameters are almost infinitely small. There are, however, vessels in plants which may easily be distinguished by the help of a small microscope, and even in many cases by the naked eye. These were seen, and distinctly described, by Grew and Malpighi. They consist of a fibre twisted round like a cork-

* Jour. de Phys. xlix. 392.

† Phil. Trans. 1801, p. 336.

‡ Physique des Arbres, i. 57.

Book IV.

screw. If we take a small cylinder of wood, and wrap round it a slender brass wire, so closely that all the rings of the wire touch each other, and if, after this, we pull out the wooden cylinder altogether, the brass wire thus twisted will give us a very good representation of these vessels. If we take hold of the two ends of the brass wire thus twisted, and pull them, we can easily draw out the wire to a considerable length. In the same manner, when we lay hold of the two extremities of these vessels, we can draw them out to a great length. Malpighi and Grew, finding them always empty, concluded that they were intended for the circulation of the air through the plant, and therefore gave them the name of *tracheæ*; which word is used to denote the *windpipe* of animals. These *tracheæ* are not found in the bark; but Hedwig has shown that they are much more numerous in the wood than was supposed; and that they are of very different diameters; and Reichel has demonstrated that they go to the minutest branches, and spread through every leaf. He has shown, too, that they contain sap; and Hedwig has proved that the notion which generally prevailed of their containing nothing but air, arose from this circumstance, that the larger *tracheæ*, which alone were attended to, lose their sap as soon as they are cut; and of course, unless they are inspected the instant they are divided, they appear empty.* Is it not probable, then, from the discoveries of that very ingenious physiologist, that the *tracheæ* are, in reality, the sap vessels of plants? Indeed it seems established, by the experiments both of Reichel and Hedwig, that almost all the vessels of plants may, if we attend only to their structure, be denominated *tracheæ*. But whether the structure of the sap vessels be similar to that of the *tracheæ* or not, it has been perfectly ascertained that the sap ascends in vessels, and that these vessels are situated in the wood, chiefly indeed in the *alburnum*. Hence Mr. Knight has given them the name of *alburnous* vessels.

Its ascension.

But by what powers is the sap made to ascend in these vessels? and not only to ascend, but to move with very considerable force; a force, as Hales has shown, sufficient to overcome the pressure of 43 feet perpendicular of water?†

Ascribed to its levity;

Grew ascribed this phenomenon to the levity of the sap; which, according to him, entered the plant in the state of a

* Fundament. Hist. Nat. Muscov Frondes. Part I. p. 54.

† Veget. Stat. i. 107.

very light vapour. But this opinion will not bear the slightest examination. Malpighi supposed that the sap was made to ascend by the contraction and dilatation of the air contained in the air vessels. But even were we to grant that the tracheæ are air vessels, the sap, according to this hypothesis, could only ascend when a change of temperature takes place; which is contrary to fact. And even if we were to waive every objection of that kind, the hypothesis would not account for the circulation of the sap, unless the sap vessels be provided with valves. Now the experiments of Hales and Duhamel show that no valves can possibly exist in them: for branches imbibe moisture nearly equally by either end; and consequently the sap moves with equal facility both upwards and downwards, which it could not do were there valves in the vessels. Besides, it is known, from many experiments, that we may convert the roots of a tree into the branches, and the branches into the roots, by covering the branches with earth, and exposing the roots to the air.* Now this would be impossible if the sap vessels were provided with valves. The same remarks overturn the hypothesis of Mr. de la Hire, which is merely that of Malpighi, expressed with greater precision, and with a greater parade of mechanical knowledge. Like Borelli, he placed the ascending power of the sap in the parenchyma. But his very experiments, had he attended to them with care, would have been sufficient to show the imperfection of his theory.

The greater number of philosophers (for it is needless to mention those who, like Perrault, had recourse to fermentation, nor those who introduced the weight of the atmosphere) have ascribed the motion of the sap to *capillary attraction*.

There exists an attraction between many solid bodies and liquids; in consequence of which, if these solid bodies be formed into small tubes, the liquid enters them, and rises in them to a certain height. But this is perceptible only when the diameter of the tube is very small. Hence the attraction has been denominated *capillary*. We know that there is such an attraction between vegetable fibres and watery liquids: for

Chap. III.
To the contraction of air;

To capillary attraction.

* Mr. Knight has lately shown that the inverted shoots by no means grow so well as when in their natural position; and has even made it probable that the vessels of the bark are furnished with valves, or with something equivalent. But no evidence has been adduced to induce us to believe that this is the case with the sap vessels.—See his Observations on the Motion of the Sap of Trees. Phil. Trans. 1801

Book IV.

such liquids will ascend through dead vegetable matter. It is highly probable, therefore, that the food of plants enters the roots, in consequence of the capillary attraction which subsists between the sap vessels and the liquid imbibed. This species of attraction, then, will account perfectly well for the entrance of moisture into the mouths of the sap vessels: but will it account also, as some have supposed, for the ascent of the sap, and for the great force with which it ascends?

The nature and laws of capillary attraction have been very much overlooked by philosophers. But we know enough concerning it to enable us to decide the present question. It consists in a certain attraction between the particles of the liquid and of the tube. It has been demonstrated, that it does not extend, or at least that it produces no sensible effect, at greater distances than $\frac{1}{1000}$ th part of an inch. It has been demonstrated, that the water ascends, not by the capillary attraction of the whole tube, but of a slender film of it; and Clairaut has shown that this film is situated at the lowermost extremity of the tube.* This film attracts the liquid with a certain force; and if this force be greater than the cohesion between the particles of the liquid, part enters the tube, and continues to enter, till the quantity above the attracting film of the tube just equals, by its weight, the excess of the capillary attraction between the tube and the liquid above the cohesion of the liquid. The quantity of water, therefore, in the tube is pretty nearly the measure of this excess; for the attracting film is probably very minute.

But without
reason.

It has been demonstrated, that the heights to which liquids rise in capillary tubes are inversely as the diameter of the tube. Consequently the smaller the diameter of the tube, the greater is the height to which the liquid will rise. But the particles of water are not infinitely small; therefore, whenever the diameter of the tube is diminished beyond a certain size, water cannot ascend in it, because its particles are now larger than the bore of the tube. Consequently the rise of water in capillary tubes must have a limit: if they exceed a certain length, how small soever their bore may be, water will either not rise to the top of them, or it will not enter them at all.

* The action of all the other films, of which the tube is composed, on the water, as far as it is measured by its effect, is nothing at all. For every particle of water in the tube (except those attracted by the undermost film) is attracted upwards and downwards by the same number of films: it is therefore precisely in the same state as if it were not attracted at all.

We have no method of ascertaining the precise height to which water would rise in a capillary tube, whose bore is just large enough to admit a single particle of water. Therefore we do not know the limit of the height to which water may be raised by capillary attraction. But whenever the bore is diminished beyond a certain size, the quantity of water which rises in it is too small to be sensible. We can easily ascertain the height which water cannot exceed in capillary tubes before this happens; and if any person calculate, he will find that this height is not nearly equal to the length of the sap vessels of many plants. But besides all this, we see in many plants very long sap vessels, of a diameter too large for a liquid to rise in them a single foot by capillary attraction, and yet the sap rises in them to very great heights.

If any person says that the sap vessels of plants gradually diminish in diameter as they ascend; and that, in consequence of this contrivance, they act precisely as an indefinite number of capillary tubes, one standing upon another, the inferior serving as a reservoir for the superior—I answer, that the sap may ascend by that means to a considerable height: but certainly not in any greater quantity than if the whole sap vessel had been precisely of the bore of its upper extremity; for the quantity of sap raised must depend upon the bore of the upper extremity, because it must all pass through that extremity.

But farther, if the sap moved only in the vessels of plants by capillary attraction, it would be so far from flowing out at the extremity of a branch, with a force sufficient to overcome the pressure of a column of water 43 feet high, that it could not flow out at all. It would be impossible in that case for any such thing as the bleeding of trees ever to happen.

If we take a capillary tube, of such a bore that a liquid will rise in it six inches, and after the liquid has risen to its greatest height, break it short three inches from the bottom, none of the liquid in the under half flows over. The tube, thus shortened, continues indeed full, but not a single particle of liquid ever escapes from it. And how is it possible for it to escape? The film, at the *upper* extremity of the tube, must certainly have as strong an attraction for the liquid as the film at the *lower* extremity. As part of the liquid is within its attracting distance, and as there is no part of the tube above to counterbalance this attraction, it must of necessity attract the liquid nearest it, and with a force sufficient to

Book IV.

counterbalance this attraction of the undermost film, how great soever we may suppose it. Of course no liquid can be forced up, and consequently none can flow out of the tube. Since, then, the sap *flows out* at the upper extremity of the sap vessels of plants, we are absolutely certain that it does not ascend in them merely by its capillary attraction, but that there is some other cause.

It is impossible, therefore, to account for the motion of the sap in plants by any mechanical or chemical principles whatever; and he who ascribes it to these principles has not formed to himself any clear or accurate conception of the subject. We know indeed that heat is an agent; for Dr. Walker found that the ascent of the sap is much promoted by heat, and that after it had begun to flow from several incisions, cold made it give over flowing from the higher orifices while it continued to flow at the lower.* . But this cannot be owing to the dilating power of heat; for unless the sap vessels of plants were furnished with valves, dilatation would rather retard than promote the ascent of the sap.

The vessels
must contract.

We must, therefore, ascribe it to some other cause: the vessels themselves must certainly act. Many philosophers have seen the necessity of this, and have accordingly ascribed the ascent of the sap to *irritability*. But the first person who gave a precise view of the manner in which the vessels probably act was Saussure. He supposes that the sap enters the open mouths of the vessels at the extremity of the roots; that these mouths then contract, and by that contraction propel the sap upwards; that this contraction gradually follows the sap, pushing it up from the extremity of the root to the summit of the plant. In the mean time the mouths are receiving new sap, which in the same manner is pushed upwards.† Whether we suppose the contraction to take place precisely in this manner or not, we can scarcely deny that it must take place; but by what means it is impossible at present to say. The agents cannot precisely resemble the muscles of animals; because the whole tube, however cut or maimed, still retains its contracting power, and because the contraction is performed with equal readiness in every direction.‡ It is

* Edin. Trans. i.

† Encycl. Meth. Phys. Veget. p. 267.

‡ Mr. Knight thinks it probable that the sap is propelled by the contraction and expansion of what is called by carpenters the *silver grain* of the wood, between the lamina of which the vessels run. (Phil. Trans. 1801, p. 544.) By *silver grain* is meant those thin longitudinal fibres, diverging

evident, however, that they must be the same in kind. Perhaps the particular structure of the vessels may fit them for their office. Does ring after ring contract its diameter? The contracting agents, whatever they are, seem to be excited to act by some stimulus communicated to them by the sap. This capacity of being excited to action is known in physiology by the name of *irritability*; and there are not wanting proofs that plants are possessed of it. It is well known that different parts of plants move when certain substances act upon them. Thus the flowers of many open at sunrise, and close again at night. Linnæus has given us a list of these plants. Des Fontaines has shown that the stamina and antheræ of many plants exhibit distinct motions.* Dr. Smith has observed, that the stamina of the barberries are thrown into motions when touched.† Roth has ascertained that the leaves of the *drosera longifolia* and *rotundifolia* have the same property. Mr. Coulomb, too, who has adopted the opinion that the motion of the sap in plants is produced by the contraction of vessels, has even made a number of experiments in order to show this contraction. But the fact is, that every one has it in his power to make a decisive experiment. Simply cutting a plant, the *euphorbia peplis* for instance, in two places, so as to separate a portion of the stem from the rest, is a complete demonstration that the vessels actually do contract. For whoever makes the experiment, will find that the milky juice of that plant flows out at both ends so completely, that if afterwards we cut the portion of the stem in the middle no juice whatever appears. Now it is impossible that these phenomena could take place without a contraction of the vessels; for the vessels in that part of the stem which has been detached cannot have been more than full; and their diameter is so small, that if it were to continue unaltered, the capillary attraction would be more than sufficient to retain their contents, and consequently not a drop would flow out. Since, therefore, the whole liquid escapes, it must be driven out forcibly, and consequently the vessels must contract.

in every direction from the pith, and composed of the lymphatic vessels of Grew and Malpighi. I do not see how the contraction of these laminae could propel the sap through the sap vessel, destitute as they are of valves, unless it were a contraction precisely similar to what Saussure supposed to take place in the sap vessels.

* Mem. Par. 1787.

† Phil. Trans. lxxviii.

Book IV.

It seems pretty plain, too, that the vessels are excited to contract by various stimuli; the experiments of Coulomb and Saussure render this probable, and an observation of Dr. Smith Barton makes it next to certain. He found that plants growing in water vegetated with much greater vigour, provided a little camphor was thrown into the water.*

SECT. IV.

OF THE FUNCTIONS OF THE LEAVES.

Peculiar
juice formed
by the
leaves.

It has been ascertained that the sap ascends to the leaves, that it there undergoes certain alterations, and is converted into another fluid called the *succus proprius*, *peculiar juice*, or *true sap*; which, like the blood in animals, is afterwards employed in forming the various substances found in plants. Now, the changes which the sap undergoes in the leaves, provided we can trace them, must throw a great deal of light upon the nature of vegetation. These changes are produced in part during the day, in part during the night. Now, as the functions of the leaves during the day are very different from what they are during the night, it will be proper to consider them separately.

Part of the
sap trans-
pires.
Its quan-
tity.

I. No sooner has the sap arrived at the leaves, than a great part of it is thrown off by evaporation.

1. The quantity thus perspired bears a very great proportion to the moisture imbibed. Mr. Woodward found that a sprig of mint, weighing 27 grains, in 77 days imbibed 2558 grains of water, and yet its weight was only increased 15 grains; therefore it must have given out 2543 grains. The same experiment was repeated by this philosopher on other plants; the following table exhibits the result: †

Plants and Water.	Weight.		Gain in 77 days.	Water wast- ed.
	When put in.	When taken out.		
	gr.	gr.	gr.	gr.
Spearmint in spring water	27	42	15	2558
Spearmint in rain water	28 $\frac{1}{2}$	45 $\frac{3}{4}$	17 $\frac{1}{2}$	3004
Spearmint in Thames water	28	54	26	2493
Common nightshade in spring water . . .	49	106	57	3708
Lathyrus in spring water	98	101 $\frac{1}{2}$	3 $\frac{1}{2}$	2501

* Ann. de Chim. xxiii. 63.

† Phil. Trans. 1699, xxix. 193.

These experiments demonstrate the great quantity of matter which is constantly leaving the plant. Dr. Hales found that a cabbage transmitted daily a quantity of moisture equal to about half its weight; and that a sun-flower, three feet high, transmitted in a day 1lb. 14oz. avoirdupois.* He showed, that the quantity of transpiration in the same plant was proportional to the surface of the leaves, and that when the leaves were taken off the transpiration nearly ceased.† By these observations, he demonstrated that the leaves are the organs of transpiration. He found, too, that the transpiration was nearly confined to the day, very little taking place during the night;‡ that it was much promoted by heat, and stopped by rain and frost:§ and Millar,|| Guettard,** and Sennebier, have shown that the transpiration is also very much promoted by sunshine. Chap. III.

The quantity of moisture imbibed by plants depends very much upon what they transpire; the reason is evident: when the vessels are once filled with sap, if none be carried off, no more can enter; and, of course, the quantity which enters must depend upon the quantity emitted.

2. In order to discover the nature of the transpired matter, Hales placed plants in large glass vessels, and by that means collected a quantity of it.†† He found that it resembled pure water in every particular, excepting only that it sometimes had the odour of the plant. He remarked, too, as Guettard and Duhamel did after him, that when kept for some time it putrefied, or at least acquired a stinking smell. Sennebier subjected a quantity of this liquid to a chemical analysis. Nature of the perspired matter.

He collected 13030 grains of it from a vine during the months of May and June. After filtration he gradually evaporated the whole to dryness. There remained behind two grains of residuum. These two grains consisted of nearly $\frac{1}{2}$ grain of carbonate of lime, $\frac{1}{7}$ grain of sulphate of lime, $\frac{1}{4}$ grain of matter soluble in water, and having the appearance of gum, and $\frac{1}{4}$ grain of matter which was soluble in alcohol, and apparently resinous. He analysed 60768 grains of the same liquid, collected from the vine during the months of July and August. On evaporation, he obtained $2\frac{1}{8}$ grains of residuum, composed of $\frac{3}{4}$ grain of carbonate of lime, $\frac{1}{7}$ grain

* Veget. Stat. i. 5 and 15.

† Veget. Stat. p. 30.

‡ Ibid. p. 5.

§ Ibid. p. 27 and 48.

|| Ibid. p. 22.

** Mem. Par. 1748.

†† Veget. Stat. i. 49.

Book IV. of sulphate of lime, $\frac{1}{2}$ grain of mucilage, and $\frac{1}{2}$ grain of resin. The liquid transpired by the *aster Novæ Angliæ* afforded precisely the same ingredients.*

Proportion
to the liquid
absorbed.

3. Sennebier attempted to ascertain the proportion which the liquid transpired bore to the quantity of moisture imbibed by the plant. But it is easy to see that such experiments are liable to too great uncertainties to be depended on. His method was as follows: he plunged the thick end of the branch on which he made the experiment into a bottle of water, while the other end, containing all its leaves, was thrust into a very large glass globe. The apparatus was then exposed to the sunshine. The quantity imbibed was known exactly by the water which disappeared from the bottle, and the quantity transpired was judged of by the liquid which condensed and trickled down the sides of the glass globe. The following table exhibits the result of his experiments:

Plants.	Imbibed.	Perspired.	Time.
Peach	100 gr.	35 gr.	—
Ditto	210	90	—
Ditto	220	120	—
Mint	200	90	2 days
Ditto	575	120	10
Rasp	725	560	2
Ditto	1232	765	2
Peach	710	295	1
Apricot	210	180	1

In some of his experiments no liquid at all was condensed. Hence it is evident that the quantity of matter transpired cannot be deduced from these experiments. The mouth of the glass globe does not seem to have been accurately closed; the air within it communicated with the external air; consequently the quantity condensed must have depended entirely upon the state of the external air, the heat, &c.

4. The first great change, then, which takes place upon the sap after it arrives at the leaves, is the evaporation of a great part of it; consequently what remains must be very different in its proportions from the sap. The leaves seem to have particular organs adapted for throwing off part of the sap by transpiration; for the experiments of Guettard,† Duhamel,‡

Peculiar
organ of
transpira-
tion.

* Encyc. Meth. Phys. Veget. 287.

† Mem. Par. 1749.

‡ Physique des Arbres, i. 158.

and Bonnet,* show that it is performed chiefly by the upper surfaces of leaves, and may be nearly stopped altogether by varnishing that surface. Chap. VII.

The leaves of plants become gradually less and less fit for this transpiration; for Sennebier found that when all other things are equal, the transpiration is much greater in May than in September.† Hence the reason that the leaves are renewed annually. Their organs become gradually unfit for performing their functions, and therefore it is necessary to renew them. Those trees which retain their leaves during the winter were found by Hales and succeeding physiologists to transpire less than others. It is now well known that these trees also renew their leaves. Why the leaves fall off.

II. Leaves have also the property of absorbing carbonic acid gas from the atmosphere. Leaves absorb carbonic acid.

1. We are indebted for this very important discovery to the experiments of Dr. Priestley. It had been long known that when a candle has been allowed to burn out in any quantity of air, no candle can afterwards be made to burn in it. In the year 1771, Dr. Priestley made a sprig of mint vegetate for ten days in contact with a quantity of such air: after which he found that a candle would burn in it perfectly well.‡ This experiment he repeated frequently, and found that it was always attended with the same result. According to the opinion at that time universally received, that the burning of candles rendered air impure by communicating phlogiston to it, he concluded from it, that plants, while they vegetate, absorb phlogiston.

Carbonic acid gas was at that time supposed to contain phlogiston. It was natural, therefore, to suppose that it would afford nourishment to plants, since they had the property of absorbing phlogiston from the atmosphere. Dr. Percival had published a set of experiments, by which he endeavoured to show that this was actually the case.

These experiments induced Dr. Priestley, in 1776, to consider the subject with more attention. But as, in all the experiments which he made, the plants confined in carbonic acid gas very soon died, he concluded that carbonic acid gas was not a food, but a poison to plants.§ Mr. Henry of Manchester was led, in 1784, probably by the contrariety of these results, to examine the subject. His experiments,

* *Traite des Feuilles*, 1. Mem.

† *Encyc. Meth. Veget.* 285.

‡ *On Air*, iii. 251.

§ *On Air*, i. 100.

Book IV. which were published in the Manchester Transactions,* perfectly coincided with those of Dr. Percival. For he found that carbonic acid gas, so far from killing plants, constantly promoted their growth and vigour. Meanwhile Mr. Sennebicr was occupied at Geneva with the same subject; and he published the result of his researches in his *Memoires Physico-Chymiques* about the year 1780. His experiments showed, in the clearest manner, that carbonic acid gas is used by plants as food. The same thing was supported by Ingenhousz in his second volume. The experiments of Saussure junior, published in 1797, have at last put the subject beyond the reach of dispute. From a careful comparison of the experiments of these philosophers, it will not be difficult for us to discover the various phenomena, and to reconcile all the seeming contradictions which occur in them. The facts are as follows:

2. Plants will not vegetate in an atmosphere of pure carbonic acid, nor if their atmosphere contains $\frac{3}{4}$ ths of its bulk of that gas. They vegetate in the sun when confined in atmospheres containing $\frac{1}{2}$, $\frac{1}{4}$ th, or $\frac{1}{8}$ th, of that gas, and the vegetation improves as the quantity of gas diminishes. When the atmosphere contains only $\frac{1}{12}$ th of carbonic acid gas, plants grow in it considerably better in the sun, than when placed in an atmosphere of common air; but when plants are placed in the shade, the presence of carbonic acid always injures their vegetation instead of promoting it.†

3. Mr. Saussure has shown, that plants will not vegetate in the sun when totally deprived of carbonic acid gas. They vegetate indeed well enough in air which has been previously deprived of carbonic acid gas; but when a quantity of lime was put into the glass vessel which contained them, they no longer continued to grow, and the leaves in a few days fell off.‡ The air, when examined, was found to contain no carbonic acid gas. The reason of this phenomenon is, that plants (as we shall see afterwards) have the power of forming and giving out carbonic acid in certain circumstances; and this quantity is sufficient to continue their vegetation for a certain time. But if this new formed gas be also withdrawn, by quicklime for instance, which absorbs it the instant it ap-

* Vol. ii. 341.

† Saussure, *Recherches Chimiques sur la Vegetation*, p. 30.

‡ *Ann. de Chim.* xxiv. 145, 148.

pears, the leaves droop, and refuse to perform their functions.* Carbonic acid gas, then, applied to the leaves of plants, is *essential* to vegetation.

Chap. III.

4. The direct contrary takes place in the shade. Plants not only continue to vegetate when deprived of all carbonic acid by means of lime; but they flourish more than if it were allowed to remain.†

5. Dr. Priestley, to whom we are indebted for many of the most important facts relative to vegetation, observed, in the year 1778, that plants, in certain circumstances, emitted oxygen gas; ‡ and Ingenhousz very soon after discovered that this gas is emitted by the *leaves* of plants, and only when they are exposed to the bright light of day. His method was to plunge the leaves of different plants into vessels full of water, and then expose them to the sun, as Bonnet, who had observed the same phenomenon, though he had given a wrong explanation of it, had done before him. Bubbles of oxygen gas very soon detached themselves from the leaves, and were collected in an inverted glass vessel.§ He observed, too, that it was not a matter of indifference what kind of water was used. If the water, for instance, had been previously boiled, little or no oxygen gas escaped from the leaves; river water afforded but little gas; but pump-water was the most productive of all.||

Emit oxygen gas.

Sennebier proved, that if the water be previously deprived of all its air by boiling, the leaves do not emit a particle of air; that those kinds of water which yield most air contain in them the greatest quantity of carbonic acid gas; that leaves do not yield any oxygen when plunged in water totally destitute of carbonic acid gas; that they emit it abundantly when the water, rendered unproductive by boiling, is impregnated with carbonic acid gas; that the quantity of oxygen emitted, and even its purity, is proportional to the quantity of carbonic acid gas which the water contains; that water impregnated with carbonic acid gas gradually loses the property of affording oxygen gas with leaves; and that whenever this happens, all the carbonic acid gas has disappeared; and on

By decomposing carbonic acid.

* Braconnot has rendered it probable, that in this experiment it was not the absence of carbonic acid, but the deleterious effects of the lime that killed the plants in Saussure's experiments. Ann. de Chim. lxi. 187.

† Saussure, Recherches, p. 36.

‡ On Air, iii. 284.

§ Ingenhousz on Veget. i. 15, &c.

|| Ibid. p. 33.

Book IV. adding more carbonic acid gas the property is renewed.* These experiments prove, in a most satisfactory manner, that the oxygen gas which the leaves of plants emit depends upon the presence of carbonic acid gas; that the leaves absorb carbonic acid gas, decompose it, give out the oxygen, and retain the carbon. They have been lately confirmed by the experiments of Dr. Woodhouse. To this gentleman likewise we are indebted for a very full set of experiments on the quantity of gas produced by exposing different plants in water to sunshine for given times.†

The decomposition performed by the parenchyma.

6. Sennebier has ascertained, that the decomposition of the carbonic acid takes place in the parenchyma. He found that the epidermis of a leaf would, when separated, give out no air, neither would the nerves in the same circumstances; but upon trying the parenchyma, thus separated from its epidermis and part of its nerves, it continued to give out oxygen as before.‡

That the decomposition is performed by peculiar organs, is evident from an experiment of Ingenhousz. Leaves cut into small pieces continued to give out oxygen as before; but leaves pounded in a mortar lost the property entirely. In the first state, the peculiar structure remained; in the other, it was destroyed. Certain experiments of Count Rumford, indeed, may seem incompatible with this conclusion; and they will naturally occur to the reader as an objection. He found that dried leaves of black poplar, fibres of raw silk, and even glass, when plunged into water, gave out oxygen gas by the light of the sun; but when Sennebier repeated these experiments, they did not succeed.§ It was probably the air contained in the water which separated in the Count's experiments.||

Carbonic acid absorbed varies in different plants.

7. From the experiments of Saussure, we learn that the quantity of carbonic acid thus absorbed and decomposed varies greatly in different plants, even when placed in the same circumstances. The *lythrum salicaria* was found to

* Encyc. Method. Physiol. Veget. 181.

† See Nicholson's Jour. ii. 154, and Ann. de Chim. xlii. 200.

‡ Encyc. Method. Physiol. Veget. 180. § Ann. de Chim. i. 115.

|| Dr. Woodhouse tried the experiments with filaments of asbestos, baked horse-hair, cotton, pannicles of rhus cotinus, cotton of asclepias syriaca, hairy plumes of clematis crispa, spikes of panicum glaucum, charcoal powder; each of these, he affirms, yielded in water a little oxygen gas, but less pure than the leaves of plants. Nicholson's Jour. ii. 158.

absorb seven or eight times its bulk of this gas in a day; while the *cactus opuntia*, and other fleshy-leaved plants, did not absorb above a fifth of that quantity. The portion absorbed, according to Saussure, depends upon the surface of the plant; and therefore thin leaved plants must absorb more than those that have fleshy leaves.*

8. It does not appear that the whole of the oxygen contained in the carbonic acid absorbed is emitted again by the plant. A considerable portion of it seems to be retained. This, at least, is the result which follows from a set of experiments made by Saussure on purpose to ascertain the point. He mixed carbonic acid with common air, in such a proportion that it occupied $7\frac{1}{2}$ hundredths of the mass. Jars, standing over mercury (covered with a thin film of water), were filled with this mixture; and plants of *vinca minor*, growing in a small vessel filled with water, were introduced into the jar. These plants, thus placed, were exposed for six successive days to the sun, from five in the morning to eleven, while the temperature of the air was 70° ; during all which time they vegetated with great vigour. The bulk of the air in the jar was not sensibly altered. No carbonic acid could be detected in it. The proportion of oxygen was $24\frac{1}{2}$ per cent. The following table exhibits the proportion of the constituents of this air, in French cubic inches, when put into the jar, and after the plants had vegetated in it six days.†

The whole oxygen not emitted.

	When put in.	When taken out.
Azote	211·92	218·95
Oxygen	56·33	71·05
Carbonic acid ..	21·75	0·00
	<hr style="width: 50px; margin: 0 auto;"/>	<hr style="width: 50px; margin: 0 auto;"/>
	290·00	290·00

Thus the whole 21·75 inches of carbonic acid were absorbed; but the oxygen emitted was only 14·72 inches, whereas the whole oxygen in the carbonic acid would have amounted to 21·75 inches. The difference, amounting to about seven inches, was made up by a quantity of azote, which had been given out by the plants along with the oxygen. The following table exhibits the result of similar experiments, made by this philosopher, on other plants. All the numbers denote cubic inches.

Some azote emitted.

* Recherches Chimiques, p. 56.

† Ibid. p. 40.

Plants.	Bulk of plants.	Air in jar.	Carbonic acid, per cent. in ditto.	Duration of experiment.	Carbonic acid, per cent. in residuary air.	Oxygen per cent. in do.	Carbonic acid absorb- ed.	Gas emitted.		Diminution of bulk of residuary gas.
								Oxygen.	Azote.	
<i>Mentha aquatica</i>	0.5	328	7.5	Days. 10	2.5	23.5	15.6	11.26	4.34	0
<i>Lythrum salicaria</i>	0.14	75	10	7	0	27.25	7.5	6.13	1.1	0.5
<i>Pinus Genevensis</i>	0.5	280	7	18	1.5	24.5	15.5	12.5	1	2
<i>Cactus opuntia</i>	1.1	155	10		4	24	9.3	6.4	2.9	0

9. Thus it appears that plants, when exposed to the light, absorb carbonic acid, decompose it, and throw out again the greatest part of the oxygen of this gas mixed (it would seem) with a little *azote*.* It is extremely probable that plants, by this process, acquire the greatest part of the carbonaceous matter which they contain; for if we compare the quantity of carbon contained in plants vegetating in the dark, where this process does not go on, with the quantity which those plants contain which vegetate in the usual manner, we shall perceive

* I presume that Saussure ascertained merely that the new portion of gas, to which he has given the name of azote, did not diminish with oxygen, and did not render lime-water turbid. Many other gases possess these properties. It might have been some inflammable gas. Hydrogen would have detonated; but some of the compound inflammable gases would not, when diluted so largely with common air.

a very conspicuous difference. Chaptal found that a byssus, which was vegetating in the dark, contained only $\frac{1}{89}$ th of its weight of carbonaceous matter; but the same plant, after being made to vegetate in the light for 30 days, contained $\frac{1}{4}$ th of its weight of carbonaceous matter.* Hassenfratz ascertained that plants growing in the dark contain much more water, and much less carbon and hydrogen, than plants growing in the light. Sennebier analysed both with the same result. Plants growing in the dark yielded less hydrogen gas and oil: their resinous matter was to that of plants growing in the light as 2 to 5.5, and their moisture as 13 to 6; they contained even one-half less of fixed matters.

The quantity of carbonic acid thus absorbed is considerable. In Saussure's experiments, the plants absorbed daily more than their own bulk of this gas: but when they grow in the open air, where the quantity of carbonic acid is much less considerable, not exceeding $\frac{1}{500}$ th part, the proportion absorbed is no doubt less.

10. Ingenhousz found that plants emit no oxygen when made to vegetate in the dark, and that in these circumstances they rather injure air than improve it. Now, as the emission of oxygen has been found to depend upon the absorption of carbonic acid, it is probable that this acid is absorbed only in the light. Saussure indeed has endeavoured to prove, that plants, even in the dark, absorb and decompose carbonic acid; but the quantity, if any, must be so extremely small that it cannot well be appreciated.

Plants emit
no oxygen
in the dark.

III. The green colour of plants has been shown, by Sennebier, to depend upon the absorption of carbonic acid. It appears only when plants vegetate in the light; for when they vegetate in the dark they are white; and when exposed to the light they acquire a green colour in a very short time, in whatsoever situation they are placed, even though plunged in water, provided always that oxygen be present; for Mr. Gough has shown, that light without oxygen has not the power of producing the green colour.†

Green colour
produced by
light.

Sennebier has observed, that when plants are made to vegetate in the dark, their etiolation is much diminished by mixing a little hydrogen gas with the air that surrounds them.‡ Ingenhousz had already remarked, that when a little hydrogen

* Mem. Par. 1786.

† Man. Mem. iv. 501.

‡ Encyc. Meth. Physiol. Veget. 75.

Book IV. gas is added to the air in which plants vegetate, even in the light, it renders their verdure deeper: * and he seems to think also, that he has proved by experiments, that plants absorb hydrogen gas in these circumstances.† Mr. Humbolt has observed that the *poa annua* and *compressa*, *plantago lanceolata*, *trifolium arvense*, *cheiri*, *lichen verticillatus*, and several other plants which grow in the galleries of mines, retain their green colour even in the dark, and that in these cases the air around them contains a quantity of hydrogen gas. This philosopher concludes, from his observations, that the white colour of etiolated plants is occasioned by their retaining an unusual proportion of oxygen, and that this is prevented by surrounding them with hydrogen gas. This may perhaps be true in certain cases; but the experiments of Mr. Gough, mentioned above, are sufficient to prove that the retention of oxygen is not the only difference between green and etiolated plants.‡

Plants supposed to decompose water.

The green colouring matter of plants has been shown by Rouelle to be of a resinous nature. From this, and from the circumstance of its being formed only in the light, Berthollet has inferred that the leaves of plants have the property of decomposing water as well as carbonic acid when exposed to the light of the sun. The oxygen emitted, according to him, is derived partly from the decomposed carbonic acid and partly from the water, while the carbon and hydrogen enter into the composition of the inflammable parts of the plant. This ingenious theory, though sufficiently probable, is not susceptible of direct proof. From the experiments of Saussure, we learn that when plants are made to vegetate in pure water, in atmospheres destitute of carbonic acid gas, the quantity of their fixed matter does not increase; but when their atmospheres contain this acid gas, the increase of weight which they receive is considerably greater than can be accounted for by the carbon and oxygen derived from the carbonic acid absorbed.§ Hence it is clear, that a portion of the water must enter into their composition. It is more likely that the elements of this portion arrange themselves in a different way, than that they still continue in the state of water. These facts certainly

* Ann. de Chim. iii. 57.

† Ibid. 61.

‡ Plants of a white colour, from vegetating in the dark, are called *etiolated*, from a French word which signifies a *star*, as if they grew by *star light*.

§ Recherches, p. 217

strengthen the hypothesis of Berthollet. Indeed, if we consider the great quantity of hydrogen contained in plants, it is difficult to conceive how they should obtain it, provided the water which they absorb does not contribute to furnish it.

Chap. III.

IV. Plants will not vegetate unless atmospheric air or oxygen gas have access to their leaves. This was rendered probable by those philosophers who, about the end of the 17th century, turned their attention particularly towards the physical properties of the air; but Dr. Ingenhousz was perhaps the first of the modern chemists who put it beyond doubt. He found that carbonic acid gas, azotic, and hydrogen gas, destroyed plants altogether, unless they were mixed with atmospheric air or oxygen gas. He found also, that plants grew very well in oxygen gas and in atmospheric air.* From these experiments, it was pretty clear that the leaves of plants absorb oxygen; and the whole series of chemical experiments on plants led to the supposition that this absorption was confined to the night. The subject has lately been very fully investigated by Saussure, who has not only confirmed these suppositions by decisive experiments, but added many new facts not previously suspected.†

Plants absorb oxygen,

1. The greater number of plants refuse to vegetate when confined in azotic gas. Those only continue to live which are abundantly supplied with green parts, as the *cactus opuntia*, &c. When plants continue to vegetate in azote, it is because they give out a portion of oxygen during the day; the absorption of which during the night preserves them.‡

2. When the leaves of plants are put in contact with common air during the night, they diminish the bulk of their atmosphere by absorbing oxygen. Some plants change, at the same time, a portion of the oxygen into carbonic acid, while others form no perceptible quantity of this last gas. The leaves, for example, of the *cactus opuntia*, *crassula cotyledon*, *sempervivum tectorum*, *agave Americana*, and *stapelia variegata*, simply absorb oxygen; while the leaves of the *quercus robur*, *sedum reflexum*, *asculus hippocastanum*, and *robinia pseudo-acacia*, absorb oxygen, and form a portion of carbonic acid, inferior in quantity to the oxygen which has disappeared.§

3. This inspiration of oxygen takes place only when the

* Ingenhousz, ii. passim.

† Recherches, p. 60.

‡ Ibid. p. 197.

§ Saussure, ibid. p. 61.

Book IV. leaves retain their organized form. If this be destroyed, by reducing them to a paste for example, all absorption of oxygen is prevented, though a portion of it is even then converted into carbonic acid by the action of the carbonaceous matter present.*

4. The oxygen thus inspired by the leaves of plants, is not separated from them again by putting them into the exhausted receiver of the air pump. By that means, indeed, they yield a little air, but always much less than the oxygen absorbed; and this air is precisely of the same nature with the atmosphere in which they were confined. Neither is the oxygen extricated by exposing the leaves to the greatest heat which they are capable of bearing without being destroyed.†

And convert it into carbonic acid.

5. There is reason to believe that the oxygen gas thus absorbed by plants is converted into carbonic acid within the plant, and that it is only after the plant is *saturated* with this acid (if the expression may be permitted), that the surrounding oxygen is partly converted into carbonic acid, by combining with the carbonaceous matter of the plant. When the leaves are exposed to the light, this carbonic acid is decomposed, and a quantity of oxygen thrown out, usually greater than what was inspired. But the oxygen given out in the light (when plants grow in atmospheres destitute of carbonic acid) is always proportional to the oxygen inspired during the night; being always greatest when the plant has absorbed the greatest quantity of oxygen.

Differ in the quantity absorbed.

6. Plants differ very much from each other in the quantity of oxygen which their leaves absorb during the night. Fleishy-leaved plants absorb the least oxygen; probably because they emit no carbonic acid gas. Hence they can vegetate in high situations where the atmosphere is rarefied. Next in order come the evergreen trees, which, though they absorb more oxygen than the fleshy-leaved plants, yet require much less than those trees which lose their leaves during winter. Those plants which flourish in marshy ground likewise absorb but little oxygen. The following tables exhibit the result of Saussure's experiments on this point. The first column contains the names of the plants whose leaves were employed; the second the month in which the experiments were made; and the third the bulk of oxygen absorbed, supposing the bulk of the leaves used in each experiment to be always 1'00.‡

* Saussure, Recherches, p. 74.

† Recherches, p. 69.

‡ Ibid. p. 99.

I. *Leaves of Evergreen Trees.*

Chap. III.
 Tables of
 the propor-
 tion absorb-
 ed.

Leaves of	Time	Oxygen absorbed.
<i>Illex aquifolium</i>	September	0·86
<i>Buxus semper virens</i>	September	1·46
<i>Prunus laurocerasus</i>	May	3·20
Ditto	September	1·36
<i>Viburnum tinus</i>	September	2·23
<i>Hedera helix</i>	September	1·00
<i>Vinca minor</i>	June	1·50
Ditto	September	0·93
<i>Pinus abies</i>	September	3·00
<i>Beupleurum fruticosum</i>	May	4·00
<i>Juniperus sabina</i>	June	2·60
<i>Juniperus communis</i>	June	2·40

II. *Leaves of Trees which lose their Leaves in Winter.*

Leaves of	Time	Oxygen absorbed.
<i>Fagus sylvatica</i>	August	8·00
<i>Carpinus betulus</i>	May	5·00
Ditto	September	6·00
<i>Quercus robur</i>	May	5·50
Ditto	September	5·50
<i>Æsculus hippocastanus</i>	September	4·80
<i>Populus alba</i>	May	6·20
Ditto	September	4·36
<i>Prunus Armeniaca</i>	September	8·00
<i>Amygdalus Persica</i>	June	6·60
Ditto	September	4·20
<i>Juglans regia</i>	May	6·60
Ditto	September	4·40
<i>Platanus occidentalis</i>	September	3·00
<i>Robinia pseudo-acacia</i>	May	5·00
Ditto	September	6·70
<i>Syringa vulgaris</i>	May	3·36
Ditto	September	2·20
<i>Fraxinus excelsior</i>	May	4·32
Ditto	September	3·71
<i>Pyrus</i>	May	5·20
Ditto	September	3·40
<i>Rosa centifolia</i>	June	5·40
<i>Fagus castanea</i>	July	5·60

III. *Leaves of Herbaceous Plants, not Aquatic.*

Leaves of	Time.	Oxygen absorbed.
Solanum tuberosum	September	2.50
Brasica oleracea	Ditto, young leaves	2.40
Ditto	Sept. old leaves ..	2.00
Urtica urens	September	2.00
Mercurialis annua	Sept. in flower	2.33
Daucus carota	September, ditto ..	1.90
Vicia faba	Before flowering ..	3.70
Ditto	In flower	2.00
Ditto	After ditto	1.60
Lilium candidum	May, in flower	0.66
Ditto	Sept. after ditto . . .	0.50
Tropæolum majus	Sept. in flower	3.00
Digitalis ambigua	July	2.00
Brassica rapa	Sept. in flower	1.25
Avena sativa	June, before ditto . .	2.70
Triticum æstivum	May, before ditto . .	5.00
Pisum sativum	May, in flower	3.72
Ruta graveolens	August	2.00

IV. *Leaves of Aquatic Plants.*

Leaves of	Time.	Oxygen absorbed.
Alisma plantago	August	0.70
Inula dysenterica	September	1.60
Epilobium molle	Sept. in flower	1.90
Sisymbrium nasturtium . .	September	1.60
Polygonum Persicaria . . .	Sept. in flower	2.00
Veronica beccabunga	September	1.70
Ranunculus reptans	September	1.50
Lythrum salicaria	May, before flower	2.30
Caltha palustris	May	1.00
Carex acuta	May	2.25

V. *Leaves of the Fleshy Plants.*

Leaves of	Time.	Oxygen absorbed.
Cactus opuntia	August	1.00
Agave Americana	August	0.80
Sempervivum tectorum . .	July	1.00
Sedum globosum	September	1.50

Leaves of	Time.	Oxygen absorbed.
Saxifraga cotyledon	September	0·60
Sedum reflexum	June	1·70
Stapelia variegata	July	0·63
Mesembryanthemum } deltoides }	.. July	1·70

7. It is not improbable, that by the absorption of oxygen, and the formation of carbonic acid, a portion of heat may be evolved, as Saussure supposes; though the quantity must be too small to be appreciated. It appears that in certain cases a very considerable degree of heat is produced by vegetables; though it has not yet been ascertained whether the appearance of it is connected with the absorption of oxygen. A very extraordinary instance of this is related by Bory de St. Vincent on the authority of Hubert. The stamina of the *arum cardifolium*, at the moment of bursting, produced so great a heat, that twelve of them placed round the bulb of a thermometer raised it from 79° to 143°. The experiment was repeated a great number of times with similar results.*

8. It appears, from the experiments of Saussure, that the roots absorb oxygen as well as the leaves, and that they transmit the carbonic acid formed to the leaves to be decomposed. The branches likewise absorb it. Flowers do not expand without its presence.

Thus it appears that during the night plants absorb oxygen; that they form with it carbonic acid; that a portion of this gas is sometimes emitted, together with a little azote; but that the greatest part is retained and decomposed by the leaves during the day. Plants will not live without this nightly inspiration, even though supplied with carbonic acid, provided the oxygen formed by them during the day be constantly withdrawn at the approach of night.

V. The leaves of plants absorb water as well as oxygen from the air. This had been suspected in all ages: the great effect which dew, slight showers, and even wetting the leaves of plants, have in recruiting their strength, and making them vegetate with vigour, are so many proofs that the leaves imbibe moisture from the atmosphere. Hales rendered this still more probable, by observing that plants increase considerably in weight when the atmosphere is moist; and Mr. Bonnet put

Jour. de Phys. lix. 231.

Book IV.

the matter beyond doubt in his Researches concerning the Use of the Leaves. He showed that leaves continue to live for weeks when one of their surfaces is applied to water; and that they not only vegetate themselves, but even imbibe enough of water to support the vegetation of a whole branch, and the leaves belonging to it. He discovered also, that the two surfaces of leaves differ very considerably in their power of imbibing moisture; that in trees and shrubs the under surface possesses almost the whole of the property, while the contrary holds in many of the other plants; the kidney bean for instance.

These facts prove not only that the leaves of plants have the power of absorbing moisture, but also that the absorption is performed by very different organs from those which emit moisture; for these organs lie on different sides of the leaf. If we consider that it is only during the night that the leaves of plants are moistened with dew, we can scarcely avoid concluding, that, except in particular cases, it is during the night that plants imbibe almost all the moisture which they do imbibe.

Thus we have seen that the leaves of plants perform very different operations at different times. During the day they are giving out moisture, absorbing carbonic acid gas, and emitting oxygen gas; during the night, on the contrary, they are absorbing moisture, giving out carbonic acid gas, and absorbing oxygen gas.

Sap converted into nutritive juices by these processes.

By these processes, and perhaps also by others which have not yet been detected, the sap of plants is new modelled, and brought to the state which is adapted for the nourishment of the plant. To describe in what manner these changes take place is impossible; because we neither know precisely the substance into which the sap has been converted by the operations performed during the day, nor the new substances formed by the operations of the night. We have reason, however, to conclude, that during the day the carbon of the sap is increased, and that during the night the hydrogen and oxygen are increased; but the precise new substances formed are unknown to us. Nor let any one suppose that the increase of the hydrogen and of the oxygen of the sap is the same thing as the addition of a quantity of water. In water, oxygen and hydrogen are already combined together in a certain proportion; and this combination must be broken before these elementary bodies can enter into those triple compounds with

carbon, of which a great part of the vegetable products consist. We have not the smallest conception of the manner in which these triple combinations are formed, and as little of the manner in which the bodies which compose vegetable substances are combined together. The combination may, for any thing we know to the contrary, be very complicated: though it consists wholly of three ingredients; and analogy leads us to suppose that it actually is very complicated: for in chemistry it may be considered as a truth, to which at present few or no exceptions are known, that bodies are decomposed with a facility inversely as the simplicity of their composition; that is to say, that those bodies which consist of the fewest ingredients are most difficultly decomposed, and that those which are formed of many ingredients are decomposed with the greatest facility.

Neither let any one suppose that the absorption of carbonic acid gas during the day is balanced by the quantity emitted during the night; and that therefore there is no increase of carbon; for Ingenhousz and Saussure have shown that the quantity of oxygen gas emitted during the day is much greater than the carbonic acid gas emitted during the night; and that in favourable circumstances, the quantity of oxygen gas in the air surrounding plants is increased, and the carbonic acid gas diminished; so much so, that both Dr. Priestley and Dr. Ingenhousz found, that air which had been spoiled by a lighted candle, or by animals, was rendered as good as ever by plants. Now we know that combustion and respiration diminish the oxygen gas, and add carbonic acid gas to air; therefore vegetation, which restores the purity of air altered by these processes, must increase the oxygen, and diminish the carbonic acid gas of that air; consequently the quantity of carbonic acid gas absorbed by plants during the day is greater than the quantity emitted by them during the night; and of course the carbon of the sap is increased in the leaves.

It is true, that when plants are made to vegetate for a number of days in a given quantity of air, its ingredients are not found to be altered. Thus Hassenfratz ascertained that the air, in which young chesnuts vegetate for a number of days together, was not altered in its properties, whether the chesnuts were vegetating in water or in earth;* and Saussure, junior, proved that peas growing for ten days in water did not

* Ann. de Chim. xiii. 325.

Book IV.

alter the surrounding air.* But this is precisely what ought to be the case, and what must take place, provided the conclusions which I have drawn be just. For if plants only emit oxygen gas by absorbing and decomposing carbonic acid gas, it is evident, that unless carbonic acid gas be present, they can emit no oxygen gas; and whenever they have decomposed all the carbonic acid gas contained in a given quantity of air, we have no longer any reason to look for their emitting any more oxygen gas; and if the quantity of carbonic acid gas emitted during the night be smaller than that absorbed during the day, it is evident that during the day the plant will constantly decompose all the acid which had been formed during the night. By these processes the mutual changes of day and night compensate each other; and they are prevented from more than compensating each other by the forced state of the plant. It is probable, that when only part of a plant is made to vegetate in this forced state, some *carburetted sap* (if I may be allowed the expression) is supplied by the rest of the plant; and that therefore the quantity of carbonic acid gas emitted during the night may bear a nearer proportion to that emitted in a state of nature than that of the absorption of fixed air can possibly do. And probably, even when the whole plant is thus confined, the nightly process goes on for a certain time at the expense of the carbon already in the sap; for Hassenfratz found, that in these cases the quantity of carbon in the plant, after it had vegetated for some time in the dark, was less than it had been when it began to vegetate.† This is the reason that plants growing in the dark, when confined, absorb all the oxygen gas, and emit carbonic acid gas: and whenever this has happened, they die; because then neither the daily nor nightly processes can go on.

SECT. V.

OF THE PECULIAR JUICES OF PLANTS.

By these changes which go on in the leaves, the nature of the sap is altogether changed. It is now converted into what is called the *peculiar juice*, or *true sap*, and is fit for being assimilated to the different parts of the plant, and for being em-

* Ann. de Chim. xxiv. 139.

† Ibid. xiii. 188.

ployed in the formation of those secretions which are necessary for the purposes of the vegetable economy.

Chap. III.

The leaves, therefore, may be considered as the digesting organs of plants, and as equivalent in some measure to the stomach and lungs of animals. The leaves consequently are not mere ornaments; they are the most important parts of the plant. Accordingly we find, that whenever we strip a plant of its leaves, we strip it entirely of its vegetating powers till new leaves are formed. It is well known that when the leaves of plants are destroyed by insects they vegetate no longer, and that their fruit never makes any farther progress in ripening, but decays and dries up. Even in germination no progress is made in the growth of the stem till the seed-leaves appear. As much food indeed is laid up in the cotyledons as advances the plant to a certain state: the root is prepared, and made ready to perform its functions; but the sap which it imbibes must be first carried to the seed-leaves, and digested there, before it be proper for forming the plumula into a stem. Accordingly if the seed-leaves are cut off, the plant refuses to vegetate.

Leaves the
digesting
organs of
plants.

It will be very natural to ask, if this be true, how come the leaves themselves to be produced? I had endeavoured to render it probable, that food for the purpose of nourishing and developing them was laid up in the buds themselves; but the experiments of Mr. Knight, formerly detailed, have shown that the alburnum is the part of the tree in which this food is deposited. After the plant has developed all the parts which are to appear during the summer, and after the buds are fabricated and rolled up, the leaves still continue to manufacture new food, and to send it to the alburnum. Here it is deposited, and employed next spring in feeding and developing those organs of plants which are absolutely necessary for enabling them to perform the functions of vegetation. From this important fact, Mr. Knight has drawn a number of curious consequences, of great importance to the practical gardener and farmer.

How they
are pro-
duced.

This deposition of food for the future supply of the plant explains to us why the branch of a vine, if it be introduced into a hot-house during the winter, puts forth leaves and vegetates with vigour, while every other part of the plant gives no signs of life. It explains to us also why the sap flows out of trees very readily in spring before the leaves appear, but after

Book IV.

that the bleeding ceases altogether. It is evident that there can be scarcely any circulation of sap before the leaves appear; for as there is no outlet, when the vessels are once full they can admit no more. It appears, however, from the bleeding, that the roots are capable of imbibing, and the vessels of circulating, the sap with vigour. Accordingly, whenever there is an outlet, they perform their functions as usual, and the tree bleeds; that is, they send up a quantity of sap to be digested as usual: but as there are no digesting organs, it flows out, because the sap that flows out would not have been imbibed at all, had it not been for the artificial opening. But when the digestive organs appear, the tree will not bleed; because these organs require all the sap, and it is constantly flowing to them.

Absolutely necessary.

If a tree be deprived of its leaves, new leaves make their appearance, because they are already prepared for that purpose. But what would be the consequence if a tree were deprived of its leaves and of all its buds for five years back? That plants do not vegetate without leaves is evident from an experiment of Duhamel. He stripped the bark off a tree in ringlets, so as to leave five or six rings of it at some distance from each other, with no bark in the intervals. Some of these rings had buds and leaves; these increased considerably in size; but one ring which had none of these remained for years unaltered. Mr. Knight found that a shoot of the vine, when deprived of its leaves died altogether.*

Form the peculiar juice,

The *succus proprius*, or *peculiar juice* of plants, may be considered as analogous to the blood of animals. It is the food altered by digestion, and rendered fit for being assimilated and converted into a part of the plant itself by the subsequent processes of vegetation. That it flows from the leaves of the plant towards the roots appears from this circumstance, that when we make an incision into the plant, into whatever position we put it, much more of the *succus proprius* flows from that side of the wound which is next the leaves and branches than from the other side; and this happens even though the leaves and branches be held undermost.† When a ligature is tied about a plant, a swelling appears above but not below the ligature.

Which flows from the leaves.

In vessels,

The vessels containing the peculiar juice have been traced

* Phil. Trans. 1801, p. 338.

† Bell, Manch. Mem. ii. 402.

by Mr. Knight from the leaves into the cortical layers of the inner bark.* Hedwig, who has examined the vessels of plants with very great care, seems to consider them as of the same structure with the tracheæ; but Mr. Knight is of a different opinion. It appears evidently from the experiments of this last gentleman, that they communicate with no part of the plant which is situated at a greater distance from the root than the leaf from which they themselves originate. For when two circular incisions are made through the bark of a branch above and below a leaf, and at some distance from it, only that part of the portion confined between the two incisions which is situated below the leaf increases in size. From the experiments of Mr. Knight, there is reason to believe that these vessels are much better calculated to carry the true sap from the leaves towards the roots, than in the contrary direction. By passing the slender shoots of a vine through the earth of a garden-pot, he made them send out roots. The shoots were then cut off from the parent tree, leaving an equal portion on each side of this new root. Each of these portions was turned up at a similar angle, and had a bud at a little distance from the cut-off extremity. Here were two stems growing from one root. But the one was obviously inverted, while the other was in its natural position. In the first, the wood between the bud and the cut-off extremity increased in size; but in the other it did not: indicating a disposition in the true sap to move in its original direction from the leaves to the root. In like manner, when shoots of gooseberry and currant bushes were planted inverted, the upper part always decayed.† These experiments are favourable to the opinion, that these vessels are furnished with valves.

When all communication between the leaves and roots is cut off, by removing a portion of the bark all round the tree, it appears, from the experiments of Knight, that the alburnous vessels acquire the property of transmitting a portion of the true sap to the roots, so as to maintain the powers of vegetating; but the quantity is greatly diminished. The surplus appears to be lodged in the alburnum, which thus becomes denser; and if the plant be allowed to vegetate, this food is expended next season upon the upper part of the plant. Thus the quantity of blossom on the branch of an apple tree is

* Phil. Trans. 1801, p. 337.

† Nicholson's Jour. x. 289.

Book IV.

greatly increased by removing a section of its bark the season before the blossom is to appear.*

Its properties.

The true sap is easily known by its colour and its consistence. In some plants it is green, in some red, in many milky. It cannot be doubted that its motion in the vessels is performed in the same way as that of the sap.

If we had any method of obtaining this peculiar juice in a state of purity, the analysis of it would throw a great deal of light upon vegetation; but this is scarcely possible, as we cannot extract it without dividing at the same time the vessels which contain the sap. In many cases, however, the peculiar juice may be known by its colour, and then its analysis may be performed with an approach towards accuracy. The facts respecting its constituents, hitherto ascertained by chemists, have, as far as I am acquainted with them, been detailed in the preceding chapter. These experiments prove, as might have been expected, that the peculiar juices differ very considerably from each other, and that every plant has a juice peculiar to itself. Hence it follows, that the processes which go on in the leaves of plants must differ at least in degree, and that we have no right to transfer the conclusions deduced from experiments on one species of plants to those of another species. It is even probable, that the processes in different plants are not the same in kind; for it is not reasonable to suppose that the phenomena of vegetation in an agaric or a boletus are precisely the same as those which take place in trees and in larger vegetables, on which alone experiments have hitherto been made.

The true sap is conveyed to every part of the plant; and all the substances which we find in plants, and even the organs themselves by which they perform their functions, are formed from it. But the thickest veil covers the whole of these processes; and so far have philosophers hitherto been from removing this veil, that they have not even been able to approach it. All these operations indeed, are evidently chemical decompositions and combinations; but we neither know what these decompositions and combinations are, nor the instruments in which they take place, nor the agents by which they are regulated.

* Knight on the Inverted Action of the Albuminous Vessels of Trees. Phil. Trans. 1806.

SECT. VI.

OF THE DECAY OF PLANTS.

SUCH, as far as I am acquainted with them, are the changes produced by vegetation. But plants do not continue to vegetate for ever; sooner or later they decay and wither, and rot, and are totally decomposed. This change indeed does not happen to all plants at the end of the same time. Some live only for a single season, or even for a shorter period; others live two seasons, others three, others a hundred or more; and there are some plants which continue to vegetate for a thousand years; but sooner or later they all cease to live: and then those very chemical and mechanical powers which had promoted vegetation combine to destroy the remains of the plant. Now, What is the cause of this change? Why do plants die?

Chap. III.
Plants decay and die.

This question can only be answered by examining with some care what it is which constitutes the *life* of plants: for it is evident, that if we can discover what that is which constitutes the life of a plant, it cannot be difficult to discover whatever constitutes its death.

Now the phenomena of vegetable life are in general *vegetation*. As long as a plant continues to vegetate, we say that it lives; when it ceases to vegetate, we conclude that it is dead.

Phenomena of vegetable life.

The life of vegetables, however, is not so intimately connected with the phenomena of vegetation that they cannot be separated. Many seeds may be kept for years without giving any symptom of vegetation; yet if they vegetate when put into the earth, we say that they possess life; and if we would speak accurately, we must say also that they possessed life even before they were put into the earth: for it would be absurd to suppose that the seed *obtained* life merely by being put into the earth. In like manner many plants decay, and give no symptoms of vegetation during winter; yet if they vegetate when the mild temperature of spring affects them, we consider them as having lived all winter. The life of plants, then, and the phenomena of vegetation, are not precisely the same thing; for the one may be separated from the other, and we can even suppose the one to exist without the other. Nay, what is more, we can, in many cases, *decide*, without hesitation, that a vegetable is not dead, even when no

Book IV. } vegetation appears; and the proof which we have for its life is, that it *remains unaltered*; for we know that when a vegetable is dead, it soon changes its appearance, and falls into decay.

Thus it appears that the *life* of a vegetable consists in two things: 1. In remaining unaltered, when circumstances are unfavourable to vegetation; 2. In exhibiting the phenomena of vegetation when circumstances are favourable. When neither of these two things happens, we may say that a vegetable is dead.

Owing to a vegetative principle. The phenomena of vegetation have been enumerated above. They consist in the formation or expansion of the organs of the plant, in the taking in of nourishment, in carrying it to the leaves, in digesting it, in distributing it through the plant, in augmenting the bulk of the plant, in repairing decayed parts, in forming new organs when they are necessary, in producing seeds capable of being converted into plants similar to the parent. The *cause* of these phenomena, whatever it may be, is the *cause* also of *vegetable life*, and may be distinguished by the name of the *vegetative principle*. But an inquiry into the nature of this principle belongs to the science of *physiology*, and would be foreign to the nature of this work.

Death of plants.

The death of plants, if we can judge from the phenomena, is owing to the organs becoming at last altogether unfit for performing their functions, and incapable of being repaired by any of the powers which the vegetative principle possesses. The changes which vegetable substances undergo after death come now to be examined. They shall form the subject of the ensuing chapter.

CHAP. IV.

OF THE DECOMPOSITION OF VEGETABLE SUBSTANCES.

Difference between minerals and living matter.

THE most striking distinction between the substances belonging to the mineral kingdom and those which make a part of animals or vegetables, is the following: mineral bodies show little or no tendency to change their nature; and when left to themselves, undergo no spontaneous decompositions: whereas animal and vegetable substances are continually al-

tering; and when left to themselves in favourable circumstances, always run through a regular set of decompositions. I propose in this chapter to treat of the spontaneous decompositions of vegetable bodies; those of animal bodies belong to the next book.

Chap. IV.

We have seen in the last chapter, that during vegetation the constituents of plants are continually going through a regular set of changes, losing the properties of one substance, and assuming those of another. Thus a substance which in the young plant has the properties of mucilage, assumes in the old the properties of starch; what in green fruit is an acid, in ripe fruit becomes sugar. Vegetable principles, then, are not fixed or stationary in the living plant; they are gradually running into one another. But this tendency to change, or rather this continual decomposition in consequence of the mutual reaction of the different simple substances of which the vegetable principles are composed, is by no means confined to the living state. It goes on with equal or with greater energy in favourable circumstances in vegetable matter after it has been completely separated from the living plants. It has been observed that this tendency to spontaneous decomposition is usually greater in animal than in vegetable bodies; and that those vegetable bodies, in which the tendency is greatest, bear the closest resemblance to animal matter. Hence the common chemical phrase, that such substances are *more animalized*. This is the case with *gluten* in particular, which undergoes spontaneous decomposition more rapidly than most other vegetable bodies.

Vegetable substances go through a series of changes.

During the spontaneous decomposition which vegetable substances undergo, it is obvious that the simple substances of which they are composed must unite together in a different manner from that in which they were formerly united, and form a new set of compounds which did not formerly exist. Now it has been observed that the specific gravity of these new compounds is almost always less than that of the old body. Some of them usually fly off in the state of gas or vapour. Hence the odour that vegetable bodies emit during the whole time that they are running through the series of their changes. When the odour is very offensive or noxious, the spontaneous decomposition is called *putrefaction*; but when the odour is not offensive, or when any of the new compounds formed is applied to useful purposes, the spontaneous decomposition is called *fermentation*. This word was first

These spontaneous alterations

Book IV. introduced into chemistry by Van Helmont.* It is supposed
 Called fer- by some to have originated from the intestine motion which
 mentation. is always perceptible while vegetable substances are ferment-
 ing; while by others it is derived from the heat which in these
 cases is always generated. The term *fermentation* is now very
 often applied to all the spontaneous changes which vegetable
 bodies undergo without regard to the products. It therefore,
 in that sense, includes *putrefaction*; and certainly there is no
 impropriety in thus extending the term so as to make it com-
 prehend every case of spontaneous decomposition. By fer-
 mentation, then, is now meant all the spontaneous changes
 which take place in vegetable substances after they are separated
 from the living plant.

All the phenomena of fermentation lay for many years con-
 cealed in the completest darkness, and no chemist was bold
 enough to hazard even an attempt to explain them. They
 were employed, however, and without hesitation too, in the
 explanation of other phenomena; as if giving to one process
 the name of another, of which we are equally ignorant, could,
 in reality, add any thing to our knowledge. The dark-
 ness which enveloped these phenomena has lately begun to
 disperse; but they are still surrounded with a very thick mist;
 and we must be much better acquainted with the composition
 of vegetable substances, and the mutual affinities of their in-
 gredients, than we are at present, before we can explain them
 in a satisfactory manner.

Fermentation never takes place unless vegetable substances
 contain a certain portion of water, and unless they are ex-
 posed to a temperature at least above the freezing point.
 When dry or freezing, many of them continue long without
 alteration. Hence we have an obvious method of preventing
 fermentation.

Tendency
 differs in
 different
 bodies.

If we take a view of all the vegetable principles described
 in the first Chapter of this Book, we shall find that they differ
 much from one another in their tendency to run into fermenta-
 tion. Gum, sarcocol, starch, indigo, wax, resins, camphor,
 caoutchouc, gum resins, wood, and suber, though mixed with
 water, and placed in the most favourable temperature, show
 scarcely any tendency to change their nature. Oils absorb
 oxygen from the atmosphere, but too slowly to produce any
 intestine motion. Tannin, some of the acids, and extractive,

* Stahl's Fundament. Chem. i. 121.

are gradually decomposed; the surface of the liquid becomes mouldy, and an insipid sediment falls to the bottom; and when the process has once begun, it goes on with greater rapidity. Albumen and fibrin putrefy very quickly, but the products have not been ascertained; gluten gradually changes into a kind of cheese.

Chap. IV.

But it is when several of the vegetable principles are mixed together, that the fermentation is most perceptible, and the change most remarkable. Thus when gluten is added to a solution of sugar in water, the liquid soon runs into vinegar, or in certain cases to alcohol and vinegar. When gluten is mixed with starch and water, alcohol and vinegar usually make their appearance; but the greatest part of the starch remains unaltered. It has been observed that certain substances are peculiarly efficacious in exciting fermentation in others. These substances have received the name of *ferments*.

Greatest in compounds.

But the phenomena of fermentation do not appear in their greatest perfection in our artificial mixture of vegetable principles. Those complicated parts of plants in which various principles are already mixed by Nature, especially the liquid parts, exhibit the finest specimens of it; such as the sap of trees, the juices of fruits, the decoctions of leaves, seeds, &c. It is from such natural mixtures that we obtain all the products of fermentation which mankind have applied to useful purposes; such as indigo, beer, bread, vinegar, wine, &c. In the present imperfect state of the subject, I shall satisfy myself with an account of the most important of these products of fermentation, and of the phenomena which take place during their formation. I shall divide this chapter into three sections: in the first, I shall treat of the fermentation which produces intoxicating liquors; in the second, of the fermentation which produces vinegar; and in the third, of that which reduces the substance to *soil*. These are usually called the *vinous*, *acetous*, and *putrefactive* fermentations.

Division.

SECT. I.

OF THE VINOUS FERMENTATION.

UNDER this name is comprehended every species of fermentation which terminates in the formation of an intoxicating liquid. Now these liquids, though numerous, may be comprehended under two general heads; namely, those which are

Book IV. obtained from the *decoctions of seeds*, and those which are obtained from the *juices of plants*. The liquids of the first class are denominated *beer* or *wash*, those of the second *wine*.

I. BEER.

The method of making beer was known in the most remote ages. The Greek writers give the honour of the invention to the Egyptians. Almost every species of corn has been employed for the purpose. In Europe it is usually made from barley, in India from rice, in the interior of Africa from the *holcus spicatus*;^{*} but whatever grain is employed, the process is nearly the same. We shall therefore select, as an example, the beer made in this country from barley.

1. As the grain in its natural state has been found unsuitable for yielding good beer, it is usual, in the first place, to convert it into *malt*.

The term *malt* is applied to grain which has been made to germinate artificially to a certain extent, after which the process is stopped by the application of heat.

Steeping of
barley.

The barley is steeped in cold water for a period which (as regulated by law) must not be less than 40 hours; but beyond that period the steeping may be continued as long as is thought proper. Here it imbibes moisture, and increases in bulk; while at the same time a quantity of carbonic acid is emitted, and a part of the substance of the husk is dissolved by the steep-water. The proportion of water imbibed depends partly on the barley, and partly on the length of time that it is steeped. From the average of a good many trials, it appears that the medium increase of weight from steeping, may be reckoned 0.47; that is to say, every 100 pounds of barley when taken out of the steep weighs 147 pounds. The average increase of bulk is about a fifth; that is to say, that 100 bushels of grain, after being steeped, swell to the bulk of 120 bushels. The carbonic acid emitted while the barley is in the steep is inconsiderable; and it is probable from the experiments of Saussure, that it owes its formation, at least in part, to the oxygen held in solution by the steep-water.

The steep-water gradually acquires a yellow colour, and the peculiar smell and taste of water in which straw has been steeped. The quantity of matter which it holds in solution varies from $\frac{1}{30}$ th to $\frac{1}{100}$ th of the weight of barley. It consists

^{*} Park's Travels, p. 63, 3vo. Edition.

chiefly of an extractive matter of a yellow colour and disagreeable bitter taste, which deliquesces in a moist atmosphere, and contains always a portion of nitrate of soda. It holds in solution most of the carbonic acid disengaged. This extractive matter is obviously derived from the husk of the barley, and is that substance to which that husk owes its colour. Accordingly grain becomes much paler by steeping.

After the grain has remained a sufficient time in the steep, Couch. the water is drained off, and the barley thrown out of the cistern upon the malt-floor, where it is formed into a rectangular heap, called the *couch*, about 16 inches deep. In this situation it is allowed to remain about 26 hours. It is then turned, by means of wooden shovels, and diminished a little in depth. This turning is repeated twice a day or oftener, and the grain is spread thinner and thinner, till at last its depth does not exceed a few inches.

When placed on the couch, it begins gradually to absorb Changes white on the malt-floor. oxygen from the atmosphere, and to convert it into carbonic acid; at first very slowly, but afterwards more rapidly. The temperature, at first the same with that of the external air, begins slowly to increase; and in about 96 hours the grain is, at an average, about 10° hotter than the surrounding atmosphere. At this time the grain, which had become dry on the surface, becomes again so moist that it will wet the hand, and exhales at the same time an agreeable odour, not unlike that of apples. The appearance of this moisture is called *sweating*. A small portion of alcohol appears to be volatilized at this period. The great object of the maltmen is to keep the temperature from becoming excessive. This they do by frequent turning. The temperature which they wish to preserve varies from 55° to 62°, according to the different modes of malting pursued.

At the period of the sweating the roots of the grains begin to appear, at first like a small white prominence, at the bottom of each seed, which soon divides itself into three rootlets, and increases in length with very great rapidity, unless checked by turning the malt. About a day after the sprouting of the roots, the rudiments of the future stem, called *acrospire* by the maltsters, may be seen to lengthen. It rises from the same extremity of the seed with the root, and advancing within the husk, at last issues from the opposite end: but the process of malting is stopped before it has made such progress.

As the acrospire shoots along the grain, the appearance of

Book IV.

the kernel, or mealy part of the corn, undergoes a considerable change. The glutinous and mucilaginous matter is taken up and removed, the colour becomes white, and the texture so loose that it crumbles to powder between the fingers. The object of malting is to produce this change: when it is accomplished, which takes place when the acrospire has come nearly to the end of the seed, the process is stopped by drying the malt upon the kiln. The temperature at first does not exceed 90° ; but it is raised very slowly up to 140° , or higher, according to circumstances. The malt is then cleaned, to separate the rootlets, which are considered as injurious.

Loss sustained in malting.

Such is a short sketch of the process of malting. Barley, by being converted into malt, generally increases two or three per cent. in bulk; and loses at an average, about a fifth of its weight, or 20 per cent. But of these 20 parts 12 are to be ascribed to kiln-drying, and consist of water, which the barley would have lost had it been exposed to the same temperature: so that the real loss does not exceed eight per cent. From a good many trials, made with as much attention to all the circumstances as possible, the following seems to be the way of accounting for this loss:

Carried off by the steep-water	1·5
Dissipated in the floor	3·0
Roots, separated by cleaning	3·0
Waste	0·5
		8·0

The loss on the floor ought to be entirely owing to the separation of carbon by the oxygen of the atmosphere; but were this the only cause, it would be much smaller than three per cent. Two other causes concur to produce this loss:—1. Many of the roots are broken off during the turning of the malt; these wither and are lost, while others grow in their place:—2. A certain portion of the seeds lose the power of germinating, by bruises or other accidents, and these lose a much greater portion than three per cent. of their real weight. From a good many trials, made with as much care as possible, I am disposed to conclude that the quantity of carbon, separated during the whole process of malting, by the formation of carbonic acid gas, does not exceed two per cent. and that the weight of the roots formed amounts often to four per cent. These two, in reality, include the whole real loss of

weight which barely sustains when malted. What is lost in the steep, being husk, need scarcely be reckoned.

The roots appear, from the process, to be formed chiefly from the mucilaginous and glutinous parts of the kernel. The starch is not employed in their formation; but undergoes a change, intended no doubt to fit it for the future nourishment of the plumula. It acquires a sweetish taste, and the property of forming a transparent solution with hot water. In short, it approaches somewhat to the nature of sugar; but is much more soluble, and much more easily decomposed, than that principle. From the experiments of Saussure on the conversion of starch into sugar, of which an account has been given in a preceding Chapter, there is reason to conclude that this change is brought about by the combination of the starch with water. The action of hot water on barley-meal seems gradually to induce a similar one.

2. Malt thus prepared is ground in a mill, and infused in a large cylindrical vessel, called the *mash-tun*, with somewhat more than its own bulk of water, of a temperature varying from 160° to 180°, according to the fancy of the brewer. The infusion is left covered for two or three hours, and then the liquid drawn off by a cock at the bottom of the vessel. More hot water is afterwards added, and the infusions repeated in a similar manner till the malt is sufficiently exhausted.

The liquid thus obtained is called *wort*. It has a brown colour, a luscious sweet taste, a peculiar smell, and when the process is properly conducted is perfectly transparent. It consists of the water employed holding the mealy part of the malt in solution. When examined by reagents, it appears to consist chiefly of four different substances held in solution together: namely, 1. A sweet-tasted substance, to which the name of *saccharine matter* has been given. It is by far the greatest in point of quality. This substance when separated has a light-brown colour; when dried at 160°, it forms a brittle mass with a glazed surface; at the temperature of 180°, or a little higher, its colour becomes darker; and if we keep it in that temperature, moistening it occasionally, it becomes at last almost black, loses its sweet taste altogether, and acquires a sharp disagreeable one. In a temperature somewhat higher, but under the boiling point, it is charred. It is very soluble in water, and if once dissolved can scarcely be recovered again by evaporation without considerable loss. Alcohol dissolves it very imperfectly cold, and when heat is ap-

Book IV. plied it deprives the alcohol of a portion of water, and forms a tough mass like turpentine, which refuses to dissolve. The specific gravity of this saccharine matter is 1.552. It appears to be the essential constituent of wort. 2. Starch. The presence of this substance is easily detected by dropping the infusion of nutgalls into wort; a precipitate appears, which is mostly redissolved by heating the liquid to 120°. 3. The insoluble part of the precipitate is a combination of gluten and tannin. The proportion of gluten in wort is very inconsiderable. That of the starch probably diminishes in proportion as the barley is more completely malted. I have detected starch in pretty old and perfectly transparent ale; but the gluten had disappeared. New ale, however, often contains traces of it. 4. The fourth ingredient in wort is *mucilage*, which precipitates in flakes when the wort is dropped into alcohol. Its quantity is more considerable in the worts last drawn off than in the first worts.

The wort is boiled with a quantity of hops, which vary considerably, but which may be stated in general at about $\frac{1}{40}$ th of the weight of the malt. When sufficiently concentrated, it is thrown into very wide flat vessels called *coolers*, placed in the most open situation that can be had. Here it is cooled down to the temperature of about 52°, and then let down into a deep round wooden vessel called the *fermenting tun*. The specific gravity of the wort when put into the fermenting tun varies very much. The wort of strong ale is sometimes as low as 1.060; sometimes as high as 1.127. In the first case it contains 14.25 per cent. of solid matter; in the second 28.2 per cent. The wort of small beer varies from 1.015 to 1.040; the first containing not quite 3.5 per cent. and the second about 9.5 per cent. of solid matter.

Hops.

The use of the hops is partly to communicate a peculiar flavour, from the oil which they contain; partly to cover the sweetness of saccharine matter by the bitter principle which they contain; and partly to counteract the tendency which wort has to run into acidity.

If the wort be let down into the fermenting tun at the temperature of 60°, or rather higher, the substances which it holds in solution begin gradually to act upon each other, and to decompose each other. The temperature rises, an intestine motion takes place, a scum collects on the surface, and carbonic acid gas is emitted. This intestine motion is called *fermentation*. The disposition of wort to ferment is not suffi-

ciently great to induce it to do so with the requisite rapidity. It takes place so slowly and imperfectly, that the liquor runs into acidity before the formation of ale has made sufficient progress. To prevent this, it is requisite to add to the wort some substance which has the property of inducing speedy fermentation. The substance made choice of is *yeast*, or the frothy matter which collects on the surface of the beer while fermenting. The quantity of yeast used by brewers is but small, generally about a gallon to every three barrels of the wort.

The yeast thus added appears to act chiefly upon the saccharine matter held in solution in the wort, to decompose it while it undergoes partial decomposition itself. By this mutual action the saccharine matter disappears, the specific gravity of the wort diminishes, and its properties alter, being converted into the intoxicating liquor called *ale*. During this mutual action the temperature of the liquid increases, and the increase depends upon the violence of the fermentation. In ale worts the rise of temperature is but small, amounting at an average to about 15°, because the quantity of yeast is small; but in the fermentation of wash, it often amounts to 50°, or even more in some cases.

Considerable exertions have been made by chemists, of late, to ascertain the nature of the substance in yeast which produces this striking effect upon wort, and to discover it if possible in other substances. Westrumb examined the yeast of beer. From 15360 parts of it he obtained the following substances :

Analysis of yeast.

Potash	13
Carbonic acid	15
Acetic acid	10
Malic acid	45
Lime	69
Alcohol	240
Extractive	120
Mucilage	240
Saccharine matter	315
Gluten	480
Water	13595
	<hr/>
	15142

Book IV.
 Gluten its
 essential
 constituents.

Besides some traces of phosphoric acid and of silica.* But it is obvious, that all these ingredients are not essential. From the experiments of Westrumb, it appears that when the yeast is filtered, a matter remains upon the filter which possesses the properties of gluten; that when this substance is separated the yeast loses the property of exciting fermentation, but recovers it again when the gluten is added. Hence it follows that this glutinous matter is the essential constituent of the yeast. When yeast is kept for some time in cylindrical glass vessels, a white substance not unlike curd separates, and swims upon the surface. If this substance be removed, the yeast loses the property of exciting fermentation. This substance possesses many of the properties of gluten, though it differs from it in others. Its colour is much whiter; it has not the same elasticity; and its particles do not adhere with the same force. It dissolves more readily in acids. I have little doubt that this is the part of yeast which is the real ferment. It may be considered as gluten somewhat altered, and much more disposed to decomposition. It existed, no doubt, in the raw grain, but underwent considerable modifications during the process of malting, and probably others during the fermentation of the beer from which it separated.

That a substance analogous to gluten is the real ferment, appears also from the experiments of Fabroni, who published a Treatise on Wine-making in 1785, which gained the prize proposed by the Florence Academy. By heating the juice of grapes, and passing it through a filter, he separated an adhesive matter which possessed the properties of gluten. The juice deprived of this substance refused to ferment, but it fermented as usual when the glutinous matter was again added. The observations of Thenard confirm those of Fabroni. He found in the juices of all fruits examined, a substance similar to that described by Fabroni, and which, according to him, is absolutely the same with pure yeast. This substance is insipid, does not change vegetable blues, is insoluble in water, loses $\frac{3}{4}$ ths of its weight when dried, and is decomposed like animal substances. When eight parts of it were distilled, they left 2.83 of charcoal, and yielded 1.61 of water, 1.81 of oil, and a quantity of ammonia, which, when saturated with muriatic acid, formed 1.46 of sal ammoniac; the gas obtained

* Crell's Annals, 1796, i. 13.

weighed 0.33, and consisted of $\frac{1}{5}$ th carbonic acid, and $\frac{4}{5}$ ths of carburetted hydrogen, requiring $1\frac{1}{2}$ time its bulk of oxygen to consume it. Nitric acid, even when diluted, converts it into a species of tallow. With potash it forms a soap, while ammonia is disengaged. When mixed with sugar and a sufficient quantity of water, fermentation takes place, carbonic acid is disengaged, and a vinous liquor formed. By this action the ferment loses the whole of its azote, and becomes incapable of exciting fermentation when mixed with a new portion of sugar.* An experiment of Kirchoff throws considerable light on the nature of yeast. Barley-meal contains both gluten and starch. If pure starch be infused in hot water it is not converted into sugar. Neither does gluten become saccharine matter when treated in the same way. But if a mixture of pure, dried, pulverized wheat gluten and potatoe starch be infused in hot water, the starch is converted into sugar. During the process an acid is evolved. The gluten is little altered in appearance, and if the liquid be filtered most of it remains upon the filter. But it cannot be employed successfully a second time to convert starch into sugar. It appears then that it is some substance connected with the gluten that acts upon the starch, and converts it into sugar.†

The essential constituent of yeast, then, may be considered as a species of gluten, differing in several respects from the gluten of wheat. When mixed with the wort, this substance acts upon the saccharine matter; the temperature rises, carbonic acid is disengaged, and the saccharine matter is converted into ale. The yeast soon collects on the top of the liquid, but the brewer occasionally mixes it again to continue the fermentation. The quantity of yeast employed being small, the saccharine matter is but imperfectly decomposed. Hence a considerable portion of it still remains in the ale, and gives it that glutinousness and body for which it is remarkable. The specific gravity of ale varies very much according to the original strength of the wort, and the extent to which the fermentation has been carried. The limits may be stated at about 1.035 and 1.012.

The properties of ale are so well known that no description is deemed necessary. It possesses intoxicating qualities, and holds a quantity of alcohol in solution; which varies consider-

* Thénard, Ann. de Chim. xvi. 303.

† Schweigger's Jour. xiv. 339.

Book IV.
Alcohol in
ale and beer.

ably according to the original strength of the wort. I distilled a quantity of London brewed ale. The specific gravity of the wort was 1·0676. The specific gravity of the ale was 1·0255. One hundred parts of it by weight yielded

9·354 parts of proof spirit of specific gravity . . . 0·91985

Or,

5·817 parts of alcohol of the specific gravity . . . 0·825

Another specimen of ale was distilled. The specific gravity of the wort was 1·0813. The specific gravity of the ale after fermentation was 1·02295. One hundred parts of it by weight yielded

11·13 parts of proof spirit.

6·92 parts of alcohol of the specific gravity 0·825.

Mr. Brande* distilled ale and brown stout porter. The quantity of alcohol which he extracted from each by measure was as follows:

Brown stout 6·80 per cent.

Ale 8·88

When reduced to weight, the quantity of alcohol of the specific gravity 0·825 obtained from each, will be as follows:

Brown stout 5·61 per cent.

Ale 7·33

II. WASH.

Ale being intended as an article of food, and its value depending in a great measure on its flavour and appearance, particular attention is paid to obtain these in as great perfection as possible. But there is another species of ale which is brewed by the distillers for the express purpose of procuring from it ardent spirits by a subsequent process. The method which they follow is in several respects different. In particular, they endeavour to carry the fermentation to as great a length as possible, because the quantity of spirits is proportional to the saccharine matter decomposed, whatever remains unaltered yielding none. It is here, therefore, that the effects of fermentation can be best observed.

* Phil. Trans. 1811, p. 346.

1. In this country the distillers do not brew from pure malt; they use chiefly raw grain. The proportion of malt varies from a third to a tenth part of the raw grain employed. This mixture they grind to meal, infuse with water at a heat considerably lower than that of the water used by the brewers, and employ much more agitation to mix it completely. The wort is drawn off and cooled in the usual way, and fresh water poured on to exhaust the grain.

Chap. IV.
Brewing.

The wort thus formed is not so transparent as that from malt, but its taste is nearly as sweet. It would appear, therefore, that the starch in the raw grain undergoes a certain change during the mashing, and is brought towards the state of saccharine matter.

In this country, where the duty is levied chiefly upon the wash,* the distillers make the specific gravity of their wort as high as from 1.084 to 1.110: this they do, not by boiling, but by *lobbing*; that is, by preparing a strong infusion of the flour of malt, or of barley and malt and hot water, and adding this almost saturated solution to the wort till it has acquired the requisite strength. But in Holland, where the duties are levied in a different way, the specific gravity of the wort is much lower.

The wort thus made is let down into the fermenting tun at a temperature varying from 55° to 70°, according to the quantity, the season, the goodness of the yeast, and the skill of the distiller. Here it is mixed successively with considerable portions of the best yeast that can be procured, and the fermentation is urged as far as it can be made to go. The process lasts about ten days, and the temperature rises usually to between 90° and 100°, and sometimes even higher. Great quantities of carbonic acid are disengaged, and the liquid becomes specifically lighter; sometimes sinking to 1.000, and usually to from 1.007 to 1.002. The success of the fermentation is estimated by this loss of specific gravity.

Fermenta-
tion.

The wash thus prepared is distilled. What comes over first is denominated *low wine*, and is concentrated by a second distillation.

This fermentation is obviously the consequence of the action of the peculiar ferment in yeast upon the saccharine matter of the wort. Even when the fermentation is carried to its greatest

Always In-
complete.

* This is the name given to the fermented wort of distillers.

Book IV.

extent it does not appear that the whole solid matter held in solution by the wort is decomposed. Nine trials were made upon the wort of pure malt; in all of them the fermentation went on pretty successfully, as may be seen from the following table:

Specific gravity of the wort.	Specific gravity of the wash.
1·0400	1·0014
1·0560	1·0016
1·0500	1·0000
1·0492	1·0012
1·0465	1·0045
1·0450	1·0047
1·0465	1·0007
1·0510	1·0007
1·0524	1·0004

When a portion of these different liquids was evaporated to dryness, the quantity of solid matter which it left was found to amount to $\frac{1}{5}$ th, at an average, of the original quantity. Thus $\frac{4}{5}$ ths had been decomposed by the fermentation, and $\frac{1}{5}$ th still remained. This matter was still capable of fermentation, when redissolved in water and mixed with fresh yeast. On comparing the quantity of alcohol of 0·825, obtained in these trials with the weight of solid matter of the malt which had been decomposed by the fermentation, the result was, that every pound weight of solid matter, so decomposed, furnished almost exactly half a pound of alcohol, of the specific gravity 0·825.

When sugar, dissolved in four times its weight of water, and mixed with yeast, is placed in the proper temperature, it ferments precisely as wort does, and yields the same products. It has been employed, therefore, by chemists as a less complicated means of ascertaining the phenomena of fermentation. Thenard mixed 60 parts of yeast with 300 of sugar, and fermented them in the temperature of 59°. In four or five days, he informs us, that all the saccharine matter had disappeared.*

Produce. The quantity of carbonic acid evolved amounted, by weight, to 94·6 parts. It was perfectly pure, being completely absorbed

* There is reason to doubt the precision of this assertion, as it never happens in the great scale, where every thing is much more favourable.

by water.* The fermented liquid, being distilled, yielded 171·5 parts of alcohol, of the specific gravity ·822. When the residue of the distillation was evaporated, 12 parts of a nauseous acid substance were obtained; and 40 parts of the yeast still remained; but, upon examination, it had lost the whole of its azote. This experiment gives us the following quantities:

1. *Substances fermented.*

Sugar	300
Yeast	60
	360

2. *Products of fermentation.*

Alcohol of ·822 ..	171·5
Carbonic acid ...	94·6
Nauseous residue.	12·0
Residual yeast ...	40·0
	318·1
Loss	41·9

But as the nauseous residue and residual yeast, nearly make up the quantity of yeast employed, let us consider only the products of decomposed sugar, supposing the loss to be proportionally divided between the carbonic acid and alcohol. Now alcohol of the specific gravity 0·822 contains one tenth of its weight of water, which can be separated from it; and if we suppose with Saussure, that absolute alcohol contains 8·3 per cent. of water, then the products of sugar decomposed by fermentation, according to the preceding experiment, are as follows:

Alcohol	47·70
Carbonic acid .	35·34
	83·04

Or in 100 parts

Alcohol	57·44
Carbonic acid .	42·56
	100·00

* The very pungent aromatic odour which carbonic acid has when disengaged from the fermenting tun shows that it contains a portion of the wash; and this has been verified by actual experiment.

Book IV.

This result approaches so nearly that of Lavoisier, that there is reason to suspect that the coincidence is more than accidental.

We have found reason to conclude that sugar is composed of

$$\begin{array}{r}
 5 \text{ atoms oxygen} \dots\dots = 5\cdot000 \\
 6 \text{ atoms carbon} \dots\dots = 4\cdot500 \\
 5 \text{ atoms hydrogen} \dots\dots = 0\cdot625 \\
 \hline
 10\cdot125 \\
 \hline
 \end{array}$$

Alcohol appears to be a compound of

$$\begin{array}{r}
 1 \text{ atom oxygen} \dots\dots = 1\cdot000 \\
 2 \text{ atoms carbon} \dots\dots = 1\cdot500 \\
 3 \text{ atoms hydrogen} \dots\dots = 0\cdot375 \\
 \hline
 2\cdot875
 \end{array}$$

And carbonic acid of

$$\begin{array}{r}
 2 \text{ atoms oxygen} = 2\cdot00 \\
 1 \text{ atom carbon} = 0\cdot75 \\
 \hline
 2\cdot75
 \end{array}$$

Now if we suppose 1 atom of sugar and 1 atom of water to be decomposed by the fermentation, it is obvious, that they may be converted into 2 atoms of alcohol and 2 atoms of carbonic acid. For an atom of sugar and an atom of water are composed of

- 6 atoms oxygen.
- 6 atoms carbon.
- 6 atoms hydrogen.

	Oxygen.	Carbon.	Hydrogen.
2 atoms of alcohol consist of ...	2 atoms,	4 atoms,	6 atoms.
2 atoms of carbonic acid of	4	2	
Making together	6	6	6

On that supposition, the weight of }
the alcohol evolved ought to be. } 5.75
And that of the carbonic acid 5.50

Or per cent.

$$\begin{array}{r}
 \text{Alcohol} \dots\dots\dots 51\cdot12 \\
 \text{Carbonic acid} \dots\dots 48\cdot88 \\
 \hline
 100\cdot00
 \end{array}$$

Fabroni found that the gluten of wheat acted but imperfectly as a ferment; but that its efficacy was much improved by the addition of tartar. Berthollet repeated these experiments successfully. He ascribes the efficacy of tartar to the property which it has of promoting the solubility of the gluten. The presence of an acid was supposed formerly to be necessary for fermentation; but this does not seem to be the case. It is true, indeed, that an acid usually makes its appearance during fermentation. The formation of this acid has been ascribed to the action of the yeast upon the mucilaginous or starchy parts of the wort: but from the experiments of Fourcroy and Vauquelin, it appears that it always makes its appearance when wort is fermented without any yeast. In these trials they obtained only vinegar, and no alcohol. When the wort, either of raw grain or of malt, is fermented at the temperature of 80°, without any yeast, the gas which comes over consists of one-half carbonic acid and one-half hydrogen; but at a lower temperature pure malt-wort does not yield any inflammable gas.*

Chap. IV.
Whether an acid be necessary for fermentation.

III. WINE.

There is a considerable number of ripe fruits from which a sweet liquor may be expressed, having at the same time a certain degree of acidity. Of such fruits we have in this country the apple, the cherry, the gooseberry, the currant, &c. but by far the most valuable of these fruits is the *grape*, which grows luxuriantly in the southern parts of Europe. From grapes, fully ripe, may be expressed a liquid of a sweet taste, to which the name of *must* has been given. This liquid is composed almost entirely of five ingredients; namely, *water*, *sugar*, *jelly*, *gluten*, and *tartaric acid* partly saturated with potash. The quantity of sugar which grapes, fully ripe, contain, is very considerable; it may be obtained in crystals by evaporating must to the consistence of syrup, separating the tartar which precipitates during the evaporation, and then setting the must aside for some months. The crystals of sugar are gradually formed.† From a French pint of must, the Marquis de Bouillon extracted half an ounce (French) of sugar, and $\frac{1}{16}$ th ounce of tartar.‡ According to Proust, the Muscadine grape contains about 30 per cent. of a peculiar species of sugar.§

Fruits affording wine.

Must.

* Ann. des Mus. d'Hist. vii. 16.

† Bouillon, Jour. de Phys. xxix. 3.

‡ Ibid. p. 5.

§ Ibid. lvi. 113.

Book IV.
 Undergoes
 the vinous
 fermenta-
 tion.

When must is put into the temperature of about 70° , the different ingredients begin to act upon each other, and what is called *vinous fermentation* commences. The phenomena of this fermentation are an intestine motion in the liquid; it becomes thick and muddy, its temperature increases, and carbonic acid gas is evolved. In short, the very same changes take place as have been remarked when describing the fermentation of ale. In a few days the fermentation ceases, the thick part subsides to the bottom, or rises to the surface, the liquid becomes clear, it has lost its saccharine taste, and assumed a new one; its specific gravity is diminished; and it has become the liquid well known under the name of *wine*.

As this fermentation takes place without adding any *ferment*, it is obvious that the requisite substance is present in the juice. This substance was separated, and found by Fabroni to be analogous to the gluten of plants; and gluten being substituted for it, the fermentation succeeded. Fabroni has shown that the saccharine part of must resides in the cells of the grapes; while the glutinous matter, or ferment, is lodged on the membranes that separate the cells. Hence it follows, that in the fruit these two substances are not in contact. It is only after the juice is squeezed out that they are mixed. All other juices which undergo a spontaneous fermentation at the requisite temperature, have been shown by Thenard and Seguin to contain a similar substance. The formation of wine, then, is owing to the action of this glutinous matter on the saccharine substance of the juice, precisely as happens in the fermentation of ale.

Gay-Lussac has shown that the juice of grapes, and indeed the juice of all fruits will not ferment, if completely excluded from the air. But if a little oxygen gas be let up to it, this gas is immediately absorbed, a brisk fermentation commences, and the quantity of carbonic acid gas evolved is 100 times as great as that of the oxygen gas absorbed.* It would be curious to know whether the same thing holds with respect to the wort of malt. It is not improbable that it does; but wort ferments so imperfectly without the addition of yeast, that it would not be easy to try the experiment.

All those juices of fruits which undergo the vinous fermentation, either with or without the addition of sugar, contain an acid. We have seen already, in the First Chapter, that the

* Ann. de Chim. lxxvi. 245

vegetable acids are obtained chiefly from fruits. The apple, for instance, contains malic acid; the lemon, citric acid; the grape, tartaric and malic acids. The Marquis de Bouillon has ascertained that must will not ferment if all the tartar which it contains be separated from it; but it ferments perfectly well on restoring that salt.* The same chemist ascertained that the strength of wine is considerably increased by adding tartar and sugar to the must.† We may conclude from these facts, that the presence of a vegetable acid is of importance in these spontaneous fermentations. It deserves attention, that Bouillon obtained more tartar from verjuice than from wine; and he observed, that the more the proportion of sugar in grapes increased, the more that of tartar diminished.‡

It seems more than probable, from the experiments of Bouillon and Chaptal, that the tartaric acid is partly decomposed during the fermentation, and that a portion of malic acid is formed. The process, therefore, is more complicated than was suspected by Lavoisier. It is obviously analogous to combustion, as is evident from the evolution of caloric and the formation of carbonic acid, which is a product of combustion. Proust has ascertained that, during the fermentation, not only carbonic acid, but azotic gas also, is disengaged. This is a demonstration that all the constituents of must are concerned; for sugar does not contain that principle.§ Thenard could detect no azote in the carbonic acid from wort.

After the fermentation has ceased, the liquor is put into casks, where the remainder of the sugar is decomposed by a slow fermentation; after which the wine, decanted off the extractive matter, is put up in bottles.

The properties of wine differ very much from each other, according to the nature of the grapes from which the must was extracted, and according to the manner in which the process was conducted. These differences are too well known to require a particular description. But all wines contain less or more of the following ingredients; not to mention water, which constitutes a very great proportion of every wine.

1. *An Acid*.—All wines give a red colour to paper stained with turnsole, and of course contain an acid. Chaptal has ascertained that the acid found in greatest abundance in wine is the malic, but he found traces also of citric acid; and it is

Component
parts of
wine.

An acid,

* Jour. de Phys. xxix. 4. But the addition of salt of wood sorrel did not restore the fermentation.

† Ann. de Chim. xxxvi. 20.

‡ Jour. de Phys. xxix. 4.

§ Ibid. lvi. 113.

Book IV. } probable that wine is never entirely destitute of tartar. All wines which have the property of frothing when poured into a glass contain also carbonic acid, to which they owe their briskness. This is the case with champagne. These wines are usually weak; their fermentation proceeds slowly, and they are put up in close vessels before it be over. Hence they retain the last portions of carbonic acid that have been evolved.

Alcohol, 2. *Alcohol*.—All wine contains less or more of this principle, to which it is indebted for its strength; but in what particular state of combination it exists in wine cannot easily be ascertained. It is undoubtedly intimately combined with the other component parts of wine; but Mr. Brande has shown by very decisive experiments that all wines contain alcohol ready formed, and that it is merely separated during the distillation of these liquids, and not formed as had been advanced by Fabroni.* These experiments were afterwards confirmed by Gay-Lussac.† When wine is distilled, the alcohol readily separates. The distillation is usually continued as long as the liquid which comes over is inflammable. The quantity obtained varies according to the wine, from a fourth to a fourteenth part of the wine distilled. The spirit thus obtained is well known under the name of *brandy*. Bouillon has observed, that when wine is distilled new it yields more alcohol than if it be allowed to get old.‡ What remains after this distillation is distinguished in France by the name of *vinasse*. It consists of tartar, &c. and when evaporated to dryness, and subjected to combustion, yields potash.

Extract, 3. *Extractive matter*.—This matter exists in all wines: but its proportion diminishes according to the age of the wine, as it gradually precipitates to the bottom.

oil, 4. Every wine is distinguished by a peculiar flavour and odour, which probably depends upon the presence of a *volatile oil*, so small in quantity that it cannot be separated.

And colouring matter. 5. The *colouring matter* of wine is originally contained in the husk of the grape, and is not dissolved till the alcohol be developed. This matter is analogous to the other colouring matters of plants: a set of bodies possessed of remarkable properties, but too little examined hitherto to be introduced with much advantage into a System of Chemistry. This colouring matter precipitates when the wine is exposed to the heat of the sun. It sometimes also precipitates in old wine,

* Phil. Trans. 1811, p. 337; and 1813, p. 82.

† Ann. de Chim. lxxxvi. 175.

‡ Jour. de Phys. xxix. 6.

and it may be easily separated by pouring lime-water into wine.*

From the experiments of Vogel it appears that the colouring matter of red wine is precipitated greenish-grey by acetate of lead. By this property we can distinguish whether a red wine be genuine or factitious.†

The following table, containing the different substances which Neumann extracted from various wines, is worth preserving.‡

A quart of	Highly rectified Spirit.			Thick, oily nunctuous, resinous matter.			Gummy and tartarous matter.			Water.			
	oz.	dr.	gr.	oz.	dr.	gr.	oz.	dr.	gr.	lb.	oz.	dr.	gr.
Aland	1	6	00	3	2	00	1	5	00	2	5	3	00
Alicant	3	6	00	6	0	20	0	1	40	2	2	6	00
Burgundy . . .	2	2	00	0	4	00	0	1	40	2	9	0	20
Carcassone . .	2	6	00	0	4	10	0	1	20	2	8	4	30
Champagne . .	2	5	20	0	6	40	0	1	00	2	8	3	00
French	3	0	00	0	6	40	0	1	00	2	8	0	20
Frontignac . .	3	0	00	3	4	00	0	5	20	2	4	6	30
Vin de Grave .	2	0	00	0	6	00	0	2	00	2	9	0	00
Hermitage . . .	2	7	00	1	2	00	0	1	40	2	7	5	20
Madeira	2	3	00	3	2	00	2	0	00	2	4	3	00
Malmsey	4	0	00	4	3	00	2	3	00	2	1	2	00
Vino de Monte } Pulciano }	2	6	00	0	3	00	0	2	40	2	8	0	20
Moselle	2	2	00	0	4	20	0	1	30	2	9	0	10
Muscadine . . .	3	0	00	2	4	00	1	0	00	2	5	4	00
Neufschatel . .	3	2	00	4	0	00	1	7	00	2	2	7	00
Palm Sec. . . .	2	3	00	2	4	00	4	4	00	2	2	5	00
Pontac	2	0	00	0	5	20	0	2	00	2	9	0	40
Old Rhenish . .	2	0	00	1	0	00	0	2	20	2	8	5	40
Rhenish	2	2	00	0	3	20	0	1	34	2	9	1	06
Salamanca . . .	3	0	00	3	4	00	2	0	00	2	3	4	00
Sherry	3	0	00	6	0	00	2	2	00	2	0	6	00
Spanish	1	2	00	2	4	00	9	4	00	1	10	6	00
Vino Tinto . . .	3	0	00	6	4	00	1	6	00	2	0	6	00
Tokay	2	2	00	4	3	00	5	0	00	2	0	3	00
Tyrol red wine } wine }	1	4	00	1	2	00	0	4	00	2	8	6	00
Red wine	1	6	00	0	4	40	0	2	20	2	9	3	20
White	2	2	00	0	7	00	0	3	00	2	7	0	00

* The most precise account of wine and of the vinous fermentation, hitherto published, is by Chaptal, and is contained in the 36th and 37th volumes of the Ann. de Chim.

† Annals of Philosophy, viii. 70.

‡ Neumann's Chemistry, p. 447.

Book IV.

To this head belong not only common wine, but all the intoxicating liquors made from vegetable juices; as *cyder* from apples, *perry* from pears, *currant wine*, &c. likewise the liquor made from the juice of the sugar cane, the sugar maple, &c.

I shall here subjoin a table of the quantity of alcohol by measure of the specific gravity 0·825 which different wines yielded to Mr. Brande in his different experiments. The results of these trials are the more valuable, because the wines examined were all genuine.*

	Alcohol per cent. by measure.		Alcohol per cent. by measure.
Lissa	26·00	Marsala	26·30
Ditto	24·00	Ditto	25·50
Port	21·40	Red Champagne	11·30
Ditto	22·30	White Champagne ..	12·80
Ditto	23·39	Burgundy	14·53
Ditto	23·71	Ditto	11·95
Ditto	24·29	White Hermitage	17·43
Ditto	25·83	Red Hermitage	12·32
Ditto	21·40	Hock	14·37
Ditto	19·00	Ditto	8·88
Madeira	19·34	Vin de Grave	12·80
Ditto	21·40	Frontignac	12·79
Ditto	23·93	Cote Roti	12·32
Ditto	24·42	Rousillon	17·26
Sherry	18·25	Cape Madeira	18·11
Ditto	18·79	Cape Muschat	18·25
Ditto	19·81	Constantia	19·75
Ditto	19·83	Tent	13·30
Claret	12·91	Sheraz	15·52
Ditto	14·08	Syracuse	15·28
Ditto	16·32	Nice	14·63
Calcavella	18·10	Tokey	9·88
Lisbon	18·94	Raisin Wine	25·77
Malaga	17·26	Grape Wine	18·11
Ditto, kept since 1666 .	18·00	Currant Wine	20·55
Bucellas	18·49	Gooseberry Wine	11·84
Red Madeira	18·40	Elder Wine	9·87
Malmsey Madeira	16·40	Cyder	9·87
Marsala	25·87	Perry	9·87
Ditto	17·26		

* Phil. Trans. 1811, p. 345; 1813, p. 87. Journal of the Royal Institution, i. 136.

SECT. II.

OF THE ACETOUS FERMENTATION.

IF wine or beer be kept in a temperature between 70° and 90°, it gradually becomes thick, its temperature augments, filaments are seen moving through it in every direction, and a kind of hissing noise may be distinguished. These intestine motions gradually disappear, the filaments attach themselves to the sides and bottom of the vessel, and the liquor becomes transparent. But it has now lost its former properties, and is converted into acetous acid. This intestine decomposition has been long distinguished by the name of *acetous fermentation*, because its product is *acetic acid*. That this fermentation may take place, certain conditions must be attended to. The most important of these will appear from the following observations:

Chap. IV.
Wine ferments and becomes sour.

1. Neither pure alcohol, nor alcohol diluted with water, is susceptible of this change. The weaker the wine or the beer is on which the experiment is made, the more readily it is converted into vinegar: the stronger they are they resist the change with the greater obstinacy. But it results from the experiments of Beecher, that strong wines, when they are made to undergo the acetous fermentation, yield a much better and stronger vinegar than weak wines. Hence it follows that alcohol, though of itself it refuses to undergo the change, yet when other bodies are present which readily ferment, is decomposed during the process, and contributes to the formation of the acetic acid.

Alcohol not susceptible of this change.

2. Wine, entirely deprived of glutinous matter, either by spontaneous deposition or by clarification, does not undergo the acetous fermentation, unless some mucilaginous matter be mixed with it. Chaptal exposed old wine destitute of this matter, in open bottles, to the greatest summer heat of Montpellier for 40 days, and yet it did not become sour: but upon adding some vine leaves to the same wine, it became acid in a few days.* When the water in which gluten of wheat has been allowed to ferment is mixed with sugar, the liquid is converted into vinegar without fermentation, without effer-

Glutinous matter necessary.

* Ann. de Chim. xxxvi. 245.

Book IV.

vescence, and without the contact of air.* The nature of this curious change has not been explained

Carbonic acid formed.

3. Wine never becomes sour, provided it be completely deprived of all access to atmospheric air. The reason is, that during the acetous fermentation the oxygen of the atmosphere is partly converted into carbonic acid. It was supposed to have been absorbed; but the late experiments of Saussure have shown that this is not the case. He found, that when wine was converted into vinegar in a given portion of air, the bulk was not diminished; the quantity of carbonic acid just compensating the oxygen which had disappeared, excepting in as far as that carbonic acid was absorbed by the liquid. The quantity of carbonic acid formed during this conversion of wine into vinegar does not seem to be great; amounting, in Saussure's trials, to about twice the bulk of the wine.† Hence the reason that wine or beer is apt to become sour after the cork has been drawn, and still more apt when part has been poured out of the bottle.

A certain temperature required.

4. A pretty high temperature is necessary for the commencement of the acetous fermentation. Wine or beer (unless very weak) scarcely becomes sour under the temperature of 65° or 70°. The fermentation is very apt to commence when the temperature suddenly rises. It is said, too, that wine and beer are more apt to become sour at certain seasons of the year than at others.

The vinegar makers in this country keep their wash in stoves, heated to about the temperature of 80°, till the change is completed. From the experiments of Fourcroy and Vauquelin, it appears that vinegar made from grain holds in solution a quantity of gluten, and that a great part of it is separated by heating the vinegar boiling hot. This separation prevents the vinegar from being so apt to spoil.

Effects of the fermentation.

5. When the acetous fermentation is completed, the whole of the malic acid originally contained in the wine has disappeared as well as the alcohol. We must conclude, therefore, that they have been both converted into acetic acid. Part of the glutinous matter has also undergone the same change, and seems indeed to have been the substance that first began the change. Part of it is deposited in the state of flakes; part

* Fourcroy and Vauquelin, Ann. de Mus. d'Hist. Nat. vii. 4.

† Recherches Chimique sur la Vegetation, p. 144.

remains in solution, and disposes the vinegar to decomposition. Vinegar also contains a little tartar, and probably also citric acid. Malic acid is also found in new vinegar; a proof that this part of the wine is the last to undergo the acetous fermentation.

Chap. IV.

6. From the experiments of Cadet it appears, that sugar is the essential constituent in liquors to be converted into vinegar, and that the quantity of vinegar formed is proportional to the sugar. But if the sugar exceed an eighth part of the liquid, the whole is not decomposed. Seven water, one sugar, and some yeast, ferment in a proper temperature, and form an excellent vinegar.*

7. Acetic acid is formed in many other cases of the decomposition of vegetables besides the acetous fermentation. These have been pointed out with much ingenuity by Vauquelin and Fourcroy. They may be reduced under three heads. First, when sugar, gum, wood, &c. are distilled in a retort, or even burnt in the open fire, acetic acid separates in combination with an empyreumatic oil, which gives it a peculiar odour. Hence it was mistaken for other acids, and distinguished by the names of *pyromucous*, *pyrolignous*, acids, till its real nature was ascertained by these distinguished chemists.† Secondly, when concentrated sulphuric acid is poured upon the same vegetable bodies, they are decomposed in a very different manner; being converted into water, charcoal, and acetic acid. Thirdly, acetic acid is evolved in considerable quantity during the spontaneous decomposition of urine and some other animal substances. Thus it appears that the component parts of acetic acid are extremely apt to combine together in those proportions which constitute that important acid.

Acetic acid formed by other processes.

SECT. III.

OF PUTREFACTION.

ALL vegetable substances, both complete plants and their component parts separately, when left entirely to themselves, are gradually decomposed and destroyed, provided moisture be present, and the temperature be not much under 45°, nor too high to evaporate suddenly all the moisture. This decomposition has obtained the name of *putrefaction*.

Nature of putrefaction.

* Ann. de Chim. lxii. 248.

† Ibid. xxxv. 83.

Book IV.

It proceeds with most rapidity in the open air; but the contact of air is not absolutely necessary. Water is in all cases essential to the process, and therefore is most probably decomposed.

Putrefaction is constantly attended with a fetid odour, owing to the emission of certain gaseous matters, which differ according to the putrefying substance. Some vegetable substances, as gluten and cruciform plants, emit ammonia; others, as onions, seem to emit phosphuretted hydrogen gas. Carbonic acid gas and hydrogen gas, impregnated with unknown vegetable matters, are almost constantly emitted in abundance. From the experiments of Süssure, we learn, that when moist wood is left exposed to the air, a portion of the oxygen of the atmosphere is converted into carbonic acid; for the quantity of this acid formed is just equal to the bulk of the oxygen which disappears. When the experiment is made under receivers, no other gas can be detected except carbonic acid; but in the open air the wood loses a much greater proportion of its weight than can be accounted for by the carbon thus abstracted. This he ascribes to the escape of oxygen and hydrogen under the form of water. Hence the quantity of carbon in a given weight of rotten wood is greater than in the same weight of fresh; but when the putrefaction goes on in water without the contact of air, the wood becomes white and light, and contains a smaller proportion of carbon than when fresh.* The nature and constituents of vegetable mould, as far as known, have been detailed in the preceding Chapter. Our chemical knowledge of vegetable compounds is still by far too limited to enable us to follow this very complicated process of putrefaction with any chance of success.

* Recherches Chimique sur la Vegetation.

BOOK V.

OF ANIMALS.

WHEN we compare animals and vegetables together, each in their most perfect state, nothing can be easier than to distinguish them. The plant is confined to a particular spot, and exhibits no mark of consciousness or intelligence; the animal, on the contrary, can remove at pleasure from one place to another, is possessed of consciousness, and a high degree of intelligence. But on approaching the contiguous extremities of the animal and vegetable kingdom, these striking differences gradually disappear, the objects acquire a greater degree of resemblance, and at last approach each other so nearly, that it is scarcely possible to decide whether some of those species which are situated on the very boundary belong to the animal or vegetable kingdom.

Book V.
Animals
and vegeta-
bles.

To draw a line of distinction, then, between animals and vegetables, would be a very difficult task: but it is not necessary at present to attempt it; for almost the only animals whose bodies have been hitherto examined with any degree of chemical accuracy, belong to the most perfect classes, and consequently are in no danger of being confounded with plants. Indeed, the greater number of facts which I have to relate apply only to the human body, and to those of a few domestic animals. The task of analysing all animal bodies is immense, and must be the work of ages of indefatigable industry.

Not easily
distinguish-
ed.

This part of the subject naturally divides itself into four Chapters. In the First Chapter, I shall give an account of the different ingredients hitherto found in animals, such of them at least as have been examined with any degree of accuracy: in the second, I shall treat of the different members of which animal bodies are composed; which must consist each of various combinations of the ingredients described in the first Chapter: in the third, I shall treat of those animal functions which may be elucidated by chemistry: and, in the fourth, of the changes which animal bodies undergo after death.

CHAP. I.

OF ANIMAL SUBSTANCES.

Book V.

THE substances which have been hitherto detected in the animal kingdom, and of which the different parts of animals, as far as these parts have been analysed, are found to be composed, may be arranged under the following heads :

- | | |
|----------------------------------|------------------|
| 1. Gelatin. | 10. Cantharidin. |
| 2. Albumen. | 11. Cochenelin. |
| 3. Fibrin. | 12. Oils. |
| 4. Colouring matter of
blood. | 13. Resins. |
| 5. Mucus. | 14. Sulphur. |
| 6. Osmazome. | 15. Phosphorus. |
| 7. Picromel. | 16. Acids. |
| 8. Urea or nephrin. | 17. Alkalies. |
| 9. Sugar. | 18. Earths. |
| | 19. Metals. |

These substances shall form the subject of the following sections.

SECT. I.

OF GELATIN.

How obtained.

1. If a piece of the fresh skin of an animal, an ox, for instance, after the hair and every impurity is carefully separated, be washed repeatedly in cold water till the liquid ceases to be coloured, or to abstract any thing; if the skin, thus purified, be put into a quantity of pure water, and boiled for some time, part of it will be dissolved. Let the decoction be slowly evaporated till it is reduced to a small quantity, and then put aside to cool. When cold, it will be found to have assumed a solid form, and to resemble precisely that tremulous substance well known to every body under the name of *jelly*. This is the substance called in chemistry *gelatin*. If the evaporation be still farther continued, by exposing the jelly to dry air, it becomes hard, semitransparent, breaks with a glassy fracture, and is in short the substance so much employed in the different arts under the name of *glue*. Gelatin, then, is precisely the same with glue; only that it must be supposed always free from those impurities with which glue is so often contaminated.

Properties.

2. Gelatin is semitransparent and colourless when pure. Its

consistency and hardness vary considerably. The best kinds are very hard, brittle, and break with a glassy fracture. Its taste is insipid, and it has no smell.

When thrown into water it swells very much, but does not readily dissolve: and when taken out, it is soft and gelatinous; but when allowed to dry, it recovers its former appearance. If it be put in this gelatinous state into warm water, it very soon dissolves, and forms a solution of an opal colour, and the more opaque according to the quantity of gelatin which it contains. Tremulous gelatin dissolves in a very small portion of hot water; but as the solution cools, it gelatinizes afresh. If this solution, as soon as it assumes the tremulous form, be mixed with cold water and shaken, a complete solution takes place.

From the experiments of Dr. Bostock, we learn that when one part of isinglass (which is nearly pure gelatin) is dissolved in 100 parts of hot water, the solution on cooling is wholly converted into a jelly. But one part of isinglass, in 150 parts of water, does not become concrete; though the solution is to a certain degree gelatinous.*

Dry gelatin undergoes no change when kept; but in the gelatinous state, or when dissolved in water, it very soon putrefies; an acid makes its appearance in the first place (probably the acetic,) a fetid odour is exhaled, and afterwards ammonia is formed.

When dry gelatin is exposed to heat, it whitens, curls up like horn, then blackens, and gradually consumes to a coal; but tremulous gelatin first melts, assuming a black colour. When distilled, it yields, like most animal substances, a watery liquid impregnated with ammonia, and a fetid empyreumatic oil; leaving a bulky charcoal of difficult incineration. It is by no means a very combustible substance.

3. Acids dissolve gelatin with facility, even when diluted, especially when assisted by heat; but we are still ignorant of the changes produced upon it by these agents except by nitric acid. When this acid is digested on it, a small quantity of azotic gas is disengaged, then abundance of nitrous gas; the gelatin is dissolved, except an oily matter which appears on the surface, and converted partly into oxalic and malic acids.†

Muriatic acid dissolves glue with great ease. The solution

* Nicholson's Jour. xi. 250.

† Scheele, Crel's Annal, ii. 47. Eng. Trans.

Book V.

is of a brown colour, and still continues strongly acid. It gradually lets fall a white powder. This solution precipitates tan in great abundance from water; and may be employed with advantage to detect tan when an alkali conceals it. Sulphuric acid acts much more slowly. The solution is brown, and gradually deepens; sulphurous acid is exhaled during the action of sulphuric acid on glue. Neither sulphuric nor muriatic acid occasion any change in the solution of glue in water.

When a current of chlorine gas is passed through a solution of gelatin in water, a white solid matter collects on the surface, and whitish filaments swim through the liquid. This solid matter, when separated by the filter and purified, possesses the following properties: its colour is white; it is specifically lighter than water; it has little or no taste; when dried in the open air it falls to powder; it is not soluble in boiling water; it dissolves in hot nitric and acetic acids, but precipitates again as the solution cools; when triturated with potash it emits the smell of ammonia; it does not affect vegetable blues.* Bouillon La Grange, to whom we are indebted for these facts, has given the gelatin thus altered the name of *oxygenized gelatin*.

Of alkalis
and earths.

Alkalies dissolve gelatin with facility, especially when assisted by heat; but the solution does not possess the properties of soap.

None of the earths seem to combine with gelatin; at least they do not precipitate it from its solution in water. The following table exhibits the effect of different earthy solutions when mixed with a pretty concentrated solution of common glue:

Substances.	Effects.
Lime water	No change.
Strontian water	No change.
Barytes water	} Became milky. Precipit. not dissolved by acetic acid.
Muriate of barytes	
Silicated potash	The same as the last.
Aluminated potash	No change.
Oxalate of ammonia	No change.
Phosphate of soda	Became milky.
	Became slightly milky.

The milkiness produced by some of these reagents was

* Bouillon La Grange, Nicholson's Jour. xiii. 209.

Of the metallic oxides.

not owing to their effect upon the gelatin, but upon the lime and the sulphuric acid which is contained.

The metals in their pure state have no effect upon gelatin; but several of the metallic oxides, when agitated in a solution of gelatin, have the property of depriving the water of the greatest part of that body, with which they form an insoluble compound. Several of the metallic salts likewise precipitate gelatin from water. The following table exhibits the result of mixing various metallic salts with a concentrated solution of gelatin, as far as my experiments have gone:

Metallic solutions.	Effects.
Nitromuriate of gold	{ A copious yellowish-white precipitate. Soluble by adding water.
Nitrate of silver	Becomes slightly milky.
Nitrate of mercury	{ A very copious curdy precipitate.
(Corrosive sublimate*	A copious white precipitate.
Supersulph. of mercury	No change.
Dry persulph. of mercury	{ Crystals become yellow, white flakes appear, and the liquid becomes transparent.
Prussiate of mercury	No change.
Pernitrate of copper	No change.
Muriate of copper	Becomes milky.
Persulphate of copper	No change.
Cuprate of ammonia	No change.
Sulphate of iron†	A few yellow flakes appear.
Persulphate of iron‡	{ Becomes slightly milky, as when alcohol is added.
Pernitrate of iron	Assumes a pink colour.
Permuriate of iron	Becomes green.
Nitromuriate of tin§	No change.
Permuriate of tin	Becomes slowly milky.

* Dr. Bostock, in his excellent paper on animal fluids, informs us, that a solution of one part of isinglass in 100 parts of water is not affected by corrosive sublimate. The result of my experiments with common glue was different; probably owing to the impurity of that substance.

† Partly in the state of persulphate.

‡ Dissolved in alcohol.

§ With excess of acid.

Book V.	Metallic solutions.	Effects.
{	Nitrate of lead	} No change.
	Acetate of lead	
	Plumbate of potash	
	Plumbate of lime	
	Muriate of zinc	No change.
	Muriate of antimony	A copious flaky precipitate.
	Tartar emetic	No change.
	Nitrate of bismuth precipitable by water	} Becomes milky.
	Ditto, not precipitable by water	
	Muriate of arsenic	No change.

Of alcohol.

Gelatin is insoluble in alcohol. When alcohol is mixed with a solution of gelatin, the mixture becomes milky; but becomes again transparent when agitated, unless the solution be concentrated, and the quantity of alcohol considerable. Gelatin is most probably equally insoluble in ether; though I believe the experiment has not been tried.

Of tannin.

When the solution of tannin is dropped into gelatin, a copious white precipitate appears, which soon forms an elastic adhesive mass, not unlike vegetable gluten. This precipitate is composed of gelatin and tannin; it soon dries in the open air, and forms a brittle resinous-like substance, insoluble in water, capable of resisting the greater number of chemical agents, and not susceptible of putrefaction. It resembles exactly overtanned leather. The precipitate is soluble in the solution of gelatin, as Davy first observed. Neither is the whole tan thrown down, unless the solutions both of tannin and gelatin be somewhat concentrated. Tremulous gelatin, as was first observed by the same chemist, does not precipitate tannin; but if we employ a solution of gelatin so strong that it gelatinizes when cold, and heat it till it becomes quite liquid, it answers best of all for throwing down tannin. It is by this property of forming a white precipitate with tannin that gelatin is usually detected in animal fluids. It is not, however, a perfectly decisive test, as *albumen* is also thrown down by tannin. Dr. Bostock has pointed out a very ingenious method of detecting and ascertaining the quantity of gelatin contained in an animal fluid. If corrosive sublimate produce no precipitate, we may be certain of the absence of albu-

men. Then the infusion of galls being mixed with the liquid, in such a proportion that the filtered liquid will neither precipitate infusion of galls, or the animal liquid under examination, a precipitate falls, composed of about two parts tannin and three parts gelatin. Hence this precipitate, dried on a steam bath, and multiplied by 0.6, gives us the weight of gelatin in the liquid examined very nearly.*

Chap. I.

Gelatin does not, properly speaking, combine with oils, but it renders them miscible with water, and forms a kind of emulsion.

4. From the effects of different reagents on gelatin, and from the decomposition which it undergoes when heated, we see that it contains carbon, hydrogen, azote, and oxygen. It was subjected to an analysis by Gay-Lussac and Thenard, who burnt it, when mixed, with chlorate of potash, and ascertained the products. According to their experiments, its constituents are as follows :

Composition.

Carbon	47.881
Oxygen	27.207
Hydrogen	7.914
Azote	16.998
	<hr/>
	100.000†

Not being possessed of any definite compound of gelatin and any other substance with the weight of which we are already acquainted, it is not in our power, from this analysis, to determine the constitution of gelatin. But the smallest number of atoms which correspond nearly with the preceding numbers, is the following :

15 atoms carbon	=	11.25	50.00
6 atoms oxygen	=	6.00	26.67
14 atoms hydrogen ..	=	1.75	7.78
2 atoms azote	=	3.50	15.55
		<hr/>		<hr/>
		22.50		100.00

* Bostock, Nicholson's Jour. xiv. 144. From the later experiments of Dr. Bostock, it appears that the compound of tannin and gelatin differs in the proportion of its constituents according to circumstances. It does not, therefore, furnish us with a method of detecting the quantity of gelatin in solution. Nicholson's Jour. xxi. 1.

† Recherches Physico-chimiques, ii. 336.

Book V.
Species.

5. Gelatin, like all other constituents of animal bodies, is susceptible of numerous shades of variations in its properties, and of course is divisible into an indefinite number of species. Several of these have been long known and manufactured for different purposes: and many curious varieties have been pointed out by Hatchett in his Dissertations on Shell, Bone, and Zoophytes, published in the Philosophical Transactions for 1797 and 1800. The most important species are the following:

Glue.

(1.) *Glue.* This well known substance has been long manufactured in most countries, and employed to cement pieces of wood together. It is extracted by water from animal substances, and differs in its qualities according to the substances employed. Bones, muscles, tendons, ligaments, membranes, and skins, all yield it; but the quality is best when skins are employed; and those of old animals yield a much stronger glue than those of young animals. English glue is considered as the best, owing to the care with which it is made. The parings of hides, pelts from furriers, the hoofs and ears of horses, oxen, calves, sheep, &c. are the substances from which it is extracted in Britain, and quantities of these substances are imported for the purpose. They are first digested in lime-water to clean them, then steeped in clean water, laid in a heap till the water runs off, and then boiled in brass caldrons with pure water. The impurities are skimmed off as they rise; and when the whole is dissolved, a little alum or finely powdered lime is thrown in. The skimming having been continued for some time, the whole is strained through baskets, and allowed to settle. The clear liquid is gently poured back into the kettle, boiled a second time, and skimmed till it is reduced to the proper consistency. It is then poured into large frames, where it concretes on cooling into a jelly. It is cut by a spade into square cakes, which are again cut by means of a wire into thin slices; these slices are put into a kind of coarse net work, and dried in the open air.* The best glue is extremely hard and brittle; it has a dark brown colour, and an equal degree of transparency without black spots. When put into cold water, it swells very much, and becomes gelatinous, but does not dissolve. When glue is soluble in cold water, it is a proof that it wants strength. Dry glue, according to Dr. Bostock, contains $10\frac{1}{2}$ per cent. of

* Clennel. See Johnson's History of Animal Chemistry, i. 315.

water. He thinks also that it contains albumen. But from the method of preparing it this can hardly be supposed.* Chap. I.

(2.) *Size*. This substance differs from glue in being colourless and more transparent. It is manufactured in the same way, but with more care; eel skins, vellum, parchment, some kinds of white leather, and the skins of horses, cats, rabbits, are the substances from which it is procured. It is commonly inferior to glue in strength. It is employed by paper-makers to give strength to that article, and likewise by linen manufacturers, gilders, polishers, painters, &c.† Size.

(3.) *Isinglass*. This substance agrees with size in being transparent, but it is much finer, and is therefore sometimes employed as an article of food. It is prepared in Russia from the air-bladders and sounds of different kinds of fish which occur in the mouths of large rivers; chiefly different species of *accipenser*, as the *sturio stellatus*, *huso ruthenus*, and likewise the *siluris glanis*. The bladder is taken from the fish, clean washed, the exterior membrane separated, cut lengthwise and formed into rolls, and then dried in the open air. When good, isinglass is of a white colour, semitransparent, and dry. It dissolves in water with more difficulty than glue, probably because it is not formed originally by solution. From the analysis of isinglass by Hatchett, we learn that it is almost completely convertible into gelatin by solution and boiling. Five hundred grains of it left by incineration 1·5 grains of phosphate of soda, mixed with a little phosphate of lime. Isinglass.

A coarse kind of isinglass is prepared from sea wolves, porpoises, sharks, cuttle fish, whales, and all fish without scales. The head, tail, fins, &c. of these are boiled in water, the liquid skimmed and filtered, and then concentrated by evaporation till it gelatinizes on cooling. At that degree of concentration it is cast on flat slabs and cut into tablets. This species is used for clarifying, stiffening silk, making sticking-plaster, and other purposes.‡

6. Gelatin exists in great abundance in animals, forming a constituent part of their solid parts. It forms an essential part of bones, ligaments, tendons, membranes, skin, muscles, hair, &c. Substances containing it.

* Nicholson's Jour. xxiv. 7.

† Clennel. See Johnson's History of Animal Chemistry, i. 315.

‡ Fabricius de Ichthyocolla, Jackson on British Isinglass, Phil. Trans. lxi. and Johnson's Animal Chemistry, i. 231.

Book V.
Uses.

7. Its uses are very numerous. In the state of jelly it constitutes one of the most nourishing and palatable species of food. It constitutes the basis of soups. The great variety of purposes to which it is applied in the state of glue, size, and isinglass, are well known.

SECT. II.

OF ALBUMEN.

Contained
in eggs.

THE eggs of fowls contain two very different substances: a yellow oily-like matter, called the *yolk*; and a colourless glossy viscid liquid, distinguished by the name of *white*. This last is the substance which chemists have agreed to denominate *albumen*.* The white of an egg, however, is not pure albumen. It contains also some mucus, soda, and sulphur: but as albumen is never found perfectly pure, and as no method is known of separating it without at the same time altering the properties of the albumen, chemists are obliged to examine it while in combination with these bodies.

Coagulates
when heated.

Albumen dissolves readily in water, and the solution has the property of giving a green colour to vegetable blues, in consequence of the soda which it contains. When albumen is heated to the temperature of 165° ,† it *coagulates* into a white solid mass; the consistency of which, when other things are equal, depends, in some measure, on the time during which the heat was applied. The coagulated mass has precisely the same weight that it had while fluid. This property of coagulating when heated is characteristic of albumen, and distinguishes it from other bodies.

The taste of coagulated albumen is quite different from that of liquid albumen: its appearance, too, and its properties, are entirely changed; for it is no longer soluble, as before, either in hot or in cold water.

The coagulation of albumen takes place even though air be completely excluded; and even when air is present there is no absorption of it, nor does albumen in coagulating change its volume.‡ Acids have the property of coagulating albumen, as Scheele ascertained.§ Alcohol also produces, in

* This is merely the Latin term for the white of an egg. It was first introduced into chemistry by the physiologists.

† Cullen.

‡ Carradori, Ann. de Chim. xxix. 98.

§ Scheele, ii. 58.

some measure, the same effect. *Heat*, then, *acids*, and *alcohol*, are the agents which may be employed to coagulate albumen.

It is remarkable, that if albumen be diluted with a sufficient quantity of water, it can no longer be coagulated by any of these agents. Scheele mixed the white of an egg with ten times its weight of water, and then, though he even *boiled* the liquid, no coagulium appeared. Acids, indeed, and alcohol, even then coagulated it; but they also lose their power if the albumen be diluted with a much greater quantity of water, as has been ascertained by many experiments. Now when water is poured into albumen, its integrant particles must be farther separated from each other, and their distance must increase with the quantity of water with which they are diluted. We see, therefore, that albumen ceases to coagulate whenever its particles are separated from each other beyond a certain distance. That no other change is produced, appears evident from this circumstance, that whenever the watery solution of albumen is sufficiently concentrated by evaporation, coagulation takes place, upon the application of the proper agents, precisely as formerly.

It does not appear that the distance of the particles of albumen is changed by coagulation; for coagulated albumen occupies precisely the same sensible space as liquid albumen.*

Now, to what is the coagulation of albumen owing? We can conceive no change to take place from a state of liquidity to that of solidity, without some change in the figure of the particles of the body which has undergone that change. Now such a change may take place three ways: 1. The figure may be changed by the addition of some new molecules to each of the molecules of the body. 2. Some molecules may be abstracted from every integrant particle of the body. 3. Or the molecules, of which the integrant particles are composed, may enter into new combinations, and form new integrant particles, whose form is different from that of the old integrant particles. Some one or other of these three things must take place during the coagulation of albumen.

1. Scheele and Fourcroy have ascribed the coagulation of albumen to the first of these causes, namely, to the addition of a new substance. According to Scheele, *caloric* is the substance which is added. Fourcroy, on the contrary, affirms that it is *oxygen*.

* Carradori.

Book V.
Ascribed to
caloric;

Scheele supported his opinion with that wonderful ingenuity which shone so eminently in every thing which he did. He mixed together one part of white of egg and four parts of water, added a little pure alkali, and then dropped in as much muriatic acid as was sufficient to saturate the alkali. The albumen coagulated. But when he repeated the experiment, and used carbonate of alkali instead of pure alkali, no coagulation ensued. In the first case, says he, there was a double decomposition: the muriatic acid separated from a quantity of caloric with which it was combined, and united with the alkali; while, at the same instant, the caloric of the acid united with the albumen, and caused it to coagulate. The same combination could not take place when the alkaline carbonate was used, because the carbonic acid gas carried off the caloric, for which it has a strong affinity.*

To oxygen.

Fourcroy observes, in support of his opinion, that the white of an egg is not at first capable of forming a hard coagulum, and that it only acquires that property by exposure to the atmosphere. It is well known that the white of a new laid egg is milky after boiling; and that if the shell be covered over with grease to exclude the external air, it continues long in that state; whereas the white of an old egg, which has not been preserved in that manner, forms a very hard tough coagulum. These facts are undoubted; and they render it exceedingly probable, that albumen acquires the property of forming a hard coagulum only by absorbing oxygen: but they by no means prove that coagulation itself is owing to such an absorption. And since coagulation takes place without the presence of air, and since no air, even when it is present, is absorbed, this opinion cannot be maintained without inconsistency.

2. The only substance which can be supposed to leave albumen during coagulation, since it does not lose its weight, is caloric. We know that in most cases where a fluid is converted into a solid, caloric is actually disengaged. It is extremely probable, then, that the same disengagement takes place here. But the opinion has not been confirmed by any proof. Fourcroy indeed says, that in an experiment made by him, the thermometer rose a great number of degrees. But as no other person has ever been able to observe any such thing, it cannot be doubted that this philosopher has been misled by some circumstance or other to which he did not attend.†

* Scheele, ii. 58.

† Thomson's Fourcroy, iii. 271.

3. The coagulation of albumen resembles exactly what takes place when concentrated silicated potash is exactly saturated with muriatic acid. The mass slowly assumes an opal colour, and at last concretes into a solid gelatinous mass. Now this jelly consists of the particles of silica combined with each other, and with a certain portion of water. These particles were formerly held in solution by the potash; that is to say, the affinity of silica for potash was superior to the cohesive force which exists between the particles of silica. The muriatic acid, by saturating the potash, diminished the force of its affinity for the silica. The cohesive force of the silica, now superior, causes it to combine in masses, consisting of a certain portion of silica and water. These masses, equally diffused through the liquid, and at such small intervals as to cohere together, gives the whole a gelatinous form. Something like this seems to take place with respect to the albumen. Its particles, combined with water and also with soda, are all kept at equal distances in the liquid; because this affinity just balances their cohesive force. But when heat is applied, this affinity is diminished by the additional elasticity, or tendency to separate, given to the water and the soda. The cohesion of the albumen, now superior, causes its particles to combine in sets, forming solid bodies, equally distant from each other, and cohering together. Hence the gelatinous form, and the solidity of the coagulum, always inversely proportional to the quantity of water present. Thus it appears, that whatever diminishes the affinity between the water, and soda, and the albumen, occasions its coagulation, by allowing its cohesive force to act.

Albumen, then, is capable of existing in two states; the one before it has been coagulated, and the other after it has undergone coagulation. Its properties are very different in each. It will be proper therefore to consider them separately.

I. Albumen in its natural state, or uncoagulated, is a glary liquid, having little taste and no smell. When dried spontaneously, or in a low heat, it becomes a brittle transparent glassy-like substance; which, when spread thin upon surfaces, forms a varnish, and is accordingly employed by bookbinders for that purpose. When thus dried it has a considerable resemblance to gum arabic, to which also its taste is similar. The white of an egg loses about $\frac{1}{3}$ ths of its weight in drying. It is still soluble in water, and forms the same glary liquid as before.

Uncoagulated albumen.

Book V.

From the experiments of Dr. Bostock, it appears, that when one part of this dry albumen is dissolved in nine parts of water, the solution becomes perfectly solid when coagulated by heat; but if the albumen amounts only to $\frac{1}{13}$ th of the liquid, then, though coagulation takes place, the liquid does not become perfectly solid, but may be poured from one vessel to another.*

When one grain of albumen is dissolved in 1000 grains of water, the solution becomes cloudy when heated.†

Its properties.

Uncoagulated albumen soon putrefies unless it be dried; in which state it does not undergo any change. It putrefies more readily when dissolved in a large quantity of water than when concentrated. The smell of white of egg, allowed to run into putrefaction, resembles that of *pus*.‡

It is insoluble in alcohol and ether, which immediately coagulate it, unless it be mixed with a very great proportion of water; in which case even acids have no effect.

Action of acids.

When acids are poured upon it, coagulation takes place equally; but several of them have the property of dissolving it again when assisted by heat. This at least is the case with sulphuric acid. The solution is of a green colour, and does not soon blacken even when boiled. It is the case also with nitric acid, and probably also with muriatic acid. Nitric acid first disengages some azotic gas; then the albumen is gradually dissolved, nitrous gas emitted, oxalic and malic acids formed, and a thick oily matter makes its appearance on the surface.§

I tried the effect of different acids upon a solution consisting of the albumen of a common-sized hen's-egg in about a pint of water, and filtered. The results may be seen in the following table:

Acids.	Effects.
1. Sulphuric acid	{ A white coagulum, which appears at the bottom first.
2. Muriatic acid	{ Becomes at once very milky, flakes fall slowly.
3. Nitric acid	{ A very copious precipitate of yellow flakes soon falls to the bottom, and a slight effervescence is perceptible.

* Nicholson's Jour. xiv. 141.

† Bostock, Nicholson's Jour. xi. 247.

‡ Ibid. xiv. 143.

§ Scheele, Crell's Annals, ii. 17. Eng. Trans.

Acids.	Effects.	Chap. I.
4. Chlorine	{ Assumes a beautiful white colour, and fine white flakes slowly separate.	}
5. Sulphurous acid		
6. A weak fluoric acid	No change.	
7. Acetic acid	No change.	
8. Distilled vinegar	No change.	

It may be proper to mention, that the solution was coagulated by alcohol but not by heat.

When alkalies are mixed with the solution of albumen in water, no apparent change takes place; but if a concentrated solution of pure potash be triturated with albumen for some time, and then allowed to remain at rest, the albumen gradually coagulates or rather gelatinizes; for the coagulum has a striking resemblance to jelly. It gradually hardens; and at a particular period of its drying it resembles very exactly the lense of the eye. When quite dry it is brittle and transparent.

The effect of the different earthy bodies on the solutions of albumen in water may be seen from the following table. The solution consisted of one white of egg dissolved in about a pint of water, and filtered to separate the opaque white films which are always mixed in the white of egg.

Substances.	Effects.
Barytes water	No change.
Strontian water	No change.
Lime water	No change.
Sulphate of magnesia	No change.
Alum	No change.
Silicated potash	No change.
Aluminated potash	No change.

Thus it appears that none of the earths form insoluble compounds with albumen: in this respect they resemble the alkalies.

The case is very different with the metallic oxides. The following table exhibits the effects of the metallic salts on the same solution of albumen, as far as my experiments go.

* The water was saturated with the acid.

Metallic Salts.	Effects.
1. Nitromuriate of gold.	{ A copious yellow flaky precipitate, redissolved by ammonia.*
2. Muriate of platinum	A yellowish-white precipitate.
3. Nitrate of silver	{ A copious reddish-brown precipitate, not redissolved by ammonia.
4. Nitrate of mercury ..	A white precipitate.
5. Corrosive sublimate ..	A light white precipitate.
6. Supersulph. of do. ..	A white precipitate.
7. Prussiate of do.	No change.
8. Pernitrate of copper ..	{ A green coagulum, redissolved by adding more of the copper, then the mixture becomes opaque.
9. Muriate of copper ...	A copious white precipitate.
10. Persulphate of copper	A greenish white precipitate.
11. Cuprate of ammonia .	{ No change. Muriatic acid renders the mixture colourless, but no precipitate falls.
12. Sulphate of iron	Brown flakes precipitate.
13. Persulphate of ditto in alcohol	{ A copious dirty white precipitate.
14. Permuriate of iron ..	No change.
15. Pernitrate of iron ...	No precipitate. Becomes green.
16. Permuriate of tin ..	Becomes slowly milky.
17. Chlorate of lead	A copious white precipitate.
18. Acetate of lead	The same.
19. Nitrate of lead	The same.
20. Plumbate of potash ..	{ No change; but muriatic acid throws down a copious white precipitate.
21. Plumbate of lime	No change.
22. Muriate of zinc	Becomes slowly very milky.
23. Nitrate of bismuth ..	{ A white precipitate, which does not appear immediately, unless there be an excess of acid.
24. Nitrate of antimony ..	A white precipitate.
25. Muriate of arsenic ..	Becomes slowly milky.
26. Arseniate of potash ..	No change.

* The salt had a slight excess of acid.

Metallic Salts.	Effects.
27. Arseniate of cobalt . .	No change.
28. Prussiate of potash . .	No change.
29. Oxalate of ammonia .	{ A slight white precipitate appears after standing an hour.

Thus we see that every metal tried, except cobalt, occasioned a precipitate; but no precipitate ever appeared when the oxide was held in solution by an alkali or earth. The effect of the metallic salts on albumen forms a striking contrast with their effect on gelatin.

From the experiments of Dr. Bostock, it appears that a drop of the saturated solution of corrosive sublimate, let fall into water containing $\frac{1}{2000}$ th part of its weight of albumen, produces an evident milkiness, and a curdy precipitate falls.* It is therefore a very delicate test of the presence of albumen in animal fluids. Dr. Bostock has proposed the following ingenious method of estimating the quantity of albumen contained in an animal fluid: Add to the fluid a quantity of corrosive sublimate, more than sufficient to saturate the albumen, then heat the mixture. By this double action a coagulum is formed, which may be separated by the filter. This precipitate, when dry, contains about $\frac{1}{4}$ ths of its weight of albumen.†

If a solution of tannin be poured into an aqueous solution of uncoagulated albumen, it forms with it a very copious yellow precipitate of the consistence of pitch, and insoluble in water. This precipitate is a combination of tannin and albumen. When dry it is brittle, like over-tanned leather, and is not susceptible of putrefaction. This property, which albumen has of precipitating with tannin, was discovered by Seguin. ‡

The infusion of galls is by no means so delicate a test of the presence of albumen as of gelatin. When an infusion of galls containing 2½ per cent. of solid matter, and water holding $\frac{1}{1000}$ th of albumen in solution, are mixed in equal quantities, no effect is produced at first, but after some time a precipitating matter appears and slowly subsides. §

II. When albumen is coagulated either by heat, alcohol, or acids, it is an opaque substance of a pearl-white colour, tough, and of a sweetish mucilaginous taste. It is no longer soluble in water, and is not nearly so susceptible of decompo-

* Nicholson's Jour. xi. 247.

† Ibid. xiv. 142.

‡ Ibid. i. 272.

§ Bostock, *ibid*, p. 141.

Book V.
Its proper-
ties.

sition as uncoagulated albumen. Mr. Hatchett kept it for a month under water, and yet it did not become putrid. It is to the experiments of this ingenious chemist that we are indebted for almost every thing at present known relative to coagulated albumen. By drying it in the temperature of 212° , he converted it into a brittle hard yellow substance, semi-transparent like horn.*

Action of
water,

When this substance was digested for some hours in water, it gradually softened, and became white and opaque like newly coagulated albumen. When water is made to act upon it long, a small portion of it is taken up. The watery liquid is not precipitated by the infusion of tan; but nitro-muriate of tin occasions a faint cloud.†

Of acids.

According to Scheele, the mineral acids, when greatly diluted with water, dissolve a portion of coagulated albumen, which is thrown down again by the same acids concentrated.‡

Changed
into gelatin.

When coagulated albumen is steeped in diluted nitric acid, the acid in about four weeks begins to acquire a yellow tinge, which becomes gradually deeper; but the albumen, though it becomes more opaque, is not dissolved. The yellow acid, when saturated with ammonia, becomes of a deep orange colour, but does not let fall any precipitate. When the albumen, thus treated, is immersed in ammonia, the liquid assumes a deep orange colour, inclining to blood red. The albumen is slowly dissolved, and the solution has a deep yellowish-brown colour. If the albumen, after being steeped in nitric acid, be washed and then boiled in water, it is dissolved, and forms a pale-yellow liquid, which gelatinizes when properly concentrated. If the gelatinous mass be again dissolved in boiling water, the solution is precipitated by tan and by nitro-muriate of tin. Hence we see that nitric acid has the property of converting coagulated albumen into gelatine. For this important fact we are indebted entirely to Mr. Hatchett.§

Concentrated nitric acid dissolves coagulated albumen with effervescence, especially when assisted by heat. It becomes orange-brown when mixed with ammonia, but no precipitate falls. ||

It is readily dissolved by a boiling lixivium of potash, am-

* Hatchett, Phil. Trans. 1800.

† Ibid.

‡ Scheele, ii. 57.

§ Hatchett, Phil. Trans. 1800.

|| Ibid.

monia is disengaged, and an animal soap is formed. This soap, when dissolved in water, and mixed with acetic, or muriatic acids, lets fall a precipitate which is of a saponaceous nature. When heated gently some oil flows from it, and a brownish viscid substance remains.* The alkalies, when diluted, and not assisted by heat, act upon it slowly and imperfectly.

Chap. I.
Action of
alkalies.

These properties indicate sufficiently that coagulated albumen is a very different substance from uncoagulated albumen. During the coagulation its component parts must arrange themselves differently.

III. From the effects of nitric acid on albumen; and its products, when subjected to destructive distillation, it has been concluded that it consists of carbon, hydrogen, azote, and oxygen. According to the experiments of Gay-Lussac and Thenard, its constituents are as follows:

Composi-
tion of al-
bumen.

Carbon	52.883
Oxygen	23.872
Hydrogen	7.540
Azote.....	15.705
	100.000 †

The smallest number of atoms that agrees with the preceding analysis nearly, is the following:

17 atoms carbon	=	12.75	53.40
6 atoms oxygen	=	6.00	25.13
13 atoms hydrogen	=	1.625	6.80
2 atoms azote	=	3.50	14.67
		23.875		100.00

According to this statement it differs from gelatin, by containing two additional atoms of carbon and having one atom fewer of hydrogen. But Dr. Prout, who has analysed albumen with great care, has found its constituents as follows:

15 atoms carbon	=	11.25	50.00
6 oxygen	=	6.00	26.67
14 hydrogen	=	1.75	7.78
2 azote	=	3.50	15.55
		22.50		100.00

* Harchett, Phil. Trans. 1800.

† Recherches Physico-chimiques, ii. 332.

Book V.

The reader will observe that this is precisely the composition of gelatin, as determined by the experiments of Gay-Lussac and Thenard. It is obvious that a mistake must exist either in the constituents of gelatin, as determined by these chemists, or in those of albumen, as determined by Dr. Prout.

Species of it.

IV. Like all other animal substances, albumen is capable of existing in various states both when coagulated and uncoagulated, forming a number of distinct species; but they have not been characterized with much precision. But if the curdy part of milk be considered as an albumen, as some chemists do, it constitutes a species essentially distinct from the albumen of eggs and blood. Coagulated albumen forms an essential part of bone and muscle; brain perhaps may be considered as a species of it, and so may the lense of the eye. Cartilage, nails, horns, hair, &c. as Hatchett has shown, are almost entirely composed of it; and it forms the membranous part of many shells, sponges, &c. In short, it is one of the most general and important of the animal substances.

V. The property which albumen has of being coagulated by heat renders it a very useful substance for clarifying liquids. The serum of blood, white of egg, or any liquid containing it, is mixed with the liquid to be clarified, and the whole is then heated. The albumen coagulates, and carries down with it the floating particles which rendered the liquid opaque.

SECT. III.

OF FIBRIN.

How obtained.

IF a quantity of blood, newly drawn from an animal, be allowed to remain at rest for some time, a thick red clot gradually forms in it, and subsides. Separate this clot from the rest of the blood, put it into a linen cloth, and wash it repeatedly in water till it ceases to give out any colour or taste to the liquid; the substance which remains after this process is denominated *fibrin*. It has been long known to physicians under the name of the *fibrous part of the blood*, but has not till lately been accurately described.

It may be procured also from the muscles of animals. Mr. Hatchett, to whom we are indebted for a very interesting set of experiments on this substance, cut a quantity of lean beef into small pieces, and macerated it in water for fifteen days, changing the water every day, and subjecting the beef to

pressure at the same time, in order to squeeze out the water. As the weather was cold, it gave no signs of putrefaction during this process. The shreds of muscle, which amounted to about three pounds, were now boiled for five hours every day for three weeks in six quarts of fresh water, which was regularly changed every day. The fibrous part was now pressed, and dried by the heat of a water bath. After this treatment it might be considered as fibrin nearly as pure as it can be obtained.*

Chap. I.

2. Fibrin is of a white colour, and has no taste nor smell. When newly extracted from blood, it is soft and elastic, and resembles the gluten of vegetables. Its colour deepens very much in drying. That which is extracted from muscle by boiling and maceration has a certain degree of transparency, and is not ductile but brittle. Its colour does not deepen nearly so much as the fibrin from blood. Properties.

It undergoes no change though kept exposed to the action of air; neither does it alter speedily though kept covered with water. Mr. Hatchett kept a quantity of the fibrin which he had prepared from beef moistened with water during the whole month of April; it acquired a musty but not a putrid smell, neither were the fibres reduced to a pulpy mass. Even when kept two months under water, it neither became putrid, nor was converted into the fatty matter obtained by macerating recent muscle.†

3. Fibrin is insoluble in cold water. In boiling water it curls up, and after the boiling has continued for some hours, the water becomes milky; but no gas is evolved. When infusion of nutgalls is dropped into the water, white flocks precipitate, which do not adhere together when heated, as those produced by gelatin. The evaporated liquid does not gelatinize, and leaves a white, dry, hard, friable residue, soluble in cold water, and having an agreeable taste similar to that of fresh broth. Fibrin by long boiling in water loses its property of softening and dissolving in acetic acid.‡ Action of water,

4. In alcohol of the specific gravity 0.81 fibrin undergoes a species of decomposition, and forms an adipocirous substance soluble in alcohol and precipitated by the addition of water. It has often a strong and unpleasant odour. The alcoholic solution leaves when evaporated a fat residue, which of alcohol,

* Hatchett, Phil. Trans. 1800.

† Ibid.

‡ Berzelius, Annals of Philosophy, ii. 20.

Book V. did not pre-exist in the fibrin. Fibrin, after being heated in alcohol, continues soluble in acetic acid.*

5. Ether acts on fibrin in the same way as alcohol. It forms an adipocirous matter in much greater abundance, and having a stronger and more disagreeable smell.†

Of acid,

6. In concentrated acetic acid fibrin becomes immediately soft and transparent, and with the assistance of heat is converted into a tremulous jelly. By adding water and warming it, this jelly is completely dissolved with the evolution of a small quantity of azotic gas. The solution is colourless and has a mawkish and slightly acid taste. During its evaporation a transparent membrane appears on the surface, and after a certain degree of concentration the gelatinous substance is again reproduced. When completely dried it is a transparent mass which reddens litmus paper, but is insoluble in water without a fresh addition of acetic acid. When ferrocyanate of potash, an alkali, or sulphuric, nitric or muriatic acid is dropped into this solution a white precipitate falls. The acid precipitate is a compound of fibrin and the acid. If it be washed, a certain portion of acid holding fibrin in solution is carried off, and the remainder is soluble in water. This solution contains a neutral compound of the acid and fibrin. The addition of a little more of the acid causes it to precipitate again. ‡

7. In weak muriatic acid fibrin shrinks and gives out a small quantity of azotic gas; but scarcely any portion is dissolved even by boiling; nor does the acid liquid afford any precipitate with ammonia or ferrocyanate of potash. The fibrin, thus treated, is hard and shrivelled. When repeatedly washed with water it is at last converted into a gelatinous mass, which is perfectly soluble in water. The solution reddens litmus paper, and yields a precipitate with acids as well as alkalies. Fibrin, therefore, combines with muriatic acid in two proportions. The one gives a neutral compound soluble in water, the other with an excess of acid is insoluble, but becomes soluble by the action of pure water.§

8. Concentrated sulphuric acid decomposes and carbonizes fibrin. When diluted with six times its weight of water, this acid acquires a red colour if it be digested with fibrin, but dissolves scarcely any thing. The undissolved portion is a

* Berzelius, *Annals of Philosophy*, ii. 20.

† *Ibid.*

‡ *Ibid.*

§ *Ibid.*

compound of fibrin and an excess of sulphuric acid. Water deprives it of this excess, and a neutral combination is obtained which is soluble in water, and possesses the same characters as neutral muriate of fibrin.*

9. Nitric acid of the specific gravity 1.25 digested with fibrin renders it yellow and diminishes its cohesion. The fluid becomes yellow, and the surface of the fibrin is covered with a small quantity of fat formed by the action of the acid. During this operation pure azotic gas is disengaged. After twenty-four hours digestion the fibrin is converted into a pulverulent mass, of a pale-citron colour, which is deposited at the bottom of the liquid. When this pulverulent matter is washed, its colour changes in proportion as the excess of acid is carried off, and it acquires a deep orange colour. This yellow matter was considered by Fourcroy and Vauquelin as a peculiar acid which they distinguished by the name of *yellow acid*. But Berzelius has shown that it is merely fibrin combined with nitric acid. It dissolves in caustic alkali communicating a yellow colour. When boiled in alcohol an adipocirous matter is taken up, but deposited again as the liquid cools. If the yellow substance, after being thus deprived of its adipocirous portion, be digested with water and carbonate of lime, it slowly decomposes and dissolves the carbonate, and forms a yellow solution. When this solution is concentrated and mixed with alcohol, malate of lime is precipitated. The portion dissolved in alcohol was nitrate and nitrite of lime.†

10. In caustic alkali fibrin increases in bulk, becomes transparent and gelatinous, and at length is completely dissolved. The solution is yellow with a shade of green. Acids occasion in it a precipitate which gradually becomes confluent. Alcohol occasions a precipitate in it. Some alteration is produced upon the fibrin by the alkali, but nothing in the least similar to a soap is formed.‡

11. When fibrin is exposed to heat, it contracts very suddenly, and moves like a bit of horn, exhaling at the same time the smell of burning feathers. In a stronger heat it melts. When exposed to destructive distillation, it yields water, carbonate of ammonia, a thick heavy fetid oil, traces of acetic acid, carbonic acid, and carburetted hydrogen gas.§

* Berzelius, *Annals of Philosophy*, ii. 20.

† *Ibid.*

‡ *Ibid.*

§ Fourcroy, ix. 158.

Book V. The charcoal, as Hatchett ascertained, is more copious than that left by gelatin or albumen. It is very difficult to incinerate, owing to the presence of phosphate of soda and some phosphate of lime, which form a glassy coat on the surface. A considerable proportion of carbonate of lime also remains after the incineration of the charcoal.*

12. Fibrin was subjected to an analysis by Gay-Lussac and Thenard, who found its constituents as follows:

Composition.	Carbon	53.360
	Oxygen	19.685
	Hydrogen	7.021
	Azote	19.934
		100.000 †

The smallest number of atoms that correspond with this analysis is the following:

18 atoms carbon ...	=	13.50	52.94
5 atoms oxygen ..	=	5.00	19.61
14 atoms hydrogen	=	1.75	6.86
3 atoms azote	=	5.25	20.59
		25.50		100.00

According to this statement it contains one more atom of carbon, hydrogen, and azote, and one atom less of oxygen than albumen.

Species. 13. Fibrin exists only in the blood and the muscles of animals: but it is a genus which includes as many species as there are varieties in the muscles of animals; and the great diversity of these substances is well known. The muscles of fish, of fowl, and of quadrupeds, bear scarcely any resemblance to each other.

SECT. IV.

OF THE COLOURING MATTER OF BLOOD.

History. THE blood was very much examined by physiologists during the last century. Leuwenhoeck described its colouring matter as consisting of globules of a red colour floating in the serum, and this description was confirmed by the micro-

* Hatchett, Fourcroy, ix. 158.

† Recherches Physico-chimiques, ii. 330.

scopical observations of Hewson and other physiologists. In 1797, Dr. Wells published a paper on the subject, in which he showed, by very ingenious and satisfactory arguments, that this colouring matter was an animal substance.* But chemists were prevented from acceding to this opinion by the experiments of Fourcroy and Vauquelin, which were intended to show that blood contains subphosphate of iron in solution, and that it is to the presence of this salt that its red colour is owing.† This opinion was generally acceded to by chemists till Mr. Brande published his *Chemical Researches on the Blood* in 1812.‡ He demonstrated by decisive experiments that the opinion of Fourcroy and Vauquelin was ill-founded, that the colouring matter of the blood is a peculiar animal substance, which he obtained in a separate state, and the properties of which he investigated.

Towards the end of the year 1812, Berzelius' *General Views of the Composition of Animal Fluids* were published in the third volume of the *Medico-chirurgical Transactions*. He details the properties of the colouring matter, and shows that it possesses very nearly the same properties as those of fibrin. These experiments had been made long before those published by Mr. Brande, and had been inserted in the second volume of his *Djur Chemien*, printed in 1808, though we only became acquainted with them in this country in consequence of the publication of the paper in question. Mr. Brande's experiments were repeated and confirmed by Vauquelin in 1815. §

The colouring matter may be procured in the following manner: Allow the blood to coagulate. Place the coagulum upon blotting paper to absorb the serum. Put the coagulum thus dried into water; that liquid dissolves the colouring matter and leaves the fibrin. Evaporate the water to dryness at a low temperature. The colouring matter remains behind in a state of tolerable purity.

Colouring matter thus obtained has a very dark-red colour. It dissolves in water, provided it has been obtained by a very low heat; but if the water which held it in solution has been boiled, the colouring matter is deprived of its solubility altogether.

* Phil. Trans. 1797, p. 416.

† Fourcroy, ix. 152.

‡ Phil. Trans. 1812, p. 90.

§ Ann. de Chim. et Phys. i. 9.

Book V.

Vauquelin and Brande found it soluble in all the acids tried; but Berzelius could not dissolve it in muriatic acid. He obtained an insoluble matter, containing an excess of acid, which became soluble in proportion as the excess of acid was removed by water.

The alkalies dissolve it and the solution has a purple colour. Berzelius found that it formed a brownish jelly in caustic fixed alkali, which was dissolved by a sufficient quantity of water. The alkaline solution is precipitated by alcohol, which however acquires a red tinge by dissolving a small quantity of the compound formed of the colouring matter with the excess of alkali.

Dilute nitric acid dissolves the colouring matter of blood without altering its colour. Muriate of barytes does not occasion any precipitate in its solution in muriatic acid. Neither is any change produced by gallic acid or ferrochyzate of potash. The infusion of nutgalls precipitates it but does not occasion any change of colour. Nitrate of silver does not render its solution in nitric acid turbid; but acetate of lead throws down a brown precipitate. It is not precipitated by ammonia, nor is the colour altered. It is soluble in sulphuric acid by heat. The solution is dark reddish-brown; and is not precipitable by water or potash. It is likewise soluble in caustic potash: the solution is green; and is precipitated in flocks by muriatic acid.

Mr. Brande tried to form a lake by precipitating its acid solution by means of different earthy and metallic salts. Neither alumina nor oxide of tin answered the purpose well. Corrosive sublimate or nitrate of mercury succeeded best. They gradually precipitated the colouring matter, and formed with it powders of a good red colour, which did not alter by exposure to the air.* Mr. Brande likewise made some attempts to employ it as a pigment in dyeing; but they were not attended with much success. Though, as blood is an ingredient employed in some dyeing processes, it is not unlikely that the colouring matter may be of some importance even as a dye stuff.

When the colouring matter of blood is heated it neither alters its form nor colour; but gives out an animal odour and furnishes carbonate of ammonia and a purple oil: but scarcely

* Phil. Trans. 1812, p. 110.

any gas. The charry residuum is as bulky as the original substance.* Chap. I.

No analysis of the colouring matter of the blood has yet been made. It is obvious, however, from the observations of Berzelius, that its properties approach very nearly those of fibrin. Hence it is probable that its constitution is nearly similar. It contains however a considerable proportion of iron in some unknown state.

SECT. V.

OF MUCUS.

No word in chemistry has been used with less precision than *mucus*. Too many experimenters have made it serve as a common name for every animal substance which cannot be referred to any other class. Dr. Bostock, in his excellent Papers on the Analysis of Animal Fluids, has endeavoured to fix the meaning of the word by ascertaining the properties of pure mucus. Fourcroy and Vauquelin have published an elaborate paper on the same subject.†

From Bostock's experiments it appears, that if the solid matter obtained by evaporating saliva to dryness be redissolved in water and filtered, the solution will contain very little except mucus. He obtained mucus, also, by macerating an oyster in water and evaporating the liquid.‡ Mucus, thus obtained, possesses the following properties:

1. It has much the appearance of gum arabic, excepting Properties. that, in general, it is rather more opaque; like it, it has little taste, dissolves readily in water, and forms an adhesive solution.

2. When evaporated to dryness it is transparent, inelastic, and has much the appearance of gum. It is insoluble in water, but dissolves readily in all the acids though very much diluted.

3. It does not dissolve in alcohol nor in ether.

4. It does not coagulate when heated; nor when concentrated by evaporation does its solution assume the form of a jelly.

5. It is not precipitated by corrosive sublimate, nor by the infusion of galls.§

* Vauquelin.

† Ann. du Mus. d'Hist. Nat. xii. 61.

‡ Nicholson's Jour. xi. 251.

§ Bostock, Nicholson's Jour. xi. 251.

Book V.

6. The subtracetate of lead occasions a copious white precipitate when dropped into solutions containing mucus; the acetate produces a much less striking effect.*

7. Nitrate of silver likewise occasions a precipitate in solutions containing mucus.

8. When heated it assumes the appearance of horn, and when distilled it yields the common products of animal substances. According to Fourcroy and Vauquelin, horn, nails, hair, feathers, the epidermis, and the scales which form on the skin, consist chiefly of mucus.

Many of the substances called *mucus* have the property of absorbing oxygen, and of becoming by that means insoluble in water. They resemble vegetable extractive in this respect.

The mucilaginous substances will be pointed out in the next Chapter. In the present state of our knowledge, any account of them here would merely be a repetition of the properties just mentioned.

SECT. VI.

OF OSMAZOME.†

History. THIS substance was first noticed by Rouelle. The name was applied to it by Thenard. It seems doubtful, whether osmazome be any thing else than fibrin slightly altered by being boiled in water and dissolved in that liquid. But as it is a substance which is often met with in the analysis of animal substances, and as its identity with fibrin, though probable, is not fully made out, I think it better, at least for the present, to give it a particular place.

Preparation. Osmazome may be obtained by the following process. Divide the muscle of beef into small fragments, and leave it in contact with twice or thrice its weight of cold water for an hour or two, taking care to squeeze it occasionally. Decant off the first portion of water, and add an additional quantity. Repeat the digestion a third time. These portions of water will dissolve the salts, the albumen, and the osmazome. Mix them all together and evaporate them in a porcelain vessel till the whole albumen has coagulated and separated. Then filter the liquor, which will now be reduced to a small quantity, and

* Bostock, Nicholson's Jour. xi. 251.

† Derived, I presume, from οσμη smell, and ζωμος broth.

of a deep colour. Evaporate it by means of a very gentle heat to the consistence of a syrup. It consists of osmazome mixed with the salts of the musele. Alcohol digested with it takes up the osmazome and leaves the salts. By evaporating the alcohol we obtain the osmazome nearly in a state of purity.*

Chap. I.

Osmazome thus obtained has a brownish-yellow colour, and the taste and smell of broth. Properties.

It is soluble both in water and alcohol. The aqueous solution does not gelatinize; but when evaporated leaves the osmazome behind it in the state of a brown matter.

The aqueous solution of it is precipitated by infusion of nutgalls, nitrate of mercury, and by the acetate and nitrate of lead.

When heated it melts, swells, and is decomposed, giving out subcarbonate of ammonia, and leaving a bulky charcoal which contains carbonate of soda, showing clearly that it was not in a state of complete purity.

SECT. VII.

OF PICROMEL.

This is a name given by Thenard to the peculiar substance which characterizes bile.† It may be obtained pure by the following process. Mix fresh bile with sulphuric acid, diluted with three or four times its weight of water. A yellow precipitate falls, which must be allowed to subside, then separated. Then continue to add fresh acid as long as any precipitate appears. Heat the mixture gently for some hours, and afterwards decant off the fluid part. A green matter remains, which chemists have been in the habit of denominating *resin of bile*. This green matter is to be thoroughly edulcorated with water. In this state it is a compound of picromel and sulphuric acid. Digest it with carbonate of barytes and water; the acid combines with the barytes, while the picromel dissolves in the water and forms a solution exactly similar in appearance to bile. Or the green matter may be dissolved in alcohol, and the liquid digested over carbonate of potash or carbonate

Preparation.

* Thenard's *Traité de Chimie*, iii. 417.

† The name picromel is derived, I presume, from *πικρος* bitter, and *μελις* honey. It is a very unfortunate name. *Cholin* (from *χολη* bile) would be much better.

Book V.

of lime till it no longer reddens litmus. It is then to be filtered and evaporated to dryness. The substance thus obtained is picromel.

Properties.

Picromel when pure resembles entirely inspissated bile. It has a green or greenish-yellow colour, and an intensely bitter taste, followed by some sweetness.

It is soluble in water, and likewise in alcohol in all proportions. Ether does not dissolve it, but converts it into an adipocirous substance, having an exceedingly disagreeable smell.

Like the albumen of blood it unites with acids, and forms compounds soluble when neutral, but insoluble, or only sparingly soluble, when there is an excess of acid. It is this last compound of picromel, and a mineral acid, which has been hitherto mistaken by chemists for the *resin* of bile.

Picromel combines likewise with many metallic oxides into a pulverulent mass.

It is not precipitated by infusion of nutgalls; but the nitrate of mercury, the subacetate of lead, and the salts of iron, occasion a precipitate when dropped into its aqueous solution.

When picromel is subjected to destructive distillation it gives out no ammonia. Hence it would seem that it contains no azote. A difference in composition sufficient to distinguish picromel from albumen and fibrin, with which it possesses many analogous properties.*

When passed through red hot peroxide of copper, it is converted into carbonic acid gas and water. One grain of picromel gives out 4.2 cubic inches of carbonic acid gas, and 0.2 grain of water. From this experiment I conclude that it is a compound of

5 atoms carbon	= 3.75	54.53
1 atom hydrogen	= 0.125	1.82
3 atoms oxygen	= 3.000	43.65
	6.875	100.00†

* Berzelius, *Annals of Philosophy*, ii. 377.

† *Annals of Philosophy*, xiv. 69.

SECT. VIII.

OF UREA OR NEPHRIN.

THE word *urea*, first applied to the substance which constitutes the subject of this section by Fourcroy, is a very unhappy one in consequence of uric acid, a substance likewise peculiar to the urine, having very nearly the same name, or at least a name that would indicate a connexion between the two substances, while in fact no such connexion exists. I conceive therefore that it would be a considerable improvement if we were henceforth to substitute the term *nephrin** for urea.

Chap. I.

This important substance may be obtained in the following manner. Evaporate healthy urine to one fourth of its volume. Then mix it with about a third part of its weight of nitric acid of the specific gravity 1.4, and set it aside for 24 hours. A great number of silky plates of a brown colour are formed, which consist of nitric acid and nephrin united together. Dry them upon blotting paper, dissolve them in water, and add as much potash as is requisite to saturate the nitric acid which they contain. Concentrate the solution and set it aside; crystals of nitre separate. These crystals are to be removed, the residual liquid is to be evaporated, and digested in alcohol. This liquid dissolves the nephrin, and the colouring matter of urine, but leaves the saltpetre undissolved. Evaporate the alcoholic solution to dryness, dissolve the residue in water, put the solution into a phial and add a quantity of fresh ivory black; agitate for some time, and then throw the whole on a filter. The ivory black separates the whole colouring matter, and the liquid which passes through is colourless as water. Evaporate this liquid slowly, and the nephrin is obtained in the state of a beautiful white substance, semitransparent and crystallized in four-sided prisms. For the above process, which I have more than once repeated, we are indebted to Dr. Prout.

Nephrin thus obtained has little or no smell; its taste is strong and acrid, resembling that of ammoniacal salts. When exposed to the open air, it attracts moisture, and is converted into a thick liquid. It is extremely soluble in

* From νεφρος the kidney, the organ by which it is separated from the blood.

Book V.

water; and during its solution a considerable degree of cold is produced.* Alcohol dissolves it with facility, but scarcely in so large a proportion as water. The alcohol solution yields crystals much more readily on evaporation than the solution in water.

Crystallizes
with nitric
acid.

When nitric acid is dropped into a concentrated solution of nephrin in water, a great number of bright pearl-coloured crystals are deposited, composed of nephrin and nitric acid. No other acid produces this singular effect except the oxalic, as has been ascertained by Dr. Prout. The infusion of nutgalls gives the aqueous solution of nephrin a yellowish-brown colour, but causes no precipitate; neither does the infusion of tannin produce any precipitate.†

Action of
heat.

When heat is applied to nephrin, it very soon melts, swells up, and evaporates with an insupportably fetid odour. When distilled, there comes over first benzoic acid, then carbonate of ammonia in crystals, some carburetted hydrogen gas, with traces of prussic acid and oil; and there remains behind a large residuum, composed of charcoal, muriate of ammonia, and muriate of soda. The distillation is accompanied with an almost insupportably fetid alliaceous odour. Two hundred and eighty-eight parts of nephrin yield by distillation 200 parts of carbonate of ammonia, 10 parts of carburetted hydrogen gas, seven parts of charcoal, and 68 parts of benzoic acid, muriate of soda, and muriate of ammonia. These three last ingredients Fourcroy and Vauquelin consider as foreign substances, separated from the urine by the alcohol at the same time with the nephrin. Hence it follows that 100 parts of nephrin when distilled, yield

Carbonate of ammonia	92·027
Carburetted hydrogen gas. . .	4·608
Charcoal	3·225
	99·860

Now 200 parts of carbonate of ammonia, according to Fourcroy and Vauquelin, are composed of 86 of ammonia, 90 carbonic acid gas, and 24 water. Hence it follows that 100 parts of nephrin are composed of

* Fourcroy and Vauquelin, *Ann. de Chim.* xxxii. 88.† *Ibid.*

Oxygen	39·5
Azote	32·5
Carbon	14·7
Hydrogen	13·3
	100·0

But it can scarcely be doubted that the water which was found in the carbonate of ammonia existed ready formed in the nephrin before the distillation.*

Dr. Prout has subjected pure nephrin to a rigid analysis, and found its constituents as follows :

2 atoms hydrogen	= 0·25	6·66
1 atom carbon	= 0·75	20·0
1 atom oxygen	= 1·00	26·66
1 atom azote	= 1·75	16·66
	3·75		100·00

These proportions differ materially from those of Fourcroy and Vauquelin, and are doubtless much more accurate.

When the solution of nephrin in water is kept in a boiling heat, and new water is added as it evaporates, the nephrin is gradually decomposed, a very great quantity of carbonate of ammonia is disengaged, and at the same time acetic acid is formed, and some charcoal precipitates.†

When a solution of nephrin in water is left to itself for some time, it is gradually decomposed. A froth collects on its surface; air bubbles are emitted which have a strong disagreeable smell, in which ammonia and acetic acid are distinguishable. The liquid contains a quantity of acetic acid. The decomposition is much more rapid if a little gelatin or lime be added to the solution. In that case more ammonia is disengaged, and the proportion of acetic acid is not so great.‡

When the solution of nephrin is mixed with one-fourth of its weight of diluted sulphuric acid, no effervescence takes place; but, on the application of heat, a quantity of oil appears on the surface, which concretes upon cooling; the liquid which comes over into the receiver contains acetic acid, and a quantity of sulphate of ammonia remains in the retort dissolved in

* Ann. de Chim. xxxii. 88.
; Ibid. p. 96.

† Ibid. p. 96.

Book V.

the undistilled mass. By repeated distillations, the whole of the nephrin is converted into acetic acid and ammonia.*

Nitric acid,

When nitric acid is poured upon crystallized nephrin, a violent effervescence takes place, the mixture froths, assumes the form of a dark-red liquid, great quantities of nitrous gas, azotic gas, and carbonic acid gas, are disengaged. When the effervescence is over, there remains only a concrete white matter, with some drops of reddish liquid. When heat is applied to this residuum, it detonates like nitrate of ammonia. Into a solution of nephrin, formed by its attracting moisture from the atmosphere, an equal quantity of nitric acid, of the specific gravity 1.460, diluted with twice its weight of water, was added; a gentle effervescence ensued: a very small heat was applied, which supported the effervescence one or two days. There was disengaged the first day a great quantity of azotic gas and carbonic acid gas; the second day, carbonic acid gas, and at last nitrous gas. At the same time with the nitrous gas the smell of chlorocyanic acid was perceptible. At the end of the second day, the matter in the retort, which was become thick, took fire, and burnt with a violent explosion. The residuum contained traces of prussic acid and ammonia. The receiver contained a yellowish acid liquor, on the surface of which some drops of oil swam. †

Muriatic acid and chlorine.

Muriatic acid dissolves nephrin, but does not alter it. Chlorine gas is absorbed very rapidly by a diluted solution of nephrin; small whitish flakes appear, which soon become brown, and adhere to the sides of the vessel like a concrete oil. After a considerable quantity of this gas had been absorbed, the solution, left to itself, continued to effervesce exceedingly slowly, and to emit carbonic acid and azotic gas. After this effervescence was over, the liquid contained muriate and carbonate of ammonia.

Action of alkalis.

Nephrin is dissolved very rapidly by a solution of potash or soda, and at the same time a quantity of ammonia is disengaged; the same substance is disengaged when nephrin is treated with barytes, lime, or even magnesia. When pure solid potash is triturated with nephrin, heat is produced, a great quantity of ammonia is disengaged; the mixture becomes brown, and a substance is deposited, having the appearance of an empyreumatic oil. One part of nephrin and two of pot-

* Ann. de Chim. xxxii. 104.

† Ibid. p. 107.

ash, dissolved in four times its weight of water, when distilled, give out a great quantity of ammoniacal water; the residuum contains acetate and carbonate of potash.*

Chap. I.

When muriate of soda is dissolved in a solution of nephrin in water, it is obtained by evaporation, not in cubic crystals, its usual form, but in regular octahedrons. Muriate of ammonia, on the contrary, which crystallizes naturally in octahedrons, is converted into cubes, by dissolving and crystallizing it in the solution of nephrin.

Such are the properties of this singular substance, as far as they have been ascertained by the experiments of Fourcroy and Vauquelin. It differs from all animal substances hitherto examined, in the great proportion of azote which enters into its composition, and in the facility with which it is decomposed, even by the heat of boiling water.

SECT. IX.

OF SACCHARINE MATTER.

THE species of saccharine matter hitherto found in the animal kingdom are the following :

1. Sugar of milk.
2. Honey.
3. Sugar of diabetic urine.

1. Sugar of milk may be obtained by the following process: How obtained. let fresh whey be evaporated to the consistence of honey, and then allowed to cool; it concretes into a solid mass. Dissolve this mass in water, clarify it with the white of eggs, filter and evaporate to the consistence of a syrup; it deposits on cooling a number of brilliant white cubic crystals, which are *sugar of milk*. Fabricius Bartholdi, an Italian, was the first European who mentioned this sugar. He described it in his *Encyclopaedia Hermetico-Dogmatica*, published at Boulognia in 1619; but it seems to have been known in India long before that period. For the best account of its properties we are indebted to Mr. Lichtenstein.

When pure it has a white colour, a sweetish taste, and no Properties. smell. Its crystals are semitransparent regular parallelopipeds, terminated by four-sided pyramids. Its specific gravity, at the

* Ann. de Chim. xxxii. 107.

Book V.

temperature of 55° , is 1.543. At the temperature of 59° , it is soluble in five times its weight of water, and in two and a half times its weight of boiling water. It is perfectly insoluble in alcohol, but the addition of a little sulphuric acid makes it soluble in that liquid. When burnt it emits the odour of caramel, and exhibits precisely the appearance of burning sugar. When distilled, it yields the same products as sugar, only the empyreumatic oil obtained has the odour of benzoic acid.* When treated with nitric acid it yields *saccharic* acid. From the experiments of Vogel and Bouillon La Grange, we learn that sugar of milk is soluble in muriatic and acetic acid. Chlorine decomposes it into water and carbonic acid. Strong potash ley decomposes it in the cold, and converts it into water, carbonic acid, acetic acid, and a colouring matter. Neither alcohol nor ether dissolves it.† From the experiment of Bucholz, it appears that sugar of milk is not susceptible of the vinous fermentation.‡

Saccolate of lead.

Sugar of milk combines with protoxide of lead in various proportions. The neutral compound has a mucous appearance when moist, but becomes yellow and semitransparent when dry. Its constituents, according to the analysis of Berzelius, are as follows :

Sugar of milk 36.471 100 8.039
Oxide of lead 63.529 174.15 14
<hr style="width: 20%; margin: 0 auto;"/>			
100.000			

Oxide of lead digested with sugar of milk appears at first not in the least altered. It is however converted into a saccolate of lead, which when well washed and dried is a yellow powder composed of

Sugar of milk 12.8 100 2.055
Oxide of lead 87.2 681 14
<hr style="width: 20%; margin: 0 auto;"/>			
100.00 §			

From these analyses we see that the equivalent number for sugar of milk is 8.22. We may take 8.25 as sufficiently near.

It has been analysed both by Gay-Lussac and Thenard, and by Berzelius. According to Gay-Lussac and Thenard, its constituents are,

* Scheele, ii. 70.

† Ibid. p. 359.

† Schweigger's Jour. ii. 342.

§ Annals of Philosophy, v. 267.

Oxygen	53·834
Carbon	38·825
Hydrogen	7·341
	<hr/>
	100·000*

Chap. I.
 Composition.

Berzelius' numbers approach very nearly to these. According to him the constituents of sugar of milk are,

Oxygen	53·359
Carbon	39·474
Hydrogen	7·167
	<hr/>
	100·000†

These analyses do not correspond well with the equivalent number for the weight of sugar of milk deduced from analysis. Were we to suppose it composed of 4 atoms oxygen, 5 atoms carbon, and 4 atoms hydrogen, its composition would be as follows :

4 Atoms oxygen .. = 4 48·4
5 Atoms carbon .. = 3·75 45·4
4 Atoms hydrogen . = 0·5 6·2
	<hr/>
	8·25 100·0

Numbers quite irreconcilable with the analyses above given. The subject therefore requires farther investigation.‡

2. Honey is prepared by bees, and perhaps rather belongs **Honey.** to the vegetable than the animal kingdom. It has a white or yellowish colour, a soft and grained consistence, a saccharine and aromatic smell. By distillation it affords an acid phlegm and an oil, and its coal is light and spongy like that of the mucilages of plants. Nitric acid extracts from it oxalic acid, precisely as it does from sugar. It is very soluble in water, with which it forms a syrup, and like sugar passes to the vinous fermentation. Mr. Cavezzali has affirmed that honey is composed of sugar, mucilage, and an acid. The sugar may

* Recherches Physico-chimiques, ii. 295.

† Annals of Philosophy, v. 266.

‡ If the reader wishes a more detailed account of sugar of milk, he may consult Lichtenstein's Abhandlung, published separately in 1772, and Ronelle's dissertation in the 39th volume of the Jour. de Medicine. An abridged view of the whole has been given by Morveau in the first volume of the Encycl. Methodique, which has been published in an English dress by Mr. Johnson in his History of Animal Chemistry, i. 123.

Book V.

be separated by melting the honey, adding carbonate of lime in powder as long as any effervescence appears, and scumming the solution while hot. The liquid thus treated gradually deposits crystals of sugar when allowed to remain in a glass vessel.* Its crystals, as far as I have observed, are quite analogous to those of sugar of grapes.

According to Proust, there are two kinds of honey; one always liquid, and the other solid and not deliquescent. They may be separated, he says, by means of alcohol.†

Sugar of
diabetic
urine.

3. The urine of persons labouring under the disease known to physicians by the name of *diabetes*, yields, when evaporated, a considerable quantity of matter which possesses properties analogous to sugar. This seems to have been first observed by Willis. Afterwards the subject was carried somewhat farther by subsequent physicians. But it is to Mr. Cruikshanks that we are indebted for the most valuable set of experiments on the subject. In this case he extracted from urine about $\frac{1}{12}$ th of its weight of a sweet-tasted extract like honey. When treated with nitric acid, it yielded the same proportion of oxalic acid as an equal quantity of common sugar would have done, making allowance for the saline substances present. No saccharic acid was formed. Hence it follows that this substance is not analogous to sugar of milk, but nearer common sugar in its properties. It crystallizes nearly in the same manner as sugar of grapes. I have seen it prepared by Dr. Wollaston in small grains having almost exactly the appearance of common white sugar. Nicolas affirms that when treated with lime it is decomposed.‡ If so, it certainly differs essentially from common sugar, which Cruikshanks first proved to be capable of uniting to lime without decomposition. Indeed Chevreul has lately ascertained that diabetic sugar possesses all the characters of sugar of grapes.§

SECT. X.

OF CANTHARIDIN.

I GIVE this name to the substance in cantharides or Spanish flies, (*meloé vesicatorius*) which occasions a blister when applied to the skin. It was obtained in a separate state, and its

* Ann. de Chim. xxxix. 410.

† Jour. de Phys. lix. 423.

‡ Ann. de Chim. xlv. 61.

§ Annals of Philosophy, ix. 55.

properties were examined by M. Robiquet in 1810.* It may be obtained by the following process. Chap. I.

Boil cantharides in water till every thing soluble in that liquid be taken up. Concentrate the solution by evaporation, and when reduced to a thick syrup boil it repeatedly in alcohol till that liquid ceases to act upon it. Evaporate the alcoholic solution gently to dryness, and put the dry residue into a phial with sulphuric ether, and agitate the mixture for a considerable time. At first the ether will seem to have no effect upon it, but after some hours it assumes a yellow colour. Decant it off and allow it to evaporate spontaneously in the open air. It deposits small crystalline plates mixed with a yellow matter. Alcohol takes up the yellow matter; but leaves the crystalline plates. These plates when dried between folds of blotting paper are cantharidin in a state of considerable purity. Preparation.

Cantharidin thus obtained is in small plates, having a shining micaceous appearance. It is insoluble in water and in cold alcohol. Boiling alcohol dissolves it, but lets it fall again in crystals as the solution cools. Ether dissolves it but not very powerfully. It dissolves very readily in oils, and when applied to the skin it acts as a vesicatory with great energy. The solution of it in oils is equally efficacious.† Properties.

It would appear probable from a set of experiments by Dr. Freeman Dana, that the *lytha vittata*, or potatoe fly, of North America, which is known to possess very powerful vesicating properties, contains likewise a quantity of cantharidin.‡

SECT. XI.

OF COCHENILIN.

THE cocheneal is an insect which inhabits different species of cactus, and which is propagated in Mexico and some other countries in order to be employed as a dye stuff. By it the scarlet, the most splendid of all colours, is given to cloth. It is to the red colouring matter of the insect that Dr. John has given the name of cochenilin. He has examined this colouring matter, and established it as a peculiar animal principle. According to him the constituents of the cocheneal insect, are the following:

* Ann. de Chim. lxxvi. 302.

† Robiquet, Ann. de Chim. lxxvi. 308.

‡ American Journals of Science of Arts, ii. 157.

Book V.

Cochenilin	50·0
Jelly	10·5
Waxy fat	10·0
Gelatinous mucus	14·0
Shining matter	14·0
Alkaline phosphate	}
Alkaline muriate	
Phosphate of lime	
Phosphate of iron	
Phosphate of ammonia	1·5
	—————
	100·0*

M. M. Pelletier and Caventou, who have examined the cochineal more lately than John, have given to the peculiar colouring matter which it contains the name of *carmine*; but I prefer John's name; both as more agreeable to the rules of the chemical nomenclature, and as more suited to our language; or at least less ambiguous. For the term *carmine* in the English language has already an appropriate signification.

The French chemists obtained cochenilin in a state of purity, by the following process. The cochineal insect was digested in alcohol as long as it gave a red colour to that liquid. The alcoholic solution when left to spontaneous evaporation let fall a crystalline matter of a fine red colour. These crystals were dissolved in strong alcohol, and the solution was mixed with its own bulk of sulphuric ether. It became muddy, and after an interval of some days the cochenilin was deposited at the bottom of the vessel forming a beautiful purplish red crust. The properties of cochenilin thus obtained are as follows:

It has a fine purple red colour. It adheres strongly to the sides of the vessel in which it is deposited. It has a granular appearance, as if it were composed of crystals. It is not altered by exposure to the air. It does not absorb any sensible quantity of moisture. When heated to the temperature of 122° it melts. If the heat be increased, it swells up and is decomposed, yielding carburetted hydrogen gas, a great deal of oil, and a little water, having a slightly acid taste. It gives out no traces of ammonia.

It is very soluble in water. The liquid may be reduced by evaporation to the consistency of a syrup; but the carmine does not crystallize. The watery solution of it has a fine car-

* John's Tabellen des Thierreichs, p. 127, and Chemische Untersuchungen, iii. 210.

mine red colour. A very small portion of this substance gives a strong colour to a great quantity of water. It is soluble likewise in alcohol; but the stronger the alcohol is, the worse a solvent does it become. It is insoluble in sulphuric ether; the weak acids dissolve it; but probably merely in consequence of the water which they contain. No acid precipitates it when pure; but they almost all throw it down when it is in combination with the peculiar animal matter of the cochineal; but all the acids produce a sensible change upon the aqueous solution of carmine. They make it assume in the first place a lively red colour, which gradually acquires a yellowish tinge, and at last becomes entirely yellow. When the acids are not very concentrated, the carmine is not altered in its nature; for when the acid is saturated, the colour resumes its former appearance.

Concentrated sulphuric acid destroys and chars carmine. Muriatic acid decomposes it without charring it, and converts it into a bitter substance, which has no resemblance to carmine. Nitric acid decomposes it with still greater rapidity. Some needle-form crystals are formed similar in appearance to oxalic acid; but they do not precipitate lime water even when mixed with ammonia. The nature of these crystals was not ascertained.

Chlorine acts with energy on carmine, giving it first a yellow colour, which it gradually destroys altogether. It occasions no precipitate in aqueous solution of carmine if no animal substance be present. It is, therefore, a useful reagent to enable us to discover the presence of animal matter in this colouring principle. Iodine acts in the same way as chlorine, but with less rapidity.

When the alkalis are poured into a solution of carmine, they give it a violet colour. If the alkali be saturated immediately, the original colour appears, and of course the carmine remains unaltered, or at least only slightly modified; but if the action of the alkali be prolonged, or if it be augmented by the application of heat, the violet colour is dissipated, the liquid becomes first red, and then yellow. The nature of the colouring matter is now completely altered.

Lime water occasions a violet coloured precipitate when dropped into the aqueous solution of carmine. Barytes and strontian occasion no precipitate, but produce the same change of colour as the alkalis. Alumina has a very strong affinity for carmine. When newly precipitated alumina is put into

Book V.

an aqueous solution of carmine, the liquid is totally deprived of its colour, and the alumina converted into a beautiful lake. If a few drops of acid be added to the aqueous solution before adding the alumina, the lake obtained has a fine red colour as before; but it becomes violet upon the application of the least heat. The same effect is produced by adding to the liquid a few grains of an aluminous salt.

Most of the saline solutions alter the colour of the aqueous solution of carmine; but few of them are capable of producing a precipitate in it. The salts of gold alter the colour merely; nitrate of silver produces no change whatever; the soluble salts of lead render the colour violet; and acetate of lead occasions a copious violet precipitate. By decomposing this precipitate by means of a current of sulphuretted hydrogen, we may obtain the carmine dissolved in water in a state of purity.

Protonitrate of mercury throws down a violet precipitate. Pernitrate of mercury does not act so powerfully, and the colour of the precipitate is scarlet. Corrosive sublimate produces no effect whatever.

Neither the salts of copper nor of iron occasion any precipitate; but the former changes the colour of the liquid to violet, the latter to brown.

Protomuriate of tin throws down a copious violet precipitate. The permuriate changes the colour to scarlet, but produces no precipitate. When gelatinous alumina is added to the mixture, we obtain a fine red precipitate, which is not altered by boiling.

None of the aluminous salts occasion a precipitate, but they change the colour to carmine. The salts of potash, soda, and ammonia, change the colour of the liquid to carmine red.

From the action of the different salts upon this colouring matter, Pelletier and Caventou have drawn as a conclusion that the metals susceptible of different degrees of oxygenation act like acids upon the colouring matter of cochineal when at a maximum of oxidation, but like alkalies when at a minimum or medium degree; and that this alkaline influence may be exercised in the midst of an acid when the oxides in question are capable of forming an insoluble precipitate with the colouring matter.

Tanmin, and astringent substances in general, do not precipitate the colouring matter of cochineal.

Pelletier and Caventou mixed a quantity of the colouring

matter with black oxide of copper, and subjected the mixture to the requisite degree of heat. The only gaseous substance obtained was carbonic acid. Hence it follows that carmine is composed of carbon, oxygen, and hydrogen, and that it contains no azote whatever.

SECT. XII.

OF OILS.

THE oily substances found in animals belong all to the class of fixed oils. They differ very much in their consistence, being found in every intermediate state from spermaceti, which is perfectly solid, to train oil, which is completely liquid. The most important of them are the following:

1. *Fat*.—This substance is found abundantly in different parts of animals. When pure it possesses the properties of the fixed oils. Its consistence varies from *tallow* or *suet*, which is brittle, to *hog's lard*, which is soft and semifluid. To obtain fat pure, it is cut in small pieces, well washed in water, and the membranous parts and vessels separated. It is then melted in a shallow vessel along with some water, and kept melted till the water is completely evaporated. Thus purified it is white, tasteless, and nearly insipid.

Different kinds of it liquify at different temperatures. Lard melts at 97° ; but the fat extracted from meat by boiling requires, according to Nicholson, a heat of 127° . When heated to about 400° , it begins to emit a white smoke, which becomes more copious and more disagreeable as the heat increases; at the same time it becomes blackish, owing doubtless to the decomposition of a portion of it and the evolution of some charcoal. If it be now cooled it becomes more brittle and solid than at first. When hog's lard is distilled in a retort, we obtain first a portion of water, and then a white oil, which concretes in the receiver into small round globules, and a black mass remains in the retort. During the whole process abundance of carburetted hydrogen and carbonic acid come over, which have a most abominable odour. When the vessels are unluted, they emit an odour so detestable and powerful that it is absolutely impossible to endure it. It was formerly supposed that this odour was exhaled by the sebacie acid formed during the process; but Thenard has shown this to be a mistake. We must therefore ascribe it to an empyreumatic oil. The smell does not suddenly disappear. During this dis-

Book V.

tillation acetic acid is disengaged, and likewise some sebacic acid, which is found in solution in the receiver mixed with the oil.

Nature.

These products demonstrate the peculiar nature of fat; at the same time its analogy to fixed oil is so strong that we cannot but consider it as belonging to the same genus of bodies. Its constituents are undoubtedly oxygen, hydrogen, and carbon, in unknown proportions. It has been supposed also to contain an acid. The opinion originated from the older chemists, that all the solid oily bodies were indebted to an acid for their solidity. The experiments of Crell were supposed to demonstrate the truth of this opinion; but Thenard has lately shown that these experiments were erroneous, and that the results obtained are to be ascribed to the impurity of the substances employed. The sebacic acid is evolved by the action of heat upon the fat.

Fat is insoluble in water, alcohol, and ether. The strong acids dissolve and gradually decompose it. With the alkalies it combines and forms soap.

When a little nitric acid is poured on fat, and a moderate heat applied, the acid deposits its oxygen, and converts the fat into a yellow-coloured ointment, which Fourcroy considers as an oxide of fat, and which is said to have been employed with great success in France in external venereal affections. A portion of sacclactic acid is formed by the action of the nitric acid. Fat dissolves a little sulphur and phosphorus. The phosphorus is speedily converted into phosphorous acid. According to Vogel the mercury in the medicinal preparations made by triturating mercury and fat is in the metallic state.*

The rancidity of fat is to be ascribed to the same cause as the rancidity of the fixed oils.

Train oil.

2. *Train oil*.—This liquid is extracted from the blubber of the whale, and from other fish. It forms a very important article of commerce, being employed for combustion in lamps, and for other purposes. It is at first thick; but on standing, a white mucilaginous matter is deposited, and the oil becomes transparent. It is then of a reddish brown colour, and has a disagreeable smell. If it be agitated with a little sulphuric acid, and then mixed with water, the oil, when allowed to settle, swims on the surface, of a much lighter colour than before, the water continues milky, and a curdy matter is ob-

* Ann. de Chim. lviii. 154.

served swimming between the oil and water. This method has been employed for purifying the oil, to make it more proper for burning in lamps. Various other methods have been employed. The most important are those recommended by Mr. Dossie.*

Train oil boils when heated to the temperature of 640° . If it be distilled over, it loses much of its viscosity, and acquires the property of boiling when heated to 410° . By passing it through an iron pipe filled with fragments of pounded brick heated to redness, it is converted into a highly volatile oil, which boils at 180° and burns as readily as ether. This oil is limpid and colourless, very fluid, has a strong smell, and is known in chemistry by the name of Dippel's oil, because it was first prepared by that chemist. The constituents of train oil according to my experiments, are as follows:

One grain of it when passed through red hot peroxide of copper, is converted into 5.43 cubic inches of carbonic acid gas, and 1.45 grains of water. Now

Carbon in 5.43 cubic inches carbonic acid gas	=	0.6887 gr.
Hydrogen in 1.45 grains of water	=	0.1610
Oxygen to make up the weight	=	0.1503

Total = 1.0000

But 0.6887 gr. carbon	=	6 volumes or 12 atoms	=	9.000	
0.1610 gr. hydrogen	=	8.5 volumes	17 atoms	=	2.125
0.1503 gr. oxygen	=	0.5 volume	2 atoms	=	2.000

Weight of an atom of whale oil = 13.125 or to a multiple of that number. From Chevreul's experiments on soap, it is probable that the true weight of an atom of whale oil is thrice the preceding number.

When whale oil is distilled over its specific gravity is reduced from 0.9191, the specific gravity of common whale oil, to 0.8680. A grain of it passed through red hot peroxide of copper is converted into 5.535 cubic inches of carbonic acid gas, and 1.3 grains of water. Hence its constituents are

Carbon	0.702 or 12 volumes or 13 atoms	=	9.000
Hydrogen	0.144 15 volumes 15 atoms	=	1.875
Oxygen	0.156 $\frac{1}{2}$ volume 2 atoms	=	2.0000

1.000

12.875

Thus it seems to have lost only two atoms of hydrogen.

* See Phil. Mag. xv. 105; and the Transactions of the Society for the Encouragement of the Arts, vol. xx.

Book V.

The specific heat of whale oil is 0.5302. When heated from 60° to 450 in glass, it expands $\frac{1}{5.65}$, not taking into account the expansion of glass.

Spermaceti oil.

3. *Spermaceti oil*.—This is the oil which separates from spermaceti during the purification of it. It is much purer than train oil, and therefore answers better for burning in lamps.

Butter.

4. *Butter* possesses all the characters of a fixed oil; but as it is not found ready formed in any animal substance, the detail of its properties belongs to the next Chapter.

Oil of ants.

5. *Oil of ants*.—Margraff obtained this oil by macerating the *formica rufa* in water, and distilling off one-half of the liquid. The oil was found swimming upon the surface of the water in the receiver. It is liquid, and dissolves completely in alcohol. Its taste is not hot, but it has a peculiar smell. It appears to be intermediate between fixed and volatile oils in its properties.*

Animal oil of Dippel.

6. Though all the oily bodies found in animal substances belong to the class of fixed oils, yet there is a peculiar volatile oil which makes its appearance, and which is doubtless formed during the distillation of different animal bodies. Though this oil has now lost that celebrity which drew the attention of the older chemists to it, yet as its properties are peculiar, a short account of it will not be improper. It is usually called the *animal oil of Dippel*, because that chemist first drew the attention of chemists to it. It is commonly obtained from the gelatinous and albuminous parts of animals. The horns are said to answer best. The product of the first distillation is to be mixed with water, and distilled with a moderate heat: the oil which is first obtained is the animal oil of Dippel.

It is colourless and transparent; its smell is strong and rather aromatic; it is almost as light and as volatile as ether; water dissolves a portion of it; and it changes syrup of violets green, owing, as is supposed, to its containing a little ammonia. The acids all dissolve it, and form with it a kind of imperfect soap. Nitrous acid sets it on fire. It forms with alkalies a soap. Alcohol, ether, and oils unite with it. When exposed to the air it becomes brown, and loses its transparency. It was formerly used as a specific in fevers.†

* Margraff, Opusc. i. 291.

† See the experiments of Model and Delue.

SECT. XIII.

OF ANIMAL RESINS.

SUBSTANCES resembling resins are found in different animal bodies; and which, for that reason, may be called *animal resins*. Their properties are somewhat different from the vegetable resins, but they have not been all examined with precision. The following are the most remarkable of these substances: Chap. II.

1. *Ambergris*.—This substance is found floating on the sea, near the coasts of India, Africa, and Brazil, usually in small pieces, but sometimes in masses of 50 or 100 pounds weight. Various opinions have been entertained concerning its origin. Some affirmed that it was the concrete juice of a tree;* others thought it a bitumen; but it is now considered as pretty well established, that it is a concretion formed in the stomach or intestines of the *physeter macrocephalus*, or spermaceti whale. † Ambergris.

Ambergris when pure is a light soft substance which swims on water. Its specific gravity varies from 0.78 to 0.92, according to Brisson: Bouillon La Grange, who has published an analysis of it, found its specific gravity from 0.849 to 0.844. ‡ Its colour is ash-grey, with brownish-yellow and white streaks. It has an agreeable smell, which improves by keeping. Its taste is insipid. Properties.

When heated to 122° it melts without frothing; if the heat be increased to 212°, it is volatilized completely in a white smoke, leaving only a trace of charcoal. When distilled, we obtain a whitish acid liquid and a light volatile oil; a bulky charcoal remains behind. §

It is insoluble in water. Acids have little action on it. Weak sulphuric acid occasions no change; but when concentrated it develops a little charcoal. Nitric acid dissolves it, giving out at the same time nitrous gas, carbonic acid, and azotic gas. A brownish liquid is formed, which leaves, when evaporated to dryness, a brittle brown substance, possessing the properties of a resin. The alkalies dissolve it by the as-

* See the New Abridgement of the Phil. Trans. ii. 89.

† See Phil. Trans. vols. xxxiii. xxxviii. and xc.

‡ Ann. de Chim. xlvii. 73.

§ Bouillon La Grange, Ann. de Chim. xlvii. 73.

Book V.

sistance of heat, and form a soap soluble in water. Both the fixed and volatile oils dissolve ambergris. It is soluble also in ether and alcohol.*

Analysis.

When 100 parts of ambergris are macerated for 24 hours in 16 parts of alcohol, the liquid assumes a deep yellow colour. By repeating the process, the alcohol takes up the whole that is soluble. Hot alcohol dissolves the rest, except a little charry matter easily separated by filtration. It weighs 5·4 parts. The cold infusions of alcohol being evaporated to dryness, there remains a brown brilliant substance, which is readily melted by heat, and which volatilizes completely when put upon hot charcoal. This substance possesses the properties of a resin. It amounts to 30·8 parts. It is precipitated from alcohol by water, leaving a portion of benzoic acid, which may be likewise separated from ambergris by heat, and by the usual process by which it is extracted from benzoin. It amounts to 11·1 parts. When the hot alcohol infusion is allowed to cool, it lets fall the portion of ambergris which it held in solution. This substance is of a pale-yellow colour, and easily softens by heat. When slowly cooled, it assumes a lamellar form. It possesses the properties of the fatty matter into which the muscles are converted by nitric acid, and which makes its appearance when dead bodies are allowed to putrefy in great numbers together. This substance has been distinguished by the name of *adipocire*, from its resemblance both to fat and wax. The quantity of it in ambergris amounts to 52·8 parts. Such is the analysis of ambergris made by Bouillon La Grange. This substance, then, according to that chemist, is composed of

Constituents.

52·7 adipocire
30·8 resin^a
11·1 benzoic acid
5·4 charcoal

100·0 †

But Bucholz, who subjected ambergris to a later examination, was unable to extract any benzoic acid from it. He considers it as a peculiar substance. ‡

Propolis.

2. *Propolis*. This is a substance collected by bees newly

* Bouillon La Grange, Ann. de Chim. xlvii. 75.

† Ann. de Chim. xlvii. 80.

‡ Ibid. lxxiii. 95.

placed in a hive. They employ it to cover the bottom of the hive, and to coat over all foreign substances which are too heavy to be removed. The bees bring it on their legs and feet; to which it adheres so strongly that the insect cannot extricate itself; but its associates pull it off with their jaws, and apply it to immediate use. It is probably a vegetable production; but as this is not known for certain, and as it is prepared by animals, I have reserved an account of its properties for this place. Vauquelin has subjected it to a chemical analysis.

It is at first soft and ductile, but it gradually hardens and becomes solid. In large masses it appears blackish; but in thin pieces it is semitransparent. The heat of the fingers is sufficient to soften it, and give it the ductility of wax; but it is more ropy and tenacious. It is tasteless, but has an aromatic odour similar to that of the balsam of Peru. When macerated in alcohol, it tinges the liquid of a deep red colour; and by repeating the maceration, the portion soluble is separated. Boiling alcohol dissolved another portion of it; and by repeated decoctions 100 parts of propolis were reduced to 14, which consisted of extraneous bodies. All the alcoholic solutions were mixed and passed through a fine strainer, which separated the portion that precipitated as the decoction cooled. The substance, when properly dried, possessed the properties of wax. The alcohol was now concentrated by distillation, then mixed with water, and boiled. A resinous mass was obtained of a red-brown colour, semitransparent, and very brittle. It weighed 57 parts. The water in which it had been boiled contained an acid, probably the benzoic. By distillation the resinous matter yielded a volatile oil, and became harder. This resin is soluble in fixed and volatile oils, and very similar in its properties to the resin of balsam of Peru. Thus 100 parts of propolis contain

57 resin
14 wax
14 impurities
15 acid and loss
<hr/>
100 *

Constitu-
ents.

3. *Castor*. This substance is obtained from the beaver. *Castor*. In each of the inguinal regions of that animal there are two

* Nicholson's Jour. v. 48.

Book V.

bags, a large and a small. The large one contains the true castor; the small one a substance which has some resemblance to it, but which is in much less estimation. We are indebted to Bouillon La Grange for a set of experiments on it.*

Properties.

Castor is of a yellow colour; and when newly taken from the animal it is nearly fluid. But by exposure to the atmosphere it gradually hardens, becomes darker coloured, and assumes a resinous appearance. Its taste is bitter and acrid, and its odour strong and aromatic. In water it softens, and tinges the liquid of a pale-yellow colour. The infusion gives a green colour to vegetable blues, and contains an alkali. By long continued maceration a deeper coloured infusion is obtained, which yields a dry extract of a tortoiseshell colour, soluble in ether and alcohol. Water precipitates a resinous substance from the alcohol. From the analysis of Bouillon La Grange, we learn that castor contains the following ingredients:

Constituents.

1. Carbonate of potash:
2. Carbonate of lime.
3. Carbonate of ammonia.
4. Iron.
5. Resin.
6. A mucilaginous extractive matter.
7. A volatile oil.

The properties of the resin are analogous to those of the resin of bile.

Civet.

4. *Civet*.—This substance, like the last, is obtained from the inguinal region of the *civet-cat*. It is squeezed out of the cavity where it is secreted every other day. It is employed as a perfume, but has not hitherto engaged the attention of chemists. Its colour is yellow; its consistence that of butter; its smell so strong as to be agreeable only when much reduced by mixture with other bodies. It unites readily with oils, but is said not to dissolve in alcohol. In that case it ought to be excluded from the class of resinous bodies.

Musk.

Musk.—This substance is secreted into a kind of bag situated in the umbilical region of the quadruped called *moschus moschifer*. Its colour is brownish-red; it feels unctuous; its taste bitter; and its smell aromatic and intensely strong. It is partially soluble in water, which acquires its smell; and in alcohol, but that liquid does not retain the odour of musk. Nitric and sulphuric acids dissolve it, but

* Jour. de Phys. xlii.

destroy the odour. Fixed alkalis develop the odour of ammonia. Oils do not act on it. At a red heat it has the same fetid smell as urine.* Its component parts have not been ascertained.

Chap. I.

6. The *formica rufa*, or red ant, contains also a peculiar resin, as has been proved by Fourcroy and Vauquelin.†

SECT. XIV.

OF SULPHUR AND PHOSPHORUS.

THESE two simple combustibles constitute a part of many animal substances, but in what state of combination they exist in them has not been ascertained. All that can be done therefore at present is to mention some of the bodies from which they may be separated.

1. Sulphur seems to be a never-failing attendant on albumen, in whatever part of the body that substance exists as a constituent. Scheele ascertained its presence in the white of an egg and in milk; ‡ Parozotier and Deyoux in the blood; § Proust in urine and feces, in the muscles, in hair, &c. and he has made it probable that sulphur is mixed with the matter which is transpired from the body. ¶ This chemist informs us, that in the blood it is combined with ammonia in the state of hydrosulphuret of ammonia; but whether it be in the same state in other parts of the body is not known.

2. Phosphorus exhales from different animal bodies during putrefaction in the state of phosphuretted hydrogen gas, and of course must have been a constituent part of these bodies. The luminous appearance which some animals exhibit seems, in several instances, to be owing to the presence of this very combustible substance: but the experiments hitherto made upon these animals are not sufficiently decisive.

SECT. XV.

OF ACIDS.

THE acids which have been discovered ready formed, and constituting a part of animal bodies, are the following:

* Nomenclature.

‡ Scheele, ii. 37.

§ Ann. de Chim. xxvii. 236.

† Phil. Mag. xv. 143.

¶ Jour. de Phys. xlv. 437.

Book V.

- | | |
|----------------|--------------|
| 1. Phosphoric. | 8. Amniotic. |
| 2. Sulphuric. | 9. Oxalic. |
| 3. Muriatic. | 10. Formic. |
| 4. Carbonic. | 11. Acetic. |
| 5. Benzoic. | 12. Malic. |
| 6. Uric. | 13. Lactic. |
| 7. Rosacic. | 14. Silica. |

Phosphoric. 1. The phosphoric acid is by far the most abundant of all the acids found in animals. Combined with lime, it constitutes the basis of bone; and the phosphate of lime is found in the muscles and almost all the solid parts of animals; neither are there many of the fluids from which it is absent. In the blood, phosphoric acid is found combined with oxide of iron; and in the urine it exists in excess, holding phosphate of lime in solution.

Sulphuric. 2. Sulphuric acid has been shown by Berzelius to be a constant ingredient of urine. It is a very common constituent of the liquid contents of the inferior animals. Thus sulphate of soda is found in the liquor of the amnios of cows, and sulphate of lime occurs usually in the urine of quadrupeds.

Muriatic. 3. Muriatic acid occurs in most of the fluid animal substances, and is almost always combined with soda, constituting common salt.

Carbonic. 4. Carbonic acid has been detected in fresh human urine by Proust, and it occurs in the urine of horses and cows abundantly, partly combined with lime.

Benzoic. 5. Benzoic acid was first discovered in human urine by Scheele; and Fourcroy and Vauquelin have found it abundantly in the urine of cows. Proust has detected it in the blood, the albumen of an egg, in glue, silk, and wool, in the sponge, different species of algæ, and even in mushrooms. This chemist even thinks it probable that it is a constituent of all substances which contain azote.* Or more probably it is formed during the action of reagents on these substances.

Uric. 6. Uric or lithic acid was discovered by Scheele in 1776. It is the most common constituent of urinary calculi, and exists also in human urine. That species of calculus, which resembles wood in its colour and appearance, is composed entirely of this substance. It was called at first lithic acid; but this name, in consequence of the remarks made by Dr. Pear-

* Ann de Chim. xxxvi. 272.

son on its impropriety, has been laid aside, and that of *uric acid* substituted in its place.* Chap. I.

7. *Rosacic*. During intermittent fevers urine deposits a very copious precipitate, which has been long known to physicians under the name of *lateritious sediment*. This sediment always makes its appearance at the crisis of fevers. In gouty people, the same sediment appears in equal abundance towards the end of a paroxysm of the disease; and if this sediment suddenly disappears after it has begun to be deposited, a fresh attack may be expected.† Schæele considered this sediment as uric acid mixed with some phosphate of lime; and the same opinion has been entertained by other chemists: but Proust affirms that it consists chiefly of a different substance, to which he has given the name of *rosacic acid* from its colour, mixed with a certain proportion of uric acid and phosphate of lime. This *rosacic acid*, he informs us, is distinguished from the uric by the facility with which it dissolves in hot water, the violet precipitate which it occasions in muriate of gold, and by the little tendency which it has to crystallize. Rosacic.

8. *Amniotic acid* has been discovered by Vanquelin and Buniva in the liquor of the amnios of the cow, and may be obtained in white crystals by evaporating that liquid slowly. Hence they have given it the name of *amniotic acid*. It is of a white and brilliant colour; its taste has a very slight degree of sourness; it reddens the tincture of turnsole; it is scarcely soluble in cold water, but very readily in hot water, from which it separates in long needles as the solution cools. It is soluble also in alcohol, especially when assisted by heat. It combines readily with pure alkalies, and forms a substance which is very soluble in water. The other acids decompose this compound; and the acid of the liquor of the amnios is precipitated in a white crystalline powder. This acid does not decompose the alkaline carbonates at the temperature of the atmosphere, but it does so when assisted by heat. It does not alter solutions of silver, lead, or mercury, in nitric acid. When exposed to a strong heat it froths, and exhales an odour of ammonia and of prussic acid. These properties are sufficient to show that it is different from every other acid. Amniotic.

* Dr. Pearson rather chooses to call it *uric oxide*.

† Cruickshanks, Phil. Mag. ii, 249.

Ann. de Chim. xxxvi. 265.

Book V.

It approaches nearest to the sacclactic and the uric acids; but the first of these does not furnish ammonia by distillation like the amniotic; the uric acid is not so soluble in hot water as the amniotic; it does not crystallize in white brilliant needles, and it is insoluble in boiling alcohol: in both which respects it differs completely from amniotic acid.*

Oxalic. 9. Oxalic acid has been found only in a species of urinary calculi first analysed by Dr. Wollaston.

10. Formic acid has been hitherto found only in the *formica rufa*, or red ant. The details have been already given in a preceding part of this work.

Acetic. 11. Acetic. This acid exists in the *formica rufa*, or red ant, as has been demonstrated by the experiments of Fourcroy and Vauquelin.

Malic. 12. Malic acid. This acid has been detected by Fourcroy and Vauquelin in the acid liquid obtained from the *formica rufa*. When this liquid is saturated with lime, if acetate of lead be dropped into the solution, a copious precipitate falls, which is soluble in acetic acid. Fourcroy and Vauquelin exposed the precipitate to the proper trials, and ascertained that it was malate of lead.

Lactic. 13. Lactic acid has been shown by Berzelius to exist in urine, milk, and all the fluids of the animal body.

Silica. 14. Silica has not hitherto been detected in any of the component parts of animals, except hair; but Fourcroy and Vauquelin found it in urinary calculi.

Other acids. Besides the acids enumerated above, several others have been mentioned by chemical writers. The truth is, that a variety of acid liquors have been observed in animals; but as these have not been examined, we do not know whether they are peculiar substances, or merely some of those already enumerated. Chaussier, in 1781, pointed out an acid in the silk worm, which has been called *bombic acid*.† But the list of its properties which he has given is altogether insufficient to constitute it a peculiar substance; and all subsequent accounts of it which I have seen are mere repetitions of the original paper. Delue obtained an acid from the insects called *meloe proscarabæus* and *majalis*. Chaussier obtained an acid from the grasshopper and the bug. In short, acids are obtained from many of the insect tribe; and from the experiments published, it is extremely probable that in most of these insects the acid is the acetic.

* Ann. de Chim. xxxiii, 279.

† Mem. Dijon, 1783, ii. 70.

SECT. XVI.

OF ALKALIES, EARTHS, AND METALS.

I. ALL the alkalies have been found in the fluids of animals. Chap. I.

1. Potash is rather uncommon in the human fluids; but it Potash. has been detected in the milk of cows, and it has been found abundantly in the urine of quadrupeds.

2. Soda exists in all the fluids, and seems always to be com- Soda. bined with albumen. Phosphate and muriate of soda are also found. It is this alkali which gives animal fluids the property of tinging vegetable blues green.

3. Ammonia has been detected by Pronst in urine; and it Ammonia. is formed in abundance during the putrefaction of most animal bodies.

II. The only earths hitherto found in animals are lime and magnesia.

1. Lime exists in great abundance in all the larger animals. Lime. Combined with phosphoric acid, it constitutes the basis of bones, while shells are composed of carbonate of lime. Phosphate of lime is found also in the muscles and other solid parts, and it is held in solution by almost all the fluids.

2. Magnesia has been detected in human urine by Four- Magnesia. croy and Vanquelin, combined with phosphoric acid and ammonia. It constitutes also sometimes a component part of urinary calculi.

The metals found in animals are two: namely, iron and manganese.

1. Iron combined with phosphoric acid is a constituent part Iron. of the blood. Its presence was first ascertained by Manghini, who proved at the same time that it does not exist in the solid parts of animals. It is said to exist also in bile.

2. Manganese has been found in human hair, but scarcely Manganese. in any other animal substance.

SECT. XVII.

OF ANIMAL PRINCIPLES IN GENERAL.

SUCH are the properties of all the animal substances hitherto detected and accurately examined. They are by no means so numerous as the vegetable principles, owing in all probability to

Book V. the chemical examination of animals having been hitherto confined to a small number, while a much more extensive range has been taken in the vegetable kingdom. The inferior animals, if we except a few insects, have attracted but little of the attention of chemists. I confess, too, that I have omitted on purpose an account of several principles which are well known to exist in the animal kingdom, merely because we know nothing so precise about their chemical properties as to enable us to class them properly. I may mention, as an example, the colouring matter of the shell-fish with which the ancients dyed their purple. The reader will find a very full and interesting account of all that we know respecting this substance in Dr. Bancroft's *Philosophy of Permanent Colours*. I may mention also the liquor emitted by the ink fish.

Substances omitted.

Almost the whole of the soft parts of animal bodies consist of albumen, fibrin, and oil; while the bases of the hard parts is either phosphate or carbonate of lime. The other animal principles occur only in small quantities, and particular parts of the animal. The oils seldom enter into the structure of the organs of animals: they serve rather to lubricate the different parts, and to fill up interstices. Gelatin is insoluble in cold water, albumen insoluble in hot, and fibrin insoluble both in cold and hot. The constituents of these three bodies, and of nephrin, according to the best analyses of them hitherto made, are as follows:

	Atoms of			
	Carbon.	Oxygen.	Hydrogen.	Azote.
Gelatin	15	6	14	2
Albumen ..	17	6	13	2
Fibrin	18	5	14	3
Nephurin . . .	1	1	2	1

The colouring matter of blood approaches albumen in many of its properties; but it seems entirely destitute of azote.

Simple substances found in animals.

The simple bodies of which the animal principles consist are the following:

- | | |
|----------------|--------------------|
| 1. Azote. | 8. Soda. |
| 2. Carbon. | 9. Potash. |
| 3. Hydrogen. | 10. Muriatic acid. |
| 4. Oxygen. | 11. Magnesia. |
| 5. Phosphorus. | 12. Silica. |
| 6. Lime. | 13. Iron. |
| 7. Sulphur. | 14. Manganese. |

Of these, magnesia and silica may in a great measure be considered as foreign bodies: for they are only found in exceedingly minute quantities. The principal elementary ingredients are the first six: animal substances may be considered as in a great measure composed of them. The first four constitute almost entirely the soft parts, and the other two form the basis of the hard parts. But we shall be able to judge of this much better after we have taken a view of the various parts of animals as they exist ready formed in the body. This shall be the subject of the next Chapter.

Chap. I.

CHAP. II.

OF THE PARTS OF ANIMALS.

THE different substances which compose the bodies of ani- Solids. mals may be arranged under the following heads:

- | | |
|---------------------------|------------------------------|
| 1. Bones and shells. | 7. Glands. |
| 2. Horns and nails. | 8. Brain and nerves. |
| 3. Muscles. | 9. Marrow. |
| 4. Skin. | 10. Hair and feathers. |
| 5. Membranes. | 11. Silk and similar bodies. |
| 6. Tendons and ligaments. | |

Besides these substances, which constitute the solid parts of the bodies of animals, there are a number of fluids, the most important of which are the *chyme* and *chyle*, liquids separated from the food in the stomach and intestines for the purpose of nourishing the animal and the *blood* into which these liquids are finally converted. The rest are known by the name of *secretions*, because they are formed or *secreted*, as the anatomists term it, from the blood. The principal animal secretions are the following:

- | | | |
|-------------------------------|----------------------------|------------------------|
| 1. Milk. | 9. Humours of the eye. | <small>Fluids.</small> |
| 2. Eggs. | 10. Mucus of the nose, &c. | |
| 3. Saliva. | 11. Sinovia. | |
| 4. Pancreatic juice. | 12. Semen. | |
| 5. Bile. | 13. Liquor of the amnios. | |
| 6. Cerumen. | 14. Poisonous secretions. | |
| 7. Tears. | 15. Air. | |
| 8. Liquor of the pericardium. | | |

Book V.

Various substances are separated either from the blood or the food, on purpose to be afterwards thrown out of the body as useless or hurtful. These are called *excretions*. The most important of them are,

1. Sweat.
2. Urine.
3. Fæces.

Besides the liquids which are secreted for the different purposes of healthy animals, there are others which make their appearance only during disease, and which may therefore be called *morbid secretions*. The most important of these are the following :

Morbid secretions.

1. Pus.
2. The liquor of dropsy.
3. The liquor of blisters.

To these we must add several solid bodies, which are occasionally formed in different cavities in consequence of the diseased action of the parts. They may be called *morbid concretions*. The most remarkable of them are the following :

Morbid concretions.

1. Salivary calculi.
2. Concretions in the lungs, liver, brain, &c.
3. Intestinal calculi.
4. Biliary calculi.
5. Urinary calculi.
6. Gouty calculi.

These different substances shall form the subjects of the following sections :

SECT. I.

OF BONES, SHELLS, AND CRUSTS.

Bones, what.

By *bones* are meant those hard, solid, well-known substances, to which the firmness, shape, and strength of animal bodies are owing ; which, in the larger animals, form as it were the ground work upon which all the rest is built. In man, in quadrupeds, and many other animals, the bones are situated below the other parts, and scarcely any of them are exposed to view ; but shell-fish and snails have a hard covering on the outside of their bodies, evidently intended for defence. As these coverings, though known by the name of *shells*, are undoubtedly of a bony matter, I shall include them in this Section. For the very same reason, it would be improper to exclude *egg-shells*, and those coverings of certain animals, the lobster for instance,

known by the name of *crusts*. The solid substances belonging to that class of bodies usually distinguished by the name *zoophites*, may be included likewise without impropriety in this Section. As these different substances, though they belong to the same genus, are notwithstanding distinguished by peculiar properties, I shall consider each set separately.

I. BONES.

The bones are the most solid part of animals. Their texture is sometimes dense, at other times cellular and porous, according to the situation of the bone. They are white, of a lamellar structure, and not flexible nor softened by heat. Their specific gravity differs in different parts. That of adults' teeth is 2.2727; the specific gravity of children's teeth is 2.0833.* Properties. It must have been always known that bones are combustile, and that when sufficiently burnt, they leave behind them a white porous substance, which is tasteless, absorbs water, and has the form of the original bone. The nature of this substance embarrassed the earlier chemists. Beccher observed, that in a very violent heat it underwent a kind of fusion, and assumed an appearance analogous to that of porcelain. It was afterwards considered as a peculiar earth, and distinguished by the name of *animal earth* or *earth of bones*. Some chemists considered it as the same with lime; but the difference between it and that earth was too well marked to allow that opinion to gain ground. It had been long known that bones yielded a gelatinous substance when boiled in water; they had even been employed in the manufacture of glue. Hence it was obvious that they contained gelatin. It had been observed, that all acids had the property of softening them by dissolving their earthy and gelatinous constituents; and that when thus softened they resembled cartilage. These facts must have been known at a very early period, and afforded chemists some data for estimating their constituents. History.

In 1771, Scheele mentioned, in his dissertation on fluor spar, that the earthy part of bones is *phosphate of lime*.† This discovery was the first and the great step towards a chemical knowledge of the composition of bones. Afterwards some new facts were made known by Berniard, Bouillon, and

* Pepys, Fox on the Teeth, p. 99.

† The discoverer of this fact was Assessor Gahn of Fahlun, who however never published it himself.

Book V.

Rouelle: but for by far the most complete analyses which have hitherto appeared we are indebted to Mr. Hatchett,* to Fourcroy and Vauquelin,† and to Berzelius.‡

Constituents.

The component parts of bones are chiefly four; namely, the *earthy salts*, *fat*, *gelatin*, and *cartilage*.

1. Earthy salts.

1. The earthy salts may be obtained either by calcining the bone to whiteness, or by steeping it for a sufficient length of time in acids. In the first case, the salts remain in the state of a brittle white substance; in the second, they are dissolved, and may be thrown down by the proper precipitants. These earthy salts are four in number: 1. Phosphate of lime, which constitutes by far the greatest part of the whole. 2. Carbonate of lime. 3. Phosphate of magnesia, discovered by Fourcroy and Vauquelin. It occurs in the bones of all the inferior animals examined by these indefatigable chemists, but could not be detected by them in human bones. Berzelius however announces its existence in human bones also. 4. Sulphate of lime, detected by Mr. Hatchett in a very minute proportion, and shown by Berzelius to be formed during the calcination.

To estimate these substances, calcined bones may be dissolved in nitric or muriatic acid. During the solution carbonic acid gas makes its escape.§ Pure ammonia dropped into the solution throws down the phosphate of lime and phosphate of magnesia in the state of a fine powder, readily soluble without effervescence in nitric and muriatic acids. Nitrate of barytes afterwards causes a small precipitate insoluble in muriatic acid, and therefore consisting of sulphate of barytes. Its weight indicates the quantity of sulphuric acid in bones, from which the sulphate of lime which they contain may be estimated.|| Carbonate of ammonia now throws down pure carbonate of lime.** These three constituents were found by Hatchett in all the bones of quadrupeds and fish which he examined. The carbonate scarcely exceeds the fifth part of the phosphate, and the proportion of sulphate is still smaller. As the proportion of phosphate of lime varies not only in different animals, but even in different bones of the same animal, and even in the same bone of different ages, no precise numbers can be given. To detect and estimate the

* Phil. Trans. 1799.

† Ann. de Chim. xlvii. 244.

‡ Afhandlingar, i. 195.

§ Hatchett, Phil. Trans. 1799, p. 327.

|| Ibid.

** Ibid.

phosphate of magnesia, Fourcroy and Vauquelin employed the following simple method: the bone, previously calcined and reduced to powder, was treated with its own weight of sulphuric acid, and the mixture left to macerate for five or six days. It was then diluted with water,edulcorated, and filtered. The filtered liquid was mixed with an excess of ammonia, which threw down a mixture of phosphate of lime and ammonio-phosphate of magnesia. This precipitate is boiled in potash, which decomposes the magnesian salt, and leaves a mixture of magnesia and phosphate of lime. This mixture is treated with boiling acetous acid, which dissolves the magnesia, and leaves the phosphate of lime. The acetate of magnesia is decomposed by carbonate of soda, and the carbonate of magnesia obtained is dissolved in sulphuric acid.*

2. The proportion of fat contained in bones is various. By breaking bones in small pieces, and boiling them for some time in water, Mr. Proust obtained their fat swimming on the surface of the liquid. It weighed, he says, one-fourth of the weight of the bones employed.† This proportion appears excessive, and can scarcely be accounted for without supposing that the fat still retained water.

3. The gelatin is separated by the same means as the fat, by breaking the bones in pieces and boiling them long enough in water. The water dissolves the gelatin, and gelatinizes when sufficiently concentrated. Hence the importance of bones in making portable soups, the basis of which is concrete gelatin; and likewise in making glue. By this process Proust obtained from powdered bones about $\frac{1}{10}$ th of their weight of gelatin. This celebrated chemist has been at considerable pains to point out the advantage of employing powdered bones in the preparation of soups and broths: and Cadet de Vaux has published a set of experiments to show, that from bones thus powdered as much good soup may be obtained as from five times their weight of meat.‡ But when Dr. Young repeated this improbable experiment it did not succeed.§

4. When bones are deprived of their gelatin by boiling them in water, and of their earthy salts by steeping them in diluted acids, there remains a soft white elastic substance, pos-

* Ann. de Chim. xvii. 247.

† Jour. de Phys. lii. 257.

‡ Jour. de Chim. iv. 95. Jour. of the Royal Institut. ii. 18.

§ Ibid. p. 25.

Book V.

sessing the figure of the bones, and known by the name of *cartilage*. From the experiments of Hatchett, it appears that this substance has the properties of coagulated albumen. Like that substance, it becomes brittle and semi-transparent when dried, is readily soluble in hot nitric acid, is converted into gelatin by the action of diluted nitric acid; for it is soluble in hot water, and gelatinises on cooling, and ammonia dissolves it and assumes a deep orange colour. Like coagulated albumen, it forms an animal soap with fixed alkalies.* Berzelius found its weight about one-third of the whole bone.

This cartilaginous substance is the portion of the bone first formed. Hence the softness of these parts at first. The phosphate of lime is afterwards gradually deposited, and gives the bone the requisite firmness. The gelatin and fat, especially the first, give the bone the requisite degree of toughness and strength; for when they are removed the bone becomes brittle. The relative proportion of phosphate of lime and cartilage differs exceedingly in different bones and in different animals.

Merat-Guillot † a good many years ago published a comparative analysis of the bones of a great number of animals. But Berzelius found these results quite inaccurate. The differences between the composition of the bones of different animals are much smaller than indicated by the table of Merat-Guillot. The following table exhibits the composition of calcined human bones according to the analysis of Berzelius:

Phosphate of lime.....	81·9
Fluate of lime	3·0
Lime	10·0
Phosphate of magnesia	1·1
Soda	2·0
Carbonic acid	2·0
	100·0 ‡

One hundred parts of fresh bones are reduced by calcination to 63 parts.

Ox bones, according to the analysis of Fourcroy and Vauquelin, are composed of

* Phil. Trans. 1800.

† Ann. de Chim. xxxiv. 71.

‡ Gehlen's Jour. Second Series, iii. 1.

Solid gelatin	51·0
Phosphate of lime	37·7
Carbonate of lime	10·0
Phosphate of magnesia	1·3
	<hr/>
	100·0 *

According to the analysis of Berzelius, ox bones are composed of

Cartilage	33·30
Phosphate of lime	55·35
Fluate of lime	3·00
Carbonate of lime	3·85
Phosphate of magnesia	2·05
Soda with a little common salt	2·45
	<hr/>
	100·00 †

From the calcined bones of horses and sheep, fowls, and fishes, Fourcroy and Vauquelin extracted about $\frac{1}{30}$ th part of phosphate of magnesia.

The only bone hitherto observed altogether destitute of ^{Enamel.} cartilage is the enamel of the teeth. When the raspings of bones are steeped in diluted acids, the cartilage alone remains undissolved. Now, when the raspings of enamel are treated in this manner, Mr. Hatchett observed that the whole was dissolved without any residuum whatever. † If we believe Fourcroy and Vauquelin, the enamel of the teeth is composed of

Phosphate of lime	72·9
Gelatin and water	27·1
	<hr/>
	100·0 §

But the most complete analysis of teeth has been made by Mr. Pepys, and his results agree exactly with those of Hatchett. He found the enamel of the teeth composed of

Phosphate of lime	78
Carbonate of lime	6
Loss and water	16
	<hr/>
	100

* Ann. de Chim. xlvij. 258.

† Djur kemien, ii. 128.

‡ Phil. Trans. 1799, p. 328.

§ Mem. de l'Institut. ii. 284.

|| Fox on the Teeth, p. 96.

Book V.

Berzelius however could not find two per cent. of combustible matter in the enamel of the teeth.*

From the analysis of Mr. Pepys, we learn that the teeth are composed as follows: †

Analysis of teeth.	Roots of the Teeth.	Teeth of Adults.	First Teeth of Children.
Phosphate of lime	58	64	62
Carbonate of lime	4	6	6
Cartilage	28	20	20
Loss	10	10	12
	— 100	— 100	— 100

From this analysis, we see that teeth contain more phosphate and less cartilage than bone, and enamel more phosphate than teeth.

Fossil
bones.

Mr. Hatchett examined likewise *fossil bones* from the rock of Gibraltar. He found them to consist of phosphate of lime without any cartilage or soft animal part. Their interstices were filled with carbonate of lime. Hence they resemble exactly bones that have been burnt. They must, then, have been acted on by some foreign agent; for putrefaction, or lying in the earth, does not soon destroy the cartilaginous part of bones. On putting a human os humeri, brought from Hythe in Kent, and said to have been taken from a Saxon tomb, into muriatic acid, he found the cartilaginous residuum nearly as complete as in a recent bone. From the recent experiments of Morichini, ‡ Klapproth, § and Fourcroy and Vauquelin, || we learn that fossil ivory and teeth of animals frequently contain a portion of fluuate of lime. Morichini and Gay-Lussac endeavoured to prove that this salt existed even in recent ivory, and that the enamel of the teeth was almost entirely composed of it.** But the experiments of Wolaston, Brande, †† Fourcroy and Vauquelin, ‡‡ have shown that there does not exist any sensible portion of fluoric acid in these substances while recent. Berzelius, however, has announced that he separated 3 *per cent.* of fluuate of lime from fresh teeth, and that he has detected it also in bones, nearly in the same proportion. He even affirms that it exists in

* View of Animal Chemistry, p. 73.

† Ibid.

‡ Phil. Mag. xxiii. 265.

§ Gehlen's Jour. iii. 625.

|| Phil. Mag. xxv. 265.

** Ibid. xxiii. 265.

†† Nicholson's Jour. xiii. 216.

‡‡ Phil. Mag. xxv. 266.

urine.* And as his experiments appear to have been made with considerable care, it is probable that the results which he obtained are correct.†

Chap. II.

II. SHELLS.

Under the name of *shells* I include all the bony coverings of the different species of shell fish. For almost all the knowledge of these substances that we possess we are indebted to the important dissertations of Mr. Hatchett. A few detached facts, indeed, had been observed by other chemists; but his experiments gave us a systematic view of the constituents of the whole class.

Shells, like bones, consist of calcareous salts united to a soft animal matter; but in them the lime is united chiefly to carbonic acid, whereas in bones it is united to phosphoric acid. In shells the predominating ingredient is carbonate of lime; whereas in bones it is phosphate of lime. This constitutes the characteristic difference in their composition.

Mr. Hatchett has divided shells into two classes. The first Divided into two classes. are usually of a compact texture, resemble porcelain, and have an enamelled surface, often finely variegated. The shells belonging to this class have been distinguished by the name of *porcelaneous shells*. To this class belong the various species of *voluta*, *cypræa*, &c. The shells belonging to the second class are usually covered with a strong epidermis, below which lies the shell in layers, and composed entirely of the substance well known by the name of *mother-of-pearl*.‡ They have been distinguished by the name of *mother-of-pearl shells*. The shell of the *fresh water muscle*, the *haliotis iris*, the *turbo olearius*, are examples of such shells. The shells of the first of these classes contain a very small portion of soft animal matter; those of the second contain a very large proportion. Hence we see that they are extremely different in their composition.

1. Porcelaneous shells, when exposed to a red heat, crackle and lose the colour of their enamelled surface. They emit no smoke or smell; their figure continues unaltered, their colour becomes opaque white, tinged partially with pale-grey. They 1. Porcelaneous.

* Gehlen's Jour. vi. 591.

† His experiments will be found in Gehlen's Jour. Second Series iii. 2. There is an abstract of them in Ann. de Chim. lxi. 256.

‡ Herissant, Mem. Par. 1766, p. 22. Hatchett, Phil. Trans. 1 p. 317.

Book V.

dissolve when fresh with effervescence in acids, and without leaving any residue; but if they have been burnt there remains always a little charcoal. The solution is transparent, gives no precipitate with ammonia or acetate of lead; of course it contains no sensible portion of phosphate or sulphate of lime. Carbonate of ammonia throws down an abundant precipitate of carbonate of lime. Porcelaneous shells, then, consist of carbonate of lime cemented together by a small portion of an animal matter, which is soluble in acids, and therefore resembles gelatin.*

Composed
of carbonate
of lime and
a little jelly.

Patellæ.

Patellæ from Madeira, examined by Mr. Hatchett, were found, like the porcelaneous shells, to consist of carbonate of lime; but when exposed to a red heat they emitted a smell like horn; and when dissolved in acids a semiliquid gelatinous matter was left behind. They contain therefore less carbonate of lime and more gelatin, which is of a more viscid nature than that of porcelaneous shells.

2. Mother-of-pearl.

2. Mother-of-pearl shells when exposed to a red heat crackle, blacken, and emit a strong fetid odour. They exfoliate, and become partly dark-grey, partly a fine white. When immersed in acids they effervesce at first strongly; but gradually more and more feebly, till at last the emission of air-bubbles is scarcely perceptible. The acids take up only lime, and leave a number of thin membranous substances, which still retain the form of the shell. From Mr. Hatchett's experiments, we learn that these membranes have the properties of coagulated albumen. Mother-of-pearl shells, then, are composed of alternate layers of coagulated albumen and carbonate of lime, beginning with the epidermis, and ending with the last-formed membrane. The animals which inhabit these shells increase their habitation by the addition of a stratum of carbonate of lime, secured by a new membrane; and as every additional stratum exceeds in extent that which was previously formed, the shell becomes stronger as it becomes larger.†

Of mem-
brane and
carbonate
of lime.

Though this in general is the structure of the mother-of-pearl shells, yet there is a considerable difference between the proportion of the component parts and the consistency of the albuminous part. Some of them, as the common oyster-shell, approach nearly to the patellæ, the albuminous portion being small, and its consistence nearly gelatinous; while in others, as the *haliotis iris*, the *turbo olearius*, the real mother-of-pearl,

* Hatchett, Phil. Trans. 1799, p. 317.

† Ibid.

and a species of fresh water muscle analysed by Hatchett, the membranes are distinct, thin, compact, and semitransparent.* Chap. II.
 Mother-of-pearl contains

Carbonate of lime	66
Membrane	34
		100†

Pearl, a well known globular concretion which is formed in Pearl. some of these shells, resembles them exactly in its structure and composition. It is a beautiful substance of a bluish-white colour, iridescent, and brilliant. It is composed of concentric and alternate coats of thin membrane and carbonate of lime. The iridescence is obviously the consequence of the lamellated structure.‡

Mr. Hatchett found that what is called the *bone* of the *cuttle fish* is exactly similar to mother-of-pearl shells in its composition. Bone of the cuttle fish.

From the comparative analysis of shells and bones, Mr. Hatchett was induced to compare them together, and has shown that porcelaneous shells bear a striking resemblance to enamel of teeth; while mother-of-pearl shells bear the same resemblance to the substance of teeth or bone: with this difference, that in enamel and bone the earthy salt is phosphate of lime, whereas in shells it is pure carbonate of lime.

III. CRUSTS.

By crusts we understand those bony coverings of which the whole external surface of crabs, lobsters, and other similar sea animals are composed. Mr. Hatchett found them composed of three ingredients: 1. A cartilaginous substance, possessing the properties of coagulated albumen; 2. Carbonate of lime; 3. Phosphate of lime. By the presence of this last substance they are essentially distinguished from shells, and by the great excess of carbonate of lime above the phosphate they are equally distinguished from bones. Thus the crusts lie intermediate Constitu-
ents. between shells and Between
shells and
bones. bones, partaking of the properties and constitution of each. The shells of the eggs of fowls must be referred likewise to the class of crusts, since they contain both

* Hatchett, Phil. Trans. 1799, p. 317.

† Merat-Guillot, Ann. de Chim. xxiv. 71.

‡ Hatchett, Phil. Trans. 1799.

Book V. phosphate and carbonate of lime. The animal cement in them, however, is much smaller in quantity. From the experiments of Berniard and Hatchett, it is extremely probable that the shells of snails are composed likewise of the same ingredients, phosphate of lime having been detected in them by these chemists.

Mr. Hatchett examined the crusts of crabs, lobsters, prawns, and cray fish. When immersed in diluted nitric acid these crusts effervesced a little, and gradually assumed the form of a yellowish-white soft elastic cartilage, retaining the form of the crust. The solution yielded a precipitate to acetate of lead, and ammonia threw down phosphate of lime. Carbonate of ammonia threw down a much more copious precipitate of carbonate of lime. On examining the crust which covers different species of *echini*, Mr. Hatchett found it to correspond with the other crusts in its composition. Some species of starfish yielded a phosphate of lime, others none: hence the covering of that genus of animals seems to be intermediate between shell and crust.

Analysis. With these observations of Mr. Hatchett the analysis of Merat-Guillot corresponds. From lobster crust he obtained

Carbonate of lime	60
Phosphate of lime	14
Cartilage	26
	<hr/>
	100*

One hundred parts of cray fish crust contain

Carbonate of lime	60
Phosphate of lime	12
Cartilage	28
	<hr/>
	100†

Egg-shells. One hundred parts of hen's egg-shells contain

Carbonate of lime	89·6
Phosphate of lime	5·7
Animal matter	4·7
	<hr/>
	100·0‡

* Ann. de Chim. xxxiv. 71.

† Merat-Guillot, Ann. de Chim. xxxiv. 71.

‡ Vauquelin, ibid. xxix. 6.

Dr. John found crab crust composed as follows:

Chap. II.

Carbonate of lime, about 2
 Membrane 1
 Phosphate of lime.
 Trace of a muriate.
 Pigment.*

IV. ZOOPHITES.

Many of the substances called *zoophites* have the hardness and appearance of shell or bone, and may therefore be included among them without impropriety. Others, indeed, are soft, and belong rather to the class of membrane or horn; but of these very few only have been examined. Indeed scarcely any chemical experiments have been published on these interesting subjects, if we except the dissertation by Hatchett, in the Philosophical Transactions for 1800, which has been so often quoted. From this dissertation, and from a few experiments of Merat-Guillot, we learn that the hard zoophites are composed chiefly of three ingredients: 1. An animal substance of the nature of coagulated albumen, varying in consistency; sometimes being gelatinous and almost liquid, at others of the consistency of cartilage. 2. Carbonate of lime. 3. Phosphate of lime.

Composition.

In some zoophites the animal matter is very scanty, and phosphate of lime wanting altogether; in others the animal matter is abundant, and the earthy salt pure carbonate of lime; while in others the animal matter is abundant, and the hardening salt a mixture of carbonate of lime and phosphate of lime; and there is a fourth class almost destitute of earthy salts altogether. Thus there are four classes of zoophites; the first resemble porcelaneous shells, the second resemble mother-of-pearl shell, the third resemble crusts, and the fourth horn.

Divided into four classes.

1. When the *madrepora virginea* is immersed in diluted nitric acid it effervesces strongly, and is soon dissolved. A few gelatinous particles float in the solution, which is otherwise transparent and colourless. Ammonia precipitates nothing; but its carbonate throws down abundance of carbonate of lime. It is composed, then, of carbonate of lime and a little animal matter. The following zoophites yield nearly the same results:

1. Composed of carbonate of lime and jelly.

* Chemische Untersuchungen, ii. 54.

Madrepora muricata.
 ——— labyrinthica.
 Millepora cerulea.
 ——— alcornis.
 Tubipora musica.

2. Of carbonate of lime and membrane.

2. When the *madrepora ramea* is plunged into weak nitric acid, an effervescence is equally produced; but after all the soluble part is taken up, there remains a membrane which retains completely the original shape of the madrepore. The substance taken up is pure lime. Hence this madrepore is composed of carbonate of lime, and a membranaceous substance which, as in mother-of-pearl shells, retains the figure of the madrepore. The following zoophites yield nearly the same results:

Madrepora fascicularis.
 Millepora cellulosa.
 ——— fascialis.
 ——— truncata.
 Iris hippuris.

The following substances, analysed by Merat-Guillot, belong to this class from their composition, though it is difficult to say what are the species of zoophites which were analysed. By *red coral* he probably meant the *gorgonia nobilis*, though that substance is known, from Hatchett's analysis, to contain also some phosphate:

	White Coral.	Red Coral.	Articulated Coraline.
Carbonate of lime	50	53·5	49
Animal matter	50	46·5	51
	100	100·0	100*

3. Of membrane, carbonate, and phosphate.

3. When the *madrepora polymorpha* is steeped in weak nitric acid, its shape continues unchanged; there remains a tough membranaceous substance of a white colour and opaque, filled with a transparent jelly. The acid solution yields a slight precipitate of phosphate of lime when treated with ammonia, and carbonate of ammonia throws down a copious precipitate of carbonate of lime. It is composed therefore of animal substance, partly in the state of jelly, partly in that of membrane, and hardened by carbonate of lime together with a little phosphate of lime.

* Merat-Guillot, Ann. de Chim. xxxiv. 71.

Flustra foliacea, treated in the same manner, left a finely reticulated membrane, which possesses the properties of coagulated albumen. The solution contained a little phosphate of lime, and yielded abundance of carbonate of lime when treated with the alkaline carbonates. The *coralina opuntia*, treated in the same manner, yielded the same constituents; with this difference, that no phosphate of lime could be detected in the fresh coraline, but the solution of burnt coraline yielded traces of it. The *iris ochracea* exhibits the same phenomena, and is formed of the same constituents. When dissolved in weak nitric acid, its colouring matter falls in the state of a fine red powder, neither soluble in nitric nor muriatic acid, nor changed by them: whereas the tinging matter of the *tubipora musica* is destroyed by these acids. The branches of this iris are divided by a series of knots. These knots are cartilaginous bodies connected together by a membranous coat. Within this coat there is a conical cavity filled with the earthy or coraline matter; so that, in the recent state, the branches of the iris are capable of considerable motion, the knots answering the purpose of joints.

When the *gorgonia nobilis*, or red coral, is immersed in weak nitric acid, its colouring matter is destroyed, an effervescence takes place, and the calcareous part is dissolved. There remains an external tubulated membrane of a yellow colour, inclosing a transparent gelatinous substance. The solution yields only carbonate of lime: but when red coral is heated to redness, and then dissolved, the solution yields a little phosphate of lime also. Red coral is composed of two parts: an internal stem, composed of gelatinous matter and carbonate of lime; and an external covering or cortex, consisting of membrane hardened by the calcareous salts, and both coloured by some unknown substance.

The *gorgonia ceratophyta* likewise consist of a stem and cortex. The stem is composed of cartilage, hardened chiefly by phosphate of lime; and containing little carbonate of lime; but the cortex consists of membrane, hardened almost entirely by carbonate of lime. The *gorgonia flabellum* is almost exactly similar. The cortex of the *gorgonia suberosa* yielded gelatine to boiling water; when steeped in acids, it left a soft yellowish membrane, and the acid had taken up a little phosphate and a large portion of carbonate of lime. The stem contained scarcely any earthy salt. When burnt, it left a little phosphate of lime. To water it yielded a little gelatin; but it consisted chiefly of a horny substance, analogous to

Book V. } coagulated albumen. The *gorgonia setosa* and *pectinata* exhibited the same phenomena.

4. Chiefly
of mem-
brane.

4. *Gorgonia antiphates*, like the other species of gorgonia, has a horny stem, but it is destitute of a cortex. To boiling water it gives out some gelatin. When steeped in nitric acid it becomes soft, and exhibits concentric coats of thin opaque brown membranes, of a ligneous aspect. It contains no earthy salt. With potash it forms an animal soap, and possesses nearly the properties of horn.

The stems of the *gorgonia umbraculum* and *verrucosa* resemble that of the *gorgonia antiphates*; but these are both provided with a cortex composed of membrane and carbonate of lime.

The *antiphates ulex* and *myriophyla* resemble almost exactly the horny stem of the *gorgonia antiphates*.

Mr. Hatchett analysed many species of sponges, but found them all similar in their composition. The *spongia cancellata*, *oculata*, *infundibuliformis*, *palmata*, and *officinalis*, may be mentioned as specimens. They consist of gelatin, which they gradually give out to water, and a thin brittle membranous substance, which possesses the properties of coagulated albumen. Hence the effect of acids and alkalies on them.

The *alcyonium ficus*, *asbestinum*, and *arboreum*, resemble very much the cortex of the *gorgonia suberosa* in their composition. They yield a little gelatin to water. In nitric acid they soften, and appear membranous. The acid takes up the carbonate of lime, and likewise a little phosphate, at least when the substance has been previously heated to redness.

Such is an abstract of Mr. Hatchett's instructive analysis of the zoophites.

SECT. II.

OF HORNS, NAILS, AND SCALES.

IN the last Section I treated of those hard parts of animals which were inflexible and incapable of being softened by heat, and which contained a great portion of calcareous salts; but there is another set of hard parts which possess considerable elasticity, which are softened by heat, and which contain but a very small portion of calcareous matter. This set comprehends the substances well known under the names of *horn*, *nails*, and *scales*. The fourth class of zoophites ought, in strict propriety, to be joined to them; but in the present state of our

knowledge I thought it better to give a view of all that is known of the zoophites in one place. Chap. II.

Almost all that we know of these bodies has been discovered by Mr. Hatchett. Many experiments, indeed, had been made by preceding chemists; but not of such a nature as to throw much light upon their composition.

1. *Horns* are well known substances that are attached to the foreheads of oxen, sheep, and various other animals. They are not very hard, as they may be easily cut with a knife or rasped with a file; but they are so tough as not to be capable of being pounded in a mortar. When in thin plates, they have a degree of transparency, and have been sometimes substituted for glass in windows. When heated sufficiently they become very soft and flexible, so that their shape may be altered considerably. Hence they may be gradually squeezed into a mould, and wrought into various forms, as is well known. When strongly heated in a Papin's digester, they are said to be converted into a gelatinous mass, which possesses the properties of gelatin.

The quantity of earthy matter which they contain is exceedingly small. Mr. Hatchett burnt 500 grains of ox horn. The residuum was only 1.5 grain, and not the half of this was phosphate of lime. Seventy-eight grains of the horn of the chamois left only 0.5 of residue, of which less than the half was phosphate of lime.* They consist chiefly of a membranous substance, which possesses the properties of coagulated albumen; and probably they contain also a little gelatin. Hence we see the reason of the products that are obtained when these substances are subjected to distillation.

Dr. John subjected the horns of black cattle to analysis, and obtained the following constituents:

Indurated albumen	90
Gelatinous albumen with osmazome	8
Lactic acid	} .. 1
Lactate of potash	
Sulphate, muriate, and phosphate of potash	
Phosphate of lime	
Trace of oxide of iron	}
Ammoniacal salt	
Fat	1
	100†

* Phil. Trans. 1799, p. 532.

† Annals of Philosophy, ix. 55.

Book V.

The horns of the buck and hart must, however, be excepted. From the experiments of Scheele and Rouelle, together with those of Hatchett, we know that these substances possess exactly the properties of bone, and are composed of the same constituents, excepting only that the proportion of cartilage is greater. They are intermediate, then, between bone and horn. The same remarks apply to a fossil horn found in France, and lately analysed by Braconnot. He found it composed of

Siliceous sand	4·0
Gelatin	4·6
Bitumen	4·4
Oxide of iron	0·5
Alumina	0·7
Phosphate of magnesia	1·0
Water	11·0
Carbonate of lime	4·5
Phosphate of lime	69·3

 100·0*

Nails.

2. The *nails*, which cover the extremities of the fingers, are attached to the epidermis, and come off along with it. Mr. Hatchett has ascertained that they are composed chiefly of a membranous substance, which possesses the properties of coagulated albumen. They seem to contain also a little phosphate of lime. Water softens but does not dissolve them; but they are readily dissolved and decomposed by concentrated acids and alkalies. Hence it appears that nails agree with horn in their nature and composition. Under the head of nails must be comprehended the talons and claws of the inferior animals, and likewise their hoofs, which differ in no respect from horn.

Tortoise-shell.

The substance called tortoise-shell is very different from shells in its composition, and approaches much nearer to the nature of *nail*; for that reason I have placed it here. When long macerated in nitric acid, it softens, and appears to be composed of membranes laid over each other, and possessing the properties of coagulated albumen. When burnt, 500 grains of it yield three of earthy matter, consisting of phosphate of lime and soda, with a little iron.†

* Gehlen's Jour. Second Series, iii. 49.

† Hatchett, Phil. Trans. 1799, p. 332.

3. The *scales* of animals are of two kinds; some, as those of serpents and other amphibious animals, have a striking resemblance to horn; while those of fish bear a greater resemblance to mother-of-pearl. The composition of these two kinds of shells is very different.

Chap. II.
Scales.

The scales of fish, as had been observed by Lewenhoeck, are composed of different membranous laminae. When immersed for four or five hours in nitric acid, they become transparent and perfectly membranaceous. The acid, when saturated with ammonia, gives a copious precipitate of phosphate of lime.* Hence they are composed of alternate layers of membrane and phosphate of lime. To this structure they owe their brilliancy. Mr. Hatchett found the spicula of the shark's skin to be similar in its composition, but the skin itself yielded no phosphate of lime.

1. Of fish.
Composed of membrane and phosphate.

The horny scales of serpents, on the other hand, are composed alone of a horny membrane, and are destitute of phosphate of lime. They yield, when boiled, but slight traces of gelatin; the horn-like crusts which cover certain insects and other animals appear, from Mr. Hatchett's experiments, to be nearly similar in their composition and nature.

2. Of serpents.

Like horn.

Thus it appears that these substances bear a striking resemblance to each other, being composed of a membrane which Hatchett considers as coagulated albumen. Vauquelin, however, who affirms that they dissolve in water, provided the temperature be raised sufficiently in a digester above the boiling point, considers them, on that account, rather as a species of concrete mucus than as coagulated albumen.†

SECT. III.

OF THE MUSCLES OF ANIMALS.

AFTER the hard parts of animals have been examined, it remains for us to consider the composition of the *soft* parts. Of these, the muscles naturally claim our attention, in the first place, as being the most important.

The muscular parts of animals are known in common language by the name of *flesh*. They constitute a considerable portion of the food of man.

* Hatchett, Phil. Trans. 1799, p. 332.

† Nicholson's Jour. xv. 147.

Book V.

History.

Muscular flesh is composed of a great number of fibres or threads, commonly of a reddish or whitish colour; but its appearance is too well known to require any description. Hitherto it has not been subjected to any accurate chemical analysis. Mr. Thouvenel, indeed, has published a valuable dissertation on the subject; and it is to him that we are indebted for almost all the facts known concerning the composition of muscle. Some curious experiments have also been made by Fourcroy and Vauquelin; Mr. Hatchett has not neglected this part of the subject in his interesting dissertations on animal substances; and Berzelius has examined muscles as well as every other part of the animal body.

It is scarcely possible to separate the muscle from all the other bodies with which it is mixed. A quantity of fat often adheres to it closely; blood pervades the whole of it: and every fibre is enveloped in a particular thin membranous matter, which anatomists distinguish by the name of *cellular* substance. The analysis of the muscle, then, cannot be supposed to exhibit an accurate view of the composition of pure muscular fibres, but only of muscular fibre not perfectly separated from other substances.

Action of
cold water.

1. When a muscle is cut in small pieces, and well washed with water, the blood and other liquids contained in it are separated, and part of the muscular substance also is dissolved. The muscle, by this process, is converted into a white fibrous substance, still retaining the form of the original body. The water assumes the colour which results from mixing water with some blood. When heated it coagulates; brown flakes swim on the surface, consisting of albumen combined with the colouring matter of the blood: some fibrin likewise precipitates. If the evaporation be continued, more albumen precipitates, and at last the whole assumes the form of a jelly. When evaporated to dryness, and treated with alcohol, the gelatin, thus formed, together with a little phosphate of soda and ammonia, remains undissolved; but the alcohol dissolves a peculiar *extractive* matter, first observed by Thouvenel. This matter may be obtained by evaporating the alcohol to dryness. It has a reddish-brown colour, a strong acrid taste, and aromatic odour. It is soluble both in water and alcohol; and when its watery solution is very much concentrated, it assumes an acid and bitter taste. It swells upon hot coals and melts, emitting an acid and penetrating smell. It attracts moisture from the air, and forms a saline

efflorescence. In a hot atmosphere it becomes sour and putrefies. When distilled, it yields an acid, partly combined with ammonia. According to Berzelius, it consists of lactic acid and lactate of soda.

2. If the muscle, after being thus treated with cold water, be boiled for a sufficient time in water, an additional portion of the same substances are separated from it. Some albumen collects on the surface in the form of scum, accompanied with some melted fat. The water, when sufficiently concentrated by evaporation, assumes the form of a jelly. When evaporated to dryness, and treated with alcohol, the gelatin and phosphoric salts remain, while the extractive matter of Thouvenel is dissolved, and may be obtained by evaporating to dryness. It is by this process that it is procured in a sufficient quantity for examination, cold water abstracting only a very small portion from the muscle.

3. The muscle, thus treated with water, is left in the state of grey fibres, insoluble in water, and becoming brittle when dry. This substance possesses all the properties of *fibrin*.

4. From these facts, ascertained by Thouvenel and Fourcroy,* it appears that the muscles are composed chiefly of fibrin, to which they owe their fibrous structure and their form, and that they contain also

- | | |
|----------------|-------------------------------------|
| 2. Albumen. | 5. Phosphate of soda. |
| 3. Gelatin. | 6. Phosphate of ammonia. |
| 4. Extractive. | 7. Phosph. of lime and carb. of do. |

For the discovery of the last ingredients we are indebted to Mr. Hatchett, who found that 500 parts of beef muscle left, after combustion, a residuum of 25.6 parts, consisting chiefly of these salts. When muscles are long boiled in water, Mr. Hatchett found that the greater part of the phosphate of lime, as well as of the alkaline phosphates, was dissolved; for the muscle, after this treatment, when dissolved in nitric acid, yielded scarcely any phosphate of lime; whereas if it was dissolved directly in nitric acid, a precipitate of phosphate of lime was thrown down by ammonia. Hence it would appear, either that the phosphate of lime is united to gelatin, or that it is rendered soluble by means of it. The carbonate of lime

* Fourcroy, ix. 242.

Book V.

still remains after the action of water, and is converted into oxalate when the muscle is treated with nitric acid.

Fourcroy and Vauquelin have likewise detected the presence of potash and sulphuric acid in muscle. To these chemists we are indebted for a very curious set of experiments on the action of nitric acid on muscular fibres. They found that 4.826 ounces of muscular fibre, mixed with its own weight of nitric acid, of the specific gravity 1.284, and heated till it boiled gently, gave 116 cubic inches of gas, composed of nine parts by bulk of azote and one part of carbonic acid. What remained in the retort consisted of three distinct substances: 1. A greasy substance of a yellow colour, which floated on the surface; 2. A yellow liquid; 3. A solid matter which still retained its fibrous appearance. This last was separated by means of alcohol into two distinct substances; namely, a fatty matter taken up by the alcohol, and a yellow substance insoluble in that liquid, which possessed the properties of an acid, and to which they gave the name of *yellow acid*, and of which the properties have been detailed in the last chapter when treating of fibrin. The yellow liquid contained oxalic and malic acid, probably formed from the cellular substance, a portion of yellow acid somewhat altered, and some bitter principle.*

5. Berthollet boiled the muscle of beef in water till the liquid ceased to be precipitated by infusion of nutgalls. He then suspended the beef in a glass cylinder filled with common air, and inverted over water. The oxygen of the air was changed into carbonic acid, the beef became putrid, and, when boiled again in water, gave out gelatin. By repeating this process a number of times the beef acquired the taste and smell of old cheese. Very little ammonia was given out during the process.†

Constituents
of muscles.

6. Berzelius subjected a portion of muscle to an elaborate analysis. About three-fourths of it were liquids. The following were the constituents which he obtained:

* Nicholson's Journal, xiii. 240.

† Mem. d'Arcueil, i. 335.

I. Solid matters.

Fibrin, vessels and nerves	15·8
Cellular matter dissolved by boiling ..	1·9
	<hr/>
	17·7 .. 17·7

II. Liquid bodies.

Muriate and lactate of soda	1·80
Albumen and colouring matter of the blood	2·20
Phosphate of soda	0·90
Extract	0·15
Albumen holding in solution phos- phate of lime	0·08
Water and loss	77·17
	<hr/>
	82·30 .. 82·3

100·0 *

7. The muscles of different animals differ exceedingly from each other in their appearance and properties, at least as articles of food; but we know little of their chemical differences. The observations of Thouvenel alone were directed to that object, and they are imperfect. The flesh of the *ox* contains, according to him, the greatest quantity of insoluble matter, and leaves the greatest residuum when dried; the flesh of the *culf* is more aqueous and mucous: the land and water *turtle* yields more matter to water than the muscle of the *ox*; but Thouvenel ascribes the difference to foreign bodies, as ligaments, &c. mixed with the muscle of the *turtle*: *snails* yield to water a quantity of matter intermediate between that given by beef and veal: with them the muscles of *frogs*, *cray fish*, and *vipers*, agree nearly in this respect; but the muscles of fresh water fish, notwithstanding their softness, yield a considerably smaller proportion.†

Different
kinds of
muscles.

8. When meat is boiled, it is obvious that the gelatin, the extractive, and a portion of the salts, will be separated, while the coagulated albumen and fibrin will remain in a solid state. Hence the flavour and the nourishing nature of soups derived from the extractive and gelatin. When meat is roasted, on

* Djur kemien, ii. 174.

† Some experiments on the gizzards or muscular stomachs of fowls have been published by Bouillon La Grange. See Nicholson's Jour. xiii. 206.

Book V.

the other hand, all these substances continue in it, and the taste and odour of the extractive is greatly heightened by the action of the fire. Hence the superior flavour of roasted meat. Fourcroy supposes that the brown crust which forms on roasted meat is composed entirely of the extractive.

SECT. IV.

OF THE SKIN.

THE skin is that strong thick covering which envelopes the whole external surface of animals. It is composed chiefly of two parts: a thin white elastic layer on the outside, which is called *epidermis* or *cuticle*; and a much thicker layer, composed of a great many fibres closely interwoven, and disposed in different directions; this is called the *cutis*, or *true skin*. The *epidermis* is that part of the skin which is raised in blisters.

Epidermis.

1. The epidermis is easily separated from the cutis by maceration in hot water. It possesses a very great degree of elasticity.

It is totally insoluble in water and in alcohol. Pure fixed alkalies dissolve it completely, as does lime likewise, though slowly.* Sulphuric and muriatic acids do not dissolve it, at least they have no sensible action on it for a considerable time; but nitric acid soon deprives it of its elasticity, and causes it to fall to pieces.† It is well known that the living epidermis is tinged yellow almost instantaneously by nitric acid; but this effect does not take place, at least so speedily, when the dead cuticle is plunged in nitric acid altogether.‡

If the cuticle be tinged with nitric acid, the application of ammonia to it is well known to give it instantaneously a deep-orange colour. Now, as Hatchett has shown that this change is also produced upon coagulated albumen in the same circumstances, and as the epidermis resembles that substance in all the properties above detailed, it can scarcely be doubted that it is any thing else than a peculiar modification of coagulated albumen.

* Chaptal, Ann. de Chim. xxii. 221. But Davy could obtain no solution.

† Cruikshanks on Insensible Perspiration, p. 32.

‡ Ibid.

Dr. John found the constituents of the epidermis of the foot, as follows :

Indurated albumen	93 to 95
Mucus, with a trace of animal matter	5
Laetic acid	}
Laetate of potash	
Phosphate of potash	
Muriate of potash	
Sulphate of lime	
Ammoniacal salt	
Phosphate of lime	
Manganese? and iron	1
Soft fat	0.05 *

2. The cutis is a thick dense membrane, composed of fibres Cutis. interwoven like the texture of a hat. When it is macerated for some hours in water, and agitation and pressure employed to accelerate the effect, the blood, and all the extraneous matter with which it was loaded, are separated from it, but its texture remains unaltered. On evaporating the water employed, a small quantity of gelatin may be obtained. No subsequent maceration in cold water has any farther effect. When distilled it yields the same products as fibrin. The concentrated alkalis dissolve it, converting it into oil and ammonia. Weak acids soften it, render it transparent, and at last dissolve it. Nitric acid converts it into oxalic acid and fat, while, at the same time, azotic gas and prussic acid are emitted.† When heated it contracts, and then swells, exhales a fetid odour, and leaves a dense charcoal, difficult to incinerate. By spontaneous decomposition in water or moist earth, it is converted into a fatty matter and into ammonia, which compose a kind of soap.‡ When allowed to remain long in water, it softens and putrefies, being converted into a kind of jelly. When long boiled in water it becomes gelatinous, and dissolves completely, constituting a viscid liquor, which, by proper evaporation, is converted into glue. Hence the cutis of animals is commonly employed in the manufacture of glue.

From these facts the cutis appears to be a peculiar modification of gelatin enabled to resist the action of water, partly A species of gelatin.

* Annals of Philosophy, ix. 55.
 † Fourcroy, ix. 251.

‡ Ibid.

Book V.

by the compactness of its texture, and partly by the viscosity of the gelatin of which it is formed; for those skins which dissolve most readily in boiling water afford the worst glue. Mr. Hatchett has observed that the viscosity of the gelatin obtained from skins is nearly inversely as their flexibility, the supplest hides always yielding the weakest glue; but this glue is very soon obtained from them by hot water. The skin of the eel is very flexible, and affords very readily a great proportion of gelatin. The skin of the shark also readily yields abundance of gelatin; and the same remark applies to the skins of the hare, rabbit, calf, and ox; the difficulty of obtaining the glue and its goodness always increasing with the toughness of the hide. The hide of the rhinoceros, which is exceedingly strong and tough, far surpasses the rest in the difficulty of solution and in the goodness of its glue. When skins are boiled, they gradually swell and assume the appearance of horn: then they dissolve slowly.

Leather.

3. It is from the skin or cutis of animals that leather is formed; and the goodness of the leather, or at least its strength, depends in some measure, on the toughness of the hides. Those easily soluble, as seal-skins, afford a weaker leather than those which are more difficultly soluble in water. The process by which the skins of animals are converted into leather is called *tanning*. It seems to have been known and practised in the earliest ages; but its nature was totally unknown till after the discovery of the *tanning principle* by Seguin. That chemist ascertained that leather is a compound of *tannin* and *skin*; that it is to the tannin that leather owes its insolubility and its power of resisting putrefaction. The subject engaged the attention of Davy, who examined it with his usual ingenuity, and added several important facts to our former knowledge.

Preparation
of the skins.

When skins are to be tanned, the first step of the process is to deprive them of their hair and cuticle. This is either done by steeping them in water till they begin to putrefy, or by steeping them in lime and water. The lime seems to combine with the cuticle, and to render it brittle and easily detachable from the hide. It produces the same effect upon the hair and the matter at its root.* When the hides have been steeped for a sufficient time, they are taken out, the hair, cuticle, &c. scraped off, and then they are washed in water.

* Davy, Journal of the Royal Instit. ii. 30.

After this preliminary process, the skins are subjected to different treatment according to the kind of leather which is to be made.

The large and thick hides are introduced for a short time into a strong infusion of bark. They are then said to be *coloured*. After this they are put into water slightly impregnated with sulphuric acid, or with the acid evolved during the fermentation of barley and rye. This renders them harder and denser than they were before, and fits them for forming sole leather. Davy thinks, that by this process a triple compound is formed of the *skin, tan, and acid*.*

The light skins of cows, those of calves, and all small skins, are steeped for some days in a lixivium made by the infusion of pigeons' dung in water. This lixivium is called the *grainer*. By this process they are rendered thinner and softer, and more proper for making flexible leather. Davy considers the effect of this lixivium to depend upon the fermentation which it undergoes; for dung that has undergone fermentation does not answer the purpose. †

After these preliminary processes, the skins are exposed to the action of the infusion of bark till they are converted into leather.

The infusion of oak bark contains two ingredients, namely, *tannin* and an *extractive*. The first is more soluble than the second. Hence, in saturated infusions, there is a much greater proportion of tannin than of extractive; whereas in weak infusions the extractive bears a greater proportion to the tannin. Davy considers it as doubtful whether the infusion of oak bark contains any gallic acid. This chemist has ascertained, that the hides extract both the tannin and extractive from the infusion, and leave nothing behind but pure water, provided they be employed in sufficient quantity. Hence it is obvious, that both the tannin and extractive must enter into the composition of leather. The extractive gives the hide a brown colour, but does not render it insoluble in boiling water; the tannin renders it insoluble, but its colour continues whitish. Hence it is likely that the lightest kinds of leather contain little else than tannin, while the brown kinds contain both tannin and extractive, and the new compound is leather. Hence the reason of the increase of its weight.

Process of tanning.

* Davy, Journal of the Royal Instit. ii. 31.

† Ibid. p. 32.

Book V.

Calf-skins, and those hides which are prepared by the *grainer*, are first steeped in weak infusions of oak bark, and gradually removed to stronger and stronger, till they are completely impregnated, which takes up from two to four months. As the weak infusions contain a greater proportion of extractive, the consequence of this process is, that the skin combines in the first place with a portion of it, and afterwards with the tannin. When saturated solutions of tannin are employed, the leather is formed in a much shorter time. This was the process recommended by Seguin; but it has been observed, that leather tanned in this way is more rigid and more liable to crack than leather tanned in the usual way. Hence it is likely, as Davy has observed, that the union of the extractive is requisite to form pliable and tough leather. Leather rapidly tanned must be less equable in its texture than leather slowly tanned, as the surface must be saturated with tannin before the liquid has time to penetrate deep. Davy has ascertained that skins, while tanning, seldom absorb more than $\frac{1}{3}$ of their weight of vegetable matter.

Skins intended for sole leather are generally kept from the first in an infusion preserved nearly saturated by means of the strata of bark with which they alternate. The full impregnation requires from ten to eighteen months. It is likely, from this process, that sole leather contains a greater proportion of tannin than soft leather. While drying, it is smoothed with a rolling-pin, and beat with a mallet, which must add considerably to its density.*

Rete mucosum.

4. As to the *rete mucosum*, or the mucous substance, situated between the cutis vera and epidermis, its composition cannot be determined with precision, because its quantity is too small to admit of examination. It is known that the black colour of negroes depends upon a black pigment, situated in this substance. Chlorine deprives it of its black colour, and renders it yellow. A negro, by keeping his foot for some time in water impregnated with that gas, deprived it of its colour, and rendered it nearly white; but in a few days the black colour returned again with its former intensity.† This experiment was first made by Dr. Beddoes on the fingers of a negro.‡

* See Davy on the Preparation of Skin for Tanning. Royal Instit. ii. 30.

† Fourcroy, ix. p. 259.

‡ Beddoes on Factitious Airs, p. 15.

SECT. V.

OF MEMBRANES, TENDONS, LIGAMENTS, AND GLANDS.

THESE substances have not hitherto been subjected to a rigid chemical analysis. But from the properties which have been observed, they appear to have a closer resemblance to the skin than to any other animal substance. Chap. II.

1. The membranes are thin semitransparent bodies which envelope certain parts of the body, especially the viscera; such as, the dura and pia mater, the pleura, the peritoneum, the periosteum, &c. These substances are soft and pliable; when macerated in water, they swell, and become somewhat pulpy; and by continued decoction in hot water they are almost completely dissolved, and the solution concretes into gelatin. They are convertible of course into the same substance as the cutis by decoction; hence we must consider their composition as similar. Like hides they may also be tanned and converted into leather. From the experiments of Mr. Hatchett, it appears that they contain no phosphate of lime as a constituent part, and scarcely any saline ingredients; for when calcined they leave but a very inconsiderable residuum. Thus 250 grains of hog's bladder left only 0.02 grain of residuum.* Membranes

2. The tendons are strong, pearl-coloured, brilliant bodies, which terminate the muscles, and attach them to the bones, and are known in common language by the name of *sinews*. When boiled they assume the form of a semitransparent gelatinous substance, of a pleasant taste, well known in boiled meat. If the decoction be continued they dissolve completely, and are converted into gelatin. From these facts we are authorized to conclude, that the composition of the tendons is similar to that of the membranes and cutis. Resemble skin.

3. The ligaments are strong bands which bind the bones together at the different joints: they are fibrous substances, very dense and strong, and somewhat elastic. When boiled they yield a portion of gelatine, but they resist the action of water with great obstinacy, and after a great deal of boiling retain their form and even their strength. The ligaments, then, differ essentially from the two last species. How far they resemble coagulated albumen remains to be ascertained. It is not unlikely that they will form a genus apart. Tendons.

* Phil. Trans. 1799, p. 353.

Book V.
Glands.

4. The glands are a set of bodies employed to form or to alter the different liquids which are employed for different purposes in the animal body. There are two sets of them: the *conglobate*, which are small, scattered in the course of the lymphatics; and the *conglomerate*, such as the liver, kidneys, &c. Fourcroy supposes the first of these to be composed of gelatin; but this is not very probable. The structure of the large glands has been examined by anatomists with great care; but we are still ignorant of their composition. Indeed the present state of chemistry scarcely admits of an accurate analysis of these complicated bodies.

SECT. VI.

OF THE BRAIN AND NERVES.

THE brain and nerves are the instruments of sensation, and even of motion; for an animal loses the power of moving a part the instant that the nerves which enter it are cut.

The brain and nerves have a strong resemblance to each other; and it is probable that they agree also in their composition. But hitherto no attempt has been made to analyze the nerves. The only chemists who have examined the nature of *brain* are Mr. Thouret,* Mr. Fourcroy,† and M. Vauquelin.‡

Properties
of brain.

The brain consists of two substances, which differ from each other somewhat in colour, but which, in other respects, seem to be of the same nature. The outermost matter, having some small resemblance in colour to wood-ashes, has been called the *cineritious* part; the innermost has been called the *medullary* part.

Brain has a soft feel, not unlike that of soap; its texture appears to be very close; its specific gravity is greater than that of water.

When brain is kept in close vessels so that the external air is excluded, it remains for a long time unaltered. Fourcroy filled a glass vessel almost completely with pieces of brain, and attached it to a pneumatic apparatus; a few bubbles of carbonic acid gas appeared at first, but it remained above a year without undergoing any farther change. §

* Jour. de Phys. xxxviii. 329.

† Annals of Philosophy, i. 332.

‡ Ann. de Chim. xvi. 232.

§ Ann. de Chim. xvi. 297.

This is very far from being the case with brain exposed to the atmosphere. In a few days (at the temperature of 60°) it exhales a most detestable odour, becomes acid, assumes a green colour, and very soon a great quantity of ammonia makes its appearance in it. Chap. II.

Cold water does not dissolve any part of the brain; but by trituration in a mortar it forms, with water, a whitish-coloured emulsion, which appears homogeneous, may be passed through a filter, and the brain does not precipitate by rest. When this emulsion is heated to 145°, a white coagulum is formed. The addition of a great quantity of water also causes a coagulum to appear, which swims on the surface, but the water still retains a milky colour. When sulphuric acid is dropped into the watery emulsion of brain, white flakes separate and swim on the surface, and the liquid becomes red. Nitric acid produces the same effects, only the liquid becomes yellow. Alcohol also separates a white coagulum from the emulsion, after it has been mixed with it for some hours. When nitric acid is added to the emulsion till it becomes slightly acid, a coagulum is also separated. This coagulum is of a white colour; it is insoluble in water and in alcohol. Heat softens, but does not melt it. When dried it becomes transparent, and breaks with a glassy fracture. It is therefore *albumen*.* Action of cold water,

When brain is triturated in a mortar with diluted sulphuric acid, part is dissolved; the rest may be separated, by filtration, in the form of a coagulum. The acid liquor is colourless. By evaporation, the liquid becomes black, sulphurous acid is exhaled, and crystals appear; and when evaporated to dryness, a black mass remains behind. When this mass is diluted with water, a quantity of charcoal separates, and the water remains clear. This brain is completely decomposed, a quantity of ammonia combines with the acid and forms sulphate of ammonia, while charcoal is precipitated. The water, by evaporation and treatment with alcohol, yields sulphates of ammonia and lime, phosphoric acid, and phosphates of soda and ammonia. Brain therefore contains Of sulphuric acid,

Phosphate of lime
 ————— soda
 ————— ammonia.

* Ann. de Chim. xvi. p. 283.

Book V.

Traces also of sulphate of lime can be discovered in it. The quantity of these salts is very small; altogether they do not amount to $\frac{1}{300}$ th part.*

Of nitric acid,

Diluted nitric acid, when triturated with brain, likewise dissolves a part, and coagulates the rest. The solution is transparent. When evaporated till the acid becomes concentrated carbonic acid gas and nitrous gas are disengaged; an effervescence takes place, white fumes appear, an immense quantity of ammonia is disengaged, a bulky charcoal remains mixed with a considerable quantity of oxalic acid.†

When brain is gradually evaporated to dryness by the heat of a water bath, a portion of transparent liquid separates at first from the rest, and the residuum, when nearly dry, acquires a brown colour; its weight amounts to about one-fourth of the fresh brain. It may still be formed into an emulsion with water, but very soon separates again spontaneously.

Of alcohol,

When alcohol is repeatedly boiled upon this dried residuum till it ceases to have any more action, it dissolves about five-eighths of the whole. When this alcohol cools, it deposits a yellowish white substance, composed of brilliant plates. When kneaded together by the fingers, it assumes the appearance of a ductile paste; at the temperature of boiling water it becomes soft, and when the heat is increased it blackens, exhales empyreumatic and ammoniacal fumes, and leaves behind it a charry matter.‡ When the alcohol is evaporated, it deposits a yellowish black matter, which reddens paper tinged with turnsole, and readily diffuses itself through water.§

Of potash.

Pure concentrated potash dissolves brain, disengaging a great quantity of ammonia.

The result of an elaborate analysis of brain by Vauquelin gave him the following constituents :

Constituents.

Water	80·00
White fatty matter	4·53
Reddish fatty matter	0·70
Albumen	7·00
Osmazome	1·12
Phosphorus	1·50
Acids, salts, and sulphur	5·15
	<hr/>
	100·00

* Ann. de Chim. xvi. p. 288.

† Ibid. 307.

‡ Ibid. 313.

§ Ibid. 317.

Dr. John has analyzed the brain of a considerable variety of animals. He found the brain of calves composed as follows: Chap. II.

Water	75 to 80
Peculiar albumen of brain	10
Osmazome	} 15 to 10
Fat	
Sulphur, trace	
Phosphate of lime	
Phosphate of soda	
Muriate of soda	
Sulphate of soda?	
Phosphate of iron	
Phosphate of magnesia?	
An ammoniacal salt	

He analyzed likewise the brain of fowls and of oxen, with nearly the same result.*

From the experiments of Vauquelin we learn that the medulla oblongata and the nerves are composed of the same materials as the brain.

Dr. John made separate experiments on the medullary and cineritious parts of the brain, on the cerebellum, the thalami nervorum opticorum, the medulla oblongata, and the medulla spinalis, but the results have added very little to our knowledge of this obscure department of chemistry.†

Vauquelin's analysis of the brain, though far more complete and satisfactory than that of any of his predecessors, throws no light upon the way in which the constituents are combined, or on what it is that gives to the brain its very peculiar appearance.

SECT. VII.

OF MARROW.

THE hollows of the long bones are, in living animals, filled with a peculiar species of fat matter, to which the name of *marrow* has been given. In some bones this matter is a good deal mixed with blood, and has a red colour; in others, as the thigh bones, it is purer, and has a yellow colour. Various experiments on this matter were made by the older chemists,

* Chemische Untersuch. iii. 228, and iv. 160.

† Ibid. iv. 168.

Book V.

showing it to be analogous to animal fats,* and pointing out some of its peculiarities. Berzelius has examined it in detail, and published the results of his experiments.† The marrow on which his trials were made was obtained from the thigh bone of an ox.

Action of cold water,

1. When marrow is digested in cold water it becomes lighter coloured, while the water acquires the colour which it would have received had it been digested on blood. When this water is boiled it becomes muddy, and a dark-brown matter precipitates. This matter consists of coagulated albumen mixed with some phosphate of lime, and phosphate of iron. A small portion of a yellow-coloured salt is dissolved by the action of alcohol or water. This matter separated from marrow by water, is obviously owing to the blood with which it was mixed. The quantity which Berzelius obtained from marrow amounted to $\frac{1}{100}$ th part of the whole. The portion of it dissolved by water and alcohol consisted partly of gelatin and common salt, and partly of the peculiar extractive matter obtained by Thouvenel from the muscles of animals, and described in the Third Section of this Chapter. The proportion of these substances obtained by Berzelius from marrow amounted to about $\frac{1}{300}$ th part of the whole.

Of hot water.

2. When marrow is boiled in water, the greatest part of it melts and swims upon the surface of the liquid. The water is at first muddy and milky, but becomes transparent on standing. When passed through the filter a substance is separated, which becomes greyish-green, and semitransparent when dry. More of this matter precipitates when the liquid is evaporated. When the water is evaporated to dryness, a substance is obtained of a sharp aromatic taste like the marrow of roasted meat. These two substances consist chiefly of extractive, gelatin, and a peculiar substance which approaches the nature of albumen in its properties.

3. When marrow, thus purified, is melted in water and passed through a cloth, a quantity of blood vessels and skins remain upon the cloth, amounting to about $\frac{1}{100}$ th part of the whole.

Properties of pure marrow.

4. Marrow, thus freed from its impurities, has a white colour with a shade of blue; its taste is insipid and rather sweetish. It softens by the heat of the hand, and melts when heated

* Neumann's Chemistry, p. 560.

† Gehlen's Jour. Second Series, ii. 287.

to 113° .* When cooled slowly it crystallizes in sphericles like olive oil. It burns with a flame like tallow. When distilled it gives first a transparent fluid yellowish oil, accompanied by carbonic acid gas, water, and heavy inflammable air. Afterwards there comes over a white solid oil, accompanied by a less copious evolution of gaseous bodies, and which does not become dark-coloured, as happens when tallow is distilled. This had already been observed by Neumann. This solid oil has a disagreeable smell, amounts to 0.3 of the marrow distilled, reddens vegetable blues, and when boiled in water gives out a portion of sebacic acid, which Berzelius considers as benzoic acid.

The empyreumatic oil combines readily with alkalies and their carbonates. With the latter it forms a snow-white soap, insoluble in water, though it increases in bulk when placed in contact with that liquid. It combines also with the earths, and forms soaps, likewise insoluble in water.

The water which comes over during the distillation of marrow is colourless, has a fetid and sour smell, and an empyreumatic taste. It contains a little acetic acid, empyreumatic oil, and probably benzoic acid; but exhibits no traces of ammonia.

The gaseous products amount to $\frac{1}{10}$ th of the marrow distilled. They contain no sulphur nor phosphorus, and consist of carbonic acid and heavy inflammable air, which burns with a white flame, and seems to contain oil in solution.

The charry matter in the retort amounts to 0.05 of the marrow distilled. It is dark-brown, heavy and brilliant. It is incinerated with difficulty, and leaves an ash consisting of phosphate of lime, carbonate of lime, and some soda.

5. Concentrated sulphuric acid dissolves marrow without the assistance of heat. The solution has the appearance of a brown syrup; and when the acid is diluted with water the marrow separates unaltered. When heat is applied, the acid decomposes the marrow and forms a resinous coal.

Action of acid.

Diluted nitric acid digested on marrow, in a moderate heat, renders it yellow, and gives it more consistence, and the smell of old bones. Concentrated nitric acid dissolves marrow without the assistance of heat, and the marrow is not precipitated by the addition of water.

* Berzelius says at 45° . This is 113° Fahrenheit, if Celsius' thermometer be meant, as is probable. If it be Reaumur's, the melting point of marrow is 133° Fahrenheit.

Book V.

6. Marrow combines with alkalies and forms soap. Boiling alcohol and ether dissolve a small portion of it, which precipitates again as the solution cools.

Constituents.

Marrow, from the thigh bone of an ox, was found by Berzelius to be composed of the following substances :

Pure marrow	0.96
Skins and blood vessels	0.01
Albumen	} 0.03
Gelatin	
Extractive	
Peculiar matter	
Water	} 1.00

From the preceding detail it appears, that pure marrow is a species of fixed oil, possessing peculiar properties, and approaching somewhat to butter in its nature. But it differs considerably in its appearance in different parts of the body, owing chiefly, in all probability, to a greater or smaller mixture of blood.

SECT. VIII.

OF HAIR AND FEATHERS.

THESE substances cover different parts of animals, and are obviously intended by Nature to protect them from the cold. For this, their softness and pliability, and the slowness with which they conduct heat, render them peculiarly proper.

Varieties.

1. *Hair* is usually distinguished into various kinds, according to its size and appearance. The strongest and stiffest of all is called *bristle*: of this kind is the hair on the backs of hogs. When remarkably fine, soft, and pliable, it is called *wool*; and the finest of all is known by the name of *down*. But all these varieties resemble one another very closely in their composition.

Properties.

Hair appears to be a kind of tube covered with a cuticle. Its surface is not smooth, but either covered with scales or consisting of imbricated cones. Hence the roughness of its feel, and the disposition which it has to entangle itself, which has given origin to the processes of *felting* and *fulling*. It is constantly increasing in length, being protruded from the roots, and seems at first to be soft or nearly gelatinous. From the

experiments which have been made on hair by Achard and Hatchett, it follows that it contains gelatin, to which it owes its suppleness and toughness. This substance may be separated by boiling the hair in water. When thus treated it becomes much more brittle than before. Indeed, if the process be continued long enough, the hair crumbles to pieces between the fingers. The portion insoluble in water possesses the properties of coagulated albumen.

Mr. Hatchett has concluded from his experiments, that the hair which loses its curl in moist weather, and which is the softest and most flexible, is that which yields its gelatin most easily; whereas strong and elastic hair yields it with the greatest difficulty, and in the smallest proportion. This conclusion has been confirmed by a very considerable hair-merchant in London, who assured him that the first kind of hair was much more injured by boiling than the second.

Vauquelin has published a curious set of experiments on the analysis of human hair of various colours. Though hair is insoluble in boiling water, he obtained a solution by raising the temperature of the liquid in a Papin's digester. If the heat thus produced was too great, the hair was decomposed, and ammonia, carbonic acid, and an empyreumatic oil formed. Sulphuretted hydrogen is always evolved, and its quantity increases with the heat. When hair is thus dissolved in water heated above the boiling point, the solution contains a kind of bituminous oil, which is deposited very slowly. This oil was black when the hair dissolved was black, but yellowish-red when red hair was employed.

When the solution is filtered to get rid of this oil, the liquid which passes through is nearly colourless. Copious precipitates are formed in it by the infusion of nutgalls and chlorine. Silver is blackened by it, and acetate of lead precipitated brown. Acids render it turbid, but the precipitate is redissolved by adding these liquids in excess. Though very much concentrated by evaporation, it does not concrete into a jelly.

Water containing only four per cent. of potash dissolves hair, while hydrosulphuret of ammonia is evolved. If the hair be black, a thick dark-coloured oil, with some sulphur and iron, remains undissolved; if the hair be red there remains a yellow oil, with some sulphur and an atom or two of iron. When acids are dropped into this solution, they throw down a white matter soluble in an excess of acid.

Book V.

Sulphuric and muriatic acids become red when first poured on hair, and gradually dissolve it. Nitric acid turns hair yellow and dissolves it, while an oil separates, which is red or black according to the colour of the hair dissolved. The solution yields a great deal of oxalic acid, and contains, besides, bitter principle, iron, and sulphuric acid. Chlorine first whitens hair, and then reduces it to a substance of the consistence of turpentine, and partly soluble in alcohol.

Colouring matter.

When alcohol is digested on black hair, it extracts from it two kinds of oil. The first, which is white, subsides in white shining scales as the liquor cools; the second is obtained by evaporating the alcohol. It has a greyish-green colour, and at last becomes solid. From red hair alcohol likewise separates two oils: the first white as from black hair, and the other as red as blood. When the red hair is deprived of this oil, it becomes of a chesnut colour. Hence its red colour is obviously owing to the red oil.

When hair is incinerated, it yields iron and manganese, phosphate, sulphate, and carbonate of lime, muriate of soda, and a considerable portion of silica. The ashes of red hair contain less iron and manganese: those of white hair still less; but in them we find magnesia, which is wanting in the other varieties of hair. The ashes of hair do not exceed 0.015 of the hair.

From the preceding experiments of Vauquelin, we learn that black hair is composed of the nine following substances:

Constituents.

1. An animal matter constituting the greatest part.
2. A white solid oil, small in quantity.
3. A greyish-green oil, more abundant.
4. Iron; state unknown.
5. Oxide of manganese.
6. Phosphate of lime.
7. Carbonate of lime, very scanty.
8. Silica.
9. Sulphur.

The colouring matter of hair appears from Vauquelin's experiments to be an oil. The oil is blackish-green in black hair, red in red hair, and white in white hair. Vauquelin supposes that sulphuretted iron contributes to the colour of dark hair; and ascribes to the presence of an excess of sulphur the property which white and red hair have of becoming black

with the oxides of the white metals. The sudden change of colour in hair from grief, he thinks, is owing to the evolution of an acid.* Chap. II.

Vauquelin considers the animal matter of which hair is chiefly composed as a variety of inspissated mucus; but some of its properties, especially its copious precipitation by tannin, do not well agree with that supposition. It seems to approach more closely to coagulated albumen, as Hatchett has shown.

Wool has not yet been subjected to a rigid examination; but from the experiments made on it by Berthollet, there is reason to conclude that its chemical qualities do not differ much from those of hair. When growing upon the sheep it is enveloped in a kind of soapy matter, which protects it from the attack of insects, and which is afterwards removed by scouring. Vauquelin has examined this matter, and found it to consist of the following ingredients: 1. A soap of potash; 2. Carbonate of potash; 3. A little acetate of potash; 4. Lime; 5. A very little muriate of potash; and, 6. An animal matter.† Wool.

2. *Feathers* seem to possess very nearly the same properties with hair. Mr. Hatchett has ascertained that the quill is composed chiefly of coagulated albumen. Though feathers were boiled for a long time in water, Mr. Hatchett could observe no traces of gelatin. Feathers.

SECT. IX.

OF SILK.

SILK is the production of different species of caterpillars. The *phalena bombyx* is most commonly propagated for that purpose in Europe; but the *phalena atlas* yields a greater quantity. A similar substance indeed is yielded by the greater number of the tribe of caterpillars. It is found inclosed in two small bags, from which it is protruded in fine threads to serve the insect for a covering during its chrysalis state. The webs of spiders are obviously of the same nature with silk; though their fibres, at least in this country, are finer and weaker. Reaumur and other naturalists ascertained, that the larger species of spiders spun cocoons sufficiently strong to be manufactured, and that the produce was neither inferior in Origin.
Spiders' webs.

* Nicholson's Jour. xv. 141.

† Ann. de Chim. xlvii. 267.

Book V.

beauty nor in strength to the silk of the silk-worm. Attempts were made in consequence to establish a manufactory of this new kind of silk; but it was soon found that the spiders could not be made to work in concert. They attacked and devoured each other without mercy, till the whole colony was exterminated to a single individual.

History.

The silkworm is a native of China, and feeds on the leaves of the white mulberry. That industrious nation was acquainted with the manufacture of silk from the most remote ages; but it was scarcely known in Europe before the time of Augustus. Its beauty attracted the attention of the luxurious Romans; and after the effeminate reign of Elagabalus, it became a common article of dress. It was brought from China at an enormous expence, manufactured again by the Phenicians, and sold at Rome for its weight of gold. In the reign of Justinian this commerce was interrupted by the conquests of the Scythian tribes, and all attempts to procure it failed, till two Persian monks had the address to convey some of the eggs of the insect from China to Constantinople, concealed in the hollow of a cane. They were hatched and the breed carefully propagated. This happened in 555; and some years after we find that the Greeks understood the art of procuring and manufacturing silk as well as the Orientals. Roger, king of Sicily, brought the manufacture to that island in 1130, forcibly carrying off the weavers from Greece, and settling them in Sicily. From that island the art passed into Italy, and thence into France: and the revocation of the edict of Nantz established the manufactory of silk in Britain.

Varnish.

Silk, as spun by the animal, is in the state of fine threads, varying in colour from white to reddish-yellow. It is very elastic, and has considerable strength, if we consider its small diameter. It is covered with a varnish, to which its elasticity is owing. This varnish is soluble in boiling water; but alcohol does not act upon it. Hence it has been compared to a gum, though it approaches much nearer to gelatin; since Berthollet has shown that it is precipitated by tannin and by muriate of tin. It differs, however, from gelatin in several particulars. Alum throws it down of a dirty white, sulphate of copper of a dark brown, and sulphate of iron of a brown colour.* When the water is evaporated, the varnish is obtained of a black colour, brittle, and of a shining fracture. Its

* Berthollet.

weight is nearly one-third of the raw silk from which it was extracted. It may be separated from silk by soap as well as water, and the soap leys containing it soon putrify.* Roard has published a set of experiments on this matter. It amounts to 23 per cent. of the silk. It is brittle and hard like resins; when pounded its colour is reddish-yellow. It has the taste and odour of extractive. It does not melt when heated. It dissolves in five times its weight in water. The solution has a brown colour, but when diluted with water becomes golden-yellow. When exposed to the air in that state it becomes greenish, is decomposed, and exhales the odour of putrid animal matter. The liquid froths like soap leys. Sulphuric and muriatic acids deepen its colour; nitric acid converts it into a fine yellow; sulphurous acid brightens the colour, and makes it greenish-yellow. Alcohol scarcely acts upon it. Chlorine throws down a copious white precipitate, which becomes reddish in the air, is no longer soluble in water, but is very soluble in alcohol. Alkalies produce no effect upon this substance. Infusion of nutgalls occasions a copious precipitate.† These properties show us that this substance is of a peculiar nature.

Besides the varnish, silk contains another substance, to which Resin. it owes its yellow colour. This substance possesses the properties of *resin*. It is yellow, soluble in alcohol and in a mixture of alcohol and muriatic acid. Baumé has ascertained, that by this last mixture it may be separated completely, and the silk deprived of it assumes a fine white colour.‡ This matter melts at 86°. Its colour is reddish-brown when in lumps, but greenish-yellow when divided. It has a strong smell, owing to a volatile oil which it contains. When exposed to the light it becomes white. It is insoluble in water but very soluble in alcohol, in the alkaline solutions and in soap.§

Roard has shown that raw silk contains also a portion of wax, which is dissolved by boiling alcohol, and by alkaline or soapy solutions. It is white and precipitates as the alcohol cools. Silk contains about $\frac{1}{100}$ th of its weight of wax, while the resinous colouring matter amounts to about $\frac{1}{60}$ th of its weight.||

* Berthollet.

‡ Ibid. xvii. 156.

|| Ibid.

† Ann. de Chim. lxxv. 59.

§ Roard, *ibid.* lxxv. 61.

Book V.
 Properties
 of silk.

The properties of silk itself have been but imperfectly examined. It is not acted on by water or alcohol, has no taste, and is but imperfectly combustible; though fire rapidly blackens and decomposes it. When distilled, it yields, according to Neumann, an uncommonly great proportion of ammonia.

The fixed alkalies dissolve it by the assistance of heat; and it is not unlikely that they form with it an animal soap.

It is dissolved likewise by sulphuric and muriatic acids, and by nitric acid. By the action of this last acid, Berthollet obtained from silk some oxalic acid, and a fatty matter which swam on the surface of the solution. By a similar treatment, Welter obtained fine yellow crystals, very combustible, to which he gave the name of *yellow brittle principle*.

Silk is very little susceptible of putrefaction. Dr. Wilson of Falkirk informs us, that a ribbon was found in the church-yard of that town wrapped round the bone of the arm. It was uninjured though it had lain eight years in the earth.* We know, at the same time, that when silk is kept in a damp place it *rots* (to use the common language) in a much shorter period.

Spiders'
 webs.

Cadet has published a chemical analysis of spiders' webs, the result of which, as these substances have a considerable analogy to silk, both in their properties and mode of being produced, may be mentioned here. He obtained from spiders' webs, by treating them with water and alcohol, and incineration, the following substances: 1. A brown extract soluble in water, and not altered by exposure to the air; 2. A resinous matter soluble in alcohol, and very deliquescent; 3. Alumina; 4. Sulphate of lime; 5. Carbonate of soda; 6. Muriate of soda; 7. Carbonate of lime; 8. Iron; 9. Silica.† He supposes, however, that some of the fixed matters obtained may have originated from accidental impurities mixed with the webs examined.

Having given the preceding account of the solids which compose animal bodies, I proceed next to the *chyle* and the fluid which circulates through living bodies, namely *blood*; and to the various *secretions* formed from the blood, either

* Nicholson's Jour. ii. 222.

† Ibid. xi. 294.

in order to answer some important purpose to the animal, or to be evacuated as useless, that the blood thus purified may be more proper for answering the ends for which it is destined. Many of these substances have been examined with more care by chemists than the animal solids.

SECT. X.

OF CHYLE.

THE name *chyle* is given to a liquid matter formed during the passage of the food along the intestinal canal. It is taken up by a set of vessels called *lacteals*, carried to the thoracic duct, from which it passes into the blood vessels and is converted into blood. It was occasionally examined and described by the physiologists and anatomists during the course of the last century. But it is only a few years since we became acquainted with its chemical properties. A set of experiments on the chyle of the horse was published by M. Emmert in 1811.* Vauquelin examined the chyle of different horses in 1812,† and during the course of the same year a set of experiments on chyle was published by Mr. Brande.‡ And in 1815, a comparative examination was made by Dr. Marcet, of the chyle of animals that live on vegetable, and that of those that live on animal food.§

Chyle when drawn from the thoracic duct, about five hours after the animal has taken food is an opaque liquid of a white colour; without smell, and having a slightly acid taste, accompanied by a perceptible sweetness. It restores the blue colour of litmus paper, reddened by acetic acid, indicating the presence of an alkali. About ten minutes after it is drawn from the animal it coagulates into a stiff jelly, which in the course of 24 hours gradually separates into two parts, producing a firm contracted coagulum surrounded by a colourless fluid.

1. The coagulum, as appears from the experiments of Vauquelin, is an intermediate substance, between albumen and fibrin. He considers it as albumen on its way to assume the nature of fibrin. It is not so stiff, nor of so fibrous a texture as fibrin; it is more easily acted on

* Ann. de Chim. lxxx. 81.

† Ibid. lxxxi. 113.

‡ Phil. Trans. 1812, p. 91.

§ Annals of Philosophy, vii. 251.

Book V.

and dissolved by caustic alkalies. In potash and soda it forms a brown solution while a little ammonia is evolved. Its solution in ammonia has a reddish hue.

It is insoluble in alcohol and ether; but readily dissolved by diluted sulphuric acid. Very dilute nitric acid gradually converts it into adipocire; when the acid is more concentrated, the coagulum assumes the appearance of gelatin; and when heat is applied oxalic acid and carbonic acid are formed. Muriatic acid does not dissolve it when in a concentrated state, but it dissolves it readily when diluted with water. Acetic and oxalic acids dissolve a small portion of it when boiled on it.

When burnt it leaves a charcoal, which is with difficulty incinerated, and which contains common salt, phosphate of lime, and gives traces of iron.

Liquid portion.

2. The liquid portion of chyle is coagulated by heat, alcohol, and acids, and the coagulated portion that falls possesses the properties of albumen. When hot alcohol is employed to produce the coagulation, it dissolves a little matter which is deposited in white flocks as the alcohol cools. Vauquelin found this matter similar to the fatty substance which he extracted from brain.

Sugar,

3. Mr. Brande, after separating the albumen by heat, evaporated the liquid portion to one half, at a temperature not exceeding 200°. On cooling, small crystals separated which he considered as of the same nature as sugar of milk. They had the shape of oblique six-sided prisms, were soluble in four times their weight of boiling water, and in 16 times their weight of cold water. They are insoluble in cold alcohol, but sparingly soluble in that liquid at the boiling temperature.

4. Dr. Marcet found the same salts in chyle, and in the same proportion as in other animal fluids.

5. The chyle differs so much at different times, that little confidence can be put in the distinctions pointed out by Dr. Marcet between the chyle of herbivorous and carnivorous animals. He found the chyle of the former much less white, it contained less solid matter, and yielded less albumen.

SECT. XI.

OF BLOOD.

BLOOD is a well-known fluid which circulates in the veins and arteries of the more perfect animals. It is of a red colour, has a considerable degree of consistency, and an unctuous feel, as if it contained a quantity of soap. Its taste is slightly saline, and it has a peculiar smell.

Chap. II.
Properties
of blood.

The specific gravity of human blood is, at a medium, 1.0527.* Mr. Fourcroy found the specific gravity of bullock's blood, at the temperature of 60°, to be 1.056.† The blood does not uniformly retain the same consistence in the same animal, and its consistence in different animals is very various. It is easy to see that its specific gravity must be equally various.

No substance has attracted a greater degree of attention than the blood. The number of philosophers who have devoted their time to develop its nature, and ascertain its properties, is immense. It was not, however, till towards the middle of the 18th century that much progress was made in its chemical analysis. ‡ The prussic acid formed during its decomposition had been observed; Hales had measured the gases exhaled during its distillation; Menghini had demonstrated the presence of iron in it; and several curious observations had been made upon its coagulation. But Ronelle junior was the first chemist who published a tolerably precise analysis of it. Since that time much additional information has been acquired by the experiments of Bucquet, Fourcroy, Deyeux, Parmentier, Brande, Berzelius, &c.

When blood, after being drawn from an animal, is allowed to remain for some time at rest, it very soon coagulates into a solid mass of the consistence of curdled milk. This mass gradually separates into two parts; one of which is fluid, and is called *serum*; the other, the coagulum, has been called *cruor*, because it alone retains the red colour which distinguishes blood. This separation is very similar to the separation of curdled milk into curds and whey.

Spontaneous
decomposition.

* Haller's Physiology, ii. 41.

† Ann. de Chim. vii. 147.

‡ A very full account of every thing that had been done before that time may be found in the second volume of Haller's Physiology.

Book V.

The proportion between the cruor and serum of the blood varies much in different animals, and even in the same animal in different circumstances. The most common proportion is about one part of cruor to three parts of serum; but in many cases the cruor exceeds or falls short of this quantity; the limits of the ratios of these substances to each other appear, from a comparison of the conclusions of most of those who have written accurately on the subject, to be 1 : 1 and 1 : 4; but the first case must be very rare indeed.*

The cause of this spontaneous decomposition of blood has not hitherto been ascertained. The coagulation takes place equally in close vessels and in the open air, whether the blood be allowed to cool, or be kept at the temperature at which it is when it issues from the animal; neither is the coagulation prevented by diluting it with water, as has been sufficiently proved by the experiments of Hewson, Hunter, Deyeux, and Parmentier.

Serum.

1. The serum is of a light greenish-yellow colour; it has the taste, smell, and feel of the blood, but its consistence is not so great. Its mean specific gravity is about 1.0287.† It converts syrup of violets to a green, and therefore contains an alkali. On examination, Rouelle found that it owes this property to a portion of soda, ‡ and this has been fully confirmed by Dr. Marcet.§ When heated to the temperature of 156°, || the serum coagulates, as Harvey first discovered.** It coagulates also when boiling water is mixed with it; but if serum be mixed with six parts of cold water, it does not coagulate by heat. †† When thus coagulated, it has a greyish-white colour, and is not unlike the boiled white of an egg. ‡‡ If the coagulum be cut into small pieces, a muddy fluid may be squeezed from it, which has been termed the *serosity*. After the separation of this fluid, if the residuum be carefully washed in boiling water and examined, it will be found to possess all the properties of coagulated *albumen*. The serum, therefore, contains a considerable proportion of albumen. Hence its coagulation by heat, and the other phenomena which albumen usually exhibits.

Contains albumen.

* Haller's Physiology, ii. 47.

+ Jurin, *ibid.* p. 41.

† Dr. Pearson considers the alkali in animal fluids as potash, and not soda.

§ Nicholson's Jour. xxxi. 230; xxxiii. 147.

|| Cullen.

** De Gener. Anim. p. 161.

†† Fourcroy, Ann. de Chim. vii. 157.

‡‡ *Ibid.* 156.

The only other constituents in the serum, besides water, are soda, lactate of soda, muriate of soda, muriate of potash, and some animal matter, the nature of which has not been precisely determined. The following table exhibits the constituents of 1000 parts of the serum of bullocks' blood according to the analysis of Berzelius :

Water	905·000	Constitu- ents.
Albumen	79·990	
Lactate of soda and extractive matter	6·175	
Muriates of soda and potash	2·565	
Soda and animal matter soluble only in } water	1·520	
Loss	4·750	
	1000·000 *	

The serum of human blood he found composed as follows:

Water	905·0
Albumen	80·0
Muriates of potash and soda	6·0
Lactate of soda with animal matter	4·0
Soda, phosphate of soda with animal matter .	4·1
Loss	0·9
	1000·0 †

Dr. Marcet found the constituents of serum as follows :

Water	900·00
Albumen	86·80
Muriates of potash and soda ..	6·60
Muco-extractive matter	4·00
Subcarbonate of soda	1·65
Sulphate of potash	0·35
Earthy phosphates	0·60
	1000·00 ‡

The *muco-extractive* matter was doubtless impure lactate of soda. Berzelius is of opinion that the sulphate of potash and the earthy phosphates which were found by Dr. Marcet in the ashes of serum, were formed during the incineration.

* Djur kemien, ii. 46 ; and Annals of Philosophy, ii. 202.

† Annals of Philosophy, *ibid.*

‡ Medico-chirurgical Transact. ii. 370.

Book V.

For phosphorus, sulphur, and the bases of lime and magnesia exist according to him as constituents of albumen.

Gelatin was considered as a constituent of the serum of blood, in consequence of the observations of De Haen. But Dr. Bostock and Professor Berzelius have shown that the opinion of its existence in blood is not well founded.

Cruor contains fibrin.

2. The cruor, or *clot*, as it is sometimes called, is of a red colour, and possesses considerable consistence. Its mean specific gravity is about 1.245.* If this cruor be washed carefully by letting a small jet of water fall upon it till the water runs off colourless, it is partly dissolved, and partly remains upon the scarce. Thus it is separated into two portions: namely, 1. A white, solid, elastic substance, which has all the properties of *fibrin*; 2. The portion held in solution by the water, which consists of the colouring matter, not however in a state of purity, for it is impossible to separate the cruor completely from the serum.

Berzelius and Brande have shown that this clot is a compound of fibrin, albumen, and colouring matter of blood. According to the analysis of Berzelius, it consists of

Constituents.

Colouring matter	64
Fibrin and albumen	36
	100

When the colouring matter is incinerated, about one-third of a per cent. of oxide of iron may be extracted from its ashes. This portion of iron is a constituent of the colouring matter, and perhaps the cause of its red colour. But in what way it is united to the albuminous portion of the colouring matter remains unknown. When incinerated the colouring matter leaves $\frac{1}{80}$ th of its weight of ashes, consisting, according to the analysis of Berzelius, of the following ingredients:

Oxide of iron	50.0
Subphosphate of iron	7.5
Phosphate of lime with traces of magnesia . . .	6.0
Pure lime	20.0
Carbonic acid and loss . . . :	16.5
	100.0

Berzelius is of opinion that none of these bodies existed in the colouring matter; but merely their bases, iron, phos-

* Jurin, Haller's Physiology, ii. 41.

phorus, calcium, &c. and that they were formed during the incineration. Chap. 11.

The albumen of blood leaves the same quantity of ashes as the colouring matter. But these ashes contain no traces of iron.

The cause of the change of colour which blood undergoes when exposed to oxygen gas or common air, is still inexplicable.

M. Vogel has shown that when blood is placed in the vacuum of an air pump, a considerable quantity of carbonic acid gas separates from it.*

Dr. Gordon has rendered it probable, that during the coagulation of blood a little heat is evolved.†

It would be of great consequence to compare together the blood of different animals, and of the same animal at different ages, and to ascertain in what particulars they differ from each other. This would probably throw light on some of the obscurest parts of the animal economy. Very little progress has hitherto been made in these researches. If we except the labours of Rouelle, who obtained nearly the same ingredients, though in different proportions, from the blood of a great variety of animals, the experiments of Fourcroy on the blood of the human fœtus are almost the only ones of that kind with which we are acquainted.

He found that it differs from the blood of the adult in three things: 1. Its colouring matter is darker, and seems to be more abundant; 2. It contains no fibrin, but probably a greater proportion of gelatin (?) than blood of adults; 3. It contains no phosphoric acid.‡ Blood of the fœtus.

The examination of diseased blood, too, would be of great consequence; because the difference of its properties from the blood of people in health might throw much light on the nature of the disease.

Deyeux and Parmentier, to whom we are indebted for an excellent analysis of blood, examined that liquid drawn from patients labouring under different diseases; but the result of their examination was not so satisfactory as might have been expected.

1. It has been long known, that blood drawn from a person labouring under inflammation is soon covered with a white Diseased blood.

* Ann. de Chim. xciii. 71.

† Annals of Philosophy, iv. 139.

‡ Fourcroy, Ann. de Chim. vii. 162.

Book V.

crust, which is distinguished by physicians by the name of the *buffy coat*, and which characterises blood during inflammation. Deyeux and Parmentier ascertained that this buffy coat consists of the fibrin. The cruor, deprived of this substance, is much softer than usual, and almost totally soluble in water. The albumen of the serum is also altered, for it does not coagulate as usual when heated, but it becomes milky when mixed with hot water.

2. The blood drawn from several patients labouring under sea scurvy afforded scarcely any remarkable properties to these chemists, except a peculiar smell, and an albumen which was not so easily coagulated as usual.

3. The blood drawn from patients afflicted with a putrid fever was also examined by these chemists. Sometimes it exhibited the inflammatory crust, but frequently it did not appear to differ from healthy blood.*

4. During that disease which is known by the name of *diabetes*, in which the urine is excessive in quantity, and contains sugar, the serum of blood often, as appears from the experiments of Dr. Dobson and Dr. Rollo, assumes the appearance of whey: Dr. Wollaston has shown that it contains no perceptible quantity of sugar, even when the urine is loaded with it.

SECT. XII.

OF MILK.

MILK is a fluid secreted by the female of all those animals denominated *mammalia*, and intended evidently for the nourishment of her offspring.

The milk of every animal has certain peculiarities which distinguish it from every other milk. But the animal whose milk is most made use of by man as an article of food, and with which, consequently, we are best acquainted, is the *cow*. Chemists, therefore, have made choice of cow's milk for their experiments. I shall, at first, confine myself to the properties and analysis of cow's milk, and afterwards point out in what respect the milk of other animals differs from it, as far at least as these differences have hitherto been ascertained.

Properties
of milk.

Milk is an opaque fluid, of a whitish colour, a slight peculiar smell, and a pleasant sweetish taste. When newly drawn

* Jour. de Phys. xliv. 454.

from the cow, it has a taste very different from that which it acquires after it has been kept for some hours. It reddens vegetable blues.

Its boiling and freezing points are nearly the same as those of water; but they vary a few degrees in different milks.* Milk is specifically heavier than water, and lighter than blood; but the precise degree cannot be ascertained, because almost every particular milk has a specific gravity peculiar to itself.

When milk is allowed to remain for some time at rest, there collects on its surface a thick unctuous yellowish-coloured substance, known by the name of *cream*. Spontaneous decomposition.

After the cream is separated, the milk which remains is much thinner than before, and it has a bluish-white colour. If it be heated to the temperature of 100°, and a little *rennet*, which is water digested with the inner coat of a calf's stomach and preserved with salt, be poured into it, coagulation ensues; and if the coagulum be broken, the milk very soon separates into two substances; a solid white part, known by the name of *curd*, and a fluid part called *whey*. Thus we see that milk may be easily separated into three parts; namely, *cream*, *curd*, and *whey*.

1. Cream is of a yellow colour, and its consistence increases gradually by exposure to the atmosphere. In three or four days it becomes so thick that the vessel which contains it may be inverted without risking any loss. In eight or ten days more, its surface is covered over with mucors and byssi, and it has no longer the flavour of cream, but of very fat cheese.† This is the process for making what in this country is called a *cream cheese*. Cream.

Cream possesses many of the properties of an oil. It is specifically lighter than water; it has an unctuous feel, stains clothes precisely in the manner of oil; and if it be kept fluid, it contracts at last a taste which is very analogous to the rancidity of oils.‡ When kept boiling for some time, a little oil makes its appearance, and floats upon its surface. § Cream is neither soluble in alcohol nor oils. || These properties are sufficient to show us that it contains a quantity of oil; but this oil is combined with a part of the curd, and mixed with some serum. Cream, then, is composed of a peculiar oil, curd,

* Jour. de Phys. xxxviii. 362.

† Parmentier and Deyeux, Jour. de Phys. xxxviii. 372.

‡ Ibid. p. 375.

§ Ibid. p. 374.

|| Ibid.

Book V.
 Converted
 into butter.

and serum. The oil may be easily obtained separate by agitating the cream for a considerable time. This process, known to every body, is called *churning*. After a certain time, the cream separates into two portions: one fluid, and resembling creamed milk; and the other solid, and called *butter*.

Butter is of a yellow colour, possesses the properties of an oil, and mixes readily with other oily bodies. When heated to the temperature of 96° it melts, and becomes transparent; if it be kept for some time melted, some curd, water, or whey, separate from it, and it assumes exactly the appearance of oil.* But this process deprives it in a great measure of its peculiar flavour.

When butter is kept for a certain time, it becomes rancid, owing in a good measure to the presence of these foreign ingredients; for if butter be well washed, and a great portion of these matters separated, it does not become rancid nearly so soon as when it is not treated in this manner. It was formerly supposed that this rancidity was owing to the development of a peculiar acid; but Parmentier and Deyeux have shown that no acid is present in rancid butter.† When butter is distilled, there comes over water, an acid, and an oil, at first fluid, but afterwards concrete. The carbonaceous residuum is but small.

Nature of
 this change.

Butter may be obtained by agitating cream newly taken from milk, or even by agitating milk newly drawn from the cow. But it is usual to allow cream to remain for some time before it is churned. Now cream, by standing, acquires a sour taste; butter, therefore, is commonly made from sour cream. Fresh cream requires at least four times as much churning before it yields its butter as sour cream does; ‡ consequently cream acquires, by being kept for some time, new properties, in consequence of which it is more easily converted into butter. When very sour cream is churned, every one who has paid the smallest attention must have perceived, that the butter-milk, after the churning, is not nearly so sour as the cream had been. The butter, in all cases, is perfectly sweet; consequently the acid which had been evolved, has in a great measure disappeared during the progress of churning. It has been ascertained that cream may be churned, and but-

* Fourcroy, Ann. de Chim. vii. 170.

† Ibid. p. 375.

‡ Ibid. p. 169.

ter obtained, though the contact of atmospheric air be excluded.* On the other hand, it has been affirmed, that when cream is churned in contact with air it absorbs a considerable quantity of it.†

In many cases there is a considerable extrication of gas during the churning of butter. From the phenomena, it can scarcely be doubted that this gas is carbonic acid. Dr. Young affirms, that during the churning there is an increase of temperature amounting to four degrees.

The affinity of the oil of cream for the other ingredients is such, that it never separates completely from them. Not only is curd and whey always found in the cream, but some of this oil is constantly found in skimmed milk and even in whey: for it has been ascertained, by actual experiment, that butter may be obtained by churning whey. Twenty-seven Scotch pints of whey yield, at an average, about a pound of butter.‡ This accounts for a fact well known to those who superintend dairies, that a good deal more butter may be obtained from the same quantity of milk, provided it be churned as drawn from the cow, than when the cream alone is collected and churned.

The butter-milk, as Parmentier and Deyeux ascertained by experiment, possesses precisely the properties of milk deprived of cream.§

Cream, of the specific gravity 1·0214, was analysed by Berzelius, who found it composed of

Butter	4·5
Cheese	3·5
Whey	92·0
	<hr/>
	100·0

2. Curd, which may be separated from creamed milk by rennet, has many of the properties of coagulated albumen. It is white and solid; and when all the moisture is squeezed out, it has a good deal of brittleness. It is precipitated by acids, and the precipitate consists of the curd combined with the acid employed to precipitate it. If we digest this precipitate over carbonate of lime, or barytes and water, the acid combines with the earth and remains undissolved (supposing it

Properties
of curd.

* Young, de Lacte, 15; and Thenard, Nicholson's Jour. xii. 518.

† Mid-Lothian Report for 1795.

‡ Ibid.

§ Jour. de Phys. xxxviii. 379.

Book V.

sulphuric acid), while the curd dissolves in the water. The aqueous solution of curd, thus obtained, is yellowish, and resembles a solution of gum. When evaporated to dryness, it leaves a yellow mass which easily redissolves in water. When the solution is boiled in an open vessel, it becomes covered with a white pellicle, precisely as milk does, and acquires the smell of boiled milk. The membrane is almost insoluble in water, and appears to be produced by the action of air on the dissolved cheese.

With the mineral acids curd forms the same compounds as albumen and fibrin do. But the neutral compounds are less soluble. A great excess of acetic acid is required to dissolve curd; and the neutral compound of curd and this acid appear to be insoluble. The alkalies easily dissolve curd. The constituents of curd, according to the analysis of Gay-Lussac and Thenard, are as follows:

Composi-
tion.

Carbon	59.781
Oxygen	11.409
Hydrogen	7.429
Azote	21.381
		100.000*

As we are ignorant of the equivalent number for curd, we are unable, from the preceding analysis, to determine the constitution of that substance. But the smallest number of atoms which accords with the analysis, is the following:

7 atoms carbon =	5.25	60.87
1 atom oxygen =	1.00	11.60
5 atoms hydrogen =	0.625	7.24
1 atom azote =	1.75	20.29
		8.625		100.00

If we double the number of atoms, it will be in our power to compare curd with gelatin, albumen, and fibrin. On that supposition its composition will be

14 atoms carbon =	10.5
2 atoms oxygen =	2.0
10 atoms hydrogen =	1.25
2 atoms azote =	3.5
		17.25

* Recherches Physico-chimiques, ii. 334.

This composition approaches nearest to that of gelatin, from which it differs by the absence of 4 atoms water and 1 atom carbon.

Curd, as is well known, is used in making *cheese*; and the cheese is the better the more it contains of cream, or of that oily matter which constitutes cream. It is well known to cheesemakers, that the goodness of it depends in a great measure on the manner of separating the whey from the curd. If the milk be much heated, the coagulum broken in pieces, and the whey forcibly separated, as is the practice in many parts of Scotland, the cheese is scarcely good for any thing; but the whey is delicious, especially the last squeezed out whey, and butter may be obtained from it in considerable quantity:—a full proof that nearly the whole creamy part of the milk has been separated with the whey. Whereas if the milk be not too much heated (about 100° is sufficient), if the coagulum be allowed to remain unbroken, and the whey be separated by very slow and gentle pressure, the cheese is excellent; but the whey is almost transparent, and nearly colourless.*

Good cheese melts at a moderate heat; but bad cheese, when heated, dries, curls, and exhibits all the phenomena of burning horn. From this it is evident, that good cheese contains a quantity of the peculiar oil which constitutes the distinguishing characteristic of cream; hence its flavour and smell. Proust has found in cheese an acid which he calls the *caseic acid*, to which he ascribes several of the peculiar properties of cheese.†

This resemblance of curd and albumen makes it probable that the coagulation of milk and of albumen depend upon the same cause. Heat, indeed, does not coagulate milk, because the curd in it is diluted with too large a quantity of water; but if milk be boiled in contact with air, a pellicle soon forms on its surface, which has the properties of curd. If this pellicle be removed, another succeeds; and by continuing the boiling, the whole of the curdy matter may be separated from milk.‡ When this pellicle is allowed to remain, it falls to the bottom of the vessel, where, being exposed to a greater

* Parmentier and Deyeux, Jour. de Phys. xxxviii. 417. Fourcroy, *ibid.* p. 173.

† Jour. de Phys. lxiv. 107.

‡ Parmentier, Jour. de Phys. xxxviii. 415.

Book V.

heat, it becomes brown, and communicates to milk that disagreeable taste which, in this country, is called a *singed* taste. It happens more readily when milk is boiled along with rice, flour, &c.

If to boiling milk there be added as much of any neutral salt as it is capable of dissolving, or of sugar, or of gum arabic, the milk coagulates, and the curd separates.* Alcohol also coagulates milk;† as do all acids, rennet, and the infusion of the flowers of artichoke and of the thistle.‡ If milk be diluted with ten times its weight of water, it cannot be made to coagulate at all.§ From the experiments of Bouillon La Grange it appears that the acid or salt employed to coagulate milk combines with the curd.||

Whey.

3. Whey, after being filtered, to separate a quantity of curd which still continues to float through it, is a thin pellucid fluid, of a yellowish-green colour and pleasant sweetish taste, in which the flavour of milk may be distinguished. It always contains some curd; but nearly the whole may be separated by keeping the whey for some time boiling; a thick white scum gathers on the surface, which in Scotland is known by the name of *float whey*. When this scum, which consists of the curdy part, is carefully separated, the whey, after being allowed to remain at rest for some hours, to give the remainder of the curd time to precipitate, is decanted off almost as colourless as water, and scarcely any of the peculiar taste of milk can be distinguished in it. If it be now slowly evaporated, it deposits at last a number of white-coloured crystals, which are *sugar of milk*. Toward the end of the evaporation, some crystals of muriate of potash and of muriate of soda make their appearance.** According to Scheele, it contains also a little phosphate of lime,†† which may be precipitated by ammonia.

Contains
sugar of
milk,

Salts.

The experiments of Fourcroy and Vauquelin, Thenard, Bouillon La Grange, and Berzelius, have added considerably to our knowledge of the constituents of whey. It always reddens vegetable blues, containing a portion of lactic acid. It contains likewise some phosphate of magnesia and phosphate of iron, as Fourcroy and Vauquelin have disco-

* Scheele, ii. 52.

† Permentier, Jour. de Phys. xxxviii. 416.

‡ Parmentier, Jour. de Phys. xxxviii. 416.

§ Scheele, ii. 54.

|| Nicholson's Jour. x. 143.

** Parmentier, Jour. de Phys. xxxviii. 417.

†† Scheele, ii. 61.

vered.* Sulphate of potash, likewise, and a peculiar extractive matter, have been separated from it.

From the analysis of Berzelius we learn that 100 parts of milk deprived of its cream is composed of the following ingredients:

Water	928·75
Curd with a little cream	28·00
Sugar of milk	35·00
Muriate of potash	1·70
Phosphate of potash	0·25
Lactic acid, acetate of potash, with } a trace of lactate of iron }	6·00
Earthy phosphates	0·30
	<hr/>
	1000·00

Milk is one of the few animal substances which may be made to undergo the vinous fermentation, and to afford a liquor resembling wine or beer, from which alcohol may be separated by distillation. This singular fact seems to have been first discovered by the Tartars; they obtain all their spirituous liquors from mare's milk. It has been ascertained that milk is incapable of being converted into wine till it has become sour; after this nothing is necessary but to place it in the proper temperature; the fermentation begins of its own accord, and continues till the formation of wine be completed.† Scheele had observed that milk was capable of fermenting, and that a great quantity of carbonic acid gas was extricated from it during this fermentation.‡ But he did not suspect that the result of this fermentation was the formation of an intoxicating liquor similar to wine. The Tartars call the vinous liquid which they prepare *koumiss*. An account of its preparation and medical uses has been published by Dr. Guthrie.§ In Orkney and Shetland also, as I have been informed by Sir John Sinclair, it is prepared nearly in a similar way.

Milk capable of the vinous fermentation.

When milk is distilled by the heat of a water-bath, there comes over water, having the peculiar odour of milk; which putrefies, and consequently contains, besides mere water, some of the other constituent parts of milk. After some time the

Distillation of milk.

* Gehlen, Ann. de Chim. l. 272. Vauquelin, Gehlen's Jour. iii. 456.

† Parmentier, Jour. de Phys. xxxviii. 365.

‡ Scheele, ii. 66.

§ Edin. Trans. vol. ii.

Book V.

milk coagulates,* as always happens when hot albumen acquires a certain degree of concentration. There remains behind a thick unctuous yellowish-white substance, to which Hoffman gave the name of *franchipann*. This substance, when the fire is increased, yields at first a transparent liquid, which becomes gradually more coloured; some very fluid oil comes over, then ammonia, an acid, and at last a very thick black oil. Towards the end of the process carburetted hydrogen gas is disengaged.† There remains in the retort a coal, which contains carbonate of potash, muriate of potash, and phosphate of lime, and sometimes magnesia, iron, and muriate of soda.‡

The milk of all other animals, as far as it has hitherto been examined, consists nearly of the same ingredients as cow's milk: but there is a very great difference in their proportion.

Woman's milk.

Woman's milk has a much sweeter taste than cow's milk. When allowed to remain at rest for a sufficient time, a cream gathers on its surface. This cream is more abundant than in cow's milk, and its colour is usually much whiter. After it is separated, the milk is exceedingly thin, and has the appearance rather of whey, with a bluish-white colour, than of skimmed milk.

None of the methods by which cow's milk is coagulated succeed in producing the coagulation of woman's milk.§ It is certain, however, that it contains curd; for if it be boiled, pellicles form on its surface, which have all the properties of curd.|| Its not coagulating, therefore, must be attributed to the great quantity of water with which the curd is diluted.

Though the cream be churned ever so long, no butter can be obtained from it; but if, after being agitated for some hours, it be allowed to remain at rest for a day or two, it separates into two parts; a fluid which occupies the inferior part of the vessel, pellucid and colourless like water, and a thick white unctuous fluid, which swims on the surface. The lowermost fluid contains sugar of milk and some curd; the uppermost does not differ from cream except in consistence. The oily part of the cream, then, cannot be separated by agitation from the curd.|| This cream contains a greater portion of curd than the cream of cow's milk.||

* Bouquet.

† Parmentier, Jour. de Phys. xxxvii. 368.

‡ Mem. Med. Par. 1787, p. 607.

§ Clarke, Irish Trans. ii. 175.

|| Parmentier, Jour. de Phys. xxxviii. 419.

When this milk, after the curd is separated from it, is slowly evaporated, it yields crystals of sugar of milk and of muriate of soda. The quantity of sugar is rather greater than in cow's milk. According to Haller, the sugar obtained from cow's milk is to that obtained from an equal quantity of woman's milk as 35 : 58, and sometimes as 37 : 67, and in all the intermediate ratios. Chap. II.

Thus it appears that woman's milk differs from that of cows in three particulars: Its peculiarities.

1. It contains a much smaller quantity of curd.
2. Its oil is so intimately combined with its curd that it does not yield butter.
3. It contains rather more sugar of milk.

Parmentier and Deyeux ascertained that the quantity of curd in woman's milk increases in proportion to the time after delivery.* Nearly the same thing has been observed with respect to cow's milk.

Ass's milk has a very strong resemblance to human milk: It has nearly the same colour, smell, and consistence. When left at rest for a sufficient time, a cream forms upon its surface, but by no means in such abundance as in woman's milk. This cream, by very long agitation, yields a butter, which is always soft, white, and tasteless; and, what is singular, very readily mixes again with the butter-milk; but it may be again separated by agitation, while the vessel, which contains it, is plunged in cold water. Skimmed ass's milk is thin, and has an agreeable sweetish taste. Alcohol and acids separate from it a little curd, which has but a small degree of consistence. The serum yields sugar of milk and muriate of lime.† Ass's milk.

Ass's milk therefore differs from cow's milk in three particulars:

1. Its cream is less abundant and more insipid.
2. It contains less curd.
3. It contains more sugar of milk: the proportion is 35 : 80.

Goat's milk, if we except its consistence, which is greater, does not differ much from cow's milk. Like that milk, it throws up abundance of cream, from which butter is easily obtained. The skimmed milk coagulates just as cow's milk, and yields a greater quantity of curd. Its whey contains sugar of milk, muriate of lime, and muriate of soda.‡ Goat's milk.

* Parmentier, Jour. de Phys. xxxviii. 422.

† Ibid. p. 423.

‡ Ibid. p. 425.

Book V.
Ewe's milk.

Ewe's milk resembles almost precisely that of the cow. Its cream is rather more abundant, and yields a butter which never acquires the consistence of butter from cow's milk. Its curd has a fat and viscid appearance, and is not without difficulty made to assume the consistence of the curd of cow's milk. It makes excellent cheese.*

Mare's milk.

Mare's milk is thinner than that of the cow, but scarcely so thin as human milk. Its cream cannot be converted into butter by agitation. The skimmed milk coagulates precisely as cow's milk, but the curd is not so abundant. The serum contains sugar of milk, sulphate of lime, and muriate of lime.†

SECT. XIII.

OF EGGS.

THE eggs of fowls consist of four distinct parts; namely, the *shell*, a thin *membrane* that lines the shell; and two liquids, the *white* and the *yolk*, which fill the whole internal cavity of the egg.

1. The shell.

1. The shell, as is well known, consists chiefly of carbonate of lime cemented together by an animal matter. But Vauquelin has shown that it contains likewise magnesia, probably in the state of carbonate, phosphate of lime, and a notable quantity of oxide of iron. The animal matter which acts as a cement to the calcareous salt contains sulphur as a constituent; for when the shell is calcined, and then dissolved in an acid, the odour of sulphuretted hydrogen gas is perceptible. This odour is mixed with that of hydrocyanic acid, which is also formed during the calcination of egg-shell.‡

2. The membrane.

2. The thin membrane which lines the inside of the shell of the egg possesses the characters of albumen. Like that substance it dissolves in caustic potash, and is precipitated by acids in white flakes.§

3. The white.

3. The white is a glary insipid liquid, often used as a kind of varnish, and to clarify turbid liquors. For the most accurate experiments on the composition of this liquid we are indebted to Dr. Bostock. It had been considered as almost pure albumen, but he has shown that it is of a more compound na-

* Parmentier, Jour. de Phys. xxxviii. 428.

† Ibid. p. 431.

‡ Ann. de Chim. lxxxi. 304.

§ Vauquelin. Ibid. p. 309.

ture. When heated it coagulates; but a small portion remains liquid, which, when evaporated to dryness, leaves a residue possessing the properties of mucus. From the experiments of Bostock, it follows that white of egg is composed of Chap. II.

Water	80·0
Albumen	15·5
Mucus	4·5

100·0 *

Besides some traces of soda, of sulphuretted hydrogen gas, and of benzoic acid.

4. The yolk has a mild oily taste; when boiled, it becomes solid, and easily separates into small particles. If it be now heated for some time in a pan, it softens; and when squeezed between the fingers, drops of oil exude. If put into linen in this state and pressed, an oil may be forced out. This oil is of a yellow colour and insipid, unless too much heat has been employed in preparing it. Its properties are those of *fixed* oil, or rather of semifluid fat. Chandelier obtained this oil without the assistance of heat. † 4. The yolk.
Contains oil,

The residue, after the separation of the oil, possesses the properties of albumen, though it is still a little coloured by the remains of the oil. Hence, when washed with water, a kind of emulsion is obtained. It is owing to the presence of albumen that the yolk hardens when heated. Albumen,

Thus it appears that the yolk consists of three constituents:

1. Water.
2. Oil.
3. Albumen.

The knowledge of its component parts enables us, without difficulty, to explain the changes produced upon the yolk by the action of different reagents.

Mr. Hatchett found that when yolk is boiled with potash, it forms a pale olive-coloured concrete animal soap, which, when dissolved in water, and saturated with muriatic acid, is thrown down in the state of fat. When burnt the yolk leaves a small residuum of phosphate of lime and phosphate of soda. † Action of potash.

When agitated with water, the yolk forms a white emulsion like milk. The principal use of the yolk is to serve for food to the chicken before it bursts the shell.

* Nicholson's Jour. xi. 246; and xiv. 142.

† Jour. de Med. xvi. 43.

‡ Hatchett, Phil. Trans. 1800.

SECT. XIV.

OF SALIVA.

Book V.

History.

THE fluid secreted in the mouth, which flows in considerable quantity during a repast, is known by the name of *saliva*. All the properties of this liquor which had been observed by philosophers before the middle of the 18th century have been collected by Haller:* but since that time several additional facts have been related by Fourcroy, Du Tennenar, and Brugnatelli; and a very numerous set of experiments have been published by Mr. Siebold in 1797, in his *Treatise on the Salivary System*. Berzelius has subjected saliva, as well as the other secretions, to an accurate analysis.†

Properties.

Saliva is a limpid fluid like water; but much more viscid: it has neither smell nor taste. Its specific gravity, according to Hamberger, is 1.0167;‡ according to Siebold, 1.080. I found it 1.0038. When agitated, it froths like all other adhesive liquids; indeed it is usually mixed with air, and has the appearance of froth.

Absorbs oxygen.

It neither mixes readily with water nor oil;§ but by trituration in a mortar it may be so mixed with water as to pass through a filter.¶ It is said to have an affinity for oxygen, to absorb it readily from the air, and to give it out again to other bodies.** Hence the reason why gold or silver, triturated with saliva in a mortar, is oxidized, as Du Tennenar has observed; and why the killing of mercury by oils is much facilitated by spitting into the mixture.†† Hence also, in all probability, the reason that saliva is a useful application to sores of the skin. Dogs, and several other animals, have constantly recourse to this remedy, and with much advantage.

Analysis.

When mixed with water, a few flakes of mucus precipitate. From the experiments of Dr. Bostock, we learn that this mucus is not in a state of solution. It is separated by the filter, and subsides of its own accord when the liquid is diluted with

* Haller's Physiology, vol. vi.

† Annals of Philosophy, ii. 379.

‡ Haller's Physiology, vol. vi. 52.

§ Narcessus, *ibid.* p. 34.

¶ Fordyce on Digest. p. 51.

** Fourcroy, *Ann. de Chim.* xxviii. 262.

†† Fourcroy, *ibid.*—Dr. Saunders ascertained long ago, that mercury soon disappears when triturated with saliva. He found the mucus of the throat still more efficacious, and the mucilage of gum arabic the most efficacious of all.

water. In his analysis, this matter amounted to 0·4 of the solid matter contained in the saliva examined.*

When saliva is evaporated, it swells exceedingly, and leaves behind it a thin brown-coloured crust: but if the evaporation be conducted slowly, small cubic crystals of muriate of soda are formed. The viscosity of saliva, the property which it has of absorbing oxygen, and of being inspissated, announce the presence of a peculiar matter as a component part. This is confirmed by the effect of acetate of lead, which produces a copious precipitate in saliva. Dr. Bostock considers this matter as constituting about $\frac{1}{2}$ of the solid contents of saliva.†

When saliva is distilled in a retort, it froths very much: 100 parts yield 80 parts of water nearly pure, then a little carbonate of ammonia, some oil, and an acid, which perhaps is the prussic. The residuum amounts to about 1·56 part, and is composed of muriate of soda, phosphate of soda and phosphate of lime.‡

The acids and alcohol inspissate saliva; the alkalies disengage ammonia; oxalic acid precipitates lime; and the nitrates of lead, mercury, and silver, precipitate phosphoric and muriatic acids.

The constituents of Saliva, according to the analysis of Berzelius, are as follows:

Water	992·9
Peculiar animal matter	2·9
Mucus	1·4
Alkaline muriates	1·7
Lactate of soda and animal matter ..	0·9
Pure soda	0·2
	1000·0§

The peculiar animal matter possesses the characters of mucus as given in the last Chapter of this work. It is soluble in water, insoluble in alcohol, and precipitated by subacetate of lead.

The substance to which Berzelius has thought proper to apply the term *mucus* is in white flakes, and has very much the appearance of coagulated albumen. It has been denominated

* Nicholson's Jour. xiv. 148.

† Ibid. p. 149.

‡ Vorbeyen, Textor, Nuck, &c. as quoted by Haller, Physiol. vi. 55. Fourcroy, ix. 365.

§ Annals of Philosophy, ii. 379.

Book V.

albumen both by Dr. Bostock and myself. It is quite insoluble in water. It becomes transparent and horny in acetic, muriatic, and sulphuric acids, but does not dissolve in them. It dissolves in caustic potash, and is again separated by acids. It is easily incinerated, and the ashes contain a considerable portion of phosphate of lime, though none of that salt can be detected in it before incineration.

It is this peculiar substance which adheres to the teeth, and gives origin to the tartar with which they are often encrusted. This tartar, according to the analysis of Berzelius, is composed of

Earthy phosphates	79·0
Undecomposed mucus	12·5
Peculiar salivary matter	1·0
Animal matter soluble in muriatic acid ..	7·5
	100·0*

Saliva of the horse.

Such are the properties of human saliva. The saliva of the horse was analysed by Hapel Delachenier in 1780. He collected 12 ounces of it in the space of 24 hours by puncturing the salivary duct. Its colour was greenish-yellow; its feel soapy; it had a weak disagreeable smell and a saline taste. Boiling water and alcohol coagulated it in part; as did the acids. When sulphuric acid was used, sulphate of soda was obtained. It putrefied in about 14 days; and when allowed to evaporate spontaneously, it left a black residuum like earth. When distilled, it yielded an insipid watery liquid, crystals of carbonate of ammonia, a thick black empyreumatic oil, carburetted hydrogen, and carbonic acid; and a charcoal remained.

It is rather surprising that no experiments have been hitherto made on the saliva of dogs; though the *hydrophobia* has been usually ascribed to the infusion of the saliva of that animal rendered morbid by disease.†

The pancreatic juice has never been examined with much attention; but it does not appear, from the experiments that have been made, to differ much from saliva.

* Annals of Philosophy, ii. 381.

† The remarks of Dr. Mead in his *Essays on Poisons* are too hypothetical and absurd to be transcribed.

SECT. XV.

OF BILE.

BILE is a liquid of a yellowish-green colour, an unctuous feel, bitter taste, and peculiar smell, which is secreted by the liver; and in most animals considerable quantities of it are usually found collected in the gall bladder. Chap. II.

Great attention has been paid to this liquid by physicians; History. because the ancients were accustomed to ascribe a very great number of diseases, and even affections of the mind, to its agency. Several important observations were made on it by Boyle, Boerhaave, Verheyen, Ramsay, and Baglivi; and some facts have since been added to our chemical knowledge of bile by Maclurg and Fourcroy. But the most accurate chemical analyses were those of Cadet, published in the *Memoirs of the French Academy* for 1767, and of Van Boehante, published in 1778, till Thenard read his very ingenious experiments on it to the National Institute in 1805.* Berzelius indeed had given a still more accurate analysis of bile in the second volume of his *Animal Chemistry*, published at Stockholm in 1808. But as that book was in the Swedish language, his results remained unknown to Chemists, till he published his "General Views of the Composition of Animal Fluids," in the third volume of the *Medico-chirurgical Transactions*, in 1812.

I shall in the first place state the nature and properties of ox bile, and then give an account of the bile of other animals as far as it has been examined.

1. Ox bile is a liquid of a yellowish-green and sometimes of a deep-green colour. Its taste is very bitter, but at the same time sweetish. Its smell is feeble, but peculiar and disagreeable. It does not act on vegetable blues. Its consistence varies very much. Sometimes it is a thin mucilage; sometimes very viscid and glutinous; sometimes it is perfectly transparent, and sometimes it contains a yellow coloured matter which precipitates when the bile is diluted with water. Properties
of ox bile.

The specific gravity of ox bile seems to vary, like that of all other animal fluids. According to Hartmann, it is 1.027; † according to Thenard, 1.026 at the temperature of 43°. When strongly agitated, it lathers like soap; and for this

* Mem. d'Arcueil, i. 23 and 46.

† Haller's Phys. vi. 516.

Book V.

reason, as well as from a medical theory concerning its use, it has been often called an *animal soap*.

Action of
heat.

It mixes readily with water in any proportion, and assumes a yellow colour; but it refuses to unite with oil; when the two fluids are agitated together, the instant that they are left at rest the oil separates and swims on the surface.* Bile, however, dissolves a portion of soap readily, and is often employed to free cloth from greasy spots.

When distilled to dryness it becomes at first slightly muddy; then it froths violently, and a colourless liquid passes into the receiver, having a smell similar to that of bile, and slightly precipitating superacetate of lead. The residue in the retort, when well dried, amounts to from $\frac{1}{8}$ th to $\frac{1}{9}$ th of the whole. It has a greenish yellow colour, a very bitter taste, slightly deliquesces in the air, and is almost totally soluble in water and alcohol. When strongly heated it is decomposed, and gives the usual products of animal substances: only the proportion of oil is greater, and that of carbonate of ammonia smaller, than what is furnished by most animal bodies. A voluminous charcoal remains behind containing soda and several saline bodies; namely, common salt, phosphate of soda, sulphate of soda, phosphate of lime, and some traces of oxide of iron. The soda does not exceed $\frac{1}{100}$ th part of the bile.

When an acid is added to bile, even in a minute quantity, it acquires the property of reddening vegetable blues. The addition of a little more acid occasions a precipitate, and sulphuric acid occasions a greater precipitate than any other acid. This precipitate consists of a yellow coloured matter often visible in bile, and which is insoluble in water. If we continue to add sulphuric acid after the yellow matter has been removed, the picromel precipitates in combination with it and constitutes the green coloured matter formerly distinguished by the name of *resin* of bile. Thenard separated the picromel by means of acetate and subacetate of lead, and he considered it as two distinct substances, to which he gave the names of *resin* and *picromel*. Berzelius has rectified this mistake. Many years ago I mixed a quantity of muriatic acid with ox bile, decanted the liquid portion from the coagulum which precipitated. The liquid was put into a phial and placed near a window through which the sun shone for several

* Ramsay, Thesaur. Med. Edin. ii. 459. Maclurg, p. 10.

hours every day. After some months I found the picromel precipitated to the bottom in the form of small spheres. Its colour was white, with a slight yellowish green tint, and its taste was sweet with an impression of bitterness. Chap. II.

By evaporating a quantity of bile to dryness, calcining it, and proceeding in the usual way, Thenard ascertained the proportion of salts which it contained. The following was the result of his analysis of 800 parts of bile. Constitu-
eats.

Water	700·0
Picromel and resin	84·3
Yellow matter	4·5
Soda	4·0
Phosphate of soda	2·0
Muriate of soda	3·2
Sulphate of soda	0·8
Phosphate of lime	1·2
Oxide of iron	Trace
	800·0

The albumen, announced as a constituent of ox bile by Cadet* and Ramsay,† was the yellow matter of Thenard.

2. Such are the properties and the constituents of ox bile, as far as they have been examined by Thenard. From the experiments of the same chemist it appears, that the bile of the calf, the dog, the sheep, and the cat, resemble that of the ox exactly, both in their properties and their constituents.‡ Bile of other
animals.

3. The bile of the sow differs entirely from that of all these animals. It contains neither albumen, nor animal matter, nor picromel, but is merely a soap, as it contains a great quantity of resin and of soda, and is decomposed with facility by all the acids, even by vinegar. It contains traces also of several salts; but Thenard did not ascertain their nature.§ Of the sow.

4. The bile of the common hen, of the turkey, and the duck, has a good deal of resemblance to that of quadrupeds. But it differs in the following particulars: 1. It contains a considerable quantity of albumen; 2. The picromel has no sensible sweet taste, but is very acrid and bitter; 3. It contains very little soda; 4. The resin is not precipitated by common acetate of lead; but acetate, boiled with $\frac{1}{4}$ th of its weight of litharge, occasions it to precipitate.|| Of fowls.

* Mem. Par. 1767, p. 340.

† Thesaur. Edin. ii. 460.

‡ Mem. d'Arcueil, i. 48.

§ Ibid. p. 49.

|| Ibid. p. 50.

Book V.
Of fishes.

5. The bile of the thornback and salmon is yellowish white. When evaporated it leaves a matter which has a very sweet and slightly acrid taste. It appears to contain no resin. The bile of the carp and the eel is very green, very bitter, contains little or no albumen, but yields soda, resin, and a sweet acrid matter similar to that which may be obtained from salmon bile.*

6. The bile of the *Ateles Arachnoides*, according to Chevallier and Lassaigne, is composed of

Albumen,
A yellow colouring matter,
Green resin,
Picromel,
Muriates of potash and soda,
Phosphate of soda,
A trace of phosphate of lime.†

7. The bile of the *Felis discolor*, according to the same chemists, contains

Albumen,
Yellow matter,
Green resin,
Picromel.‡

Human
bile.

8. Human bile differs considerably from that of all other animals examined. Its colour is sometimes green, sometimes yellowish brown; sometimes it is nearly colourless. Its taste is not very bitter. It is seldom completely liquid, but usually contains some yellow matter suspended in it. When evaporated to dryness it leaves a brown matter amounting to about $\frac{1}{11}$ th of the original weight. When this matter is calcined it yields all the salts which are to be found in ox bile. All the acids decompose human bile, and throw down a copious precipitate consisting of albumen and resin. One part of nitric acid is capable of saturating 100 parts of bile. The acetate of lead throws down the resin, and leaves nothing but a yellowish liquid consisting of the salts of bile, and some atoms of a peculiar animal matter, the nature of which was not ascertained. The following were the proportions of these substances obtained by Thenard from 1100 parts of human bile:

* Mem. d'Arcueil, i. 53.

† Annals of Philosophy, xvi. 95.

‡ Ibid.

- 1000·0 water
 - from 2 to 10 yellow insoluble matter.
 - yellow matter in solution, a trace.
- 42·0 albumen.
- 41·0 resin.
- 5·6 soda.
- 4·5 phosphate of soda, sulphate of soda, muriate of soda, phosphate of lime, oxide of iron.

The following are the constituents which Berzelius found in human bile

Water	908·4
Picromel	80·0
Albumen	3·0
Soda	4·1
Phosphate of lime	0·1
Common salt	3·4
Phosphate of soda with some lime ..	1·0
	1000·0*

SECT. XVI.

OF THE CERUMEN OF THE EAR.

CERUMEN is a viscid yellow-coloured liquid secreted by the glands of the auditory canal, which gradually becomes concrete by exposure to the air. For all the facts respecting its composition at present known, we are indebted to Fourcroy and Vauquelin, especially the last, who analysed a considerable portion of it.

It has an orange-yellow colour and a bitter taste. When slightly heated upon paper, it melts, and stains the paper like an oil; at the same time it emits a slightly aromatic odour. On burning coals it softens, emits a white smoke, which resembles that given out by burning fat; it afterwards melts, swells, becomes dark-coloured, and emits an ammoniacal and empyreumatic odour. A light coal remains behind. Properties.

When agitated in water, cerumen forms a kind of emulsion, which soon putrefies, depositing at the same time white flakes.

Alcohol, when assisted by heat, dissolves $\frac{2}{3}$ of the cerumen; the $\frac{1}{3}$ which remain behind have the properties of albumen, contains albumen,

* Djur kemien, ii. 48.

Book V. mixed however with a little oily matter. When the alcohol is evaporated, it leaves a deep-orange residuum of a very bitter taste, having a smell and a consistence analogous to turpentine. It melts when heated, evaporates in a white smoke without leaving any residuum, and in short resembles very strongly the *resin of bile*. Ether also dissolves this oily body; but it is much less bitter and much lighter coloured. When the albuminous part of cerumen is burnt, it leaves traces of soda and of phosphate of lime. From these facts Vauquelin considers cerumen as composed of the following substances:

And resin.
Composi-
tion.

- | | |
|------------------------|-----------------------|
| 1. Albumen. | 4. Soda. |
| 2. An inspissated oil. | 5. Phosphate of lime. |
| 3. A colouring matter. | |

The presence of the colouring matter is inferred from the portion taken up by ether being less coloured than that taken up by alcohol.*

SECT. XVII.

OF TEARS AND MUCUS.

1. THAT peculiar fluid which is employed in lubricating the eye, and which is emitted in considerable quantities when we express grief by weeping, is known by the name of *tears*. For an accurate analysis of this fluid we are indebted to Messrs. Fourcroy and Vauquelin. Before their dissertation, which was published in 1791, appeared, scarcely any thing was known about the nature of tears.

Properties.

The liquid called tears is transparent and colourless like water; it has scarcely any smell, but its taste is always perceptibly salt. Its specific gravity is somewhat greater than that of distilled water. It gives to paper stained with the juice of the petals of mallows or violet a permanently green colour, and therefore contains a fixed alkali.† It unites with water, whether cold or hot, in all proportions. Alkalies unite with it readily, and render it more fluid. The mineral acids produce no apparent change upon it.‡ Exposed to the air, this liquid gradually evaporates, and becomes thicker. When nearly reduced to a state of dryness, a number of cubic crystals form in the midst of a kind of mucilage. These crystals

* Fourcroy, ix. 373.

† Fourcroy and Vauquelin, Jour. de Phys. xxxix. 256.

‡ Ibid. 257.

possess the properties of muriate of soda; but they tinge vegetable blues green, and therefore contain an excess of soda. The mucilaginous matter acquires a yellowish-colour as it dries.*

This liquid boils like water, excepting that a considerable froth collects on its surface. If it be kept a sufficient time at the boiling temperature, $\frac{96}{100}$ parts of it evaporate in water, and there remain about $\cdot 04$ parts of a yellowish-matter, which by distillation in a strong heat yields water and a little oil: the residuum consists of different saline matters.†

When alcohol is poured into this liquid, a mucilaginous matter is precipitated in the form of large white flakes. The alcohol leaves behind it, when evaporated, traces of muriate of soda and soda. The residuum which remains behind, when inspissated tears are burnt in the open air, exhibits some traces of phosphate of lime and phosphate of soda.‡

Thus it appears that tears are composed of the following ingredients:

1. Water.	4. Soda.	Composi- tion.
2. Mucus.	5. Phosphate of lime.	
3. Muriate of soda.	6. Phosphate of soda.	

The saline parts amount only to about 0·01 of the whole, or probably not so much.

The *mucus* contained in the tears has the property of absorbing oxygen gradually from the atmosphere, and of becoming thick and viscid, and of a yellow colour. It is then insoluble in water, and remains long suspended in it without alteration. When a sufficient quantity of chlorine is poured into tears, a yellow flaky precipitate appears, absolutely similar to this inspissated mucilage. The chlorine loses its peculiar odour; hence it is probable that oxygen has united to the mucilage. The property which this mucilage has of absorbing oxygen, and of acquiring new qualities, explains the changes which take place in tears which are exposed for a long time to the action of the atmosphere, as is the case with those persons who labour under a fistula lachrymalis.§

2. The mucus of the nose has also been examined by Fourcroy and Vauquelin. They found it composed of precisely the same ingredients with the tears. As this fluid is more

* Fourcroy and Vauquelin, Jour. de Phys. xxxix. 257.

† Ibid.

‡ Ibid. 259.

§ Ibid. 257.

Book V. exposed to the action of the air than the tears, in most cases its mucilage has undergone less or more of that change which is the consequence of the absorption of oxygen. Hence the reason of the greater viscosity and consistence of the mucus of the nose; hence also the great consistence which it acquires during colds, where the action of the atmosphere is assisted by the increased action of the parts.*

According to the analysis of Berzelius, the constituents of the mucus of the nose are as follows :

Water	933·7
Mucus	53·3
Muriate of potash and soda	5·6
Lactate of soda with animal matter	3·0
Soda	0·9
Albumen and animal matter soluble in water but insoluble in alcohol, with a trace of phosphate of soda	} 3·5
1000·0 †	

The mucus of the nose when immersed in water imbibes so much as to become transparent, and when dried on blotting paper loses nearly all the moisture it had imbibed. This may be repeated as often as we please; but the mucus gradually assumes a yellow colour. Though boiled in water it does not lose its mucilaginous nature. It dissolves in diluted sulphuric acid. Nitric acid at first coagulates it; but if the digestion be continued, the mucus is at last dissolved into a yellow clear liquid. Acetic acid hardens and does not dissolve it even in a boiling heat. Caustic alkali renders it at first more viscid, but finally dissolves it into a clear liquid. Tannin coagulates it. ‡

Expectorated matter.

The matter expectorated from the lungs and bronchiæ might be supposed to be mucilaginous; but from the experiments on it by Dr. Pearson, who has examined it with much care and ingenuity, § this does not appear to be the case. Dr. Pearson distinguishes seven different kinds of expectorated matter. 1. The jelly-like semitransparent kind of a bluish hue, excreted in a state of health; 2. The thin mucilage-like transparent matter so copiously expectorated in bronchial catarrhs; 3. The thick opaque straw-coloured, or white and

* Fourcroy and Vauquelin, Jour. de Phys. xxxix. 259.

† Annals of Philosophy, ii. 382.

‡ Berzelius, *ibid.*

§ Pearson on Expectorated Matter. Phil. Trans. 1809.

very tenacious matter, coughed up in a great variety of bronchial and pulmonary affections, especially in that of tubercles; 4. Puriform matter, secreted without any division of continuity or breach of surface of the bronchial membrane, very commonly occurring in pulmonary consumptions: 5. The matter which consists of opaque viscid masses, together with transparent fluid; or the second sort above stated with nodules of the third and fourth kind; 6. Pus from the vomicæ of tubercles; 7. Pus from vomicæ by simple inflammation of the lungs and without tubercles.

From the experiments of Dr. Pearson, it appears that the constituents of the first five of these species are nearly the same. The chief variation is in the proportions of the ingredients. They all consist of water holding in solution a quantity of albumen or animal oxide (as Dr. Pearson terms it), coagulable by heat and by the usual chemical agents. This albumen is combined with potash* which it neutralizes. This albumen sometimes amounts to the twelfth or even the tenth of the expectorated matter; and when smallest it is never less than $\frac{1}{3}$ th. The usual proportion varies from $\frac{1}{10}$ th to $\frac{1}{4}$ th. The potash varies between $\frac{1}{4}$ and $\frac{1}{3}$ ths of a part in 1000 of the expectorated matter.

Expectorated matter yields traces of sulphur, and perhaps also of phosphorus, and it contains the following saline substances: 1. Muriate of soda, varying from 1 $\frac{1}{2}$ to 2 $\frac{1}{2}$ in the 1000 of expectorated matter; 2. Phosphate of lime about $\frac{1}{4}$ a part in the 1000; 3. Ammonia united probably to phosphoric acid; 4. A phosphate probably of magnesia; 5. Carbonate of lime; 6. A sulphate; 7. Vitriifiable matter probably silica; 8. Oxide of iron. The whole of these last six substances scarcely amounts to one part in the 1000 of expectorated matter.

The proportion of saline matter and of albumen present in expectorated matter varies much in different circumstances. The thicker it is in general the smaller is the quantity of saline matter; whereas when very thin it is often impregnated with salts, especially with muriate of soda to a great degree, and tastes distinctly salt and hot.

* Dr. Pearson thinks that the alkali in animal fluids is not soda, as has hitherto been supposed, but potash. He has found this to be the case in blood, dropsy fluid, pus, liquid from blisters, urine, and mucus of the nose.

SECT. XVIII.

LIQUOR OF THE PERICARDIUM.

Book V.

THIS is a liquor which lubricates the heart. It has been examined by Dr. Bostock, having been obtained from the pericardium of a boy who had died suddenly.*

It had the colour and appearance of the serum of the blood. Evaporated to dryness, it left a residue amounting to $\frac{1}{13}$ th of its weight. When exposed to the heat of boiling water, it became opaque and thready. It was copiously precipitated by corrosive sublimate before boiling; but when boiled, evaporated to dryness, and redissolved, the solution was not affected by this salt. These experiments show us that it contained albumen. When saturated with corrosive sublimate, infusion of galls produced no effect, indicating the absence of gelatin. It was copiously precipitated by subacetate of lead, even after being boiled to dryness and the residue redissolved in water. Nitrate of silver indicated the presence of muriatic acid. Dr. Bostock, from his experiments, considers it as composed of

Water	92·0
Albumen	5·5
Mucus	2·0
Muriate of soda	0·5
	100·0

SECT. XIX.

OF THE HUMOURS OF THE EYE.

Structure of
the eye.

THE eye is one of the most delicate and complicated organs in the body; at the same time its structure, and the uses of its parts, are better understood than almost any of the other instruments of sensation. It is composed of several concentric coats, which have not been chemically examined; but, from the experiments of Hatchett on similar substances, we may consider it as probable that they possess the properties of coagulated albumen. The internal part of the eye is chiefly filled with three transparent substances, which have been called

* Nicholson's Jour. xiv. 147.

humours by anatomists; namely, 1. The *aqueous humour*, immediately behind the cornea; 2. The *crystalline humour* or *lens*; and, 3. The *vitreous humour*, behind the lens, and occupying the greatest part of the eye. We knew scarcely any thing of the chemical properties of these humours till Mr. Chenevix published an analysis of them in the Philosophical Transactions for 1802. Since that time an analysis of them has likewise been published by Nicholas,* and by Berzelius.†

I. Most of the experiments of Chenevix were made upon the eyes of sheep, as most easily procured. He afterwards tried the properties of humours from the eyes of other animals.

1. The aqueous humour of the eye of the sheep is a clear transparent liquid like water, which has very little smell or taste when fresh. Its specific gravity is 1.0090 at the temperature of 60°.

It scarcely alters vegetable blues when fresh. † When exposed to the air it evaporates slowly, and becomes slightly putrid. When boiled a very slight coagulum is formed. When 100 parts are evaporated to dryness, they leave eight parts of residuum. Tannin occasions a precipitate in it both before boiling and after it. Nitrate of silver occasions a precipitate of muriate of silver; but no precipitate is produced by other metallic salts. Hence it appears that the aqueous humour is water slightly impregnated with the following substances:

1. Albumen.
2. Gelatin.
3. Muriate of soda.

For Mr. Chenevix found that the muriatic acid was in combination with soda. Nicholas has likewise detected a little phosphate of lime in it.

Its constituents, according to the analysis of Berzelius, are,

Water	98.10
Albumen	Trace
Muriates and lactates	1.15
Soda with animal matter soluble only in water	0.75
	100.00

* Ann. de Chim. liii. 307.

† Annals of Philosophy, ii. 385.

‡ Nicholas found that it turned paper stained with the petals of the mallow to green.

Book V.
 Vitreous
 humour.

2. The vitreous matter possesses the very same properties as the aqueous; even its specific gravity is the same, or only a very little greater. Its constituents, according to the analysis of Berzelius, are

Water	98·40
Albumen	0·16
Muriates and lactates	1·42
Soda with animal matter soluble only in water	0·02
	<hr/>
	100·00

Lens.

3. The crystalline lens is solid: densest in the centre, and becoming less solid towards the circumference. It is composed of concentric coats, and is transparent. Its specific gravity is 1·1000. When fresh it has little taste. It putrefies very rapidly.

It is almost completely soluble in water. The solution is partly coagulated by heat, and gives a copious precipitate with tannin both before the coagulation and after it. It gives no traces of muriatic acid. Its composition, according to the analysis of Berzelius, is as follows:

Water	58
Peculiar matter	35·9
Muriates, lactates, and animal matter, soluble } in alcohol	} 2·4
Animal matter soluble only in water with some } phosphates	
Portions of the remaining insoluble cellular } membrane	} 2·4
	100·0

The peculiar matter of the lens possesses all the chemical characters of the colouring matter of the blood, except colour. When burnt it leaves a little ash containing a very small portion of iron. When its solution in water is coagulated by boiling, the liquid in which the coagulum was formed reddens litmus, has the smell of the humours of the muscles, and like them contains free lactic acid.*

Human eye.

II. The humours of the human eye are composed of the same ingredients as those of the sheep; the only perceptible difference consists in their specific gravity. The specific gravity of the human aqueous and vitreous humours is 1·0053; that of the crystalline 1·0790.

* Berzelius, *Annals of Philosophy*, ii. 385.

III. The humours of the eyes of oxen resemble those of the sheep in their composition. The specific gravity of the aqueous and vitreous humours is 1·0088; that of the crystalline 1·0765. Chap. II.
Eye of
oxen.

From these three sets of experiments Mr. Chenevix has drawn, as a probable conclusion, that the difference between the density of the aqueous and crystalline humours, and that of the crystalline, is in the inverse ratio of the diameter of the eye, taken from the cornea to the optic nerve.

The crystalline of the ox weighed 30 grains. When the whole of it was pared away except the six grains in the centre, the specific gravity was found to be 1·194.*

IV. Mr. Chenevix, in his paper, makes no mention of the eyes of birds; but Sir H. Davy published an abstract of it in the first volume of the Journal of the Royal Institution. He there states that Mr. Chenevix found the humours of the eyes of birds composed of the same constituents as the eyes of other animals; but that he found at the same time the specific gravity of the vitreous humour in these animals greater than the specific gravity of the crystalline.† Eye of
birds.

V. A curious set of experiments has been made by Leopold Gmelin on the black pigment, which lines the choroid coat of the eye. From 500 eyes of oxen and calves he collected 75 grains of this substance. Its colour is blackish brown. It is tasteless, and adheres to the tongue like clay. It is insoluble in water, alcohol, ether, oils, lime water, and distilled vinegar. It dissolves in potash and ammonia when assisted by heat, and is again precipitated by acids. Sulphuric acid dissolves it and acquires a black colour. Muriatic acid forms only an imperfect solution. Nitric acid dissolves it, and changes its colour to reddish-brown. When distilled it yields water, a brown oil, and carbonate of ammonia. It gives out at the same time carburetted hydrogen, carbonic oxide, azotic and oxygen gases. The coal remaining in the retort consists almost entirely of charcoal.‡

SECT. XX.

OF SINOVIÆ.

WITHIN the capsular ligament of the different joints of the body there is contained a peculiar liquid, intended evidently to

* See Phil. Trans. 1802, and Phil. Mag. xvi. 268.

† Jour. of the Royal Instit. i. 297.

‡ Schweigger's Jour, x. 507.

Book V. lubricate the parts, and to facilitate their motion. This liquid is known among anatomists by the name of *sinovia*.

Whether it be the same in different animals, or even in all the different joints of the same animal, has not been determined, as no accurate analysis of the *sinovia* of different animals has been attempted. The only analyses of *sinovia* hitherto published are, 1. that by Mr. Margueron, in the 14th volume of the *Annales de Chimie*. He made use of *sinovia* obtained from the joints of the lower extremities of oxen. 2. That of the *sinovia* of an elephant by M. Vauquelin.

Sinovia of the ox.

The *sinovia* of the ox, when it has just flowed from the joint, is a viscid semitransparent fluid, of a greenish-white colour, and a smell not unlike frog-spawn. It very soon acquires the consistence of jelly; and this happens equally whether it be kept in a cold or a hot temperature, whether it be exposed to the air or excluded from it. This consistence does not continue long; the *sinovia* soon recovers again its fluidity, and at the same time deposits a thready-like matter.*

Sinovia mixes readily with water, and imparts to that liquid a great deal of viscosity. The mixture froths when agitated; becomes milky when boiled, and deposits some pellicles on the sides of the dish; but its viscosity is not diminished.†

Contains albumen,

When alcohol is poured into *sinovia*, a white substance precipitates, which has all the properties of albumen. One hundred parts of *sinovia* contain 4.52 of albumen. The liquid still continues as viscid as ever; but if acetic acid be poured into it, the viscosity disappears altogether, the liquid becomes transparent, and deposits a quantity of matter in white threads, which possesses the following properties: 1. It has the colour, smell, taste, and elasticity of vegetable gluten. 2. It is soluble in concentrated acids and pure alkalies. It is soluble in cold water; the solution froths. Acids and alcohol precipitate the fibrous matter in flakes. One hundred parts of *sinovia* contain 11.86 of this matter.‡

And a fibrous matter,

When the liquid, after these substances have been separated from it, is concentrated by evaporation, it deposits crystals of acetate of soda. *Sinovia*, therefore, contains *soda*. Margueron found that 100 parts of *sinovia* contained about 0.71 of *soda*.

* Margueron, Ann. de Chim. xiv. 124.

† Ibid. p. 126.

‡ Ibid. p. 126, 130.

When strong sulphuric, muriatic, nitric, acetic, or sulphurous acid is poured into sinovia, a number of white flakes precipitate at first, but they are soon redissolved, and the viscosity of the liquid continues. When these acids are diluted with five times their weight of water, they diminish the transparency of sinovia, but not its viscosity; but when they are so much diluted that their acid taste is just perceptible, they precipitate the peculiar thready matter, and the viscosity of the sinovia disappears.*

And salts.

When sinovia is exposed to a dry atmosphere, it gradually evaporates, and a scaly residuum remains, in which cubic crystals, and the white saline efflorescence are apparent. The cubic crystals are muriate of soda. One hundred parts of sinovia contain about 1.75 of this salt. The saline efflorescence is carbonate of soda.†

Sinovia soon putrefies in a moist atmosphere, and during the putrefaction ammonia is exhaled. When it is distilled in a retort, there comes over first, water, which soon putrefies; then water containing ammonia; then empyreumatic oil and carbonate of ammonia. From the residuum muriate and carbonate of soda may be extracted by lixiviation. The coal contains some phosphate of lime.‡

From the analysis of Mr. Margueron, it appears that sinovia is composed of the following ingredients:

Fibrous matter	11.86
Albumen	4.52
Muriate of soda	1.75
Soda	0.71
Phosphate of lime	0.70§
Water	80.46
	100.00

Composition.

Vauquelin found the properties of the sinovia of the elephant very similar to that of oxen.||

How far the sinovia of the human body resembles that of oxen it is impossible to say. The following fact renders it

Gouty concretions.

* Margueron, Ann. de Chim. xiv. 127.

† Ibid. p. 125.

‡ Ibid. p. 128.

§ Mr. Hatchett found only 0.208 of phosphate of lime in the sinovia which he examined. He found, however, traces of some other phosphate; probably phosphate of soda. Phil. Trans. 1799, p. 216.

|| Annals of Philosophy, xii. 120.

Book V. probable that it either contains different ingredients, or is liable to alteration from disease. Concretions often form in the joints of gouty people, and make their way through the skin. These, in common language, are called *chalk-stones*. Dr. Wollaston, Dr. Pearson, and Mr. Tennant analysed some of them, and found them composed of *urate of soda*.* This analysis has been repeated by Fourcroy, and confirmed.† It is doubtless this fact that has induced Fourcroy to conjecture that uric acid forms a constituent part of sinovia.‡

SECT. XXI.

OF SEMEN.

THE peculiar liquid secreted in the testes of males, and destined for the impregnation of females, is known by the name of *semen*. The human semen and the melt of fresh water fishes alone have hitherto been subjected to chemical analysis. Nothing is known concerning the seminal fluid of other animals. 1. Vauquelin published an analysis of the human semen in 1791.

Properties. Semen, when newly ejected, is evidently a mixture of two different substances: the one fluid and milky, which is supposed to be secreted by the prostate gland; the other, which is considered as the secretion of the testes, is a thick mucilaginous substance, in which numerous white shining filaments may be discovered.§ It has a slight disagreeable odour, an acrid irritating taste, and its specific gravity is greater than that of water. When rubbed in a mortar it becomes frothy, and of the consistence of pomatum, in consequence of its enveloping a great number of air-bubbles. It converts paper stained with the blossoms of mallows or violets to a green colour, and consequently contains an alkali.||

As the liquid cools, the mucilaginous parts become transparent, and acquire a greater consistency; but in about twenty minutes after its emission, the whole becomes perfectly liquid. This liquefaction is not owing to the absorption of moisture from the air, for it loses instead of acquiring weight during its exposure to the atmosphere; nor is it owing to the action of the air, for it takes place equally in close vessels.**

* Jour. de Phys. xlv. 399.

† Fourcroy, x. 267.

‡ Ibid. ix. 224.

§ Vauquelin, Ann. de Chim. ix. 64.

|| Ibid. p. 65.

** Ibid. p. 66.

Semen is insoluble in water before this spontaneous liquefaction, but afterwards it dissolves readily in it. When alcohol or chlorine is poured into this solution, a number of white flakes are precipitated.* Concentrated alkalies facilitate its combination with water. Acids readily dissolve the semen, and the solution is not decomposed by alkalies; neither indeed is the alkaline solution decomposed by acids.†

Lime disengages no ammonia from fresh semen; but after that fluid has remained for some time in a moist and warm atmosphere, lime separates a great quantity from it. Consequently ammonia is formed during the exposure of semen to the air.‡

When chlorine is poured into semen, a number of white flakes precipitate, and the acid loses its peculiar odour. These flakes are insoluble in water, and even in acids. If the quantity of acid be sufficient, the semen acquires a yellow colour. Thus it appears that semen contains a mucilaginous substance analogous to that of the tears, which coagulates by absorbing oxygen. Mr. Vauquelin obtained from 100 parts of semen six parts of this mucilage.

When semen is exposed to the air about the temperature of 60°, it becomes gradually covered with a transparent pellicle, and in three or four days deposits small transparent crystals, often crossing each other in such a manner as to represent the spokes of a wheel. These crystals, when viewed through a microscope, appear to be four-sided prisms, terminated by very long four-sided pyramids. They may be separated by diluting the liquid with water, and decanting it off. They have all the properties of phosphate of lime.§ If, after the appearance of these crystals, the semen be still allowed to remain exposed to the atmosphere, the pellicle on its surface gradually thickens, and a number of white round bodies appear on different parts of it. These bodies also are phosphate of lime, prevented from crystallizing regularly by the too rapid abstraction of moisture. Mr. Vauquelin found that 100 parts of semen contain three parts of phosphate of lime.¶ If at this period of the evaporation the air becomes moist, other crystals appear in the semen, which have the properties of carbonate of soda. The evaporation does not go on to complete exsiccation, unless at the temperature of 77°, and

Contains
mucilage.Phosphate
of lime,

And soda.

* Vauquelin, Ann. de Chim. ix. 70.

† Ibid. p. 71.

‡ Ibid. p. 71.

§ Ibid. p. 67 and 73.

¶ Ibid. p. 68.

Book V. when the air is very dry. When all the moisture is evaporated, the semen has lost 0.9 of its weight; the residuum is semitransparent like horn, and brittle.*

When semen is kept in very moist air, at the temperature of about 77°, it acquires a yellow colour, like that of the yolk of an egg; its taste becomes acid, it exhales the odour of putrid fish, and its surface is covered with abundance of the *byssus septica*.†

Action of heat.

When dried semen is exposed to heat in a crucible, it melts, acquires a brown colour, and exhales a yellow fume, having the odour of burnt horn. When the heat is raised, the matter swells, becomes black, and gives out a strong odour of ammonia. When the odour of ammonia disappears, if the matter be lixiviated with water, an alkaline solution may be obtained, which, by evaporation, yields crystals of carbonate of soda. Mr. Vauquelin found that 100 parts of semen contain one part of soda.‡ If the residuum be incinerated, there will remain only a quantity of white ashes, consisting of phosphate of lime.

Thus it appears that semen is composed of the following ingredients :

Composition.

Water	90
Mucilage	6
Phosphate of lime . .	3
Soda	1

100

Milt of the carp.

II. Fourcroy and Vauquelin published a set of experiments on the milt of the carp in the year 1807,§ from which it appears that the nature and composition of this substance is different from that of every other hitherto examined. The milt of this fish, as is well known, has a whitish colour, a soft consistence, a greasy feel, and a smell similar to that of fish. It is neither acid nor alkaline. When triturated with potash no ammoniacal odour is exhaled, and it forms with the alkali a thick magma. Thirty parts of milt mixed with six parts of potash, and a sufficient quantity of water, and distilled, yielded only traces of ammonia, coming obviously from some muriate of ammonia, which exists naturally in the milt. When milt is dried slowly in a moderate heat it loses $\frac{3}{4}$ ths of its weight, be-

* Vauquelin, Ann. de Chim. ix. 68.

† Ann. de Chim. ix. 75.

‡ Ibid. p. 68.

§ Ibid. lxiv. 5.

comes yellow and brittle. When heated in a platinum crucible it softens and then melts, exhaling yellow vapours having the smell of animal oil. The charcoal formed contains a notable quantity of uncombined phosphoric acid, together with some phosphate of lime and phosphate of magnesia. As the acid did not exist in the milt it must have been formed during the combustion; and hence it follows that milt contains a notable quantity of phosphorus as a constituent.

Contains
phosphorus.

One hundred and twenty-three parts of fresh milt, cautiously distilled in an earthen-ware crucible gradually heated to whiteness, furnished the following products: 1. A great deal of colourless water holding in solution carbonate of ammonia, a good deal of prussiate of ammonia, and traces of muriate of ammonia; 2. A transparent oil slightly yellow; 3. A fluid blood-red oil; 4. A thick blackish-brown oil; 5. Crystals of carbonate and prussiate of ammonia; 6. A quantity of phosphorus; 7. A small quantity of carbonic acid and heavy inflammable air. The charcoal remaining in the retort amounts to 7½ parts, and contains no disengaged phosphoric acid.

When milt is triturated in distilled water a white opaque liquid is obtained, which does not become transparent though passed through the filter. When the liquid is boiled an albuminous matter coagulates; and if the residuary liquid be evaporated sufficiently it gelatinizes; a proof that it contains gelatin. Alcohol digested on milt dissolves a substance which possesses the properties of animal soap. When it is separated the milt becomes dry and harsh to the feel; a proof that its unctuousity was owing to the presence of the animal soap.

Thus it appears that milt contains albumen, gelatin, phosphorus, phosphate of lime, phosphate of magnesia, and muriate of ammonia.

Constituents.

SECT. XXII.

OF THE LIQUOR OF THE AMNIOS.

THE fœtus in the uterus is enveloped in a peculiar membranous covering, to which anatomists have given the name of *amnios*. Within this *amnios* there is a liquid, distinguished by the name of the *liquor of the amnios*, which surrounds the fœtus on every part. This liquid, as might have been expected, is very different in different animals; at least the liquor amnii in women and in cows, which alone have hitherto been analysed, have not the smallest resemblance to each other. These

Book V. two liquids have been analysed by Vauquelin and Buniva, and the result of their analysis has been published in the 33d volume of the *Annales de Chimie*.

Liquor of the human amnios.

1. The liquor of the amnios of women is a fluid of a slightly milky colour, a weak pleasant odour, and a saltish taste. The white colour is owing to a curdy matter suspended in it, for it may be obtained quite transparent by filtration.*

Its specific gravity is 1·005. It gives a green colour to the tincture of violets, and yet it reddens very decidedly the tincture of turnsole. These two properties would indicate at once the presence of an acid and of an alkali. It froths considerably when agitated. On the application of heat it becomes opaque, and has then a great resemblance to milk diluted with a large quantity of water. At the same time it exhales the odour of boiled white of egg. †

Acids render it more transparent. Alkalies precipitate an animal matter in small flakes. Alcohol likewise produces a flaky precipitate, which, when collected and dried, becomes transparent and very like glue. The infusion of nutgalls produces a very copious brown-coloured precipitate. Nitrate of silver occasions a white precipitate, which is insoluble in nitric acid, and consequently is muriate of silver. ‡

When slowly evaporated, it becomes slightly milky, a transparent pellicle forms on its surface, and it leaves a residuum which does not exceed 0·012 of the whole. By lixiviating this residuum, and evaporating the ley, crystals of muriate and carbonate of soda may be obtained. The remainder, when incinerated, exhales a fetid and ammoniacal odour, resembling that of burning horn; the ashes consist of a small quantity of carbonate of soda, and of phosphate and carbonate of lime.§

Composition.

Thus we see that the liquor of the human amnios is composed of about

Water	98·8
Albumen	} 1·2
Muriate of soda, soda	
Phosphate of lime, lime	
	100·0

Curdy matter deposited on the fœtus.

While the fœtus is in the uterus, a curdy-like matter is deposited on the surface of its skin, and in particular parts of its body. This matter is often found collected in considerable

* Ann. de Chim. xxxiii. 270.

† Ibid. p. 271.

‡ Ibid. p. 271.

§ Ibid. p. 272.

quantities. It is evidently deposited from the liquor of the amnios; and consequently the knowledge of its peculiar nature must throw considerable light upon the properties and use of that liquor. For an analysis of this substance we are also indebted to Vauquelin and Buniva.

Its colour is white and brilliant; it has a soft feel, and very much resembles newly prepared soap. It is insoluble in water, alcohol, and oils. Pure alkalis dissolve part of it, and form with it a kind of soap. On burning coals it decrepitates like a salt, becomes dry and black, exhales vapours which have the odour of empyreumatic oil, and leaves a residuum which is very difficultly reduced to ashes. When heated in a platinum crucible it decrepitates, lets an oil exude, curls up like horn, and leaves a residuum, consisting chiefly of carbonate of lime.*

These properties show that this matter is different from every one of the component parts of the liquor of the amnios, and that it has a great resemblance to the *fat*. It is probable, as Vauquelin and Buniva have conjectured, that it is formed from the albumen of that liquid, which has undergone some unknown changes. It has been long known that the parts of a fœtus, which has lain for some time after it had been deprived of life in the uterus, are sometimes converted into a kind of fatty matter. It is evident that this substance, after it is deposited upon the skin of the fœtus, must preserve it in a great measure from being acted upon by the liquor of the amnios.

2. The liquor of the amnios of the cow † has a viscosity similar to mucilage of gum arabic, a brownish-red colour, an acid and bitter taste, and a peculiar odour, not unlike that of some vegetable extracts. Its specific gravity is 1.028. It reddens the tincture of turnsole, and, therefore, contains an acid. Muriate of barytes causes a very abundant precipitate, which renders it probable that it contains sulphuric acid. Alcohol separates from it a great quantity of a reddish-coloured matter. ‡

Liquor of
the amnios
of the cow.

When this liquid is evaporated, a thick frothy scum gathers on the surface, which is easily separated, and in which some

* Ann. de Chim. xxxiii. 274.

† A minute account of the liquor amnii, and of the liquor alantoidis, of the cow, has been published by Dzondi; for an abstract of which see Gehlen's Jour. Second Series, ii. 652.

‡ Ann. de Chim. xxxiii. 275.

Book V.

white acid-tasted crystals may be discovered. By continuing the evaporation, the matter becomes thick and viscid, and has very much the look of honey. Alcohol boiled upon this thick matter, and filtered off, deposits upon cooling brilliant needle-form crystals nearly an inch in length. These crystals may be obtained in abundance by evaporating the liquor of the amnios to a fourth part of its bulk, and then allowing it to cool. The crystals soon make their appearance. They may be separated and purified by washing them in a small quantity of cold water. These crystals are *amniotic acid*.*

If after the separation of this acid the liquor of the amnios be evaporated to the consistence of a syrup, large transparent crystals appear in it, which have all the properties of sulphate of soda. The liquid of the amnios of cows contains a considerable quantity of this salt.

Composi-
tion.

Thus it appears that the liquor of the amnios of cows contains the following ingredients :

- | | |
|------------------------------|----------------------|
| 1. Water. | 3. Acid. |
| 2. A peculiar animal matter. | 4. Sulphate of soda. |

The animal matter possesses the following properties : it has a reddish-brown colour and a peculiar taste; it is very soluble in water, but insoluble in alcohol, which has the property of separating it from water. When exposed to a strong heat it swells, exhales first the odour of burning gum, then of empyreumatic oil and of ammonia, and at last the peculiar odour of prussic acid becomes very conspicuous. It differs from gelatin in the viscosity which it communicates to water, in not forming a jelly when concentrated, and in not being precipitated by tannin. It must be therefore ranked among the very undefined and inaccurate class of *animal mucilages*. When burnt, it leaves a very large coal, which is readily incinerated, and leaves a little white ashes, composed of phosphate of magnesia, and a very small proportion of phosphate of lime. †

Dr. Prout examined the liquor amnii, taken from the uterus of a cow, slaughtered in an early period of gestation. Its colour was yellow, its smell like that of new milk, its taste like that of fresh whey. He could not detect in it any amniotic acid, but it contained a sensible quantity of sugar of milk. Its constituents were as follows :

* Ann. de Chim. xxxiii. 276

† Ibid. 278.

Water	977·0	} Chap. II.
Albumen	2·6	
Substances soluble in alcohol	16·6	
Substances soluble in water, chiefly sulphate of soda? and other salts, also sugar of milk	3·8	
—————		
1000·0 *		

The portion soluble in alcohol consisted partly of the lactates, but chiefly of a peculiar substance, having considerable external resemblance to the external brown parts of roasted veal.

SECT. XXIII.

OF ANIMAL POISONS.

SEVERAL animals are furnished with liquid juices of a poisonous nature, which, when poured into fresh wounds, occasion the disease or death of the wounded animal. Serpents, bees, scorpions, spiders, are well-known examples of such animals. The chemical properties of these poisonous juices deserve peculiar attention; because it is only from such an investigation that we can hope to explain the fatal changes which they induce on the animal economy, or to discover an antidote sufficiently powerful to counteract their baneful influence. Unfortunately the task is difficult, and perhaps surpasses our chemical powers. For the progress already made in the investigation we are indebted almost entirely to the labours of Fontana.

1. The poison of the viper is a yellow liquid, which lodges Poison of the viper. in two small vesicles in the animal's mouth. These communicate by a tube with the crooked fangs, which are hollow, and terminate in a small cavity. When the animal bites, the vesicles are squeezed, and the poison forced through the fangs into the wound. This structure was partly observed by Redi, an Italian Philosopher; and his discoveries were completed and confirmed by the experiments and observations of Francini,† Tyson,‡ Mead,§ and Fontana.

This poisonous juice occasions the fatal effects of the viper's

* Annals of Philosophy, v. 416.

† New Abridg. of the Phil. Trans. ii. 8.

‡ Phil. Trans. vol. xii.

§ Mead on Poisons, p. 35

Book V.

bite. If the vesicles be extracted, or the liquid prevented from flowing into the wound, the bite is harmless. If it be infused into wounds made by sharp instruments, it proves as fatal as when introduced by the viper itself. Some of the properties of this liquid were pointed out by Mead; but it was Fontana who first subjected it to a chemical examination, sacrificing many hundred vipers to his experiments. The quantity contained in a single vesicle scarcely exceeds a drop.

Resembles gum.

It has a yellow colour, has no taste; but when applied to the tongue occasions numbness. It has the appearance of oil before the microscope, but it unites readily with water. It produces no change on vegetable blues.

When exposed to the open air, the watery part gradually evaporates, and a yellowish-brown substance remains, which has the appearance of gum arabic. In this state it feels viscid like gum between the teeth; it dissolves readily in water, but not in alcohol; and alcohol throws it down in a white powder from water. Neither acids nor alkalies have much effect upon it. It does not unite with volatile oils nor sulphuret of potash. When heated it does not melt, but swells, and does not inflame till it has become black. These properties are similar to the properties of gum, and indicate the gummy nature of this poisonous substance. Fontana made a set of experiments on the dry poison of the viper, and a similar set on gum arabic, and obtained the same results.

From the observations of Dr. Russel, there is reason to believe that the poisonous juices of the other serpents are similar in their properties to those of the viper.

This striking resemblance between gums and the poison of the viper, two substances of so opposite a nature in their effects upon the living body, is a humiliating proof of the small progress we have made in the chemical knowledge of these intricate substances. The poison of the viper, and of serpents in general, is most hurtful when mixed with the blood. Taken into the stomach it kills if the quantity be considerable. Fontana has ascertained that its fatal effects are proportional to its quantity, compared with the quantity of the blood. Hence the danger diminishes as the size of the animal increases. Small birds and quadrupeds die immediately when they are bitten by a viper; but to a full-sized man the bite seldom proves fatal.

Ammonia proposed as an antidote.

Ammonia has been proposed as an antidote to the bite of the viper. It was introduced in consequence of the theory of

Dr. Mead, that the poison was of an acid nature. The numerous trials of that medicine by Fontana robbed it of all its celebrity; but it has been lately revived and recommended by Dr. Ramsay as a certain cure for the bite of the rattle-snake.* Chap. 11.

2. The venom of the bee and the wasp is also a liquid contained in a small vesicle forced through the hollow tube of the sting into the wound inflicted by that instrument.† From the experiments of Fontana we learn that it bears a striking resemblance to the poison of the viper. That of the bee is much longer in drying when exposed to the air than the venom of the wasp. Venom of the bee and wasp,

3. The poison of the scorpion resembles that of the viper also. But its taste is hot and acrid, which is the case also with the venom of the bee and the wasp. Of the scorpion,

4. No experiments upon which we can rely have been made upon the poison of the spider tribe. From the rapidity with which these animals destroy their prey, and even one another, we cannot doubt that their poison is sufficiently virulent. ‡ Of spiders.

SECT. XXIV.

OF THE AIR CONTAINED IN THE SWIMMING BLADDERS OF FISHES.

MANY fish are furnished with a bladder filled with air, by means of which they are supposed to rise or sink in the water. When they wish to rise they are supposed to dilate their air bladder; when they wish to sink they compress it. Whether this be the use of the air bladder of fishes is somewhat doubtful. Most fish have a peculiar depth at which they almost always remain. Thus the flat fish constantly affect the bottom of the sea, while there are others that as constantly affect the surface. From the late observations of Biot it appears, that when a fish is suddenly brought from a great depth towards the surface, the air bladder swells so much that the fish cannot History.

* Phil. Mag. xvij. 125. The reader will find an interesting dissertation on the different remedies applied to the cure of the rattlesnake in the Amer. Trans. vol. iii. p. 100, by Dr. Smith Barton. The observations of Fontana in his treatise on poisons deserve particular attention.

† See a curious account of the structure of the sting by Dr. Hooke in his Micrographia.

‡ Dr. Mead's romantic account of the bite of the tarantula will entertain the reader. See Mead on Poisons, p. 57.

Book V. again sink; nay, it often bursts; and the air making its way into the stomach, swells it up and forces it into the mouth or œsophagus. The air with which these bladders is filled was first examined by Dr. Priestley in 1774. From his observations it appears that it varies in its nature. The roach was the fish the air bladder of which he examined. At first he found it filled with azote, but afterwards he got a mixture of oxygen and azote.*

Fourcroy long after examined the air in the air bladder of the carp, and found it almost pure azote; and similar results were obtained by other chemists. But by far the most complete analysis of this kind of air has been made by Biot, while in Yviza and Formentera, two islands a little to the south of Majorca and Minorca. He was employed by the French government to prolong the meridian of France to the Balearian islands, and embraced the opportunity which presented itself to examine the air in the bladders of the different species of fish caught in the neighbourhood of these islands. Next season he returned to the same islands with Mr. Laroche, who repeated and confirmed his preceding experiments.†

Experiments of Biot.

Biot found the air in the air-bladders a mixture of azotic and oxygen gas in very variable proportions. No traces of hydrogen gas could be detected; nor was there any sensible quantity of carbonic acid. The proportion of oxygen gas was very various, being sometimes very minute, and sometimes constituting almost the whole of the gas. The air bladders of those fish which live near the surface contained least oxygen gas, and the bladders of those which were brought up from a great depth contained the most. The following table exhibits the proportion of oxygen in 100 parts of the air in the different fish examined.

Names of the Fish.	Proportion of Oxygen.
Mugil cephalus (<i>Linn.</i>)	Quantity insensible.
Ditto	Ditto.
Murœnophis helena (<i>Lacepede</i>)	Very little.
Sparus annularis (<i>Linn.</i>) female	0·09
Ditto, male	0·08
Sparus sargus (<i>Linn.</i>) female	0·09
Ditto, male	0·20

* Priestley on Air, ii. 462.

† Biot's Memoirs are printed in the Mem. D'Arcueil, i. 252, and ii. 8.

Names of the Fish.	Proportion of Oxygen.
Holocentrus marinus (<i>Lacepede</i>)	0·12
Labrus turdus (<i>Linn.</i>)	0·16
Sparus melanurus (<i>Linn.</i>)	0·20
Labrus turdus (<i>Var. Linn.</i>)	0·24
Sciaena nigra, female	0·27
Ditto, male	0·25
Labrus turdus (<i>Linn.</i>) female	0·24
Ditto, male	0·28
Sparus dentex (<i>Linn.</i>) female	0·40
Sphyræna spet. (<i>Lacepede</i>)	0·44
Sparus argenteus	0·50
Sparus crythrinus	Much.
Holocentrus gigas	0·69
Gadus merluccius (<i>Linn.</i>)	0·79
Trygla lyra (<i>Linn.</i>)	0·87

The depth at which the fish in the preceding table are caught increases gradually, as well as the proportion of oxygen, from the beginning to the end of the table. The last-mentioned fish, the trygla lyra is always caught at a very great depth. The experiments of Laroche confirm the accuracy of this curious fact. The mean result, furnished by all the fishes taken at a depth greater than 150 feet, was 0·70 of oxygen; while the mean result, furnished by the fish caught at less depths, was 0·29. This superior purity is not owing to any superior purity in the air of the water of the sea at great depths. The air obtained from sea water, brought up from a great depth, yielded 0·265 of oxygen, while that from water taken at the surface was purer.

It is very remarkable that the air in the bladder of fishes, taken near the surface, should be almost pure azote. But this holds also with respect to fresh water fish. Thus Biot found the air in the air bladder of a carp to contain 0·03 of oxygen, while that of a tench contained 0·16; and Geoffroy and Vauquelin found the air in the air bladder of pikes, roaches, and perches, to contain 0·05 of oxygen. Humboldt likewise found very little oxygen in the air bladder of the gymnotus electricus.

SECT. XXV.

OF SWEAT.

Book V.

A QUANTITY of matter is constantly emitted from the skin; this matter is invisible, and is distinguished by the name of *perspiration*. Several experiments were made by Lavoisier and Seguin to ascertain its amount. Mr. Cruickshanks* made numerous trials to determine its nature, and it has been subjected to a chemical examination by Thenard.† Berzelius also has made some experiments on it.

1. Water.

1. Mr. Cruickshanks put his hand into a glass vessel, and luted its mouth at his wrist by means of a bladder. The interior surface of the vessel became gradually dim, and drops of water trickled down. By keeping his hand in this manner for an hour, he collected 30 grains of a liquid, which possessed all the properties of pure water.‡ On repeating the same experiment at nine in the evening (thermometer 62°), he collected only 12 grains. The mean of these is 21 grains. But as the hand is more exposed than the trunk of the body, it is reasonable to suppose that the perspiration from it is greater than that from the hand. Let us therefore take 30 grains per hour as the mean; and let us suppose, with Mr. Cruickshanks, that the hand is $\frac{1}{10}$ th of the surface of the body: The perspiration in an hour would amount to 1880 grains, and in 24 hours to 43,200 grains, or seven pounds six ounces troy. This is almost double of the quantity ascertained by Lavoisier and Seguin. Hence we may conclude that more matter is perspired through the hand than the other parts of the body, provided Mr. Cruickshanks's estimate of the ratio between the surface of the hand and body be not erroneous.

He repeated the experiment again after hard exercise, and collected in an hour 48 grains of water.§ He found also, that this aqueous vapour pervaded his stocking without difficulty; and that it made its way through a shamoy leather glove, and even through a leather boot, though in a much smaller quantity than when the leg wanted that covering.||

It is not difficult to see why the quantity of watery vapour diminishes with cold. When the surface of the body is ex-

* Cruickshanks on Insensible Perspiration.

+ Ann. de Chim. lix. 263.

† On Insensible Perspiration, p. 68.

§ On Insensible Perspiration, p. 70.

|| Ibid. p. 82.

posed to a cold temperature, the capacity of the cutaneous vessels diminishes, and consequently the quantity which flows through them must decrease.

When the temperature, on the other hand, is much increased, either by being exposed to a hot atmosphere, or by violent exercise, the perspired vapour not only increases in quantity, but even appears in a liquid form. This is known by the name of *sweat*. In what manner sweat is produced, is not at present known; but we can see a very important service which it performs to the animal. No sooner is it thrown upon the surface of the skin than it begins to evaporate. But the change into vapour requires heat; accordingly a quantity of heat is absorbed, and the temperature of the animal is lowered. This is the reason that animals can endure to remain for some time in a much higher temperature without injury than could have been supposed.

The experiments of Tillet, and the still more decisive experiments of Fordyce and his associates, are well known. These gentlemen remained a considerable time in a temperature exceeding the boiling point of water.

2. Besides water, it cannot be doubted that *carbon* is also emitted from the skin; but in what state, the experiments hitherto made do not enable us to decide. Mr. Cruickshanks found that the air of the glass vessel in which his hand and foot had been confined for an hour contained carbonic acid gas; for a candle burned dimly in it, and it rendered lime-water turbid.* And Mr. Jurine found that air which had remained for some time in contact with the skin consisted almost entirely of carbonic acid gas.† The same conclusion may be drawn from the experiments of Ingenhousz and Milly.‡ Troussset has observed that air was separated copiously from a patient of his while bathing.§

Now it is evident that the carbonic acid gas which appeared during Mr. Cruickshanks's experiment did not previously exist in the glass vessel; consequently it must have either been transmitted ready formed through the skin, or formed during the experiment by the absorption of oxygen gas, and the consequent emission of carbonic acid gas. The experiments of Mr. Jurine do not allow us to suppose the first of these to be

* On Insensible Perspiration, p. 70 and 81.

† Encyc. Meth. Med. i. 515.

‡ Ibid. p. 511.

§ Ann. de Chim. xlv. 73.

Book V.

true; for he found that the quantity of air allowed to remain in contact with the skin did not increase. Consequently the appearance of the carbonic acid gas must be owing either to the emission of carbon, which forms carbonic acid gas by combining with the oxygen gas of the air, or to the absorption of oxygen gas, and the subsequent emission of carbonic acid gas; precisely in the same manner, and for the same reason, that these substances are emitted by the lungs. The last is the more probable opinion; but the experiments hitherto made do not enable us to decide.

3. An oily matter.

3. Besides water and carbon, or carbonic acid gas, the skin emits also a particular odorous substance. That every animal has a peculiar smell, is well known: the dog can discover his master, and even trace him to a distance, by the scent. A dog, chained some hours after his master had set out on a journey of some hundred miles, followed his footsteps by the smell, and found him on the third day in the midst of a crowd.* But it is needless to multiply instances of this fact; they are too well known to every one. Now this smell must be owing to some peculiar matter which is constantly emitted; and this matter must differ somewhat either in quantity or some other property, as we see that the dog easily distinguishes the individual by means of it. Mr. Cruickshanks has made it probable that this matter is an oily substance; or at least that there is an oily substance emitted by the skin. He wore repeatedly, night and day, for a month, the same vest of fleecy hosiery during the hottest part of the summer. At the end of this time he always found an oily substance accumulated in considerable masses on the nap of the inner surface of the vest, in the form of black tears. When rubbed on paper, it makes it transparent, and hardens on it like grease. It burns with a white flame, and leaves behind it a charry residuum.†

4. An acid, and

4. Berthollet has observed the perspiration acid;‡ and he has concluded that the acid which is present is the phosphoric: but that has not been proved. Thenard obtained *acetic acid*; but Berzelius has rendered it probable that the true acid of sweat is the *lactic*, and that in Thenard's experiments the acetic was formed at the expense of the lactic. He evaporated a few drops of sweat to dryness in a watch glass. The yellow-

* Cruickshanks on Insensible Perspiration, p. 93.

† Ibid. p. 92.

‡ He observed that a blue paper, applied to a part labouring under a gouty paroxysm, became always red. Jour. de Phys. xxviii. 275.

ish residue had all the appearance under the microscope of the usual mixture of the muriates of potash and soda, with lactic acid, lactate of soda and its accompanying animal matter. It reddened litmus and dissolved in alcohol.*

Chap. II.

5. The small quantity of animal matter which Thenard found in the perspired matter possessed characters which induced him to consider it as similar to gelatin in its nature.

5. Animal matter.

Fourcroy and Vauquelin have ascertained that the scurf which collects on the skins of horses consists chiefly of phosphate of lime, and nephrin is even sometimes mixed with it.

Such are the constituents of *perspired matter*. Though the sweat has not been directly analysed, there is every reason to believe that it is composed of similar constituents.

SECT. XXVI.

OF URINE.

No animal substance has attracted more attention than urine, both on account of its supposed connection with various diseases, and on account of the very singular products which have been obtained from it. Mr. Boyle, and the other chemists who were his contemporaries, were induced to attend particularly to this liquid, by the discovery of a method of obtaining phosphorus from it. Boerhaave, Haller, Haupt, Margraff, Pott, Rouelle, Pronst, and Klaproth, successively improved the method of obtaining the phosphoric salts from urine, or added something to our knowledge of the component parts of these salts. Scheele added greatly to our knowledge of urine by detecting several new substances in it which had not been suspected. Cruickshanks has given us a very valuable paper on urine in the second edition of *Rollo's Diatesis*. Fourcroy and Vauquelin have published an elaborate analysis of it; and Mr. Proust has published another, no less ingenious, in which he points out the existence of several substances which had been overlooked by other chemicall philosophers. But the most elaborate and exact analysis of urine has been made by Professor Berzelius.†

History.

Fresh urine differs considerably in its appearance according to the state of the person and the time at which it is voided. In general, healthy urine is a transparent liquid of a light-

Proper:

* Annals of Philosophy, ii. 415.

† Djour kemich, ii. 278.

Book V.

amber colour, an aromatic odour resembling that of violets, and a disagreeable bitter taste. Its specific gravity varies, according to Mr. Cruickshanks, from 1.005 to 1.033.* When it cools, the aromatic smell leaves it, and is succeeded by another, well known by the name of *urinous smell*. This smell is succeeded in two or three days by another, which has a considerable resemblance to that of sour milk. This smell gradually disappears in its turn, and is succeeded by a fetid alkaline odour.

Contains an acid,

1. Urine reddens paper stained with turnsole and with the juice of radishes, and therefore contains an acid. This acid has been generally considered as the phosphoric; Thenard affirms that it is in reality the acetic.† According to Berzelius it is the lactic.‡

Phosphate of lime,

2. If a solution of ammonia be poured into fresh urine, a white powder precipitates, which has the properties of phosphate of lime. The presence of this substance in urine was first discovered by Scheele.§ The phosphate of lime is obviously kept in solution by the acetic acid. Scheele affirms that it is most abundant in the urine of the sick. Berthollet has observed, that the urine of gouty people is more acid than that of people in perfect health. The average quantity of phosphate of lime in healthy urine is, as Cruickshanks has ascertained, about $\frac{1}{900}$ of the weight of the urine. ||

Phosphate of magnesia,

3. If the phosphate of lime precipitated from urine be examined, a little magnesia will be found mixed with it. Fourcroy and Vauquelin have ascertained that this is owing to a little phosphate of magnesia which urine contains, and which is decomposed by the alkali or lime employed to precipitate the phosphate of lime.**

4. Berzelius informs us that there is a little fluuate of lime mixed with the phosphate of lime of urine. He mixed the phosphates of urine with their own weight of sulphuric acid, and exposed a piece of glass to the fumes prepared for etching. It was sensibly corroded, and thus showed the evolution of fluoric acid.††

* Dr. Henry informs us that he never found it higher than 1.020. *Annals of Philosophy*, i. 28. I have repeatedly found that of diabetic urine 1.048.

† *Ann. de Chim.* lix. 269.

‡ *Annals of Philosophy*, ii. 418.

§ Scheele, i. 208.

|| *Phil. Mag.* ii. 241.

** *Ann. de Chim.* xxxi. 66. The same discovery had been previously made by Dr. Wollaston. See his examination of calculi, *Phil. Trans.* 1797.

†† *Annals of Philosophy*, ii. 416.

5. Urine contains likewise a notable proportion of sulphuric acid united both to potash and to soda. This was ascertained by Berzelius. If urine be acidulated with nitric acid to prevent the precipitation of the phosphates, and then mixed with muriate of barytes, a copious precipitate of sulphate of barytes falls. If the urine, after the removal of this precipitate, be supersaturated with ammonia, the phosphates fall down, and along with them there is an additional quantity of sulphate of barytes.*

Chap. 11.
Sulphate of
potash and
soda.

6. Proust informs us that carbonic acid exists in urine, and that its separation occasions the froth which appears during the evaporation of urine.† Foureroy and Vauquelin, on the other hand, consider this acid as formed during the evaporation by the decomposition of the nephrin. The observations of Proust confirm those that had been made by Priestley and Perceval. Vogel has confirmed the assertion of Proust. He put a quantity of fresh urine under the exhausted receiver of an air pump. Many air bubbles were extricated. This air was made to pass through lime water, which became milky and carbonate of lime was deposited. Thus showing that the air evolved was carbonic acid gas.‡

Carbonic
acid,

6. When fresh urine cools, it often lets fall a brick-coloured precipitate, which Scheele first ascertained to be crystals of uric acid. All urine contains this acid, even when no sensible precipitate appears when it cools. For, if a sufficient quantity of clear and fresh urine be evaporated to $\frac{1}{10}$ of its weight, a subtle powder precipitates to the bottom, and attaches itself in part very firmly to the vessel. This part may be dissolved in pure alkali, and precipitated again by acetic acid. It exhibits all the properties of uric acid.§ The fact is, that the precipitate which usually falls when urine cools consists chiefly of phosphate of lime and uric acid. It may be dissolved in diluted nitric acid. If the solution be heated and evaporated to dryness, it assumes a fine rose colour if uric acid be present.|| The proportion of uric acid varies considerably in urine. It crystallizes in small red prisms, partly on the surface, if urine be mixed with some nitric acid, and left exposed to the air. Indeed, from the experiments of Dr. Egan, we learn that all acids when mixed with urine throw down uric acid in crystals.**

Uric acid,

* Berzelius, *Annals of Philosophy*, ii. 417.

† *Ann. de Chim.* xxxvi. 260.

‡ Scheele, i. 207.

** *Phil. Mag.* xxiii. 298.

‡ *Ibid.* xciii. 71.

|| Cruickshanks.

Book V.
Rosacic
acid,

8. During intermittent fevers, and especially during diseases of the liver, a copious sediment of a brick-red colour is deposited from urine. This sediment contains the rosacic acid of Proust.

But Dr. Prout has ascertained, by very careful experiments, that the pink sediment consists sometimes of urate of ammonia, and sometimes of urate of soda mixed with phosphate of lime.*

Benzoic
acid,

9. If fresh urine be evaporated to the consistence of a syrup, and muriatic acid be then poured into it, a precipitate appears which possesses the properties of benzoic acid. Scheele first discovered the presence of benzoic acid in urine. He evaporated it to dryness, separated the saline part, and applied heat to the residuum. The benzoic acid was sublimed, and found crystallized in the receiver. The above method was first proposed by Fourcroy and Vauquelin.† By it very considerable quantities of benzoic acid may be obtained from the urine of horses and cows, where it is much more abundant than in human urine. In human urine it varies from $\frac{1}{1000}$ to $\frac{1}{10000}$ of the whole.‡ Proust affirms that the acid obtained by Scheele's process is not the benzoic, but another possessed of similar properties; but differing in this circumstance, that nitric acid decomposes it, whereas it only whitens benzoic acid.§

Albumen
and gelatin,

10. When an infusion of tannin is dropped into urine, a white precipitate appears, having the properties of the combination of tannin and albumen. Urine, therefore, contains albumen or gelatin. These substances had been suspected to be in urine, but their presence was first demonstrated by Seguin, who discovered the above method of detecting them. Their quantity in healthy urine is very small, often indeed not sensible. Cruickshanks found that the precipitate afforded by tannin in healthy urine amounted to $\frac{1}{240}$ th part of the weight of urine.|| In many diseases the quantity of these matters is very much increased. The urine of dropsical people often contains so much albumen, that it coagulates not only on the addition of acids, but even on the application of heat.** In all cases of impaired digestion, the albuminous part of urine

* Annals of Philosophy, xv. 155.

† Ann. de Chim. xxxi. 62.

§ Ibid. xxxvi. 278.

** Cruickshanks, Phil. Mag. ii. 243.

‡ Ibid. p. 63.

|| Phil. Mag. ii. 243.

is much increased. This forms one of the most conspicuous and important distinctions between the urine of those who enjoy good and bad health.*

11. If urine be evaporated by a slow fire to the consistence of a thick syrup, it assumes a deep brown colour, and exhales a fetid ammoniacal odour. When allowed to cool, it concretes into a mass of crystals, composed of all the component parts of urine. If four times its weight of alcohol be poured upon this mass, at intervals, and a slight heat be applied, the greatest part of it is dissolved. The alcohol, which has acquired a brown colour, is to be decanted off, and distilled in a retort in a sand heat, till the mixture has boiled for some time, and acquired the consistence of a syrup. By this time the whole of the alcohol has passed off, and the matter, on cooling, crystallizes in quadrangular plates which intersect each other. This substance is *nephrin*, which composes $\frac{2}{100}$ of the urine, provided the watery part be excluded. To this substance the taste of urine is owing. It is a substance which characterizes urine, and constitutes it what it is, and to which the greater part of the very singular phenomena of urine are to be ascribed.

It may be detected by evaporating urine to the consistence of a syrup, and pouring into it concentrated nitric acid. Immediately a great number of white shining crystals appear in the form of plates, very much resembling crystallized boracic acid. These crystals are *nephrin* combined with nitric acid.

The quantity of *nephrin* varies exceedingly in different urines. In the urine voided soon after a meal, very little of it is to be found, and scarcely any at all in that which hysterical patients void during a paroxysm.

12. If urine be slowly evaporated to the consistence of a syrup, a number of crystals make their appearance on its surface: these possess the properties of muriate of soda. Urine therefore contains muriate of soda. It is well known that muriate of soda crystallizes in cubes; but when obtained from urine it has the form of octahedrons. This singular modification of its form is owing to the action of *nephrin*.† It has been long known that urine saturated with muriate of soda deposits that salt in regular octahedrons.

13. The saline residuum which remains after the separation

* Fourcroy and Vauquelin, Ann. de Chim. xxxi. 61.

† Fourcroy and Vauquelin.

Phosphates
of ammonia
and of soda,

Book V.

of nephrin from crystallized urine by means of alcohol has been long known by the names of *fusible salt of urine* and *microcosmic salt*. Various methods of obtaining it have been given by chemists, from Boerhaave, who first published a process, to Rouelle and Chaulnes, who gave the method just mentioned. If this saline mass be dissolved in a sufficient quantity of hot water, and allowed to crystallize spontaneously in a close vessel, two sets of crystals are gradually deposited. The lowermost set has the figure of flat rhomboidal prisms; the uppermost, on the contrary, has the form of rectangular tables. These two may be easily separated by exposing them for some time to a dry atmosphere. The rectangular tables effloresce and fall to powder, but the rhomboidal prisms remain unaltered.

When these salts are examined, they are found to have the properties of phosphates. The rhomboidal prisms consist of phosphate of ammonia united to a little phosphate of soda; the rectangular tables, on the contrary, are phosphate of soda united to a small quantity of phosphate of ammonia. Urine, then, contains phosphate of soda and phosphate of ammonia.

Muriate of ammonia,

14. When urine is cautiously evaporated, a few cubic crystals are often deposited among the other salts; these crystals have the properties of muriate of ammonia. Now the usual form of the crystals of muriate of ammonia is the octahedron. The change of its form in urine is produced also by nephrin. This salt is obtained in greater abundance when the crystals of nephrin obtained from the alcohol solution are distilled.*

Sulphur.

15. When urine is boiled in a silver basin, it blackens the basin; and if the quantity of urine be large, small crusts of sulphuret of silver may be detached. Hence we see that urine contains sulphur. This sulphur exhales along with the carbonic acid when the urine putrefies; for the fumes which separate from urine in that state blacken paper stained with acetate of lead.†

16. The *cloud* which appears when urine cools, and which may always be detected in greater or smaller quantity in urine, is the *mucus* of the bladder. It may be separated by the filter. This mucus facilitates the deposition of crystals of uric acid. Hence when it happens to be detained in the bladder it may serve as a nucleus for the formation of calculus.

* Fourcroy and Vauquelin.

† Proust, Ann. de Chim. xxxvi. 258.

17. Dr. Prout has several times detected nitric acid in the pink sediments of urine. His method was to combine the acid with potash and convert it into saltpetre. He is of opinion that the red colour of this sediment is owing to the action of the nitric acid on the uric acid, and its consequent conversion into purpuric acid.*

18. We are at present ignorant of the nature and properties of the substance which gives urine its yellow colour; and likewise of the substance to which it owes its odour.

The constituents of healthy urine according to the analysis of Berzelius are as follows:

Water	933.00	Constitu- ents.
Nephrin	30.10	
Sulphate of potash	3.71	
Sulphate of soda	3.16	
Phosphate of soda	2.94	
Muriate of soda	4.45	
Phosphate of ammonia	1.65	
Muriate of ammonia	1.50	
Free lactic acid	} 17.14	
Lactate of ammonia		
Animal matter soluble in alcohol		
Nephrin not separable from the preceding		
Earthy phosphates with a trace of fluuate of lime .	1.00	
Uric acid	1.00	
Mucus of the bladder	0.32	
Silica	0.03	
	1000.00†	

No substance putrefies sooner, or exhales a more detestable odour, during its spontaneous decomposition, than urine; but there is a very great difference in this respect in different urines. In some, putrefaction takes place almost instantaneously as soon as it is voided; in others, scarcely any change appears for a number of days. Fourroy and Vauquelin have ascertained that this difference depends on the quantity of gelatin and albumen which urine contains. When there is very little of these substances present, urine remains long unchanged; on the contrary, the greater the quantity of gelatin or albumen, the sooner does putrefaction commence. The putrefaction of urine, therefore, is in some degree the test of

* Annals of Philosophy, xv. 155.

† Ibid. ii. 423.

Book V.

the health of the person who has voided it: for a superabundance of albumen in urine always indicates some defect in the power of digestion.*

The rapid putrefaction of urine, then, is owing to the action of albumen on nephrin. We have seen already the facility with which that singular substance is decomposed, and that the new products into which it is changed, are ammonia, carbonic acid, and acetic acid. Accordingly, the putrefaction of urine is announced by an ammoniacal smell. Mucilaginous flakes are deposited, consisting of part of the gelatinous matter. The lactic acid is saturated with ammonia, and the phosphate of lime, in consequence, is precipitated. Ammonia combines with the phosphate of magnesia, forms with it a triple salt, which crystallizes upon the sides of the vessel in the form of white crystals, composed of six-sided prisms, terminated by six-sided pyramids. The uric and benzoic acids are saturated with ammonia; the acetic acid, and the carbonic acid, which are the products of the decomposition of the nephrin, are also saturated with ammonia; and notwithstanding the quantity which exhales, the production of this substance is so abundant, that there is a quantity of unsaturated alkali in the liquid. Putrefied urine, therefore, contains chiefly the following substances, most of which are the products of putrefaction:

- Ammonia.
- Carbonate of ammonia.
- Phosphate of ammonia.
- Phosphate of magnesia and ammonia.
- Urate of ammonia.
- Acetate of ammonia.
- Benzoate of ammonia.
- Muriate of soda.
- Muriate of ammonia.

Besides the precipitated albumen and phosphate of lime.†

Distillation
of urine.

The distillation of urine produces almost the same changes; for the heat of boiling water is sufficient to decompose nephrin, and to convert it into ammonia, carbonic and acetic acids. Accordingly, when urine is distilled, there come over water, containing ammonia dissolved in it, and carbonate of ammonia in crystals; the acids contained in urine are saturated with ammonia, and the gelatin and phosphate of lime precipitate.‡

* Fourcroy, Ann. de Chim. xxxi. 61.

† Ann. de Chim. xxxi. 70.

‡ Ibid. xxxi. 55.

Such are the properties of human urine in a state of health : Chap. II.
 but this excretion is singularly modified by disease ; and the Changes in
 changes to which it is liable have attracted the attention of it by dis-
 physicians in all ages, because they serve in some measure to ease.
 indicate the state of the patient, and the progress of the dis-
 ease under which he labours. The following are the most
 remarkable of these changes that have been observed.*

1. In *inflammatory* diseases the urine is of a red colour, and Inflamma-
 peculiarly acrid ; it deposits no sediment on standing, but with tion.
 corrosive sublimate it yields a copious precipitate.

2. During *jaundice* the urine has an orange-yellow colour, Jaundice.
 and communicates the same tint to linen. Muriatic acid
 renders this urine green, and thus detects the presence of a
 little bile.

From the experiments of Fourcroy and Vauquelin, we learn
 that urine sometimes in these cases contains a substance ana-
 logous to the *yellow* matter which they formed by the action
 of nitric acid on muscular fibres.†

3. About the end of *inflammatory* diseases the urine be- End of in-
 comes abundant, and deposits a copious pink-coloured sedi- flammation.
 ment, composed of rosacic acid, a little phosphate of lime,
 and uric acid.

4. During *hysterical* paroxisms the urine usually flows Hysteria.
 abundantly. It is limpid and colourless, containing much
 salt, but scarcely any nephrin or gelatin.

5. Mr. Berthollet observed, that the urine of *gouty* persons Gout
 contains usually much less phosphoric acid than healthy urine.
 But during a gouty paroxism it contains much more phos-
 phoric acid than usual ; though not more than constantly
 exists in healthy urine.‡

6. In general *dropsy*, the urine is loaded with albumen, and Dropsy
 becomes milky, or even coagulates when heated, or at least
 when acids are mixed with it. In dropsy from diseased liver,
 no albumen is present ; the urine is scanty, high coloured,
 and deposits the pink-coloured sediment.

In certain cases, females have been observed to pass urine
 which had the appearance of milk, and which upon examina-
 tion, proved to differ from common urine in containing a
 notable proportion of the curdy part of milk.§

* See Cruickshanks, Phil. Mag. ii. 240 ; and Fourcroy, x. 166.

† Phil. Mag. xxiii. 331.

‡ Jour. de Phys. xxviii. 275.

§ See the analysis of a specimen of this urine by Caballe, Ann. de Chim.
 lv. 61.

Book V.
Dyspepsia.

7. In *dyspepsia* the urine always yields a copious precipitate with tan, and putrefies rapidly.

Rickets.

8. The urine of *rickety* patients is said to be loaded with phosphate of lime, or, according to others, with oxalate of lime.

Hepatitis.

9. Mr. Rose has ascertained that in chronic hepatitis the urine is destitute of nephrin.* This curious fact has been confirmed by the experiments of Dr. Henry.†

Diabetes.

10. In *diabetes* the urine is sweet-tasted, and often loaded with saccharine matter. In one case, the urine emitted daily by a diabetic patient, according to the experiments of Cruickshanks, contained 29 ounces of sugar.‡

Urine of inferior animals.

The urine of other animals differs considerably from that of man. For the analysis of the urine of quadrupeds hitherto made we are chiefly indebted to Rouelle junior. The following facts have been ascertained by that chemist, and by the experiments of Fourcroy, Vauquelin, Brande, and Chevreul.

The horse.

I. The urine of the horse has a peculiar odour: after exercise it is emitted thick and milky; at other times it is transparent, but becomes muddy soon after its emission. When exposed to the air, its surface becomes covered with a crust of carbonate of lime. It gives a green colour to syrup of violets, and has the consistence of mucilage. The following are its constituents, as estimated by Fourcroy and Vauquelin from their experiments:

Carbonate of lime	0·011
Carbonate of soda	0·009
Benzoate of soda	0·024
Muriate of potash	0·009
Nephrin	0·007
Water and mucilage	0·940

1·000 §

The urine of the horse has been analyzed by Mr. Brande. The salts which it contains are, according to him, the following:

* Annals of Philosophy, v. 42.

† Ibid. vi. 392.

‡ The reader may consult a set of comparative experiments on healthy and diabetic urine by Nicolas, Ann. de Chim. xlv. 32; the analysis of diabetic urine by Sorg (Gehlen's Jour. vi. 9), and by Dupuytren and Thenard (Ann. de Chim. lix. 41.)

§ Mem. de l'Institut, ii. 431.

Carbonate of lime.
Carbonate of soda.
Sulphate of soda.

Muriate of soda.
Benzoate of soda.
Phosphate of lime.

These saline substances amount to about one-eighth of the urine.* The urine of the horse has been recently subjected to a new examination by Chevreul expressly to ascertain whether it contained phosphate of lime announced in it by Brande. His experiments confirmed those of Fourcroy and Vauquelin. The urine yielded no traces of phosphate of lime; but it contained magnesia and sulphate of potash, neither of which substances had been observed by Mr. Brande.†

From the late experiments of Mr. Giese, we learn that the quantity of benzoate of soda varies considerably in the urine of horses. In some specimens he found it in abundance, and easily precipitated by muriatic acid. In others there was little or none. He could detect no benzoic acid in the food of horses. Hence he considers it as formed within the animal, and he thinks that it appears only in cases of disease.‡

Dr. Prout has lately shown that the urine of the horse contains a much greater proportion of nephrin than had been detected in it by Fourcroy. When fresh urine of the horse is mixed with nitric acid, crystals of nitrate of nephrin are deposited without any previous concentration. §

II. The urine of the ass was likewise examined by Mr. Brande. It is transparent, but mucilaginous. It gives a green colour to syrup of violets, but no carbonate of lime is deposited from it on standing. It contains, according to Brande, nephrin more phosphate of lime than the urine of the horse, carbonate of soda, sulphate of soda, muriate of soda, and probably muriate of potash. It contains no ammonia.||

III. The urine of the cow has a strong resemblance to that of the horse; it has nearly the same odour and the same mucilaginous consistence. It tinges syrup of violets green, and deposits a gelatinous matter. On standing, small crystals are formed on its surface. It contains, according to Rouelle,

- | | |
|-------------------------|------------------|
| 1. Carbonate of potash. | 4. Benzoic acid. |
| 2. Sulphate of potash. | 5. Nephrin. |
| 3. Muriate of potash. | |

* Ann. de Chim. lxxvii. 276.

† Ibid. 303.

‡ Phil. Mag. xvii. 151.

§ Annals of Philosophy, xvi. 150.

|| Ann. de Chim. lxxvii. 277.

Book V.

The urine of the cow has been more recently examined by Mr. Brande. He found it composed of the following constituents:

Water	65
Phosphate of lime	3
Muriate of potash	}
Muriate of ammonia	
Sulphate of potash	6
Carbonate of potash	}
Carb. of ammonia	
Nephrin	4
Loss	3
	100

He obtained from it likewise a quantity of benzoic acid, but conceives that this acid was formed during the process.*

The camel.

IV. The urine of the camel was also examined by Rouelle. Its odour resembles that of the urine of the cow; its colour is that of beer; it is not mucilaginous, and does not deposit carbonate of lime. It gives a green colour to syrup of violets, and effervesces with acids like the urine of the horse and cow. Rouelle obtained from it,

- | | |
|-------------------------|-----------------------|
| 1. Carbonate of potash. | 3. Muriate of potash. |
| 2. Sulphate of potash. | 4. Nephrin. |

This urine has been more recently examined by Mr. Brande.† He obtained from it the following substances:

Water	75
Phosphate of lime	}
Muriate of ammonia	
Sulphate of potash	
Urate of potash	
Carbonate of potash	6
Muriate of potash	8
Nephrin	6
Loss	5
	100

Chevreul subjected the urine of the camel to a rigid examination, on purpose to ascertain whether it really contained the phosphate of lime announced as a constituent of it by Mr. Brande. He could detect none of it whatever, but separated from the urine of the camel the following substances:

* Ann. de Chim. lxxvii. 268.

† Ibid. 266.

Albumen.	A little sulphate of soda.
Carbonate of lime.	Much sulphate of potash.
Carbonate of magnesia.	A little carbonate of potash.
Silica.	Benzoic acid.
An atom of sulphate of lime.	Nephrin.
A trace of iron.	A red oil, to which the urine
Carbonate of ammonia.	owes its smell and colour.
A little muriate of potash.	

The last substance, the oil, seems to have been taken by Brande for uric acid.*

V. The urine of the rabbit has been analyzed by Vauquelin. The rabbit. When exposed to the air, it becomes milky, and deposits carbonate of lime. It gives a green colour to syrup of violets, and effervesces with acids. That chemist detected in it the following substances:

- | | |
|---------------------------|-----------------------|
| 1. Carbonate of lime. | 6. Muriate of potash. |
| 2. Carbonate of magnesia. | 7. Nephrin. |
| 3. Carbonate of potash. | 8. Gelatin. |
| 4. Sulphate of potash. | 9. Sulphur. |
| 5. Sulphate of lime. | |

VI. Vauquelin has also made some experiments on the urine of the guinea pig; from which it appears that it resembles the urine of the other quadrupeds. It deposits carbonate of lime, gives a green colour to syrup of violets, and contains carbonate and muriate of potash, but no phosphate nor uric acid.†

VII. The urine of the sow has been examined by M. The sow. Lassaigne. It is transparent, slightly yellow, without smell, and having a disagreeable, but not a saline taste. It contained

- Nephrin.
- Muriate of ammonia.
- Muriate of potash.
- Muriate of soda.
- Sulphate of potash.
- A little sulphate of soda.
- A trace of sulphate and carbonate of lime. ‡

VIII. The urine of the lion and tiger is alkaline at the moment of its emission. It contains a quantity of ammonia. The lion and tiger. It is destitute of uric acid and of phosphate of lime. Its constituents, as determined by Vauquelin, are the following: §

* Ann. de Chim. lxxvii. 294.

† Fourcroy, x. 181.

‡ Annals of Philosophy, xiv. 146.

§ Ann. de Chim. lxxxii. 198.

Book V.

- | | |
|--------------------------|--------------------------------|
| 1. Nephirin. | 5. Muriate of ammonia. |
| 2. Animal mucus. | 6. Trace of phosphate of lime. |
| 3. Phosphate of soda. | 7. Much sulphate of potash. |
| 4. Phosphate of ammonia. | 8. A trace of muriate of soda. |

The castor.

IX. The urine of the castor has a striking resemblance to that of herbivorous animals in general. Vauquelin obtained from it the following substances:*

- | | |
|-------------------------------------|---------------------------------|
| 1. Nephirin. | 5. Acetate of magnesia? |
| 2. Animal mucus. | 6. Sulphate of potash. |
| 3. Benzoate of potash. | 7. Muriates of potash and soda. |
| 4. Carbonates of lime and magnesia. | 8. Vegetable colouring matter. |
| | 9. Trace of iron. |

Fowls.

X. From the experiments of Fourcroy and Vauquelin, we learn that the urine of domestic fowls contains uric acid.†

The existence of that acid in the urine of fowls was called in question by Gehlen; but it has been fully confirmed by the experiments of Chevreul.‡

XI. The urine of several amphibious animals has been examined by Dr. Prout and Dr. John Davy. Dr. Prout found the urine of the boa constrictor composed almost entirely of uric acid. Dr. John Davy observed that the urine of different species of serpents when voided, was of a butyracious consistence, but became quite hard on exposure to the air. It was always uric acid, nearly pure. The urine of lizards consists likewise of nearly pure uric acid. That of the alligator, besides uric acid, contains a large portion of carbonate and phosphate of lime. The urine of turtles is a liquid containing flakes of uric acid, and holding in solution a little mucus and common salt; but no sensible portion of nephirin.§

Urinary excretions of insects.

XII. Some experiments on the urinary excretions of certain insects have been made by M. Scholz. He obtained from them the following constituents:

Uric acid	94.00
Ammonia	2.00
Phosphate of lime	3.33
	<hr/>
	99.33

* Ann. de Chim. lxxxii. 201.

† Jour. de Phys. lix. 66.

‡ Ann. de Chim. lxxvii. 305.

§ Annals of Philosophy, xiii. 209.

|| Gilbert's Annalen, xliii. 83.

SECT. XXVII.

OF FECES.

THE excrementitious matter of animals, evacuated *per anum*, consists of all that part of the food which cannot be employed for the purposes of nutrition, considerably altered, at least in part, and mixed or united with various bodies employed during digestion to separate the useless parts of the food from the nutritious. An accurate examination of these matters has long been wished for by physiologists, as likely to throw much new light on the process of digestion. For if we knew accurately the substances which were taken into the body as food, and all the new substances which were formed by digestion; that is to say, the component parts of chyle and of excrement, and the variation which different kinds of food produce in the excrement, it would be a very considerable step towards ascertaining precisely the changes produced on food by digestion.

Chap. II.

Some of the older chemists had turned their attention to the excrements of animals;* but no discovery of importance rewarded them for their disagreeable labour. Vauquelin has ascertained some curious facts respecting the excrementitious matter of fowls. In the summer of 1806, a laborious set of experiments on human feces was published by Berzelius, undertaken, as he informs us, chiefly with a view to elucidate the function of digestion.† About two years before, Thær and Einhof had published a similar set of experiments on the excrements of cattle; made chiefly to discover, if possible, how they act so powerfully as manure.‡ I shall in this Section give a view of the results obtained by these different chemists.

I. The appearance of human feces requires no particular detail. Their colour is supposed to depend upon the bile mixed with the food in the alimentary canal. When too light, it is supposed to denote a deficiency of bile; when too dark, there is supposed to be a redundancy of that secretion. The smell is fetid and peculiar, which after some time gradually changes into a sourish odour. The taste is sweetish bitter. The colour of vegetable blue infusions is not altered

Human feces.

* Van Helmont's *Custos Errans*. Sect. vi. Opera Helmont, p. 147. Neumann's Works, p. 585.

† Gehlen's Jour. vi. 509.

‡ Ibid. iii. 276.

Book V. by fresh feces, indicating the absence of any uncombined acid or alkali.*

1. The consistency of human feces varies considerably in different circumstances; but at a medium, they may be stated to lose three-fourths of their weight when dried upon a water bath.†

Constitu-
ents.

2. They do not mix readily with water; but by sufficient agitation and maceration, they may be diffused through it. The liquid, in this state, being strained through a linen cloth, leaves a matter of a greyish-brown colour, retaining a peculiar odour, which adheres long and obstinately to all those substances that come in contact with this residue. When dried, this substance exhibits the remains of vegetable matters used in food, and perhaps also of some animal matters. Its quantity amounts to about seven per cent. of the feces.‡

3. The strained liquid deposited, on standing, a yellowish-green slimy matter, which was separated by the filter. It amounted, when dry, to 14 per cent. of the feces employed. From the numerous experiments of Berzelius upon this matter, it appears to be composed chiefly of three substances: 1. A fatty matter, separated by means of alcohol, which possesses many properties in common with picromel, and which Berzelius considers as that substance a little altered. 2. A peculiar yellow-coloured substance, dissolved by water after the fatty matter is removed. This substance Berzelius compares to gelatin; but it appears to be rather more closely allied to mucus, or, at least, to contain mucus as a constituent. It dissolves in water, but not in alcohol; tannin makes its solution muddy, but occasions no precipitate; acetate of lead occasions a copious white precipitate, but does not deprive the solution of its yellow colour. It soon runs to putrefaction, exhaling the odour of putrid urine. 3. A greenish-grey residue, insoluble both in water and alcohol, and leaving, when incinerated, some silica and phosphate of potash.§

4. The liquid which passed through the filter was at first light-yellow; but by exposure to the air it became brown, which gradually deepened in colour, till the solution grew at last muddy. When concentrated by evaporation, small transparent crystals made their appearance: which proved, on examination, to be crystals of ammonio-phosphate of magnesia. The solution, on examination, was found to contain

* Gehlen's Jour. vi. 512.

† Ibid. 535.

‡ Ibid. 513.

§ Ibid. 526—531.

the following substances: 1. Albumen, which was obtained by mixing the concentrated solution with alcohol. The precipitate consisted of a mixture of albumen and phosphoric salts. The albumen obtained from 100 parts of feces amounted only to 0·9 parts. 2. Bile. By this Berzelius understood a mixture of picromel and soda. The presence of this substance was inferred from the nature of the precipitate obtained by acids, and the salt of soda obtained by evaporating the residue. The quantity contained in 100 parts of feces was 0·9. 3. A peculiar substance, of a reddish-brown colour, soluble both in water and alcohol. Acids give it an intense brown colour. A small quantity of tannin throws it down of a red colour and in a pulverulent form; a large quantity throws it down in greyish-brown flakes. It is precipitated by muriate of tin, nitrate of silver, and acetate of lead. When heated it melts and emits the smell of ammonia. It leaves behind it when burnt, traces of soda and of phosphoric salts. Berzelius supposes that this substance is formed from the picromel, by some change which it undergoes after the feces are exposed to the air. The quantity of it obtained from 100 parts of feces was 2·7 parts. 4. Various salts: these in all, from 100 parts of feces (including the ammonio-phosphate of magnesia), amounted to 1·2 parts. Their relative proportions were as follows:

Carbonate of soda	35
Muriate of soda	4
Sulphate of soda	2
Amm.-phosphate of magnesia....	2
Phosphate of lime	4

Such are the constituents of human feces, according to the experiments of Berzelius. The following table exhibits the result of his analysis.*

Water	73·3
Vegetable and animal remains	7·0
Bile	0·9
Albumen	0·9
Peculiar and extractive matter	2·7
Salts	1·2
Slimy matter; consisting of picromel, peculiar animal matter, and insoluble residue	} 14·0

100·0

* Gehler's Jour. vi. 536.

Book V.
 }
 Feces of
 cattle.

II. The excrementitious matter examined by Thacr and Einhof was that of cattle fed at the stall, chiefly on turnips. It had a yellowish-green colour, a smell somewhat similar to that of musk, and but little taste. Its specific gravity was 1.045. It did not alter vegetable blues, and of course contained no uncombined acid or alkali.

1. Sulphuric acid, when mixed with this matter, develops the odour of acetic acid; but Thacr and Einhof have shown that this acid does not exist in the feces, but is formed by the action of the sulphuric acid. The pure alkalies, nitric and muriatic acids, produce little change on the feces of cattle, at least when not assisted by heat.

2. When 100 parts are dried on a steam bath, they leave $28\frac{1}{8}$ of solid matter.

3. When eight ounces, or 3840 grains, were diffused through water, they let fall a quantity of sand, weighing 45 grains.

4. The watery solution, being strained through a linen cloth, left 600 grains of a yellowish fibrous matter, which possessed the properties of the fibrous matter of plants.*

5. The liquid, on standing, deposited a slimy substance, which was separated by filtration. It weighed when dry 480 grains. To this matter the feces owe their peculiar colour and smell. It was insoluble in water and alcohol. When heated it smelled like ox bile. It burnt like vegetable matter. Alkalies scarcely affected it. Sulphuric acid developed the odour of acetic acid. Chlorine rendered it yellow. Thacr and Einhof considered this substance as the remains of the vegetable matter employed as food by the cattle; but it is extremely probable that it might contain also a portion of picromel, as Berzelius detected that substance in similar matter from the human feces.

6. The filtered solution passed through colourless, but on exposure to the air became in a few minutes wine-yellow and then brown. When evaporated to dryness, it left a brownish matter, of a bitterish taste, and weighing 90 grains. It was soluble in water, insoluble in alcohol, and precipitated from water by that liquid. It was not precipitated by infusion of galls. The solution was found to contain some phosphoric salts. The 90 grains of residue, when heated, burnt like ani-

* Gehlen, iii. 286.

mal matter. They soon ran into putrefaction, exhaling ammonia.*

7. When evaporated to dryness and burnt, this excrementitious matter left behind it an ash, which was found (not reckoning the sand) to consist of the following salts and earths in the proportion stated: †

Lime	12
Phosphate of lime	12·5
Magnesia	2
Iron	5
Alumina with some manganese	14
Silica	52
Muriate and sulphate of potash	1·2

8. Thaer and Einhof made numerous experiments on the putrefaction of cow dung, both in close vessels and in the open air, from which it would appear that the process resembles closely the putrefaction of vegetable matter; the oxygen of the air being abundantly changed into carbonic acid. †

III. To Vauquelin we are indebted for an analysis of the ^{Excrements} fixed parts of the excrements of fowls, and a comparison ^{of fowls.} of them with the fixed parts of the food; from which some very curious consequences may be deduced.

He found that a hen devoured in ten days 11111·843 grains troy of oats. These contained

Phosphate of lime	126·509 grains
Silica	219·548
	356·057

During these ten days she laid four eggs; the shells of which contained 98·779 grains phosphate of lime, and 453·417 grains carbonate of lime. The excrements emitted during these ten days contained 175·529 grains phosphate of lime, 58·194 grains of carbonate of lime, and 185·266 grains of silica. Consequently the fixed parts thrown out of the system during these ten days amounted to

* Gehlen, iii. 287.
 † Ibid. 295, 313.

† Ibid. 321.

Book V.

Phosphate of lime	274·305 grains.
Carbonate of lime	511·911
Silica	185·266
	<hr style="width: 100%;"/>
Given out	971·482
Taken in	356·057
	<hr style="width: 100%;"/>
Surplus	615·425

Consequently the quantity of fixed matter given out of the system in ten days exceeded the quantity taken in by 615·425 grains.

The silica taken in amounted to	219·548 grains.
That given out was only	185·266
	<hr style="width: 100%;"/>
Remain	34·282

Consequently there disappeared 34·282 grains of silica.

The phosphate of lime taken in was	136·509 grains.
That given out was	274·305
	<hr style="width: 100%;"/>
	137·796

Fowls seem to form lime and phosphorus.

Consequently there must have been formed, by digestion in this fowl, no less than 137·796 grains of phosphate of lime, besides 511·911 grains of carbonate. Consequently lime (and perhaps also phosphorus) is not a simple substance, but a compound, and formed of ingredients which exist in oat-seed, water, or air, the only substances to which the fowl had access. Silica may enter into its composition, as part of the silica had disappeared; but if so, it must be combined with a great quantity of some other substance.*

These consequences are too important to be admitted without a very rigorous examination. The experiment must be repeated frequently, and we must be absolutely certain that the hen has no access to any calcareous earth, and that she has not diminished in weight; because in that case some of the calcareous earth, of which part of her body is composed, may have been employed. This rigour is the more necessary, as it seems pretty evident, from experiments made long ago, that *some* birds at least cannot produce eggs unless they have access to calcareous earth. Dr. Fordyce found, that if the canary bird was not supplied with lime at the time of her lay-

* Ann. de Chim. xxix. 61.

ing, she frequently died, from her eggs not coming forward properly.* He divided a number of these birds at the time of their laying eggs into two parties: to the one he gave a piece of old mortar, which the little animals swallowed greedily; they laid their eggs as usual, and all of them lived; whereas many of the other party, which were supplied with no lime, died.†

Vauquelin also ascertained, according to Fourcroy, that pigeons' dung contains an acid of a peculiar nature, which increases when the matter is diluted with water; but gradually gives place to ammonia, which is at last exhaled in abundance.‡

IV. The white matter voided by dogs who feed chiefly on bones, was formerly used in medicine under the name of *album græcum*. It has not been examined by modern chemists; but is supposed to consist in a great measure of the earthy part of the bones used as food.§

SECT. XXVIII.

OF MORBID CONCRETIONS.

HARD substances occasionally make their appearance in History. different parts of the animal body, both in the solids and in the cavities destined to contain the fluids. In the first case, they are usually denominated *concretions* or *ossifications*; in the second, *calculi*. The knowledge of these bodies is of importance both to the physiologist and physician. The formation is an irregularity in the animal economy, and is likely therefore to throw light upon its functions; for we succeed best in detecting the secrets of Nature when she deviates from that regularity which she usually follows. They often produce the most excruciating diseases; and the sufferings of the patient can only be effectually relieved by removing their cause. The different animal concretions at present known Divisible into five classes. may be divided into five classes:

- | | |
|----------------------------|---------------------|
| 1. Ossifications. | 4. Urinary calculi. |
| 2. Intestinal concretions. | 5. Gouty calculi. |
| 3. Biliary calculi. | |

We shall take a view of each of these classes in order.

* On Digestion, p. 25.

† Ibid. p. 26.

‡ Fourcroy, s. 70.

§ Neumann's Chemistry, p. 585.

I. OSSIFICATIONS.

Book V.

All the concretions which make their appearance in the solids of the animal body may be comprehended under this title with propriety; because they have all a close resemblance to bone, being composed of similar constituents. The following are the most remarkable of these concretions.

1. Pineal concretions.

1. *Pineal concretions*.—It is well known to anatomists that small concretions like sand are often found lodged in that part of the brain called the *pineal gland*. It was suspected from analogy that they consisted chiefly of phosphate of lime; but Dr. Wollaston was the first who examined them chemically, and proved the truth of this opinion. He dissolved some of the sand in nitric acid, and evaporated the solution; small crystalline needles made their appearance, indicating the presence of phosphate of lime.*

Sometimes fatty looking concretions occur both in the human brain and in that of inferior animals. M. Lassaigne examined a concretion of this kind from the brain of a horse. It was white, slightly soft, and of the size of a nut. It dissolved partially in boiling alcohol, and the alcohol on cooling let fall a white matter in plates of a pearly lustre. It melted at 276° , did not stain paper like tallow, and did not combine with caustic alkalies. It is therefore the same with the peculiar matter of biliary calculi called *cholesterin* by Chevreul.† I lately received from Dr. Corkindale of Glasgow, a number of similar concretions which he extracted from the brain of a human subject.

2. Salivary concretions.

2. *Salivary concretions*.—Small concretions occasionally make their appearance in the salivary glands, especially the parotid and sublingual. From the experiments of Wollaston and of Fourcroy, we learn that the basis of these concretions is phosphate of lime, united to a membranous substance, which retains the shape of the concretion after the solution of the phosphate. This at least was the case in a small salivary calculus which I examined.‡ In a salivary concretion weighing $1\frac{1}{3}$ grain, examined by Dr. Bostock, the whole consisted

* See Dr. Wollaston's important paper on Urinary and Gouty Concretions, Phil. Trans. 1797, p. 386. I shall have occasion to refer to it very frequently in the subsequent part of this Section.

† Annals of Philosophy, xiii. 75.

‡ Wollaston, Phil. Trans. 1797, p. 386. Fourcroy, ix. 367.

of phosphate of lime, excepting a few films of matter which was considered as coagulated albumen.*

M. Lassaigne examined a salivary concretion from a cow. It was white, very hard, capable of being polished, about the size of a pigeon's egg, and its nucleus was an oat seed. It consisted of carbonate of lime, mixed with a little phosphate of lime and some animal matter. But a salivary concretion from a horse, examined by the same chemist, was of quite a different nature. It was white, soft, elastic, and had exactly the form of the canal in which it was lodged. Cold water extracted from it a little albumen with some carbonate and muriate of soda. It dissolved with facility in alkaline leys and burnt like horn. Hence it seems to have consisted chiefly of a substance possessing the properties of coagulated albumen.†

3. *Pancreatic concretions.*—The hard substances sometimes found in the pancreas are said to be of the same nature with those of the salivary glands. The assertion, I presume, is merely from analogy.

3. Pancreatic concretions.

4. *Pulmonary concretions.*—Many people subject to cough, or threatened with consumption, occasionally cough up small rounded white concretions. It is very common to find the lungs of such persons filled with similar bodies. I examined some of these bodies coughed up by a consumptive patient, and found them composed of phosphate of lime united to a thick membranous substance, which retained the form of the concretion. The same result had been obtained long before by Fourcroy.‡ It is probable that pulmonary concretions very frequently contain both phosphate and carbonate of lime. This, Dr. Henry informs me, was the case with several of these bodies examined by him. In some cases they appear to contain no phosphate. Thus a pulmonary calculus examined by Mr. Crumpton§ was composed of

4. Pulmonary concretions.

Carbonate of lime	82
Animal matter and water	18

100

Dr. Prout has examined many of these concretions, and has always found them composed of phosphate of lime, with some carbonate of lime, and an animal matter retaining the

* Nicholson's Jour. xiii. 374.

† Annals of Philosophy, xiii. 75.

‡ Ann. de Chim. xvi. 91.

§ Phil. Mag. xiii. 287.

Book V.

shape of the concretion. He is of opinion that pulmonary concretions composed of pure carbonate of lime do not exist.*

Pulmonary concretions from a cow labouring under plithisis pulmonaris examined by M. Lassaigne, were found to consist of phospliate of lime, with a little carbonate of lime cemented by means of a mucous membrane.†

Hepatic concretions.

5. *Hepatic concretions*.—The liver also is sometimes full of similar bodies. The shape of the hepatic concretions, as far as my observations go, is more irregular, and I have seen them of greater size than the pulmonary concretions. By my analysis, they are composed of phosphate of lime and a tough animal membranous matter.

6. *Calculi from the spleen*.—Some of these rare calculi were given to Dr. Henry, by the late Mr. Allan Burns, of Glasgow. They were small, pear-shaped, yellowish white, and composed of bone earth.‡

7. Concretions in the prostate.

7. *Concretions in the prostate*.—From the experiments of Dr. Wollaston, we learn that the concretions which sometimes form in the prostate gland have likewise phosphate of lime for their basis.

8. Ossifications.

8. The extremities of the muscles and the larger blood vessels sometimes harden, and assume the appearance of bone. It is believed that this change is a real ossification, or that these bodies are converted into real bone. I do not know whether any such ossifications have been analysed.

II. INTESTINAL CONCRETIONS.

History.

Concretions of very considerable size are sometimes found lodged in the stomach and intestines; seldom indeed in the human body, but more frequently in some of the inferior animals. Some of these bodies have acquired great celebrity under the name of *bezoars*; but their medical virtue, at least in this country, has long ago sunk to its just level. It is to Fourcroy and Vauquelin that we are indebted for almost all the knowledge of these concretions which we possess. They have distinguished no fewer than seven species of them, all composed of different ingredients. These species, distinguished by the constituents of each, are as follows.§

* Annals of Philosophy, xiv. 233.

† Ibid. xviii. 76.

‡ Ibid. xv. 117.

§ Ann. de Mus. d'Hist. Nat. iv. 331.

- | | |
|--------------------------------------|--------------|
| 1. Superphosphate of lime. | 4. Biliary. |
| 2. Phosphate of magnesia. | 5. Resinous. |
| 3. Ammonio-phosphate of
magnesia. | 6. Fungous. |
| | 7. Hairy. |

1. *Superphosphate of lime*.—The intestinal concretions belonging to this species are composed of concentric layers, easily separable from each other and very brittle. They redden vegetable blues, and are partially soluble in water. The layers are unequally thick, and differ in their colour.* They were found in the intestines of different mammalia.

2. *Phosphate of magnesia*.—These concretions are uncommon. They are semitransparent, and have usually a yellowish colour. Their specific gravity is 2.160. They are formed of layers less numerous and not so easily separated as those of the preceding species.†

3. *Phosphate of ammonia-and-magnesia*.—This species is the most common of the intestinal concretions. Its colour is grey or brown, and it is composed of crystals diverging like rays from a centre. It has some resemblance to calcareous spar. It contains abundance of animal matter. This species occurs frequently in the intestines of herbivorous animals, as the horse, the elephant, &c.

4. *Biliary*.—This is a species of concretion found frequently in the intestines of oxen, and likewise in their gall-bladder, and employed by painters as an orange-yellow pigment. Its colour is reddish-brown. It is not composed of layers, but is merely a coagulated mass, and appears to be but little different from the matter of bile. When heated it melts. It dissolves readily in alkalis. Alcohol dissolves it partially, and acquires a very bitter taste.‡

5. *Resinous*.—To this species belong many of the oriental bezoars, formerly so celebrated, obtained from the intestines of animals with which we are unacquainted. They are fusible and combustible, composed of concentric layers, smooth, soft, and finely polished. Fournroy and Vauquelin have distinguished two varieties: the first of a pale-green colour, a slightly bitter taste, almost completely volatile; giving by heat a solid tenacious matter, soluble in alcohol, and separating in crystals as the solution cools. This matter consists partly of bile, partly of resin. The second variety has a brown or

* Ann. de Mus. d'Hist. Nat. i. 102, and iv. 331.

† Ibid. iv. 332.

‡ Ibid. iv. 333.

Book V.

violet colour; its taste is not bitter; it does not dissolve in alcohol, but is soluble in alkalies. The solution becomes purple-red when allowed to dry in the open air. When distilled it yields a yellow sublimate, having the smell and taste of soot, and insoluble in water and alcohol.*

6. *Fungous*.—This species consists of concretions composed of pieces of the *boletus igniarius*, disposed in layers, and cemented by an animal matter. These pieces had been doubtless swallowed by the animals in whose intestines they were found.†

7. *Hairy*.—Balls of hair felted together, sometimes pure, sometimes covered with animal matter, and sometimes mixed with vegetable remains, occur very frequently in the intestines of animals.‡

8. *Ligniform*.—This eighth species must be added in consequence of the experiments of Berthollet. Among the presents sent to Bonaparte by the king of Persia were three bezoars, which were consigned to Berthollet for analysis. They all belonged to this species. They had an oval shape, and a very smooth surface. Their colour externally was greenish-black, internally brown. They were formed of irregular concentric layers. In the centre of one was found a collection of straws and other vegetable fragments; in that of the others, small pieces of wood about the size of a pin. Their specific gravity was 1.463. They were insoluble in water, alcohol, and diluted muriatic acid. Potash ley dissolved them readily, and they were thrown down unaltered by muriatic acid. When distilled they yielded the products of wood, and left a quantity of charcoal in the retort, which, when incinerated, gave traces of sulphate of soda, muriate of soda, lime, and silica. Thus it appears that they possessed all the properties of pure woody fibre. They must have been formed in the stomach of the animals, and not in the alimentary canal.§

It may be worth while to mention here a collection of intestinal concretions, made by the late Dr. Monro, of Edinburgh, and which may be yet seen in the collection belonging to the anatomical class, in the University of Edinburgh. These concretions were from human subjects. They were almost all of a large size, and had in every case proved fatal. They have a brown colour, and very much the appearance of

* Ann. de Mus. d'Hist. Nat. iv. 334.

† Ibid. 335.

‡ Ibid. 336.

§ Mem. d'Arcueil, ii. 448.

felt. I extracted from them small quantities of different animal matters; but the great body of the concretion is composed of the small tuft of hairs which terminate the kernel of the oat-seed. Hence their origin is obvious.

Chap. II.

III. BILIARY CALCULI.

Hard bodies sometimes form in the gall-bladder, or in the duct through which the bile passes into the intestinal canal, and stop up the passage altogether. These concretions have got the name of *biliary calculi* or *gall-stones*. They naturally drew the attention of physicians, because it was soon ascertained that they occasioned the disease called *jaundice*. Accordingly they were examined and described by different chemists; and the facts ascertained before 1764 were collected by Haller in the sixth volume of his *Physiology*. Since that period, a treatise has been published on them by Vicq d'Azyr; and several new facts have been discovered by Poulletier de la Sale, Fourcroy, Gren, and Saunders. Thenard has lately subjected them to examination.*

History.

Such of the biliary calculi as have been hitherto examined with attention may be arranged under four classes.

Divisible
into four
classes.

1. The first kind comprehends those which have a white colour, and a crystallized, shining, lamellated structure. They consist of adipocire.

2. The second species are polygonal, of a light greyish-brown colour. Externally they have a covering composed of thin concentric layers, within a matter either crystallized or having the appearance of coagulated honey. They are composed chiefly of adipocire; but contain a small portion of brown matter, considered as matter of bile a little altered.

3. The third kind are of a brown colour, and are supposed to be composed of altered matter of bile.

4. The fourth comprehends those gall-stones which do not flame, but gradually waste away at a red heat.

Let us take a view of each of these classes of biliary calculi.

1. The first species of biliary calculi was pointed out for the first time by Haller in a dissertation published in 1749. Walther afterwards added several new facts; and at last it was accurately described by Vicq d'Azyr. It is almost always of an oval shape, sometimes as large as a pigeon's egg, but commonly about the size of a sparrow's; and for the most part,

1st Species
resembles
spermacell.

* Mem. d'Arcueil, i. 59.

Book V. only one calculus (when of this species) is found in the gall-bladder at a time. It has a white colour; and when broken, presents crystalline plates or striæ, brilliant and white like mica, and having a soft greasy feel. Sometimes its colour is yellow or greenish; and it has constantly a nucleus of inspissated bile.* Its specific gravity is inferior to that of water: Gren found the specific gravity of one 0·803.†

When exposed to the temperature of about 278°,‡ this crystallized calculus softens and melts, and crystallizes again when the temperature is lowered.§ It is altogether insoluble in water; but hot alcohol dissolves it with facility. Alcohol, of the temperature of 167° dissolves $\frac{1}{10}$ of its weight of this substance; but alcohol, at the temperature of 60° scarcely dissolves any of it.|| As the alcohol cools, the matter is deposited in brilliant plates resembling talc or boracic acid.** ‡ It is soluble in oil of turpentine.†† When melted it has the appearance of oil, and exhales the odour of melted wax: when suddenly heated it evaporates altogether in a thick smoke. It is soluble in pure alkalies, and the solution has all the properties of a soap. Nitric acid also dissolves it; but it is precipitated altered by water.‡‡ Indeed the greatest part separates as the liquid cools, and swims on the surface like drops of oil, and seems by the action of the acid to be brought nearly to the state of a resin.§§ Fourcroy, who first examined this peculiar matter of biliary calculi, has given it the name of *adipocire*, from its resemblance both to fat and to wax. Chevreul could not succeed in converting it into soap, by boiling it with potash ley. He has given it the name of *cholesterin*.

2d Species
nearly simi-
lar.

2. The second kind of biliary calculi is of a polygonal shape, and a number of them almost always exist in the gall-bladder together. To this, probably, they owe their peculiar form. Most commonly they have three blunt angles, and have some resemblance to two low tetrahedrons, applied base to base, with their edges and angles rounded off. They vary considerably in their specific gravity. A specimen seemingly belonging to this species, examined by Dr. Bostock, was of the specific gravity 0·900.|||| The mean specific gravity of six which I analysed was 1·061, and they all sunk in water. Their outer

* Fourcroy, Ann. de Chim. iii. 245.

† Chevreul, Ann. de Chim. xcv. 7.

‡ Ann. de Chim. ii. 180.

†† Gren, ibid. v. 187.

§§ Bostock, Nicholson's Jour. iv. 138.

† Ann. de Chim. v. 186.

§ Ann. de Chim. ii. 123.

** Ibid. iii. 256.

†† Fourcroy, ibid. iii. 247.

|||| Ibid. 136.

surface is smooth, and has a soft feel. When broken they exhibit a thin outer crust, composed of concentric layers, alternately crystallized in small rays inclining towards the centre. The nucleus, in all the specimens which I have examined, had the appearance of granulated honey.

These calculi, in their composition, differ but little from the last species, since they consist almost entirely of adipocire. In six gall-stones which I analysed, this matter amounted to at least $\frac{1}{8}$ ths of the whole. The residue was a reddish-brown substance, insoluble in alcohol. Nitric acid dissolved it readily, and formed a pink-coloured liquid, from which ammonia threw down no precipitate. Pure potash ley dissolved most of it readily when assisted by heat. From the solution, muriatic acid threw down a dark-green matter, which had a bitter taste, dissolved in alcohol, melted when heated, and exhibited most of the properties of picromel. The residue, insoluble in potash, was in grey flakes, and resembled albumen in such of its properties as could be traced. But as it never exceeded $\frac{1}{8}$ th of a grain, it was not possible to ascertain its nature with precision.

3. I have never had an opportunity of seeing any gall-stones from the human species belonging to the third kind, or composed entirely of altered yellow matter. But in the gall-stones examined by Thenard several were of this kind.

Sd. of altered yellow matter.

4. Concerning the fourth species of gall-stone very little is known with accuracy. Dr. Saunders tells us that he has met with some gall-stones insoluble both in alcohol and oil of turpentine; some of which do not flame, but become red, and consume to an ash like a charcoal.* Haller quotes several examples of similar calculi.† Probably they do not differ from the third kind.

5. The gall-stones of oxen are always of a yellow colour, and consist of the yellow matter of bile mixed with minute traces of bile which may be separated by water. They are tasteless, when thus washed, and insoluble in water and alcohol. They are used by painters, though their colour is not permanent, but soon changes to a brown.‡

* On the Liver, p. 112.

† *Physiol.* vi. 567

‡ Thenard, *Mem. d'Arcueil*, i. 59.

IV. URINARY CALCULI.

Book V.

It is well known that concretions not unfrequently form in the bladder, or the other urinary organs, and occasion one of the most dismal diseases to which the human species is liable.

History.

These concretions were distinguished by the name of *calculi*, from a supposition that they are of a stony nature. They have long attracted the attention of physicians. Chemistry had no sooner made its way into medicine than it began to exercise its ingenuity upon the urinary calculi; and various theories were given of their nature and origin. According to Paracelsus, who gave them the ridiculous name of *duelech*, urinary calculi were composed of a mucilaginous tartar which floated in the blood vessels. The schoolmen, on the other hand, considered them as a peculiar mucilage concocted and petrified by the heat of the body. These opinions were ably refuted by Van Helmont in his Treatise De Lithiasi, which contains the first attempt towards an analysis of urine and urinary calculi; and, considering the period at which it was written, is certainly possessed of uncommon merit. He demonstrates that the materials of calculi exist in urine: he considers them as composed of a volatile earthy matter and the saline spirit of urine which coagulates instantaneously when they combine together; but which are prevented from combining in healthy people by the presence of what he calls *scoria*, which saturates the salt of urine.* Boyle extracted from calculi, by distillation, oil and a great quantity of volatile salts. Boerhaave supposed them compounds of oil and volatile salts. Slare attempted a chemical analysis of them.† Hales extracted from them a prodigious quantity of air. He gave them the name of *animal tartar*, pointed out several circumstances in which they resemble common tartar, and made many experiments to find a solvent of them.‡ Drs. Whytt and Alston pointed out alkalies as solvents of calculi. It was an attempt to discover a more perfect solvent that induced Dr. Black to make those experiments which terminated in the discovery of the nature of the alkaline carbonates.

Such was the state of the chemical analysis of calculus, when in 1776, Scheele published a dissertation on the subject in the Stockholm Transactions; which was succeeded by some

* De Lithiasi, p. 21.

† Phil. Trans. xvi. 140.

‡ Veget. Stat. ii. 189.

remarks of Bergman. These illustrious chemists completely removed the uncertainty which had hitherto hung over the subject, and ascertained the nature of the calculi which they examined.

Since that time considerable additional light has been thrown upon the nature of these concretions by the labours of Austin, Walther, Brugnatelli, Pearson, &c. But the most important additions to our knowledge of calculi were made by Dr. Wollaston. That skilful chemist distinguished them into species, and ascertained the composition of each, pointing out several new constituents which had not been suspected before.* The dissertation on calculi by Fourcroy and Vauquelin is no less important. Most of their results indeed had been anticipated by Wollaston: but they analyzed about 500 calculi with precision; a number sufficient to enable them to make a more exact arrangement of them than could be obtained, till the frequency of the occurrence of every ingredient was known. They detected also some substances in calculi which had not been observed in them before. A valuable dissertation on the subject has been published by Mr. Brande. He examined 150 calculi in the Hunterian museum, and has rectified some mistakes of his predecessors, and added several new facts.† Dr. Henry, more lately, has examined the constituents of 187 calculi.‡

Urinary calculi are usually spheroidal or egg-shaped; sometimes they are polygonous, or resemble a cluster of mulberries; and in that case they are distinguished by the epithet *mulberry*. Their size is various; sometimes they are very small, and sometimes as large as a goose-egg, or even larger. The colour of some of them is a deep-brown resembling that of wood. In some cases they are white, and not unlike chalk; in others, of a dark-grey, and hard. These different colours are often intermixed, and occur of various degrees of intensity. Their surface in some cases is polished like marble; in others, rough and unequal; sometimes they are covered with semitransparent crystals. Their specific gravity varies from 1.213 to 1.976.§

The substances hitherto discovered in urinary calculi are the following:

* Phil. Trans. 1797.

† Ibid. 1808.

‡ Annals of Philosophy, xv. 107.

§ Fourcroy, x. 213.

1. Uric acid.
2. Urate of ammonia.
3. Phosphate of lime.*
4. Phosphate of magnesia and ammonia.
5. Oxalate of lime.
6. Muriate of ammonia.
7. Magnesia.
8. Phosphate of iron.
9. Silica.
10. Nephryn.
11. Cystic oxide.
12. Mucus.
13. Lime.

1. Uric acid.

1. Uric acid was first discovered in calculi by Scheele. The greater number of those concretions hitherto analyzed consisted of it. All those analyzed by Scheele were composed of it entirely. Of 300 calculi analyzed by Dr. Pearson, scarcely one was found which did not contain a considerable quantity of it, and the greater number manifestly were formed chiefly of it. Fourcroy and Vauquelin found it also in the greater number of the 500 calculi which they analyzed. Of 150 examined by Brande only 16 were composed entirely of this acid; but almost the whole of them contained this acid. Of 187 examined by Dr. Henry 71 were composed chiefly of uric acid.

The presence of this acid may be easily ascertained by the following properties: The calculi formed of it are brown, polished, and resemble wood. A solution of potash or soda dissolves it readily, and it is precipitated by the weakest acids. The precipitate is soluble in nitric acid; the solution, when evaporated to dryness, is of a pink colour, and tinges the skin red.

2. Urate of ammonia.

2. *Urate of ammonia.* The existence of calculi composed of this constituent was announced by Fourcroy; but stated to be rare. Mr. Brande did not find any such substance in any of the calculi which he examined. On that account he denied its existence in calculi; and his opinion has been generally adopted by British chemists. But Dr. Prout has lately found a calculus, extracted from a boy about two years of age, composed of it. It had a greenish-clay colour and was composed

* Brugnatelli found also phosphate of lime, with excess of acid in calculi. See Ann. de Chim. xxxii. 183.

of concentric layers; it was soluble in water, gave out abundance of ammonia when treated with the fixed alkalies and decrepitated strongly before the blow-pipe. It occasioned great irritation.*

3. Phosphate of lime was observed in calculi by Bergman; afterwards it was found in abundance by Pearson, and more lately by Fourcroy and Vauquelin. Dr. Wollaston was the first who observed calculi composed of it entirely. The calculi which he examined were brown, and so smooth externally as to appear polished. They were composed of laminae easily separable into concentric crusts. In the calculi observed by Fourcroy and Vauquelin, the phosphate of lime was white, without lustre, friable, stained the hands, paper, and cloth. It had very much the appearance of chalk, broke under the forceps, was insipid and insoluble in water. It is soluble in nitric, muriatic, and acetic acids, and is again precipitated by ammonia, fixed alkalies, and oxalic acid. It is intimately mixed with a gelatinous matter, which remains under the form of a membrane when the earthy part is dissolved by very diluted acids.†

3. Phosphate of lime.

4. The calculi containing phosphate of magnesia and ammonia were first distinguished by Mr. Tennant; but for their first analysis we are indebted to Dr. Wollaston. He first pointed out the presence of the triple phosphate, ascertained its characters, and showed how it might be imitated by art. ‡ Fourcroy and Vauquelin, who detected it soon after, though their experiments were not published before the year 1800, suppose that it owes its existence to a commencement of putrefaction of the urine in the bladder. It occurs in white, semitransparent, lamellar layers; sometimes it is crystallized on the surface of the calculi in prisms, or what are called *dog-tooth* crystals. Its powder is of a brilliant white; it has a weak sweetish taste; it is somewhat soluble in water, and very soluble in acids, though greatly diluted. Fixed alkalies decompose it, leaving magnesia undissolved.

4. Phosphate of magnesia and ammonia.

It never forms entire calculi. Sometimes it is mixed with phosphate of lime, and sometimes layers of it cover uric acid or oxalate of lime. It is mixed with the same gelatinous matter as phosphate of lime.§

* Annals of Philosophy, xv. 436.

† Fourcroy, Ann. de Chim. xxxii. 218.

‡ Phil. Trans. 1797.

§ Fourcroy, Ann. de Chim. xxxii. 219.

Book V.
5. Oxalate
of lime.

5. Oxalate of lime was first detected in calculi by Dr. Wollaston.* The calculi containing it had been distinguished by the name of *mulberry*; but scarcely any of their characters was known till they attracted his attention. In all the calculi which he examined, the oxalate was mixed with phosphate of lime, and usually also with uric acid; but Fourcroy and Vauquelin found several in which it was united only to animal matter. It commonly forms a very hard calculus, of a dark-green colour, difficult to saw asunder, admitting a polish like ivory, exhaling, when sawed, an odour like that of semen. Insoluble and indecomposable by alkalies; soluble in very diluted nitric acid, but slowly and with difficulty. It may be decomposed by the carbonates of potash and soda. When burnt, it leaves behind a quantity of pure lime, amounting to one-third of its weight, which may be easily recognised by its properties. †

6. Muriate
of ammonia.

6. Muriate of ammonia was first detected in calculi by Brande. Fourcroy and Vauquelin had announced the existence of urate of ammonia in calculi. They founded their opinion chiefly on the easy solubility of the supposed urate in alkaline leys, and the smell of ammonia given out during the solution. Brande found that if such calculi were treated with water, a portion was dissolved, and the residue consisted of pure uric acid. The matter dissolved was partly nephryn and partly muriate of ammonia. To the first of these bodies he ascribed the easy solubility of the calculus, and to the second the smell of ammonia which it gave out during its solution. ‡

7 and 8. The presence of magnesia and phosphate of iron in calculi, has been announced by Alemanni, an Italian apothecary. He analyzed an urinary calculus, and obtained from it

Magnesia	51·00
Silica	20·00
Phosphate of iron	21·84
Carbonate of magnesia	4·00
Loss	3·16
	100·00

As no details are given, we have no means of ascertaining how far these results are to be considered as correct. §

* Phil. Trans. 1797.

† Fourcroy, Ann. de Chim. xxxii. 220.

‡ Phil. Mag. xxxii. 171.

§ Ann. de Chim. lxxv. 222.

9. Silica has only been found in two instances by Fourcroy and Vauquelin, though they have analyzed about 600 calculi. It must therefore be considered as a very uncommon ingredient of these concretions. In the two instances in which it occurred, it was mixed with phosphate of lime.* The two calculi containing it were mulberry ones, but of a lighter colour than usual, extremely hard, and difficult to saw or reduce to powder. The presence of silica may be easily detected by its fusibility into glass with fixed alkalies, and its other well known properties. Professor Wurzer has detected silica, to the amount of one per cent. in a calculus which he analyzed. It was composed of

Chap. II.
9. Silica.

Phosphate of lime	17.33
Uric acid	75.34
Animal matter	6.33
Silica	1.00
	100.00 †

Silica was found also in an excessive proportion in the calculus analyzed by Alemani, the result of which has been stated above.

10. Nephryn was suspected to exist in calculi by Fourcroy and Vauquelin, but was first detected by Mr. Brande. He obtained it by digesting the calculi supposed to contain it in water or alcohol, and evaporating the solution. The proportion of it seems sometimes to be great. In one instance Mr. Brande found 36 per cent. of animal matter, the greatest part of which no doubt was nephryn. †

10. Nephryn.

11. The substance called cystic oxide was discovered, and its nature ascertained, by Dr. Wollaston. A small calculus, passed from the bladder of a man, was entirely composed of it. It was white, dense, brittle, and had pretty much the appearance of magnesian lime-stone. The characters of this new and peculiar animal substance, as ascertained by Dr. Wollaston, are the following: 1. It dissolves and combines equally with acids and alkalies, and crystallizes with both. 2. It is precipitated from nitric acid by alcohol. 3. It does not become red when treated with nitric acid. 4. It produces no change on vegetable blues. 5. It is insoluble in water,

11. Cistic oxide.

* Fourcroy, *ibid.* xxxii. 221.
 † Gehlen's *Jour.* Second Series, ii. 265.
 ‡ *Phil. Mag.* xxxii. 171.

Book V:

alcohol, and ether. 6. When distilled it yields carbonate of ammonia and oil, and leaves only a very small fixed residue, which is phosphate of lime. This substance has only been found hitherto in a few calculi.

Dr. Wollaston met with two composed of it. Two of those in Dr. Henry's collection consist of it; and a third contains a nucleus of it. This peculiar substance, according to the analysis of Dr. Prout, is composed of

6 atoms hydrogen	= 0.75
6 atoms carbon	= 4.50
8 atoms oxygen	= 8.00
1 atom azote	= 1.75
21	15.00

So that it contains at least 21 atoms, and the weight of an integral particle of it is 15.

12. Animal matter.

12. All calculi contain an animal matter, which serves as a cement to glue the other ingredients together. This matter is often very small in quantity, though sometimes it is very considerable. Fourcroy and Vauquelin at first considered it as albumen, but they have lately announced that they have ascertained it to be mucus; an opinion much more likely to be correct, as mucus is secreted by the internal surface of the bladder, and of course is always present in more or less quantity.

13. Lime.

13. *Lime*. I have obtained from a calculus extracted from a gouty person, by simply exposing it to a red heat and then digesting it in water, a solution of quicklime.* Dr. Prout has more lately made a similar observation.† I do not know in what state of combination the lime existed before the heat, unless it was united to uric acid.

Such are the component parts of urinary calculi; but, as these substances often occur mixed together in various ways, it is a point of some consequence to be able to ascertain the component parts of the different calculi from their appearance, and to know what substances usually associate together; as this knowledge may lead hereafter to the true theory of their formation, and thus perhaps enable us to prevent that most dreadful of all diseases. Fourcroy and Vauquelin, after an examination of more than 600 calculi, have divided them into

* Annals of Philosophy, xiv. 468.

† Ibid. xv. 437.

three genera and twelve species. The following is a view of their arrangement: Chap. II.

GENUS I. *Calculi composed of one ingredient.*

- Sp. 1. Uric acid.
- Sp. 2. Urate of ammonia.
- Sp. 3. Oxalate of lime.

Arrange-
ment of the
calculi.

GENUS II. *Calculi composed of two ingredients.*

- Sp. 1. Uric acid and the phosphates in layers.
- Sp. 2. Ditto, mixed together.
- Sp. 3. Urate of ammonia and the phosphates in layers.
- Sp. 4. Ditto, mixed together.
- Sp. 5. The phosphates mixed or in layers.
- Sp. 6. Oxalate of lime and uric acid in layers.
- Sp. 7. Oxalate of lime and the phosphates in layers.

GENUS III. *Calculi containing more than two ingredients.*

- Sp. 1. Uric acid, the phosphates, and oxalate of lime.
 - Sp. 2. Uric acid, urate of ammonia, the phosphates, silica.
- Let us take a view of each of these species.

Sp. 1. *Uric acid.* Colour that of wood, various shades of yellow or red. Texture laminar and radiated, compact and fine. Specific gravity from 1.5 to 1.786; sometimes so low as 1.276. Surface usually smooth and polished. Completely soluble in fixed alkaline leys, without emitting any odour of ammonia. This species is the most common. One fourth of the 600 calculi examined by Fourcroy and Vauquelin belonged to it. Description
of the spe-
cies.

Sp. 2. *Urate of ammonia.* Colour brownish white. Texture laminar; laminae easily separated from each other. Specific gravity from 1.225 to 1.720. Surface often crystallized. Soluble in hot water, especially when reduced to powder. Soluble in fixed alkalis, while ammonia is evolved. This species is uncommon. It seems confined to persons below the age of puberty. This calculus, as appears from Dr. Prout's observations, is peculiarly irritating.

Sp. 3. *Oxalate of lime.* Colour soot brown. Texture dense and hard, resembling ivory. Surface unequal, and full either of pointed or rounded protuberances. Hence the epithet *mulberry*, by which the calculi of this species are distinguished. Specific gravity from 1.428 to 1.976. When sawn, exhales the odour of semen. When calcined, leaves a

Book V. residuum of carbonate of lime. Insoluble in alkalies; soluble with difficulty in acids. This species occurs frequently.

Sp. 4. *Uric acid and the phosphates in layers.* Surface white like chalk, and friable or sparry and semitransparent, according as the outermost coat is phosphate of lime or of magnesia. Often large. When cut, they present a nucleus of uric acid. Sometimes this nucleus is covered with alternate layers of the two phosphates. Specific gravity very variable. This species is not uncommon. About $\frac{1}{30}$ th of the calculi examined by Fourcroy and Vauquelin belonged to it.

Sp. 5. *Uric acid and the phosphates mixed together.* This species varies in its appearance. Sometimes the component parts alternate in visible layers; at others they are too thin to be perceived by the eye, and can only be ascertained by analysis. Specific gravity from 1.213 to 1.739. This species is not uncommon. About $\frac{1}{40}$ th of the calculi analyzed by Fourcroy and Vauquelin belonged to it.

Sp. 6. *Urate of ammonia and the phosphates in layers.* This species resembles the fourth in its appearance; but its nucleus, instead of uric acid, is composed of urate of ammonia. It is not very common, and the calculi belonging to it are smaller than those of the fourth species.

Sp. 7. *Urate of ammonia and the phosphates mixed.* This species resembles the fifth, but may be distinguished by a colour less yellow, and by the ammonia emitted when the calculi are treated with potash. The calculi belonging to it are small and rather uncommon.

Sp. 8. *The two phosphates mixed or in layers.* Colour white like chalk. Texture laminar; friable, easily separated, and staining the surface of other bodies like chalk. Often mixed with thin layers of phosphate of magnesia and ammonia, sparry and semitransparent. Specific gravity from 1.138 to 1.471. Soluble in acids, insoluble in alkalies. About $\frac{1}{15}$ th of the calculi examined by Fourcroy and Vauquelin belonged to this species.

Sp. 9. *Oxalate of lime and uric acid in layers.* The calculi belonging to this species have a nucleus of oxalate of lime covered with a layer, more or less thick, of uric acid; or they consist of a small calculus of the third species covered with a layer belonging to the first species. They are easily distinguished when sawn in two. About $\frac{1}{30}$ th of the 600 calculi analyzed by Fourcroy and Vauquelin belonged to this species.

Sp. 10. *Oxalate of lime and the phosphates in layers.* The

calculi belonging to this species have a nucleus of oxalate of lime covered with a coat of the phosphates. Externally, therefore they are the same with the fourth and eighth species; while the internal nucleus belongs to the third species. They are easily distinguished when sawn in two. The calculi of this species are the most numerous of all, next to those of the first species. About $\frac{1}{3}$ th of those examined by Fourcroy and Vauquelin belonged to it.

Sp. 11. *Uric acid or urate of ammonia, the phosphates, oxalate of lime.* The calculi of this species have a nucleus of oxalate of lime, over this a coat of uric acid or urate of ammonia, or of a mixture of both; while the outermost coat is composed of the phosphates.

Sp. 12. *Uric acid, urate of ammonia, the phosphates, silica.* The calculi of this species bear a considerable resemblance to those of the last. The nucleus is composed of silica and phosphate of lime, then there is a coat of uric acid and urate of ammonia, and over all a coat of the phosphates. Only four calculi of this kind have been observed.

Such is the arrangement of Fourcroy and Vauquelin, certainly the most complete which has hitherto appeared. Dr. Wollaston had previously given another, which in many respects deserves the preference in a practical view. It consists of a small number of genera, admirably well distinguished by their characters and constituents. It certainly deserves the peculiar attention of surgeons. He divides calculi into four kinds. They are as follows:

Wollaston's
arrangement.

1. *Uric calculi.* These are the calculi composed of uric acid, or which consist chiefly of that substance. They are either completely soluble in pure alkaline leys or nearly so.

2. *Fusible calculi.* Composed chiefly of phosphate of lime and phosphate of magnesia and ammonia. Before the blow-pipe they melt into enamel. They are completely soluble, or nearly so, in muriatic acid.

3. *Mulberry calculi.* Composed of oxalate of lime or of oxalate and phosphate of lime. In the state of powder they are slowly dissolved by muriatic acid. The residuum is uric acid.

4. *Bone earth calculi.* These, as the name implies, are composed chiefly of phosphate of lime. They are soluble in muriatic acid.

From the observations and experiments of Mr. Brande, it appears that the calculi formed in the kidneys almost always

Book V.

consist of uric acid and animal matter. Sometimes, but very rarely, they consist of oxalate of lime; and when calculi remain in the kidney a considerable time after they have been formed, their external coat consists sometimes of the triple phosphate of magnesia and phosphate of lime.*

Of 150 calculi found in the bladder, and analyzed by Mr. Brande, the following was the composition :

16	Composed of uric acid.
45 uric acid, with a small relative proportion of the phosphates.
66 the phosphates with a small proportion of uric acid.
12 the phosphates entirely.
5 uric acid with the phosphates and nuclei of oxalate of lime.
6 chiefly oxalate of lime.

150 †

Of 187 calculi examined by Dr. Henry

71	were composed chiefly of uric acid.
22 earthy phosphates.
11 oxalate of lime.
8 compound.
2 cystic oxide.
39 alternate layers of uric acid and phosphates.
16 oxalate and phosphate.
11 alternate layers of oxalate and uric acid.
7 oxalate, uric acid, and phosphates.

187 ‡

From the preceding account of the different urinary calculi, it appears that most of their component parts exist in urine. But little satisfactory is known concerning the manner in which these concretions are formed, or of the cause of their formation. Whenever any solid body makes its way into the bladder, it has been observed that it is soon encrusted with

* Phil. Mag. xxxii. 167.

† Ibid. 171.

‡ Annals of Philosophy, xv. 108.

a coat of phosphate of lime; and this first nucleus soon occasions a calculus. Concretions of uric acid seldom or never form in the bladder, unless a primitive nucleus has originated in the kidneys. The gravel which is so frequently emitted by persons threatened with the stone consists always of this acid. As oxalic acid does not exist in urine, some morbid change must take place in the urine when calculi composed of oxalate of lime are deposited. Brugnatelli's discovery of the instantaneous conversion of uric acid into oxalic acid by chlorine, which has been confirmed by the experiments of Fourcroy and Vanquelin, throws considerable light upon the formation of oxalic acid in urine, by showing us that uric acid is probably the basis of it; but in what manner the change is actually produced, it is not so easy to say.

As our ignorance of the cause of urinary concretions puts it out of our power to prevent their formation, the ingenuity of physicians has been employed in attempting to discover substances capable of dissolving them after they have formed, and thus to relieve the human race from one of the most dreadful diseases to which it is subject. These attempts must have been in vain, or their success must have entirely depended upon chance, till the properties of the concretions themselves had been discovered, and the substances capable of dissolving them ascertained by experiment. I shall therefore pass over the numerous lithontriptics* which have been recommended in all ages, and satisfy myself with giving an account of the experiments made by Fourcroy and Vanquelin to dissolve stones by injections through the urethra, made after their analysis of the urinary calculi.

Attempts to discover a solvent of calculi.

The component parts of urinary calculi, as far as solvents are concerned, may be reduced under three heads:

1. Uric acid and urate of ammonia.
2. The phosphates.
3. Oxalate of lime.

Solvents.

1. A solution of pure potash and soda, so weak that it may be kept in the mouth, and even swallowed without pain, soon dissolves calculi composed of uric acid, or urate of ammonia, provided they be kept plunged in it.

2. The phosphates are very quickly dissolved by nitric or

* There is no room to believe that any of these medicines hitherto proposed, whether alkalies or acids, have any effect as solvents of calculi.

Book V

muriatic acid, so weak that it may be swallowed without inconvenience, and possessed of no greater acridness than urine itself.

3. Oxalate of lime is much more difficult of solution than the preceding substances. Calculi composed of it are slowly dissolved by nitric acid, or by carbonate of potash or soda, weak enough not to irritate the bladder: but the action of these substances is slow, and scarcely complete.

How applied.

These solvents injected into the bladder repeatedly, and retained in it as long as the patient can bear their action without inconvenience, ought to act upon the stone and gradually dissolve it. The difficulty, however, is to determine the composition of the calculus to be acted upon, in order to know which of the solvents to employ. But as no method of deciding this point with certainty is at present known, we must try some one of the solvents for once or twice, and examine it after it has been thrown out of the bladder. Let us begin, for instance, with injecting a weak solution of potash; and after it has remained in the bladder half an hour, or longer if the patient can bear it, let the liquid, as soon as passed, be filtered and mixed with a little muriatic acid; if any uric acid has been dissolved, a white solution will make its appearance. This precipitate is a proof that the calculus is composed of uric acid. If it does not appear, after persevering in the alkaline solution for some days, then there is reason to expect the presence of the phosphates; of course a weak muriatic acid solution should be injected. After this solution is emitted, let it be mixed with ammonia, and the phosphate of lime will precipitate, if the calculus be composed of it. If neither of these solutions take up any thing, and if the symptoms are not alleviated, we must have recourse to the action of nitric acid, on the supposition that the calculus is composed of oxalate of lime. These different solutions must be persisted in, and varied occasionally as they lose their efficacy, in order to dissolve the different coats of the calculus. Such are the methods pointed out by Fournroy and Vauquelin. It is scarcely necessary to observe, that the bladder should be evacuated of urine previous to the injections, and that the injections should be previously heated to the temperature of the body.

Calculi of inferior animals.

The calculi found in the bladder of the inferior animals have been examined by different chemists, especially by Fournroy and Vauquelin, and by Dr. Pearson, and Mr. Brande.

As far as experiment has hitherto gone, they consist chiefly of the three following species :

1. *Carbonate of lime.* These calculi have a white colour, are opaque, and are easily recognized by the action of acids. They occur most frequently in the bladder of the graminivorous animals, and seem to have been first detected by Fourcroy and Vauquelin. The lime is cemented by an animal matter. Calculi belonging to this species have been found in the bladder of the horse,* the sow,† the rabbit,* the ox.‡

2. *Earthy phosphate.* These calculi belong chiefly to the carnivorous animals, and seem first to have been examined by Dr. Pearson. They consist sometimes of phosphate of lime and animal matter, sometimes of a mixture of phosphate of lime and phosphate of ammonia and magnesia and animal matter. They have been found in the bladder of the dog,§ the hog,|| the rat,** and the cat:** and a calculus from a horse, examined by Dr. Pearson, was obviously composed of the two phosphates. I found the calculus from a hog composed entirely of phosphate of lime.††

3. *Oxalate of lime.* Calculi composed of this substance have been recently detected by Fourcroy and Vauquelin from the bladder of the dog and the rat.‡‡

The following table exhibits a view of the constituents of the calculi of different animals analysed by Mr. Brande :

	Horse.§§			Sheep.	Dog.		Hog.	Rabbit.
Phosphate of lime	76	45	60	72	61	80		39
Carbonate of lime	22	10	40	20		20	90	42
Phosphate of magnesia and ammonia		28			30			
Animal matter		15		8	6			19
Loss	2	2					10	
	100	100	100	100	100	100	100	100

Several calculi from the bladder of the ox, likewise ex-

* Pearson. † Brugnatelli and Fourcroy. ‡ Fourcroy.
 § Pearson. || Bartholdi. ** Fourcroy and Vauquelin.
 †† Annals of Philosophy, ii. 59.
 ‡‡ Ann. de Mus. d'Hist. Nat. iv. 338.
 §§ The first calculus in the table was from the kidney, all the rest were from the bladder.

Book V.

amined by Mr. Brande, were composed of carbonate of lime and animal matter.*

Thus it appears that the calculi of carnivorous animals resemble the human in their composition, excepting that uric acid has never been detected in them.

V. GOUTY CONCRETIONS.

It is well known that concretions occasionally make their appearance in joints long subject to gout. These concretions, from their colour and softness, have received the name of *chalk-stones*. They are usually small, though they have been observed of the size of an egg.† It had long been the opinion of physicians that these concretions were similar to the urinary calculi. Of course, after the discovery of uric acid by Scheele, it was usual to consider the gouty chalk stones as collections of that acid. They were subjected to a chemical analysis by Dr. Wollaston in 1797, who found them composed of uric acid and soda.

Chalk stones.

Consist of urate of soda.

Gouty concretions are soft and friable. Cold water has little effect upon them; but boiling water dissolves a small portion. If an acid be added to this solution, small crystals of uric acid are deposited on the sides of the vessel.

These concretions are completely soluble in potash when the action of the alkaline solution is assisted by heat.

When treated with diluted sulphuric or with muriatic acid, the soda is separated; but the uric acid remains, and may be separated by filtration. The liquid, when evaporated, yields crystals of sulphate or muriate of soda, according to the acid employed. The residuum possesses all the characters of uric acid. When distilled it yields ammonia, prussic acid, and the acid sublimate of Scheele. When dissolved in a little nitric acid, it tinges the skin of a rose colour, and when evaporated leaves a rose-coloured deliquescent residuum. It is soluble in potash, and may be precipitated by any acid, and by ammonia; first in the state of a jelly, and then breaking down into a white powder.

When uric acid, soda, and a little warm water, are triturated together, a mass is formed, which, after the surplus of soda is washed off, possesses the chemical properties of gouty concretions.‡

* Phil. Mag. xxxii. 175.

† Severinus.

‡ See Dr. Wollaston's analysis, Phil. Trans. 1797.

SECT. XXIX.

MORBID SECRETIONS.

IN different diseases to which the animal body is subject, various fluids make their appearance which did not exist previously, at least under the form which they assume. Thus in dropsy, the cellular substance, and frequently also the cavities of the head, breast, or abdomen, are filled with a whitish liquid. When any part of the skin is irritated into a blister, the interval between the cutis and cuticle is filled with a transparent fluid: and when any part of the muscles or skin is wounded, the ulcer is soon covered with a thick matter called *pus*. A thin sanies exudes from cancers and carious bones. A chemical knowledge of these secretions cannot but be attended with advantage both to the physiologist and physician. Unfortunately little precise information respecting these fluids can be at present given. It will be worth while, however, to state the observations that have been made.

Chap. II.

Various.

I. PUS.

The liquid called *pus* is secreted from the surface of an inflamed part, and usually moderates and terminates the inflammation. It assumes different appearances according to the state of the sore. When it indicates a healing sore, it is called *healthy* or *good-conditioned* pus. The liquid possesses the following properties:

It is of a yellowish-white colour, and of the consistence of cream. Its taste is insipid, and it has no smell when cold. Before the microscope it exhibits the appearance of white globules swimming in a transparent fluid.* Its specific gravity varies from 1·031 to 1·033.†

Properties
of healthy
pus.

It produces no change on vegetable blues.‡ When exposed to a moderate heat it gradually dries, and assumes the appearance of horn. When exposed to destructive distillation, Bergman obtained first about one-fourth of the pus in the state of insipid water. On increasing the fire, a liquid came over, containing abundance of ammonia, and accompanied by gaseous bodies, which were not examined. Some concrete

* Home on Ulcers.

† Pearson, Nicholson's Jour. xxx. 17.

‡ Cruickshanks.

carbonate of ammonia sublimed, accompanied by empyreumatic oil. A light brilliant coal remained of difficult incineration. The ashes gave traces of iron.*

When pus is left exposed to the air, it gradually becomes acid, according to Hildebrant; and Haller affirms that it sometimes gives a red colour to litmus even when recent. When thrown into water it sinks to the bottom. When agitated the mixture becomes milky; but the pus separates again when allowed to remain undisturbed. However, by repeated agitation, and especially by the application of heat, a milky liquid is obtained, which passes in that state through the filter.

Alcohol thickens pus, but does not dissolve it; neither does pus unite with oils.

Sulphuric acid dissolves it, and forms a purple coloured solution. When diluted with water, the dark colour disappears, and the pus separates; either sinking to the bottom, or rising to the surface, according to the quantity of water added, and the time that the solution has been allowed to stand. Diluted sulphuric acid does not act upon it.

Concentrated nitric acid effervesces with it, and forms a yellow solution, which when recent is decomposed by water, the pus subsiding in the state of grey flakes.

Muriatic acid also dissolves it when heated, and the pus is separated by water.

With the fixed alkaline leys it forms a whitish ropy fluid, which is decomposed by water; the pus precipitating. Pure ammonia reduces it to a transparent jelly, and gradually dissolves a considerable portion of it.†

When nitrate of silver is dropped into the solution of pus in water, a white precipitate separates. Nitrate of mercury and corrosive sublimate occasion a much more copious flaky precipitate.‡

How distinguished
from mucus.

Such are the properties of healthy pus hitherto observed by chemists. They indicate a considerable analogy with albumen. Various observations have been made to enable physicians to distinguish pus from the mucus of the internal cavities, especially of the lungs. In cases of copious expectoration, it is sometimes of consequence to know whether the matter thrown out of the lungs is pus or mucus. Mr. Charles Darwin made a set of experiments on the subject, and pointed out three criteria which distinguished pus. 1. Sulphuric acid

* Gren's Handbuch, ii. 426.

† Cruickshanks.

‡ Ibid.

dissolves it. When the solution is diluted the pus precipitates; but mucus treated in the same manner swims. But this distinction depends upon the quantity of water added, and is therefore ambiguous. 2. Pus is diffusible through diluted sulphuric acid, through water, and through brine; but mucus is not. 3. Alkaline leys dissolve pus; water precipitates pus thus dissolved, but not mucus. How far these two last distinctions can be trusted, is rather doubtful. Grasmeyer has proposed the following method, which he considers as complete. Triturate the substance to be tried with an equal quantity of warm water; then add to it an equal portion of a saturated solution of carbonate of potash, and set the mixture aside. If it contains pus, a transparent jelly subsides in a few hours; but this does not happen if only mucus be present.*

2. When the ulcer is ill-conditioned, the pus secreted in it possesses different properties. It has usually a fetid smell, is much thinner, and to a certain degree acrid. We are in possession of two sets of experiments on this unhealthy pus: one by Mr. Cruickshanks on the pus discharged from what is called the *hospital sore*; another by Dr. Crawford on the matter of *cancers*.

The pus from the hospital sore possesses most of the properties of healthy pus; but is distinguished by its odour, and by some shades of difference when exposed to the action of the metallic precipitates. Lime-water changes its fetid odour, but does not destroy it; sulphuric acid increases it, as do alcohol and the solution of oxide of arsenic in potash. Bark has no effect upon it; but it is destroyed by the nitrate of mercury and corrosive sublimate, by nitric acid, and by chlorine. Nitrate of silver does not destroy it. Mr. Cruickshanks supposes that the fetid smell is occasioned by the alteration of some part of true pus. He considers the pus of the hospital sore as a matter *sui generis*, which is capable of generating more, and even of producing an alteration in the system. Hence, to heal the sore, the matter must be destroyed, and prevented from appearing again. This was done by washing the sores with nitrate of mercury, diluted nitric acid, and chlorine at every dressing. This method constantly succeeded with Dr. Rollo, except when the sore was too large to admit it to be put in practice completely.

Pus of the
hospital
sore.

* Gren's Handbuch, ii. 433. Dr. Pearson's experiments on pus deserve the attention of the reader. See Nicholson's Jour. xxx. 17.

Book V.
Matter of
cancer.

3. The matter of cancer, examined by Dr. Crawford, gave a green colour to syrup of violets. Potash produced no change; but sulphuric acid extricated a gas which possessed many of the properties of sulphuretted hydrogen. This gas he supposes to exist in the matter united to ammonia. The presence of this compound explains the effects of the matter of cancer, and virulent matter in general, upon metallic salts. Dr. Crawford found that the odour of this matter was completely destroyed by chlorine; and therefore recommends it as a proper substance for washing cancerous ulcers.

4. Besides the species mentioned above, there are many others which we know from their effects to be peculiar, though we cannot find any chemical distinctions between them sufficiently well marked. But that they are specifically different cannot be doubted, if we consider that every one of them produces a disease peculiar to itself. The matter of small-pox, of venereal ulcers, of cow-pox, &c. may be mentioned as instances.

II. LIQUOR OF DROPSY.

The liquor which fills the cavities of the body in dropsy has a yellowish-green colour, and is sometimes turbid, sometimes nearly transparent. A few experiments on a colourless liquor extracted from a dropsical patient have been published by Wurzer: from which it appears to have contained albumen, mucus, sulphuretted hydrogen, phosphate of lime, muriate of soda, and soda.* To Dr. Bostock we are indebted for an examination of the colourless liquid obtained by puncturing a tumor in the spine, formed in the disease called *spina bifida*. It was slightly opaque, and did not alter vegetable blues; heat increased its opacity, but did not coagulate it. Its constituents were found to be as follows:

Water	97.8	
Muriate of soda	1.0	
Albumen	0.5	
Mucus	0.5	} proportions conjectural.
Gelatin	0.2	
Lime, a trace		

100.0†

A fluid of the same kind from the head of a child about ten years of age has been more lately examined by Dr. Prout. It

* Gehlen's Jour. v. 663.

† Nicholson's Jour. xiv. 145.

very faintly reddened litmus paper, and restored the yellow colour of turmeric paper rendered brown by ammonia. Its specific gravity was 1.0085. Its constituents were as follows:

Water	987.18	
Albumen, precipitated by nitric acid and heat . . .	1.66	
Substances soluble in alcohol (fatty adipociferous matter, lactate of soda?)	1.65	
Substances soluble in water, viz.		
Muriates of potash and soda	6.80	} 9.51
Sulphate of soda and some animal matter not coagulated by heat	2.71	
		1000.00*

III. LIQUOR OF BLISTERS.

The liquid which makes its appearance when the epidermis is raised into blisters is perfectly transparent and liquid. When the blisters are artificial, it is usually yellow, and has the odour of the blistering plaster. From the experiments of Margueron, we learn that it is composed of the same constituents as the serum of the blood. From 200 parts of this liquid he obtained,

Albumen	36	Constituents.
Muriate of soda	4	
Carbonate of soda	2	
Phosphate of lime	2	
Water	156	
	200†	

Thus I have given an account of all those secretions which have been attentively examined by chemists. The remainder have been hitherto neglected; partly owing to the difficulty of procuring them, and partly on account of the multiplicity of other objects which occupied the attention of chemical philosophers. It remains for us now to examine by what processes these different secretions are formed, how the constant waste of living bodies is repaired, and how the organs themselves are nourished and preserved. This shall form the subject of the following Chapter.

* Annals of Philosophy, xvi. 151.

† Ann. de Chim. xiv. 225.

CHAP. III.

OF THE FUNCTIONS OF ANIMALS.

Book V.

THE intention of the two last Chapters was to exhibit a view of the different substances which enter into the composition of animals, as far as the present limited state of our knowledge puts it in our power. But were our inquiries concerning animals confined to the mere ingredients of which their bodies are composed, even supposing the analysis as complete as possible, our knowledge of the nature and properties of animals would be imperfect indeed.

How are these substances arranged? How are they produced? What purposes do they serve? What are the distinguishing properties of animals, and the laws by which they are regulated?

Animals resemble vegetables.

Animals resemble vegetables in the complexness of their structure. Like them, they are machines nicely adapted for particular purposes, constituting one whole, and continually performing an infinite number of the most delicate processes. But neither an account of the structure of animals, nor of the properties which distinguish them from other beings, will be expected here: these topics belong entirely to the anatomist and physiologist. I mean in the present Chapter to take a view of those processes only that are concerned in the *production* of animal substances, which alone properly belong to Chemistry. The other functions are regulated by laws of a very different nature, which have no resemblance or analogy to the laws of Chemistry or Mechanics.

SECT. I.

OF DIGESTION.

Food necessary.

EVERY body knows that animals require food, and that they die sooner or later if food be withheld from them. There is indeed a very great difference in the quantity of food which animals require, and in the time which they can pass without it. In general, those animals which are most active require most, and those which are most indolent require least food. The cause of this is pretty obvious; the bodies of animals do not remain stationary, they are constantly wasting; and the

waste is proportional to the activity of the animal. Hence the body must receive, from time to time, new supplies, in place of what has been carried off. The use of food answers this purpose.

2. We are much better acquainted with the food of animals its nature. than of vegetables. It consists of almost all the animal and vegetable substances which have been treated of in this and the preceding Book: for there are but very few of them which some animal or other does not use as food. Man uses as food chiefly the muscles of animals, the seed of certain grasses, and a variety of vegetable fruits. Almost all the inferior animals have particular substances on which they feed exclusively. Some of them feed on animals, others on vegetables. Man has a greater range; he can feed on a very great number of substances. To enumerate these substances would be useless; as we are not able to point out with accuracy what it is which renders one substance more nourishing than another.*

Many substances do not serve as nourishment at all; † and not a few, instead of nourishing, destroy life. These last are called *poisons*. Some poisons act chemically, by decomposing the animal body. The action of others is not so well understood.

* Mr. Holt has ascertained a very curious fact, with respect to the *arana scenica*. It devours sulphate of zinc, and appears to deprive that salt of its acid. See *Annals of Philosophy*, xii. 454.

† M. Magendie, of Paris, made a curious set of experiments, which he considers as demonstrating the necessity of the presence of azote as a constituent of food. He fed dogs on sugar, allowing them no other article of nourishment except water. The animals swallowed the food with avidity; but they gradually became emaciated, and in about six weeks died of starvation. An ulcer always broke out in the corner of the eyes, and they lost their sight some days before their death. The same experiments were made upon other dogs, substituting gum and butter for sugar. The result in both cases was exactly the same.

But I do not consider these experiments as affording a decisive proof that sugar, gum, and butter are incapable of supporting the life of dogs. I think it not improbable that dogs, having been accustomed to a very different kind of food, may not be at first capable of digesting such articles. But if the change were brought about gradually instead of all at once, is it not probable that dogs might be made to live upon sugar, gum, and butter? A sheep is accustomed to live upon vegetable food, and it is incapable at first of digesting animal food; but it may be brought by degrees to relish animal food, and even to live upon it. It is said that in such cases the animal becomes incapable of digesting vegetable food. This effect of habit renders it very difficult to draw unexceptionable consequences from experiments on living animals.

Book V.
 Converted
 into chyme
 in the sto-
 mach.

3. The food is introduced into the body by the mouth, and almost all animals reduce it to a kind of pulpy consistence. In man, and many other animals, this is done in the mouth by means of teeth and the saliva with which it is there mixed; but many other animals grind their food in a different manner. After the food has been thus ground, it is introduced into the stomach, where it is subjected to new changes. The stomach is a strong soft bag, of different forms in different animals: in man it has some resemblance to the bag of a *bag-pipe*. In this organ the food is converted into a soft pap, which has no resemblance to the food when first introduced. This pap has been called *chyme*.

4. Since chyme possesses new properties, it is evident that the food has undergone some changes in the stomach, and that the ingredients of which it was composed have entered into new combinations. Now, in what manner have these changes been produced?

This change
 ascribed to
 the mecha-
 nical action
 of the sto-
 mach.

At first they were ascribed to the mechanical action of the stomach. The food, it was said, was still farther triturated in that organ: and being long agitated backwards and forwards in it, was at last reduced to a pulp. But this opinion, upon examination, was found not to be true. The experiments of Stevens, Reaumur, and Spallanzani, demonstrated that the formation of chyme is not owing to trituration; for on inclosing different kinds of food in metallic tubes and balls full of holes, in such a manner as to screen them from the mechanical action of the stomach, they found that these substances, after having remained a sufficient time in the stomach, were converted into chyme, just as if they had not been inclosed in such tubes. Indeed the opinion was untenable, even independent of these decisive experiments, the moment it was perceived that chyme differed entirely from the food which had been taken: that is to say, that if the same food were triturated mechanically out of the body, and reduced to pap of precisely the same consistence with chyme, it would not possess the same properties with chyme; for whenever this fact was known, it could not but be evident that the food had undergone changes in its composition.

Ascribed to
 fermenta-
 tion,


The change of food into chyme therefore was ascribed by many to *fermentation*. This opinion is indeed very ancient, and it has had many zealous supporters among the moderns. When the word *fermentation* was applied to the change produced on the food in the stomach, the nature of the process

called *fermentation* was altogether unknown. The appearances, indeed, which take place during that process had been described, and the progress and the result of it were known: but no attempt had been made to explain the cause of fermentation, or to trace the changes which take place during its continuance. All that could be meant, then, by saying that the conversion of food into chyme in the stomach is owing to fermentation, was merely, that the unknown cause which acted during the conversion of vegetable substances into wine or acid, or during their putrefaction, acted also during the conversion of food into chyme, and that the result in both cases was precisely the same. Accordingly, the advocates for this opinion attempted to prove, that air was constantly generated in the stomach, and that an acid was constantly produced; for it was the vinous and acetous fermentations which were assigned by the greater number of physiologists as the cause of the formation of chyme. Some indeed attempted to prove that it was produced by the putrefactive fermentation; but their number was inconsiderable, compared with those who adopted the other opinion.

Our ideas respecting fermentation are now somewhat more precise. It signifies a slow decomposition, which takes place when certain animal or vegetable substances are mixed together at a given temperature; and the consequent production of particular compounds. If, therefore, the conversion of the food into chyme be owing to fermentation, it is evident that it is totally independent of the stomach any farther than as it supplies temperature; and that the food would be converted into chyme exactly in the same manner if it were reduced to the same consistence, and placed in the same temperature out of the body. But this is by no means the case; substances

But without
reason.

are reduced to the state of chyme in a short time in the stomach, which would remain unaltered for weeks in the same temperature out of the body. This is the case with bones; which the experiments of Stevens and Spallanzani have shown to be soon digested in the stomach of the dog. Further, if the conversion of food into chyme were owing to fermentation, it ought to go on equally well in the stomach and œsophagus. Now, it was observed long ago by Ray and Boyle, that when voracious fish had swallowed animals too large to be contained in the stomach, that part only which was in the stomach was converted into chyme, while what was in the œsophagus re-

Book V.  remained entire; and this has been fully confirmed by subsequent observations.

Still farther, if the conversion were owing to fermentation, it ought always to take place equally well, provided the temperature be the same, whether the stomach be in a healthy state or not. But it is well known that this is not the case. The formation of chyme depends very much on the state of the stomach. When that organ is diseased, digestion is constantly ill performed. In these cases, indeed, fermentation sometimes appears, and produces flatulence, acid eructations, &c. which are the well-known symptoms of indigestion. These facts, which have been long known, are incompatible with the supposition, that the formation of chyme is owing to fermentation. Accordingly that opinion has been for some time abandoned, by all those at least who have taken the trouble to examine the subject.

Owing to
the action
of the gas-
tric juice.

The formation of chyme, then, is owing to the stomach; and it has been concluded, from the experiments of Stevens, Reaumur, Spallanzani, Scopoli, Brugnatelli, Carimini, &c. that its formation is brought about by the action of a particular liquid secreted by the stomach, and for that reason called *gastric juice*.

That it is owing to the action of a liquid is evident; because, if pieces of food be inclosed in close tubes, they pass through the stomach without any farther alteration than would have taken place at the same temperature out of the body; but if the tubes be perforated with small holes, the food is converted into chyme.

Nature of
the gastric
juice.

This liquid does not act indiscriminately upon all substances: for if grains of corn be put into a perforated tube, and a granivorous bird be made to swallow it, the corn will remain the usual time in the stomach without alteration; whereas if the husk of the grain be previously taken off, the whole of it will be converted into chyme. It is well known, too, that many substances pass unaltered through the intestines of animals, and consequently are not acted upon by the gastric juice. This is the case frequently with grains of oats, when they have been swallowed by horses entire with their husks on. This is the case also with the seeds of apples, &c. when swallowed entire by man; yet these very substances, if they have been previously ground sufficiently by the teeth, are digested. It appears, therefore, that it is chiefly the husk

or outside of these substances which resists the action of the gastric juice. We see, also, that trituration greatly facilitates the conversion of food into chyme. Chap. III.

The gastric juice is not the same in all animals; for many animals cannot digest the food on which others live. The *conium maculatum* (hemlock), for instance, is a poison to man instead of food, yet the goat often feeds upon it. Many animals, as sheep, live wholly upon vegetables; and if they are made to feed on animals, their stomachs will not digest them: others, again, as the eagle, feed wholly on animal substances, and cannot digest vegetables.

The gastric juice does not continue always of the same nature, even in the same animal; it changes gradually according to circumstances. Graminivorous animals may be brought to live on animal food; and after they have been accustomed to this for some time, their stomachs become incapable of digesting vegetables. On the other hand, those animals which naturally digest nothing but animal food may be brought to digest vegetables.

5. What is the nature of the gastric juice which possesses these singular properties? It is evidently different in different animals; but it is a very difficult task, if not an impossible one, to obtain it in a state of purity. Various attempts have indeed been made by very ingenious philosophers to procure it; but their analyses of it is sufficient to show us that they have never obtained it in a state of purity. Methods of procuring it.

The methods which have been used to procure gastric juice are, *first*, to kill the animal whose gastric juice is to be examined after it has fasted for some time. By this method Spallanzani collected 37 spoonfuls from the two first stomachs of a sheep. It was of a green colour, undoubtedly owing to the grass which the animal had eaten. He found also half a spoonful in the stomach of some young crows which he killed before they had left their nest.

Small tubes of metal pierced with holes, and containing a dry sponge, have been swallowed by animals; and when vomited up, the liquid imbibed by the sponge is squeezed out. By this method, Spallanzani collected 481 grains of gastric juice from the stomachs of five crows.

A *third* method consists in exciting vomiting in the morning, when the stomach is without food. Spallanzani tried this method twice upon himself, and collected one of the times 1 oz. 32 gr. of liquid; but the pain was so great that he did

Book V.

not think proper to try the experiment a third time. Mr. Gosse, however, who could excite vomiting whenever he thought proper, by swallowing air, has employed that method to collect gastric juice.

Spallanzani has observed that eagles throw up every morning a quantity of liquid, which he considers as gastric juice; and he has availed himself of this to collect it in considerable quantities.

It is almost unnecessary to remark how imperfect these different methods are, and how far every conclusion drawn from the examination of such juices must deviate from the truth. It is impossible that the gastric juice obtained by any one of these processes can be pure; because in the stomach it must be constantly mixed with large quantities of saliva, mucus, bile, food, &c. It may be questioned, indeed, whether any gastric juice at all can be obtained by these methods; for, as the intention of the gastric juice is to convert the food into chyme, in all probability its only secreted, or at least thrown into the stomach, when food is present.

Attempts to
analyse it.

We need not be surprised, then, at the contradictory accounts concerning its nature, given us by those philosophers who have attempted to examine it; as these relate not so much to the gastric juice, as to the different substances found in the stomach. The idea that the gastric juice can be obtained by vomiting, or that it is thrown up spontaneously by some animals, is, to say the least of it, very far from being probable.

According to Brugnatelli, the gastric juice of carnivorous animals, as hawks, kites, &c. has an acid and resinous odour, is very bitter, and not at all watery; and is composed of an uncombined acid, a resin, an animal substance, and a small quantity of muriate of soda.* The gastric juice of herbivorous animals, on the contrary, as goats, sheep, &c. is very watery, a little muddy, has a bitter saltish taste, and contains ammonia, an animal extract, and a pretty large quantity of muriate of soda.† Mr. Carminati found the same ingredients; but he supposes that the ammonia had been formed by the putrefaction of a part of their food, and that in reality the gastric juice of these animals is of an acid nature.‡

The accounts which have been given of the gastric juice of man are so various, that it is not worth while to transcribe

* Scopoli, Macquer's Dict.

† Ibid.

‡ Sennebier's Observations on Gastric Juice.

them. Sometimes it has been found of an acid nature, at other times not. The experiments of Spallanzani are sufficient to show that this acidity is not owing to the gastric juice, but to the food. He never found any acidity in the gastric juice of birds of prey, nor of serpents, frogs, and fishes. Crows gave an acidulous gastric juice only when fed on grain; and he finds that the same observation holds with respect to dogs, herbivorous animals, and domestic fowls. Carnivorous birds threw up pieces of shells and coral without alteration; but these substances were sensibly diminished in the stomachs of hens, even when inclosed in perforated tubes. Spallanzani himself swallowed calcareous substances inclosed in tubes; and when he fed on vegetables and fruits, they were sometimes altered and a little diminished in weight, just as if they had been put into weak vinegar; but when he used only animal food, they came out untouched. According to this philosopher, whose experiments have been by far the most numerous, the gastric juice is naturally neither acid nor alkaline. When poured on the carbonate of potash, it causes no effervescence.

Such are the results of the experiments on the juices taken from the stomach of animals. No conclusion can be drawn from them respecting the nature of the gastric juice. But from the experiments which have been made on the digestion of the stomach, especially by Spallanzani, the following facts are established:

The gastric juice attacks the surface of bodies, unites to the particles of them, which it carries off, and cannot be separated from them by filtration. It operates with more energy and rapidity the more the food is divided, and its action is increased by a warm temperature. The food is not merely reduced to very minute parts; its taste and smell are quite changed; its sensible properties are destroyed, and it acquires new and very different ones. This juice does not act as a ferment; so far from it, that it is a powerful antiseptic, and even restores flesh already putrefied. There is not the smallest appearance of such a process; indeed, when the juice is renewed frequently, as in the stomach, substances dissolve in it with a rapidity which excludes all idea of fermentation. Only a few air bubbles make their escape, which adhere to the elementary matter, and buoy it up to the top, and which are probably extricated by the heat of the solution.

Its effect on
the food.

Book V.
 Substances
 found in the
 stomach.

With respect to the substances contained in the stomach, only two facts have been perfectly ascertained: the first is, that the juice contained in the stomach of oxen, calves, sheep, invariably contains uncombined phosphoric acid, as Macquart and Vanquelin have demonstrated: the second, that the juice contained in the stomach, and even the inner coat of the stomach itself, has the property of coagulating milk and the serum of blood. Dr. Young found that seven grains of the inner coat of a calf's stomach, infused in water, gave a liquid which coagulated more than 100 ounces of milk; that is to say, more than 6857 times its own weight: and yet, in all probability, its weight was not much diminished.

What the substance is which possesses this coagulating property has not yet been ascertained; but it is evidently not very soluble in water: for the inside of a calf's stomach, after being steeped in water for six hours, and then well washed with water, still furnishes a liquor on infusion which coagulates milk: * And Dr. Young found that a piece of the inner coat of the stomach, after being previously washed with water, and then with a diluted solution of carbonate of potash, still afforded a liquid which coagulated milk and serum.

It is evident, from these facts, that this coagulating substance, whatever it is, acts very powerfully; and that it is scarcely possible to separate it completely from the stomach. But we know at present too little of the nature of coagulation to be able to draw any inference from these facts. An almost imperceptible quantity of some substances seems to be sufficient to coagulate milk: For Mr. Vaillant mentions, in his Travels in Africa, that a porcelain dish which he procured, and which had lain for some years at the bottom of the sea, possessed, in consequence, the property of coagulating milk when put into it; yet it communicated no taste to the milk, and did not differ in appearance from other cups.

It is probable that the saliva is of service in the conversion of food into chyme as well as the gastric juice. It evidently serves to dilute the food; and probably it may be serviceable also by communicating oxygen.

Chyme converted into chyle and excrement.

6. The chyme, thus formed, passes from the stomach into the intestines, where it is subjected to new changes, and at last converted into two very different substances, chyle and excrementitious matter.

* Dr. Young.

The *chyle* is a white-coloured liquid, very much resembling blood in its properties; with this difference that the globules which float in it instead of being red as is the case with the colouring globules of the blood, are white. But the nature and constituents of chyle have been described in the preceding chapter. Dr. Charles Smith, of New Jersey, relates an instance of a dropsy of the abdomen, in which the liquid accumulated appears to have been chyle. The patient, a boy twelve years of age, was tapped twice, and each time between seven and eight quarts of liquid abstracted. Its colour was chalky-white, and resembled milk pretty nearly, both in its taste, smell, and appearance. On standing a night, it threw up a good cream, though not so much in proportion as cow's milk usually does.*

Chap. III.
Nature of
chyle.

7. Concerning the process by which chyle is formed from chyme, scarcely any thing is known. It does not appear that the chyme is precisely the same in all animals; for those which are herbivorous have a greater length of intestine than those which are carnivorous. It is certain that the formation of the chyle is brought about by a chemical change, although we cannot say precisely what that change is, or what the agents are by which it is produced. But that the change is chemical, is evident, because the chyle is entirely different, both in its properties and appearance, from the chyme. The chyme, by the action of the intestines, is separated into two parts, chyle and excrement: the first of which is absorbed by a number of small vessels called *lacteals*; the second is pushed along the intestinal canal, and at last thrown out of the body altogether.

Its formation a chemical process.

After the chyme has been converted into chyle and excrement, although these two substances remain mixed together, it does not appear that they are able to decompose each other; for persons have been known seldom or never to emit any excrementitious matter *per anum* for years. In these, not only the chyle, but the excrementitious matter also, was absorbed by the lacteals; and the excrement was afterwards thrown out of the body by other outlets, particularly by the skin: in consequence of which, those persons have constantly that particular odour about them which distinguishes excrement. Now, in these persons it is evident that the chyle and excrement, though mixed together, and even absorbed to-

* Phil. Mag. ix. 16.

Book V. } gether, did not act on each other; because these persons have been known to enjoy good health for years, which could not have been the case had the chyle been destroyed.

Use of the bile.

It has been supposed by some that the decomposition of the chyme, and the formation of chyle, is produced by the agency of the bile, which is poured out abundantly, and mixed with the chyme, soon after its entrance into the intestines. If this theory were true, no chyle could be formed whenever any accident prevented the bile from passing into the intestinal canal: but this is obviously not the case; for frequent instances have occurred of persons labouring under jaundice, from the bile ducts being stopped, either by gall-stones or some other cause, so completely, that no bile could pass into the intestines; yet these persons have lived for a considerable time in that state. Consequently digestion, and therefore the formation of chyle, must be possible, independent of bile.

The principal use of the bile seems to be to separate the excrement from the chyle, after both have been formed, and to produce the evacuation of the excrement out of the body. It is probable that these substances would remain mixed together, and that they would perhaps even be partly absorbed together, were it not for the bile, which seems to combine with the excrement, and by this combination to facilitate its separation from the chyle, and thus to prevent its absorption. Fourcroy supposes that the bile, as soon as it is mixed with the contents of the intestinal canal, suffers a decomposition: that its alkali and saline ingredients combine with the chyle, and render it more liquid, while its albumen and picromel combine with the excrementitious matter, and gradually render them less fluid.* From the experiments of Berzelius on feces, detailed in the preceding chapter, it cannot be doubted that the constituents of the bile are to be found in the excrementitious matter: so that the ingenious theory of Fourcroy is so far probable. The bile also stimulates the intestinal canal, and causes it to evacuate its contents sooner than it otherwise would do; for when there is a deficiency of bile the body is constantly costive.

Chyle mixes with lymph,

8. The chyle, after it has been absorbed by the lacteals, is carried by them into a pretty large vessel, known by the name of the *thoracic duct*. Into the same vessel likewise is dis-

* Fourcroy, x. 46.

charged a transparent fluid, conveyed by a set of vessels which arise from all the cavities of the body. These vessels are called *lymphatics*, and the fluid which they convey is called *lymph*. In the thoracic duct, then, the chyle and the lymph are mixed together.

Very little is known concerning the nature of the *lymph*, as it is scarcely possible to collect it in any quantity. It is colourless, has some viscosity, and is said to be specifically heavier than water. Mr. Brande found that it did not coagulate when heated, but he separated albumen by means of the galvanic battery. He found in it also some common salt; but no iron.* From the observations of Berzelius it is probable that it contains the lactates. It is said to be coagulable by heat: if so, it contains albumen. Its quantity is certainly considerable, for the lymphatics are very numerous.

9. The chyle and lymph being thus mixed together, are conveyed directly into the blood vessels. The effect produced by their union in the thoracic duct is not known, but neither the colour nor external properties of the chyle are altered. In man, and many other animals, the thoracic duct enters at the junction of the left subclavian and carotid veins, and the chyle is conveyed directly to the heart, mixed with the blood, which already exists in the blood vessels. From the heart, the blood and chyle thus mixed together are propelled into the lungs, where they undergo farther changes.

10. The stomach and intestines usually contain some gaseous matters. Indeed it is probable that common air is swallowed along with the food. But from the experiments of Magendie and Chevreul there is reason to believe, that the oxygen gas which usually exists in the stomach in a considerable proportion when compared with the other gases, gradually diminishes, and at last disappears as we proceed along the course of the alimentary canal; while at the same time, the proportion of carbonic acid gas and of hydrogen gas increases. They examined the gaseous contents of the stomach and intestines of four criminals executed in Paris, very soon after death. The results were as follows:

* Phil. Trans. 1812, p. 97.

Book V.

1. Gases in the stomach.

Oxygen	11·00
Carbonic acid	14·00
Hydrogen	3·55
Azote	71·45
	100·00

2. Gases in the small intestines.

Oxygen	0·00	0·00	0·0
Carbonic acid ..	24·39	40·00	25·0
Hydrogen	55·53	51·15	8·4
Azote	20·08	8·85	66·6
	100·00		100·00		100·0

3. Gases in the large intestines.

Oxygen	0·00	0·0
Carbonic acid ..	43·50	70·0
Hydrogen and carburetted hydrogen	5·47	11·6
Azote	51·03	18·4
	100·00		100·0

4. Gases in the cœcum.

Oxygen	0·0
Carbonic acid	12·5
Hydrogen	7·5
Carburetted hydrogen.	12·5
Azote	67·5
	100·0

5. Gases in the rectum.

Oxygen	0·00
Carbonic acid	42·86
Carburetted hydrogen	11·18
Azote	45·96
	100·00 *

Such are the phenomena of digestion, as far as they have been traced. The food is first conveyed to the stomach,

* Ann. de Chim. et Phys. ii. 292.

where, by means of the gastric juice, it is converted into chyme. The chyme passes into the intestinal canal, where it is subjected to a new process, being gradually decomposed and converted into chyle and excrementitious matter, which, by means of the bile, are separated from each other. The excrementitious matter is evacuated, but the chyle is absorbed by the lacteals, and conveyed to the blood vessels and lungs. Let us now endeavour to trace the changes produced on it by these organs.

Chap. III.

SECT. II.

OF RESPIRATION.

THE absolute necessity of respiration, or of something analogous, is known to every one; and few are ignorant that in man, and hot blooded animals, the organ by which respiration is performed is the lungs. Now respiration consists in drawing a certain quantity of air into the lungs, and throwing it out again alternately. Whenever this function is suspended, even for a very short time, the animal dies.

Respiration necessary.

The fluid respired by animals is common atmospherical air; and it has been ascertained by experiment, that no other gaseous body with which we are acquainted can be substituted for it. All the known gases have been tried; but they all prove fatal to the animal which is made to breathe them. Gaseous bodies, as far as respiration is concerned, may be divided into two classes: 1. Unrespirable gases; 2. Respirable gases.

I. The gases belonging to the first class are of such a nature that they cannot be drawn into the lungs of animals at all; the epiglottis closing spasmodically whenever they are applied to it. To this class belong carbonic acid, and probably all the other acid gases, as has been ascertained by the experiments of Pilatre de Rozier.* Ammoniacal gas belongs

Unrespirable gases.

* Jour. de Phys. xxviii. 418. Pilatre de Rozier went into a brewer's tub while full of carbonic acid gas evolved by fermentation. A gentle heat manifested itself in all parts of his body, and occasioned a sensible perspiration. A slight itching sensation constrained him frequently to shut his eyes. When he attempted to breathe, a violent feeling of suffocation prevented him. He sought for the steps to get out; but not finding them readily, the necessity of breathing increased, he became giddy, and felt a tingling sensation in his ears. As soon as his mouth reached the air he

Book V.

to the same class; for the lungs of animals suffocated by it were found by Pilatre not to give a green colour to vegetable blues.*

Respirable gases, of four kinds.

1. Suffocate.

2. Kill by occasioning changes in the blood.

3. Support life imperfectly.

4. May be breathed without injury.

Oxygen necessary in respiration.

II. The gases belonging to the second class may be drawn into the lungs, and thrown out again without any opposition from the respiratory organs; of course the animal is capable of respiring them. They may be divided into four subordinate classes: 1. The first set of gases occasion death immediately, but produce no visible change in the blood. They occasion the animal's death merely by depriving him of air, in the same way as he would be suffocated by being kept under water. The only gases which belong to this class are *hydrogen* and *azotic*. 2. The second set of gases occasion death immediately, but at the same time they produce certain changes in the blood, and therefore kill not merely by depriving the animal of air, but by certain specific properties. The gases belonging to this class are *carburetted hydrogen*, *sulphuretted hydrogen*,† *carbonic oxide*, and perhaps also *nitrous gas*. 3. The third set of gases may be breathed for some time without destroying the animal, but death ensues at last, provided their action be long enough continued. To this class belong the *nitrous oxide* and *oxygen gas*.‡ 4. The fourth set may be breathed any length of time without injuring the animal. *Air* is the only gaseous body belonging to this class.§

It has been long known that an animal can only breathe a certain quantity of air for a limited time: after which it becomes the most deadly poison, and produces suffocation as effectually as the most noxious gas, or a total absence of air. It was suspected long ago that this change is owing to the

breathed freely, but for some time he could not distinguish objects; his face was purple, his limbs weak, and he understood with difficulty what was said to him. But these symptoms soon left him. He repeated the experiment often; and always found, that as long as he continued without breathing, he could speak and move about without inconvenience; but whenever he attempted to breathe, the feeling of suffocation came on. Ibid. p. 422.

* Jour. de Phys. xxviii. 424.

† See Chausier's experiments, *ibid.* lvi. 35.

‡ Perhaps also nitrous gas might have the same effect, if it could be breathed by an animal whose lungs contained no oxygen.

§ For by far the completest account of all the experiments hitherto made on the respiration of animals, the reader is referred to Dr. Bostock's excellent Essay on Respiration. Nor are the clearness, impartiality, and solid judgment of this author inferior to the extent of his information.

absorption of a part of the air; and Mayow made a number of very ingenious experiments in order to prove the fact. In 1757 Dr. Black, by breathing through lime-water, ascertained that the air, when thrown out of the lungs, contained carbonic acid.* This discovery was still farther confirmed by Lavoisier in his first dissertation on respiration, published in the Memoirs of the French Academy for 1777. Priestley and Scheele demonstrated that the quantity of oxygen in atmospherical air is diminished by respiration. It was afterwards proved by Lavoisier, and many other philosophers, who confirmed and extended his facts, that no animal can live in air totally destitute of oxygen. Even fish, which do not sensibly respire, die very soon if the water in which they live be deprived of oxygen gas. Frogs, which can suspend their respiration at pleasure, die in about forty minutes, if the water in which they have been confined be covered over with oil.† Insects and worms, as Vauquelin has proved, exhibit precisely the same phenomena. They require air as well as other animals, and die like them if they be deprived of it. They diminish the quantity of oxygen in the air in which they live, and give out, by respiration, the very same products as other animals. Worms, which are more retentive of life than most other animals, or at least not so much affected by poisonous gases, absorb every particle of the oxygen contained in the air in which they are confined before they die. Mr. Vauquelin's experiments were made on the *gryllus viridissimus*, the *limax flavus*, and *helix pomatia*.‡

The quantity of air respired differs very much in different animals. Man and hot-blooded animals are under the necessity of breathing constantly; whereas amphibious animals have a certain power over respiration, and can suspend the function altogether for a limited time. Dr. Barclay has ascertained that these animals acquire a much greater command over their respiratory organs by habit. Fish do not breathe at all, and consume so little air, that the small portion of it held in solution by the water in which they swim is sufficient for them. It appears that the number of respirations made in a given time differ considerably in different men. Dr. Hales reckons them at 20 in a minute. A man on whom

* Black's Lectures, ii. 87.

† Carradéri, Ann. de Chim. xxix. 171.

‡ Ann. de Chim. xii. 278.

Book V. Dr. Menzies made experiments, breathed only 14 times in a minute. Sir H. Davy informs us that he makes between 26 and 27 in a minute. I myself make about 19 at an average. The average of all is 20. Now 20 in a minute make 58,800 in 24 hours.

Quantity of air respired. The quantity of air drawn in and emitted at every respiration must differ considerably with the size of the man and the capacity of his lungs. Dr. Menzies found that a man draws in at a medium 40 cubic inches of air at every inspiration. Dr. Goodwin has concluded, from his experiments, that, after a complete expiration, the mean quantity of air which remains in the lungs amounts to 109 cubic inches; and Menzies has endeavoured to prove that, after an ordinary expiration, there remains 179. Davy has concluded that his lungs, after a forced expiration, still retain 41 cubic inches of air; after a natural expiration they contain 118 cubic inches.
 After a natural inspiration 135
 After a forced inspiration 254

By a full forced inspiration, after a forced inspiration, he threw out 190 cubic inches.
 After a natural inspiration 78·5
 After a natural expiration 67·5

Messrs. Allen and Pepys have calculated, that, in an ordinary inspiration, 16·5 cubic inches of air are drawn into the lungs. In their experiments the average quantity of air thrown out of the lungs in an expiration amounted to 61 cubic inches, but the breathing was much fuller and slower than usual. In one case a forced expiration amounted to 166 cubic inches; in another to 204 cubic inches. From the experiments of the same gentlemen it appears, that the lungs of a stout man, about five feet ten inches high after death, contain nearly 100 cubic inches of air.*

Dr. Menzies' estimate of an ordinary inspiration seems to approach nearest the average; but Dr. Bostock has shown that his estimate of the capacity of the lungs is too low. Perhaps we shall not err very much if we suppose, with him, that the ordinary quantity of air contained in the lungs is 280 cubic inches; and that 40 inches, or $\frac{1}{7}$ th of the whole, is drawn in and thrown out at every ordinary respiration. This, supposing 20 respirations in a minute, will make the

* Phil. Trans. 1809.

quantity of air drawn in and thrown out of the lungs every minute amount to 800 inches; to 48,000 inches in the hour; and to 1,152,000 inches in 24 hours, which amounts to rather more than 52½ lbs avoirdupois. If this estimate is too high, it is probably at least as near the truth as that of Allen and Pepys, which appears as much too low. Chap. III.

Let us now endeavour to trace the changes produced by respiration. These are of two kinds, namely, 1. The changes produced upon the air respired; 2. Changes produced upon the blood exposed to this air. Each of these naturally claims our attention.

I. For our knowledge of the changes produced upon the air by respiration, we are chiefly indebted to Priestley, Cigna, Menzies, Lavoisier, and Seguin, Davy, Allen, and Pepys. Changes produced on the air respired. These changes are the following: 1. Part of the oxygen of the air respired disappears; 2. Carbonic acid gas is found in its place; 3. It is loaded with water in the state of vapour.

1. A considerable number of experiments have been made to determine the change of bulk which air undergoes by being respired. Change of bulk. According to Davy, air, by a single inspiration and expiration, is diminished from $\frac{1}{700}$ th to $\frac{1}{1000}$ th part of its bulk.* In the numerous and accurate experiments made by Allen and Pepys on a very large scale, the average diminution was little more than half a per cent., and even this seems to have been owing rather to unavoidable inaccuracy than to real absorption. In the experiments of Berthollet, conducted also with very great care, the diminution varied from 0.69 to 3.70 per cent.† If, with Dr. Bostock, we take $\frac{1}{100}$ th as the average diminution, and 40 inches as the quantity of air drawn into the lungs at each inspiration, then it will follow that half a cubic inch of the air disappears each time we respire. This in a day would amount to 14,400 cubic inches, which is rather more than eight cubic feet.

When air is breathed till the animal can support it no longer, the diminution which it undergoes is variously stated. Davy found it to amount to $\frac{1}{100}$ th of the whole air; † Lavoisier and Goodwyn to $\frac{1}{80}$ th; § Allen and Pepys to $\frac{1}{100}$ th; || and Crawford,** on the other hand, found no diminution at all. These numbers vary so much from each other, and from the estimate from a single respiration, that it is impossible to avoid

* Davy's Researches, p. 435.

† Mem. d'Arcueil, ii. 461.

‡ Davy's Researches, p. 431.

§ Bostock on Respiration, p. 87.

|| Phil. Mag. xxxiii. 254.

** Crawford on Heat, p. 146.

Book V. concluding the diminution to be different at different times. I was induced by a letter from Mr. Dalton, in the summer of 1806, to make some experiments on the subject. In some cases I could perceive no diminution at all; in others it was perceptible. It was greatest when the animal was taken out repeatedly during the experiment, or when air was employed which was purer than that of the atmosphere. I am disposed to consider this diminution as accidental, and owing to some absorption of air altogether independent of respiration, and exceedingly various in different circumstances.

Oxygen
consumed.

2. When the air drawn into the lungs is thrown out again, it is found deprived of a portion of its oxygen. Various experiments have been made to ascertain how much of this principle is lost by respiration in a given time; but they by no means correspond with one another. Indeed, it is extremely probable, if not absolutely certain, that the degree of effect which the same animal produces upon the air respired differs materially at different times, and in consequence of different circumstances. Nothing therefore beyond an approximation can be expected from our experiments on this function.

Dr. Menzies was the first who attempted to ascertain the quantity of oxygen consumed by a man in a day. According to him, 36 inches are consumed in a minute, and of course 51,840 inches in 24 hours.* This estimate exceeds that obtained by Lavoisier and Davy from their experiments. Lavoisier and Seguin estimate the quantity of oxygen consumed by a man in 24 hours at 46,037 cubic inches, and this nearly coincides with the result which Lavoisier obtained from his last experiments, in which he was occupied when he was dragged to the place of execution. With this also the experiments of Davy coincide very well. He calculates that 31.6 inches of oxygen are consumed in a minute, which, in 24 hours, makes 45,504 inches.† This coincidence would dispose us to embrace this result as affording a near approximation to the truth. We may conclude, then, that in a day a man consumes rather more than 25 cubic feet of oxygen. Now, since the oxygen amounts only to about one-fifth of the atmosphere, it follows, that in a day, a man destroys, or renders unfit for supporting combination and respiration, no less than 125 cubic feet of air. The result obtained by

* Bostock on Respiration, p. 81.

† Davy's Researches, p. 433.

Messrs. Allen and Pepys is somewhat less, or about 27.5 cubic inches per minute. And they think that in ordinary respiration the proportion consumed is much smaller. Chap. III.

3. The air which is thrown out of the lungs contains in it a quantity of carbonic acid which did not exist in it previous to its being used for respiration. Dr. Menzies conceived that the bulk of this gas is precisely equivalent to that of the oxygen consumed. This also was the result of the experiments of Dr. Crawford. Lavoisier, in his experiments on the Guinea pig, found it somewhat less. In his first experiment he found the oxygen consumed, to the carbonic acid formed, as 20 to 16.5; in his second, as 20 to 17.3.* In his experiments on Seguin, in 1789, the oxygen consumed was to the bulk of carbonic acid formed nearly as 20 to 16.6; but in those that were made afterwards, the proportion of carbonic acid is diminished by nearly one-half. In Davy's experiments, the bulk of carbonic acid formed corresponded very nearly with that of the oxygen consumed: † so that in this respect they coincided with those of Crawford and Menzies. I was informed by Mr. Dalton, in the summer of 1805, that he had satisfied himself, by a variety of experiments, that the bulk of carbonic acid gas formed was exactly equal to that of the oxygen gas consumed. On repeating the experiment with that particular view, I found that in some cases this took place very nearly; but upon the whole, the bulk of oxygen which disappeared was somewhat greater than that of the carbonic acid formed; but the difference varied considerably, and kept pace with the diminution of the bulk of air by respiration. Hence I consider it as owing to the abstraction of a part of the air by some other way than respiration. If this abstraction be allowed for, I have no doubt, from my own experiments, that the bulk of the carbonic acid formed by respiration is precisely equal to that of the oxygen which has disappeared. The absolute quantity it is difficult to state, as it depends upon a variety of circumstances. I am disposed to consider it, at an average, as approaching to 40,000 cubic inches in 24 hours, though probably somewhat under that quantity. Now, this quantity of carbonic acid contains little Carbonic acid formed.

* See the details, *Mém. Par.* 1780, p. 401; *Ann. de Chim.* v. 301; and in *Hastock on Respiration*, p. 72.

† Davy's *Researches*, p. 431.

Book V.

less than three quarters of a pound avoirdupois of carbon. Messrs. Allen and Pepys found the carbonic acid formed exactly equal in bulk to the oxygen which had disappeared. It amounted in their experiments to about $27\frac{1}{2}$ cubic inches per minute, or 39,534 cubic inches in 24 hours, a quantity which contains about 11 ounces troy of solid carbon. Air thrown out of the lungs contained in their experiments about 8 per cent. of carbonic acid gas. In Berthollet's experiments, the carbonic acid gas formed varied from 5.53 to 13.82 per cent. But the animal was confined for several hours in the same air.*

Azote.

4. Dr. Priestley concluded from the experiments, that not only the oxygen, but the azote also, of the air respired was diminished.† This opinion was still farther confirmed by Davy, who found the consumption of azote to amount to about $\frac{1}{7}$ th of that of the oxygen.‡ Dr. Henderson has made experiments with the same result, though the proportion of azote absorbed was rather less.§ Upon repeating these experiments, I found likewise a loss of azote: but it was extremely inconstant, sometimes being scarcely perceptible, and at other times considerable. It kept pace with the diminution of the bulk of the air respired, and with the difference between the bulk of the oxygen consumed and the carbonic acid formed. Hence I am disposed to ascribe all these differences to the same cause. I conceive that a portion of the air respired disappears without undergoing any change, and that this portion occasions the diminution of the azote, and the difference between the bulk of the carbonic acid formed and the oxygen consumed. What comes of this portion of air it is difficult to say; but I think it conceivable that the disappearing of such a portion may be confined to the unnatural circumstances occasioned by the experiment; that the difficulty of throwing out the air from the lungs in these circumstances may be such as to induce absorbents to act, and remove a portion which, in the ordinary situation of the lungs, would have been thrown out by expiration. Messrs. Allen and Pepys, in all their experiments, found no change in the azote, and no absorption of it whatever. Hence it is probable that the diminution formerly observed may have been owing to errors in the experi-

* Mem. d'Arcueil, ii. 461.

† Davy's Researches, p. 433.

† Priestley, ii. 380.

§ Nicholson's Jour. viii. 44.

ments. Berthollet found the proportion of azote to increase a little instead of diminishing, which he ascribes to the increased bulk of the air from moisture and heat.

5. Dr. Prout has shown by a number of well-conducted experiments on himself, that the proportion of carbonic acid formed at each inspiration is different at different periods of the day. It is at its maximum nearly about noon, and is at its minimum about midnight. It appears farther from his trials that the quantity of carbonic acid gas in expired air begins to increase nearly at twilight. The following table exhibits the proportion per cent. of carbonic acid in the air expired from his lungs during every hour of the day. The experiments from which it was deduced were made in August.*

Hour A. M.	Carbonic acid per cent.	Hour P. M.	Carbonic acid per cent.
6	3.43	6	3.40
7	3.48	7	3.35
8	3.56	8	3.32
9	3.66	9	3.30
10	3.78	10	3.30
11	3.92	11	3.30
12	4.10	12	3.30
1	3.98	1	3.30
2	3.80	2	3.30
3	3.65	3	3.30
4	3.54	4	3.33
5	3.46	5	3.38

Mean 3.45

Dr. Prout found that alcohol and all fermented liquors diminished the proportion of carbonic acid formed by respiration, and this was confirmed by the experiments of Dr. A. Fyfe. They found likewise that when the constitution is affected by mercury, the proportion of carbonic acid gas in the air expired is diminished. Dr. Fyfe found that the quantity was likewise diminished by a course of nitric acid, and by a vegetable diet.†

6. It is not so easy to determine the proportion of water emitted from the lungs mixed with the air expired, as it is that of the carbonic acid. According to the experiments of

Water
emitted.

* Annals of Philosophy, ii. 328, and iv. 331.

† Ibid. iv. 334.

Book V.

Dr. Hales, it amounts in a day to 20·4 oz;* but his method was not susceptible of great accuracy. Mr. Lavoisier, on the other hand, estimates it rather higher; but the proportion seems rather to have been the result of calculation than of any direct measurement. The result of a few trials which I made during the summer of 1806, gave me very nearly 19 ounces per day as the quantity emitted from my own lungs; but I do not lay much stress upon the results, as they were not sufficiently varied to give a fair average; but I think it worth while to mention the method I followed, on account of its simplicity. I filled a glass with water, and then by breathing on it, and gradually raising its temperature till the vapour of the breath just ceased to be condensed, I ascertained the force of the vapour contained in the air respired from Mr. Dalton's table. Then, ascertaining the force of vapour in the atmosphere at the time, and subtracting it from the force of the vapour in the breath, I obtained the force of the vapour actually emitted from the lungs. Supposing 40 inches of air to be drawn in at each inspiration, it was easy to calculate, from Dalton's table, the weight of the vapour which it contained when emitted.

7. In ordinary cases of respiration, the oxygen, which disappears, is just balanced by the carbonic acid formed, so that the bulk of the air continues unaltered: but it appears from the experiments of Messrs. Allen and Pepys, that when the same quantity of air is breathed backwards and forwards as long as possible, a greater quantity of oxygen disappears than can be accounted for by the carbonic acid formed. This oxygen is absorbed by the system. It diminishes the bulk of the air respired, and amounts to about $\frac{1}{14}$ th of the whole air respired at an average.

Respiration
of oxygen
gas.

8. When oxygen gas, nearly pure, is breathed, rather a greater quantity of carbonic acid is given out; for the air expired contains, at an average, above ten per cent. of that gas. It appears also from the experiments of Messrs. Allen and Pepys, that a portion of the oxygen gas disappears, and then an equal bulk of azotic gas is found in its place. This substitution of azote is greater at the commencement of the experiment, and diminishes as the respiration continues. It is difficult at present to account for this substitution. The

* Veget. Stat. ii. 327.

quantity observed, amounting at an average to nearly 80 cubic inches, is too great to be ascribed to an error in the experiment. It can only be fully accounted for when we become acquainted with the composition of azote, which is still unknown. Chap. III.

9. When an animal is made to breathe a mixture of 79 hydrogen gas and 21 oxygen gas, the respiration goes on without any inconvenience; carbonic acid gas is formed as usual, and at the same time a portion of the oxygen disappears, and azote is found in its place. This mixture has a sedative effect, and makes the animal sleepy. It appears probable, from the observations of Messrs. Allen and Pepys, that during sleep the quantity of carbonic acid formed is less than when the animal is awake. Of a mixture of oxygen and hydrogen.

10. A great number of experiments have lately been made on the respiration of fishes by Provençal and Humboldt.* It is well known that these animals require oxygen gas as well as other animals, and that if the water in which they are be deprived of the whole of its air, they die very speedily. Provençal and Humboldt employed for their experiments the water of the Seine. They separated the air from a quantity of it by boiling, and subjected it to a chemical analysis. Into another quantity of the same water, tenches were put and confined for several hours till they began to suffer; they were then withdrawn, and the air separated from the water in which they had lived, and subjected it to chemical analysis. In every case a portion, both of oxygen and azote had disappeared, and a quantity of carbonic acid had been formed. The following table exhibits the results of a variety of their experiments: Respiration of fishes.

* Mem. d'Arcueil, ii. 259.

Nature of the gases.	Gases before experiment.	Ditto after ditto.	Difference.	The fishes have		Azote absorbed, the oxygen being 100.	Carb. acid produced, the oxygen absorbed being 100.	Number of fishes and time.
				Absorbed.	Produced.			
Total.....	175.0	135.1	39.9	—	—	—	—	Three tenches during 5 hours, 15 minutes.
Oxygen	52.1	5.6	—	46.5	—	—	—	
Azote	115.9	95.8	—	20.1	—	43	57	
Carbonic acid..	7.0	33.7	—	—	26.7	—	—	
Total.....	524.0	404.0	119.6	—	—	—	—	Seven tenches during 6 hours.
Oxygen	155.9	44.0	—	111.9	—	—	—	
Azote	347.1	249.5	—	97.6	—	87	80	
Carbonic acid..	21.0	110.9	—	—	89.9	—	—	
Total.....	524.0	45.0	71.0	—	—	—	—	Seven tenches during 8 hours.
Oxygen	155.9	10.5	—	145.4	—	—	—	
Azote	347.1	289.5	—	57.6	—	40	91	
Carbonic acid..	21.0	153.0	—	—	132.0	—	—	
Total.....	483.0	345.5	137.5	—	—	—	—	One tench during 17 hours.
Oxygen	143.7	4.2	—	139.5	—	—	—	
Azote	320.0	294.1	—	25.9	—	19	—	
Carbonic acid..	19.8	47.2	—	—	27.9	—	20	
Total.....	483.0	408.0	75.0	—	—	—	—	Three tenches during 7½ hours.
Oxygen	143.7	62.6	—	81.1	—	—	—	
Azote	320.0	285.4	—	34.6	—	43	—	
Carbonic acid..	19.3	60.0	—	—	40.7	—	50	
Total.....	483.0	398.6	84.4	—	—	—	—	Three tenches during 5 hours.
Oxygen	143.7	40.0	—	103.7	—	—	—	
Azote	320.0	246.6	—	73.4	—	71	—	
Carbonic acid..	19.3	112.0	—	—	92.7	—	89	
Total.....	483.0	372.5	110.5	—	—	—	—	Two tenches during 7 hours.*
Oxygen	143.7	37.8	—	105.9	—	—	—	
Azote	320.0	252.9	—	67.1	—	63	—	
Carbonic acid..	19.3	81.8	—	—	62.5	—	59	

The quantity of gas obtained from the Seine water, was at an average 0.0275 of its bulk, or not quite $\frac{1}{36}$ part; the average quantity of oxygen which this gas contained was 0.310.

From these experiments it appears, that the respiration of fishes differs very much from that of other animals. The oxygen is not merely converted into carbonic acid, as happens during the respiration of men and the larger animals;

* The numbers in this table indicate cubic centimetres. A cubic centimetre is equal to 0.0610 of a cubic inch.

but a portion of it is absorbed and introduced into the system. A portion also of azote is absorbed. The quantity of air consumed by fishes is extremely small, when compared with that consumed by terrestrial animals. This will appear from the following table, in which the bulk of the air consumed, and of the carbonic acid formed in an hour, is stated in cubic inches.

Time.	Oxygen in the air after the experiments.	Number of fish.	Hours the experiment lasted.	Absorption in 1 hour in cubic inches.		Carbonic acid produced, in cubic inches.
				Oxygen.	Azote.	
28 Feb.	0·056	3	5½	0·0245	0·0106	0·0140
3 March	0·151	7	6	0·0221	0·0192	0·0177
7 March	0·034	7	8½	—	—	0·0185
11 March	0·017	1	17	0·0679	0·0126	0·0136
28 Feb.	0·178	3	7½	0·0298	0·0123	0·0150
24 Feb.	0·141	3	5	0·0575	0·0405	0·0512
20 Feb.	0·130	2	7	0·0635	0·0397	0·0370

From this table, compared with the facts stated in the preceding part of this section, it follows, that in a given time a man consumes 50,000 times as much oxygen gas as a tench. Yet the presence of this principle is equally necessary for the existence of both.

II. Let us now endeavour to ascertain the changes produced on the blood by repiration. The whole of the blood is propelled from the heart to the lungs, circulates through the vessels of that organ; and during that circulation it is exposed to the influence of the air which the animal is constantly drawing into the lungs. Now, certain changes are produced upon it by this action, which have been partly traced by the experiments of Priestley, Cigna, Fourcroy, Hassenfratz, Beddoes, Watt, and above all by those of Davy. These changes, as far as we are acquainted with them, are the following: 1. It acquires a florid red colour, and the chyle disappears; 2. It loses a portion of carbon; 3. It emits water.

Changes produced on the blood.

1. It has been long known that the blood which flows in the veins is of a dark reddish purple colour, whereas the arte-

Book V.

rial blood is of a florid scarlet colour. Lower observed that the colour of the venous blood was converted into that of arterial during its passage through the lungs. No chyle can be distinguished by its white colour in the blood after it has passed through the lungs. The changes, then, which take place upon the appearance of the blood are two: 1. It acquires a florid red colour: 2. The chyle totally disappears. Lower himself knew that the change was produced by the air, and Mayo attempted to prove that it was by absorbing a part of the air. But it was not till Dr. Priestley discovered that venous blood acquires a scarlet colour when put in contact with oxygen gas, and arterial blood a dark red colour when put in contact with hydrogen gas; or, which is the same thing, that oxygen gas instantly gives venous blood the colour of arterial; and hydrogen, on the contrary, gives arterial blood the colour of venous blood—it was not till then that philosophers began to attempt any thing like an explanation of the phenomena of respiration.

Action of
different
gases on
blood.

The blood is a fluid of so complex a nature that it is not easy to ascertain the changes produced in it by exposure to different gases out of the body; and even if that could be done, we have no method of proving that the effects of these gaseous bodies upon the coagulated blood are the same as they would be on the blood in its natural state, circulating in the vessels of a living animal. The facts which have been ascertained are the following:

1. Oxygen,

1st. It appears from the experiments of Priestley, Girtanner, and Hassenfratz, that when venous blood is exposed to oxygen gas confined over it, the blood instantly assumes a scarlet colour. Davy could not perceive any sensible diminution of the bulk of the gas.

2. Air,

2d. The same change of colour takes place when blood is exposed to common air. In this case a quantity of carbonic acid gas is formed, and a quantity of oxygen gas, exactly equal to it in bulk, disappears; making allowance for the small quantity of carbonic acid, which we may suppose to be absorbed by the blood itself.

3d. Venous blood exposed to the action of azotic gas continues unaltered in colour; neither does any perceptible diminution of the gas ensue.

3. Nitrous,

4th. Venous blood exposed to the action of nitrous gas becomes of a deep purple, and about $\frac{1}{8}$ th of the gas is absorbed.

5th. Venous blood exposed to nitrous oxide becomes of a brighter purple, especially on the surface, and a considerable portion of the gas is absorbed. Chap. III.
5. Nitrous oxide,

6th. Venous blood exposed to carbonic acid gas becomes of a brownish-red colour, much darker than usual, and the gas is slightly diminished in bulk. 6. Carbonic acid,

7th. Carburetted hydrogen gas gives venous blood a fine red colour, a shade darker than oxygen gas does, as was first observed by Dr. Beddoes, and at the same time a small portion of the gas is absorbed. This gas has the property of preventing, or at least greatly retarding, the putrefaction of blood, as was first observed by Mr. Watt.* 7. Carbu-
retted hy-
drogen.

8th. When arterial blood is put in contact with azotic gas, or carbonic acid gas, it gradually assumes the dark colour of venous blood, as Dr. Priestley found.† The same philosopher also observed, that arterial blood acquired the colour of venous blood when placed *in vacuo*.‡ Consequently this alteration of colour is owing to some change which takes place in the blood itself, independent of any external agent. Arterial
blood dark-
ened by
azotic and
carbonic
acid gas;

The arterial blood becomes much more rapidly and deeply dark coloured when it is left in contact with hydrogen gas placed above it.§ We must suppose therefore that the presence of this gas accelerates and increases the change, which would have taken place upon the blood without any external agent.

9th. If arterial blood be left in contact with oxygen gas, it gradually assumes the same dark colour which it would have acquired *in vacuo*, or in contact with hydrogen; and after this change oxygen can no longer restore its scarlet colour.|| Therefore it is only upon a part of the blood that the oxygen acts; and after this part has undergone the change which occasions the dark colour, the blood loses the power of being affected by oxygen. By rest,

10th. Mr. Hassenfratz poured into venous blood a quantity of chlorine; the blood was instantly decomposed, and assumed a deep and almost black colour. When he poured common muriatic acid into blood, the colour was not altered.** And by
chlorine.

2. Dr. Priestley, the first of the modern chemists who turned his attention to respiration, concluded from some of Theory of
Priestley.

* Davy's Researches, p. 380.

† Priestley, iii. 363.

‡ Priestley, iii. 363, and Ann. de Chim. ix. 269.

§ Fourcroy, *ibid.* vii. 149.

|| *Ibid.* ix. 268.

** *Ibid.*

Book V.

his earliest experiments, that the blood, as it passed through the lungs, gave out phlogiston to the air, which was expired loaded with that substance; and of course that the purpose of respiration was to free the blood of phlogiston. Lavoisier soon after ascertained with more precision the changes which the air undergoes during respiration; and he formed a theory in order to explain that function, assuming as its basis that all the changes on the air inspired are produced in the lungs; and of course, that all the new substances expired are formed in the lungs. According to him, the blood absorbs no air in the lungs; but it gives out hydrogen and carbon, which, combining with the oxygen of the air inspired, form water and carbonic acid. This theory was adopted by Laplace, Crawford, Gren, and Girtanner, with a small variation. Indeed it does not differ, except in detail, from the original hypothesis of Dr. Priestley, that the use of respiration is to rid the blood of phlogiston; for if we substitute carbon and hydrogen for phlogiston, the two theories precisely agree. Mr. Lavoisier attempted not to prove its truth; he only tried to show that the oxygen absorbed corresponds exactly with the quantity of oxygen contained in the carbonic acid and the water emitted.

Modified by
Lavoisier.Theory of
La Grange.

A different theory was afterwards proposed by Mr. De La Grange. According to this philosopher, the oxygen gas, which disappears, combines with the blood as it passes through the lungs; and at the instant of this combination there is set free from the blood a quantity of carbonic acid gas and of water, which are thrown out along with the air expired. This theory was adopted and illustrated by Mr. Hassenfratz. But as the carbonic acid formed is exactly equal to the bulk of the oxygen which disappears, this oxygen must be changed into carbonic acid in the lungs; for oxygen, when changed into carbonic acid, does not sensibly alter its bulk. The blood, then, must emit carbon in the lungs to the extent of about $\frac{3}{4}$ ths of a pound in the day. This, I conceive, is all that happens in the lungs. The watery vapour, indeed, is doubtless secreted from the blood to mix with the air inspired; but probably the secretion takes place in appropriate organs. By the loss of carbon the venous blood is changed into arterial. The very same thing happens when venous blood is presented to the air.

Chyle converted into
blood.

The change of the chyle into blood, I presume, takes place in the blood-vessels. Indeed the blood is a liquid which is

constantly running through a suite of changes. Hence it is kept of a similar nature only by the constant influx of new matter, which is as constantly manufactured into blood. It appears, from the most accurate observations hitherto made, that neither chyle nor lymph contain fibrin, which forms a very conspicuous part of the blood. This fibrin is employed to supply the waste of the muscles; the most active parts of the body, and therefore in all probability, requiring the most frequent supply. Nor can it be doubted that it is employed for other useful purposes. The quantity of fibrin in the blood, then, must be constantly diminishing, and therefore new fibrin must be constantly formed. But the only substances out of which it can be formed are the chyle and lymph, neither of which contain it. There must therefore be a continual decomposition of the chyle, and lymph going on in the blood-vessels, and a continual new formation of fibrin. Other substances also may be formed; but we are certain that this *must* be formed there, because it does not exist previously. Now, one great end of respiration must undoubtedly be to assist this decomposition of chyle and complete formation of blood.

In what manner the chyle, or a part of it, is converted into fibrin, it is impossible to say: we are not sufficiently acquainted with the subject to be able to explain the process. But we can see, at least, that carbon must be abstracted from that part of the chyle which is to be converted into fibrin. Hence, as the process of blood-making advances, there must be a greater and greater redundancy of carbon in the liquid. Unless this redundancy were removed, the process could not go on, and probably the whole would run into putrefaction. We may conclude, then, that one great use of respiration is to abstract this carbon, by forming with it carbonic acid. How this is performed, indeed, it is impossible at present to explain; but the fact is undoubted.

But the abstraction of carbon is not the only advantage gained by respiration: the *temperature* of all animals depends upon it. It has been long known, that those animals which do not breathe have a temperature but very little superior to the medium in which they live. This is the case with fishes and many insects. Man, on the contrary, and the quadrupeds which breathe, have a temperature considerably higher than the atmosphere: that of man is 98°. Birds, which breathe in proportion a still greater quantity of air than man, have a temperature equal to 103° or 104°. It has been proved,

Animal
heat.

Book V. that the temperature of all animals is proportional to the quantity of air which they breathe in a given time.

These facts seem sufficient to demonstrate that the heat of animals depends upon respiration. But it was not till Dr. Black's doctrine of latent heat became known to the world, that any explanation of the cause of the temperature of breathing animals was attempted. That illustrious philosopher, whose discoveries form the basis upon which all the scientific part of chemistry has been reared, saw at once the light which his doctrine of latent heat threw upon this part of physiology, and he applied it very early to explain the temperature of animals.

Dr. Black's theory.

According to him, part of the latent heat of the air inspired becomes sensible; and of course the temperature of the lungs, and the blood that passes through them, must be raised; and the blood, thus heated, communicates its heat to the whole body. This opinion was ingenious, but it was liable to an unanswerable objection: for if it were true, the temperature of the body ought to be greatest in the lungs, and to diminish gradually as the distance from the lungs increases; which is not true. The theory, in consequence, was abandoned even by Dr. Black himself; at least he made no attempt to support it.

Theory of Crawford.

Dr. Crawford, who considered all the changes operated by respiration as taking place in the lungs, accounted for the origin of the animal heat almost precisely in the same manner with Dr. Black. According to him, the oxygen gas of the air combines in the lungs with the carbon emitted by the blood. During this combination, the oxygen gives out a great quantity of caloric, with which it had been combined; and this caloric is not only sufficient to support the temperature of the body, but also to carry off the new-formed water in the state of vapour, and to raise considerably the temperature of the air inspired. According to this philosopher, then, the whole of the caloric which supports the temperature of the body is evolved in the lungs. His theory accordingly was liable to the same objection with Dr. Black's; but Dr. Crawford obviated it in the following manner: He found that the specific caloric of arterial blood was 1.0300, while that of venous blood was only 0.8928. Hence he concluded, that the instant venous blood is changed into arterial blood, its specific caloric increases; consequently it requires an additional quantity of caloric to keep its temperature as high as it had been while

venous blood. This addition is so great that the whole new caloric evolved is employed: therefore the temperature of the lungs must necessarily remain the same as that of the rest of the body. During the circulation, arterial blood is gradually converted into venous; consequently its specific caloric diminishes, and it must give out heat. This is the reason that the temperature of the extreme parts of the body does not diminish.

But this theory, ingenious and plausible as it is, has not been able to stand the test of subsequent investigations. Dr. John Davy has shown that the difference between the specific caloric of arterial and venous blood is much smaller than Dr. Crawford had estimated it. According to Dr. Davy, the specific caloric of these liquids is as follows:

Arterial blood	0·913
Venous blood	0·903

Delaroche and Berard have shown that the specific caloric of oxygen gas and carbonic acid gas differs much less than Crawford had supposed. According to these philosophers, the specific heats of these gases are as follows

Oxygen gas	0·8848
Carbonic acid gas	0·8280

Now such minute differences would hardly admit the application of Dr. Crawford's explanation. But the most formidable objections result from the experiments of Mr. Brodie. He found that when artificial respiration is kept up in the lungs of animals, after decapitation, the usual proportion of carbonic acid gas is formed, and the circulation continues nearly as usual. Yet in these animals the heat diminishes more rapidly than in a dead animal, in which artificial respiration is not kept up.* From these experiments Mr. Brodie concludes that the production of animal heat is owing to the action of the brain, and not to respiration.

I think it must be admitted that these experiments have entirely destroyed the foundations on which Dr. Crawford's theory was built. But I am disposed to think that Mr. Brodie has gone rather too far in his consequences when he conceives that respiration has no connection whatever with the production of animal heat. The fact that all animals which respire

* Phil. Trans. 1812, p. 378.

Book V.

regularly are hot-blooded, while amphibious animals and fishes are cold-blooded, seems to me to establish a connection between respiration and heat. Though we are unable to explain in what way the heat is evolved.

SECT. III.

OF THE ACTION OF THE KIDNEYS.

THUS we have reason to suppose that chyle and lymph are converted into blood during the circulation: but besides the lungs and arteries, there is another organ, the sole use of which is also to produce some change or other in the blood, which renders it more complete, and more proper for the various purposes to which it is applied. This organ is the KIDNEY.

A very great proportion of blood passes through the kidneys; indeed we have every reason to conclude that the whole of the blood passes through them very frequently. These organs separate the urine from the blood, to be afterwards evacuated without being applied to any purpose useful to the animal.

Action of
the kidneys
necessary.

The kidneys are absolutely necessary for the continuance of the life of the animal; for it dies very speedily when they become by disease unfit to perform their functions: therefore the change which they produce in the blood is a change necessary for qualifying it to answer the purposes for which it is intended.

It changes
the nature
of the
blood.

As the urine is immediately excreted, it is evident that the change which the kidneys perform is intended solely for the sake of the blood. It is not merely the abstraction of a quantity of water and of salts, accumulated in the blood, which the kidney performs. A chemical change is certainly produced, either upon the whole blood, or at least on some important part of it; for there are two substances found in the urine which do not exist in the blood. These two substances are nephirin and uric acid. They are formed, therefore, in the kidneys; and as they are thrown out, after being formed, without being applied to any useful purpose, they are certainly not formed in the kidneys for their own sake. Some part of the blood, then, must be decomposed in the kidney, and a new substance, or new substances, must be formed; and the urea and uric acid must be formed at the same time, in consequence of the combined action of the affinities which produce the

change on the blood; and being useless, they are thrown out together with a quantity of water and salts, which, in all probability, were useful in bringing about the changes which take place in the arteries and in the kidneys, but which are no longer of any service after these changes are brought about.

The changes operated upon the blood in the kidneys are at present altogether unknown; but they must be important. Provided the method of analyzing animal substances were so far perfected as to admit of accurate conclusions, considerable light might be thrown upon this subject, by analyzing with care a portion of blood from the emulgent vein and artery separately, and ascertaining precisely in what particulars they differ from each other.

SECT. IV.

OF PERSPIRATION.

Thus we have seen that the principal changes which the blood undergoes, as far at least as we are at present acquainted with them, take place in the lungs, in the kidneys, and in the arteries. In the lungs, a quantity of water and carbonic acid gas is emitted from the blood; and in the kidneys, the urine is formed and separated from it. There seems also to be something thrown out from the blood during its circulation in the arteries, at least through those vessels which are near the surface of the body: for it is a fact, that certain substances are constantly emitted from the skins of animals. These substances are known in general by the name of *perspirable matter*, or *perspiration*. They have a great resemblance to what is emitted in the lungs; which renders it probable that both excretions are owing to the same cause; namely, to the decomposition produced in the blood by the effects of respiration.

Many experiments have been made to ascertain the quantity of matter perspired through the skin. For the first set, and not the least remarkable, we are indebted to Sanetorins; who continued them for no less than 30 years. He ascertained his own weight and the weight of his food; and whatever weight he lost over and above that of his excrements, he ascribed to perspiration. A similar set of experiments was afterwards made in France by Dodart; in England by Keil; in Ireland by Bryan Robertson and Rye; and in Carolina by

Quantity
perspired.

Lining. The result of all these experiments has been collected by Haller; but it gives us no precise estimate of the amount of the transpiration, since these philosophers have not distinguished between what is lost by the skin and by the lungs. Lavoisier and Seguin alone have attempted to ascertain the amount of the matter perspired through the skin. A bag composed of varnished silk, and perfectly air-tight, was procured, within which Seguin, who was usually the subject of experiment, was enclosed, and the bag was closed exactly over his head. There was a slit in the bag opposite to his mouth, and the edges of this slit were accurately cemented round the mouth by means of a mixture of turpentine and pitch. Thus every thing emitted by the body was retained in the bag, except what made its escape from the lungs by respiration. By weighing himself in a delicate balance at the commencement of the experiment, and again after he had continued for some time in the bag, the quantity of matter carried off by respiration was ascertained. By weighing himself without this varnished covering, and repeating the operation after the same interval of time had elapsed as in the former experiment, he ascertained the loss of weight occasioned by perspiration and respiration. By subtracting from this sum the loss of weight indicated by the first experiment, he obtained the quantity of matter which made its escape by perspiration in a given time. The following facts were ascertained by these experiments:

1. The maximum of matter perspired in a minute amounted to 26.25 grains troy; the minimum to nine grains: which gives 17.63 grains at a medium in the minute, or 52.89 ounces in the 24 hours. This quantity differs less than might have been expected from the result of former experiments made by Dodart, Keil, Rye, &c.

2. The quantity perspired is increased by drink, but not by solid food.

3. Perspiration is at its minimum immediately after a repast. It reaches its maximum during digestion.*

The quantity of matter perspired differs very considerably according to circumstances. It has been shown to be greatest in hot weather, and in hot climates, and after great exercise; and its relation to the quantity of urine has been long known.

* Fourcroy, ix. 198.

When the matter perspired is great, the quantity of urine is small, and *vice versa*. Chap. III.

To ascertain the substance thus emitted by perspiration is a difficult task, because it passes off invisibly, and in small quantities at a time. It has, notwithstanding, been ascertained, that water, carbon, and an oily matter are emitted; and that an acid, phosphate of lime, and even urea, are sometimes emitted through the skin.* Substances perspired.

It has been supposed that the skin has the property of *absorbing moisture* from the air; but this opinion has not been confirmed by experiments, but rather the contrary.

The chief arguments in favour of absorption of the skin have been drawn from the quantity of moisture discharged by urine being, in some cases, not only greater than the whole drink of the patient, but even than the whole of his drink and food. But it ought to be remembered that, in diabetes, the disease here alluded to, the weight of the body is continually diminishing, and therefore part of it must be constantly thrown off. Besides, it is scarcely possible in that disease to get an accurate account of the food swallowed by the patients; and in those cases where very accurate accounts have been kept, and where deception was not so much practised, the urine was found not to exceed the quantity of drink.† Whether the skin absorbs moisture. In a case of diabetes, related with much accuracy by Dr. Gerard, the patient was bathed regularly during the early part of the disease in warm water, and afterwards in cold water: he was weighed before and after bathing, and no sensible difference was ever found in his weight.‡ Consequently, in that case, the quantity absorbed, if any, must have been very small.

It is well known that thirst is much alleviated by cold bathing. By this plan Captain Bligh kept his men cool and in good health during their very extraordinary voyage across the South Sea. This has been considered as owing to the absorption of water by the skin. But Dr. Currie had a patient who was wasting fast for want of nourishment, a tumor in the œsophagus preventing the possibility of taking food, and whose thirst was always alleviated by bathing; yet no sensible increase of weight, but rather the contrary, was perceived after bathing. It does not appear, then, that in either of these cases water was absorbed.

* See the preceding Chapter, Sect. 25.

† See Rollo on Diabetes.

‡ Ibid. ii. 73.

Book V.

Farther, Seguin has shown that the skin does not absorb water during bathing, by a still more complete experiment: He dissolved some mercurial salt in water, and found that the mercury produced no effect upon a person that bathed in the water, provided no part of the cuticle was injured; but upon rubbing off a portion of the cuticle, the mercurial solution was absorbed, and the effects of the mercury became evident upon the body. Hence it follows irresistibly, that water, at least in the state of *water*, is not absorbed by the skin when the body is plunged into it, unless the cuticle be first removed.

Whether it
absorbs
other sub-
stances.

This may perhaps be considered as a complete proof that no such thing as absorption is performed by the skin; and that therefore the appearance of carbonic acid gas, which takes place when air is confined around the skin, must be owing to the emission of carbon. But it ought to be considered, that although the skin cannot absorb water, this is no proof that it cannot absorb other substances; particularly that it cannot absorb oxygen gas, which is very different from water. It is well known that water will not pass through bladders, at least for some time: yet Dr. Priestley found that venous blood acquired the colour of arterial blood from oxygen gas, as readily when these substances were separated by a bladder as when they were in actual contact. He found, too, that when gases were confined in bladders, they gradually lost their properties. It is clear from these facts, that oxygen gas can pervade bladders: and if it can pervade them, why may it not also pervade the cuticle? Nay, farther, we know from the experiments of Cruickshanks, that the vapour perspired passes through leather, even when prepared so as to keep out moisture, at least for a certain time. It is possible, then, that water, when in the state of vapour, or when dissolved in air, may be absorbed, although water, while in the state of water, may be incapable of pervading the cuticle. The experiments, therefore, which have hitherto been made upon the absorption of the skin, are insufficient to prove that air and vapour cannot pervade the cuticle, provided there be any facts to render the contrary supposition probable.

Now, that there are such facts cannot be denied. I shall not indeed produce the experiment of Van Mons as a fact of that kind, because it is liable to objections, and at best is very indecisive. Having a patient under his care who, from a wound in the throat, was incapable for several days of taking any nourishment, he kept him alive during that time by ap-

plying to the skin, in different parts of the body, several times a-day, a sponge dipped in wine or strong soup.* A fact mentioned by Dr. Watson is much more important, and much more decisive. A lad at Newmarket, who had been almost starved in order to bring him down to such a weight as would qualify him for running a horse race, was weighed in the morning of the race day; he was weighed again an hour after, and was found to have gained 30 ounces of weight; yet in the interval he had only taken half a glass of wine. Here absorption must have taken place, either by the skin or lungs, or both. The difficulties in either case are the same; and whatever renders absorption by one probable, will equally strengthen the probability that absorption takes place by the other.†

Chap. III.

SECT. V.

OF ASSIMILATION.

WE have now seen the progress of digestion, and the formation of blood, as far at least as we are acquainted with it. But to what purposes is this blood employed, which is formed with so much care, and for the formation of which so great an apparatus has been provided? It answers two purposes. The parts of which the body is composed, bones, muscles, ligaments, membranes, &c. are continually changing. In youth they are increasing in size and strength, and in mature age they are continually acting, and consequently continually liable to waste and decay. They are often exposed to accidents, which render them unfit for performing their various functions; and even when no such accident happens, it seems necessary for the health of the system that they should be now and then renewed. Materials therefore must be provided for repairing, increasing, or renewing all the various organs of the body; phosphate of lime and gelatin for the bones, fibrin for the muscles, albumen for the cartilages and membranes, &c. Accordingly all these substances are laid up in the blood; and

Blood answers two purposes:
1. Supplies the waste of the system.

* Phil. Mag. vi. 95.

† Watson's Chemical Essays, iii. 101. The Abbé Fontana also found that, after walking in moist air for an hour or two, he returned home some ounces heavier than he went out, notwithstanding he had suffered considerable evacuation from a brisk purge purposely taken for the experiment. This increase, indeed, might be partly accounted for by the absorption of moisture by his clothes.

Book V.
Assimila-
tion.

they are drawn from that fluid, as from a storehouse, whenever they are required. The process by which the different ingredients of the blood are made part of the various organs of the body is called ASSIMILATION.

Over the nature of assimilation the thickest darkness still hangs: there is no key to explain it, nothing to lead us to the knowledge of the instruments employed. Facts, however, have been accumulated in sufficient numbers to put the existence of the process beyond the reach of doubt. The healing, indeed, of every fractured bone, and every wound of the body, is a proof of its existence, and an instance of its action.

Every assi-
milating
organ pro-
duces a spe-
cific change.

Every organ employed in assimilation has a peculiar office; and it always performs this office whenever it has materials to act upon, even when the performance of it is contrary to the interest of the animal. Thus the stomach always converts food into chyme, even when the food is of such a nature that the process of digestion will be retarded rather than promoted by the change. If warm milk, for instance, or warm blood, be thrown into the stomach, they are always decomposed by that organ, and converted into chyme; yet these substances are much more nearly assimilated to the animal before the action of the stomach than after it. The same thing happens when we eat animal food.

And no
other.

On the other hand, a substance introduced into an organ employed in assimilation, if it has undergone precisely the change which that organ is fitted to produce, is not acted upon by that organ, but passed on unaltered to the next assimilating organ. Thus it is the office of the intestines to convert chyme into chyle. Accordingly, whenever chyme is introduced into the intestines, they perform their office, and produce the usual change; but if chyle itself be introduced into the intestines, it is absorbed by the lacteals without alteration. The experiment, indeed, has not been tried with true chyle, because it is scarce possible to procure it in sufficient quantity; but when milk, which resembles chyle pretty accurately, is thrown into the jejunum, it is absorbed unchanged by the lacteals.*

Again, the office of the blood-vessels, as assimilating organs, is to convert chyle into blood. Chyle, accordingly, cannot be introduced into the arteries without undergoing that change; but *blood* may be introduced from another animal without any injury, and consequently without undergoing any change.

* Fordyce on Digestion, p. 189.

This experiment was first made by Lower, and it has since Chap. III.
 been very often repeated.

Also, if a piece of fresh muscular flesh be applied to the muscle of an animal, they adhere and incorporate without any change, as has been sufficiently established by the experiments of Mr. J. Hunter; and Buvina has ascertained, that fresh bone may, in the same manner, be engrafted on the bones of animals of the same or of different species.*

Foreign substances may be incorporated with the body.

In short, it seems to hold, at least as far as experiments have hitherto been made, that foreign substances may be incorporated with those of the body, provided they be precisely of the same kind with those to which they are added, whether fluid or solid. Thus chyle may be mixed with chyle, blood with blood, muscle with muscle, and bone with bone. The experiment has not been extended to the other animal substances, the nerves, for instance; but it is extremely probable that it would hold with respect to them also.

On the other hand, when substances are introduced into any part of the body which are not the same with that part, nor the same with the substance upon which that part acts, provided they cannot be thrown out readily, they destroy the part, and perhaps even the animal. Thus foreign substances introduced into the blood very soon prove fatal; and introduced into wounds of the flesh or bones, they prevent these parts from healing.

Although the different assimilating organs have the power of changing certain substances into others, and of throwing out the useless ingredients, yet this power is not absolute, even when the substances on which they act are proper for undergoing the change which the organs produce. Thus the stomach converts food into chyme, the intestines chyme into chyle, and the substances which have not been converted into chyle are thrown out of the body. If there happen to be present in the stomach and intestines any substance which, though incapable of undergoing these changes, at least by the action of the stomach and intestines, yet has a strong affinity, either for the whole chyme and chyle, or for some particular part of it, and no affinity for the substances which are thrown out, that substance passes along with the chyle, and in many cases continues to remain chemically combined with the substance to

Powers of the assimilating organs limited.

* Phil. Mag. vi. 303.

Book V.

which it is united in the stomach, even after that substance has been completely assimilated, and made a part of the body of the animal. Thus there is a strong affinity between the colouring matter of madder and phosphate of lime. Accordingly, when madder is taken into the stomach, it combines with the phosphate of lime of the food, passes with it through the lacteals and blood vessels, and is deposited with it in the bones, as was proved by the experiments of Bechier* and Duhamel.† In the same manner, musk, indigo, &c. when taken into the stomach, make their way into many of the secretions.

Assimilation
a chemical
process;

These facts show us that assimilation is a chemical process from beginning to end; that all the changes are produced according to the laws of chemistry; and that we can even derange the regularity of the process by introducing substances whose mutual affinities are too strong for the organs to overcome.

It cannot be denied, then, that the assimilation of food consists merely in a certain number of chemical decompositions which that food undergoes, and the consequent formation of certain new compounds. But are the *agents* employed in assimilation merely chemical agents? We cannot produce any thing like these changes on the food out of the body, and therefore we must allow that they are the consequence of the action of the animal organs. But this action, it may be said, is merely the secretion of particular juices, which have the property of inducing the wished-for change upon the food: and this very change would be produced out of the body, provided we could procure these substances, and apply them in proper quantity to the food. If this supposition be true, the specific action of the vessels consists in the secretion of certain substances; consequently the cause of this secretion is the *real* agent in assimilation. Now, can the *cause* of this secretion be shown to be merely a chemical agent? Certainly not. For in the stomach, where only this secretion can be shown to exist, it is not always the same, but varies according to circumstances. Thus eagles at first cannot digest grain, but they may be brought to do it by persisting in making them use it as food. On the contrary, a lamb cannot at first digest animal

* Phil. Trans. 1736, p. 287.

† Ibid. 1740, p. 390. The fact was mentioned by Mizaldus in a book published in 1566, entitled, *Memorabilum, utilium ac jucundorum Centuriæ novem.*

food, but habit will also give it this power. In this case, it is evident that the gastric juice changes according to circumstances. Chap. III.

The presence of some agent, different from a mere chemical power, will be still more evident, if we consider the immunity of the stomach of the living animal during the process of digestion. The stomach of animals is as fit for food as any other substance. The gastric juice, therefore, must have the same power of acting on it, and of decomposing it, that it has of acting on other substances; yet it is well known that the stomach is not affected by digestion while the animal retains life; though, as Mr. Hunter ascertained, the very gastric juice which the living stomach secretes, often dissolves the stomach itself after death.* Now what is the power which prevents the gastric juice from acting on the stomach during life? Certainly neither a chemical nor mechanical agent, for these agents must still retain the same power after death. We must, then, of necessity conclude, that there exists in the animal an agent very different from chemical and mechanical powers, since it controuls these powers according to its pleasure. These powers, therefore, in the living body, are merely the servants of this superior agent, which directs them so as to accomplish always one particular end. This agent seems to regulate the chemical powers, chiefly by bringing only certain substances together which are to be decomposed, and by keeping at a distance those substances which would interfere with, or diminish, or spoil the product, or injure the organ; and we see that this separation is always attended to even when the substances are apparently mixed together: For the very same products are not obtained, which would be obtained by mixing the same substances together out of the body, that are produced by mixing them in the body; consequently all the substances are not left at full liberty to obey the laws of their mutual affinities. The superior agent, however, is not able to exercise an unlimited authority over the chemical powers sometimes they are too strong for it: some substances, accordingly, as madder, make their way into the system; while others, as arsenic, decompose and destroy the organs of the body themselves.

Not the
agent not
chemical.

But it is not in digestion alone that this superior agent makes the most wonderful display of its power; it is in the last part

* Phil. Trans. 1772, p. 447.

Book V.

of assimilation that our admiration is most powerfully excited. How comes it that the precise substances wanted are always carried to every organ of the body? How comes it that fibrin is always regularly deposited in the muscles, and phosphate of lime in the bones? And, what is still more unaccountable, how comes it that prodigious quantities of some one particular substance are formed and carried to a particular place, in order to supply new wants which did not before exist? A bone, for example, becomes diseased and unfit for the use of the animal; a new bone therefore is formed in its place, and the old one is carried off by the absorbents. In order to form this new bone, large quantities of phosphate of lime are deposited in a place where the same quantity was not before necessary. Now, who informs this agent that an unusual quantity of phosphate of lime is necessary, and that it must be carried to that particular place? Or, granting, as is most probable, that the phosphate of lime of the old bone is partly employed for this purpose, who taught this agent that the old bone must be carried off, new-modelled, and deposited and assimilated anew? The same wonders take place during the healing of every wound, and the renewing of every diseased part.

But neither in this case is the power of this agent over the chemical agents which are employed absolute. We may prevent a fractured bone from healing, by giving the patient large quantities of acids. And unless the materials for new-wanted substances be supplied by the food, they cannot in many cases be formed at all. Thus the canary-bird cannot complete her eggs unless she be furnished with lime.

1. Nature of this agent.

As this agent which characterises living bodies does not appear to act according to the principles of chemistry, any inquiry into its nature would be foreign to the subject of this work. Physiologists have given it the name of the *living* or *animal* principle; and to them I beg leave to refer the reader.

2. Blood furnishes the materials of secretion.

Besides the different organs of the body, the blood is also employed in forming all the different secretions which are necessary for the purposes of the animal economy. These have been enumerated in the last chapter. The process is similar to that of assimilation, and undoubtedly the agents in both cases are the same: but we are equally ignorant of the precise manner in which secretion is performed as we are of assimilation.

After these functions have gone on for a certain time,

which is longer or shorter according to the nature of the animal, the body gradually decays, at last all its functions cease completely, and the animal dies. The cause of this must appear very extraordinary, when we consider the power which the animal has of renewing decayed parts; for it cannot be doubted the death proceeds, in most cases at least, from the body becoming incapable of performing its functions. But if we consider that this power is limited, and that it must cease altogether when those parts of the system begin to decay which are employed in preparing materials for future assimilation, our surprise will, in some measure, cease. It is in these parts, in the organs of digestion and assimilation, accordingly, that this decay usually proves fatal. The decay in other parts destroys life only when the waste is so rapid that it does not admit of repair.

What the reason is that the decay of the organs causes death, or, which is the same thing, causes the living principle either to cease to act, or to leave the body altogether, it is perfectly impossible to say, because we know too little of the nature of the living principle, and of the manner in which it is connected with the body. The last is evidently above the human understanding; but many of the properties of the living principle have been discovered: and were the facts already known properly arranged, and such general conclusions drawn from them as their connection with each other fully warrant, a degree of light would be thrown upon the animal economy, which those who have not attended to the subject are not aware of.

No sooner is the animal dead, than the chemical and mechanical agents, which were formerly servants, usurp the supreme power, and soon decompose and destroy that very body which had been in a great measure reared by their means. But the changes which take place upon animal bodies after death, are too important to be passed over slightly. They shall therefore form the subject of the next chapter.

CHAP. IV.

OF THE DECOMPOSITION OF ANIMAL BODIES.

THE rapidity with which animal bodies undergo decomposition, and the disgusting fetor which accompanies this

Book V.

decomposition, have long been considered as some of their most striking peculiarities. This spontaneous destruction is denominated *putrefaction*. Considerable attention has been paid to it by chemists. Becchar and Stahl have described with fidelity the phenomena with which it is attended, and the circumstances necessary for its taking place. Several curious remarks on it were made by Boyle and Beale.* To Sir John Pringle we are indebted for some important experiments on the method of retarding putrefication; † neither are the experiments of Dr. Macbride less valuable, though the consequences which he drew from them were erroneous. We are indebted also to Crell and Priestley for many valuable facts; and to Berthollet and Lavoisier for the first attempts to determine the real changes which take place, and the manner in which the new products which appear during putrefaction are formed. But notwithstanding the labours of these philosophers, and of many others, much is still wanting to enable us to trace the complicated changes which take place during putrefaction, and to account for them in a satisfactory manner.

Conditions
requisite for
putrefac-
tion.

It has been ascertained long ago, that putrefaction never takes place in those animal substances which contain only two or three ingredients, such as oils, resin, sugar: they must always be more complicated in their texture: and perhaps in all cases a mixture of two or more compound bodies is necessary for speedy decomposition. But however complicated the animal substance may be, it does not putrefy unless moisture be present; for dry animal substances are not susceptible of alteration. A certain degree of heat is also necessary. Animal bodies may be kept without decomposing for any length of time at the freezing temperature. In general, the higher the temperature the more rapid is the putrefaction, provided the heat be not great enough to reduce the animal body to dryness. It has been observed, too, that putrefaction advances with more rapidity in the open air; but exposure to the air is not necessary, though it modifies the decomposition.

Putrefaction
in air.

When these conditions are observed, and dead animal matter is left to itself, its colour becomes gradually paler, and its consistence diminishes: if it be a solid part, such as flesh, it softens, and a serous matter sweats out, whose colour quickly changes; the texture of the part becomes relaxed, and its

* Phil. Trans. iv. 1135.

† Ibid. xvi. 480, &c.

organization is destroyed; it acquires a disagreeable smell; the substance gradually sinks down, and is diminished in bulk; its smell becomes strongly ammoniacal. If the subject be contained in a close vessel, the progress of putrefaction at this stage seems to slacken; no other smell but that of a pungent alkali is perceived; the matter effervesces with acids, and converts syrup of violets to a green. But if the communication with the air be admitted, the urinous exhalation is dissipated, and a peculiar putrid smell is spread around with a kind of impetuosity; a smell of the most insupportable kind, which lasts a long time, and pervades every place, affecting the bodies of living animals after the manner of a ferment, capable of altering the fluids: this smell is corrected, and as it were confined by ammonia. When the latter is volatilized, the putrefaction process becomes active a second time, and the substance suddenly swells up, becomes filled with bubbles of air, and soon after subsides again. Its colour changes, the fibrous texture of the flesh being then scarcely distinguishable; and the whole is changed into a soft brown or greenish matter, of the consistence of a poultice, whose smell is faint, nauseous, and very active on the bodies of animals. The odorant principle gradually loses its force, the fluid portion of the flesh assumes a kind of consistence, its colour becomes deeper, and it is finally reduced into a friable matter, rather deliquescent: which, being rubbed between the fingers, breaks into a coarse powder like earth. This is the last state observed in the putrefaction of animal substances; they do not arrive at this term but at the end of a considerable time.*

During this decomposition a variety of gaseous bodies are emitted: these vary according to the substance exposed to putrefaction; but they consist chiefly of hydrogen gas, holding sulphur, phosphorus, and carbon in solution; of ammonia, water, and carbonic acid, and perhaps also of azotic gas. Nitric acid seems in some cases to be formed and emitted. The earthy-like residuum, which remains after the decomposition is completed, consists of the fixed parts of the animal substance, mixed with charcoal, oil, and ammonia. Thus it appears that putrefaction consists in a total decomposition of the animal body; the elements of which combine together two and two, and thus form a new set of less complicated bodies. But any attempt to explain the manner in

* Fourcroy.

Book V.

which these changes take place would be exceedingly imperfect indeed; not only because we are ignorant of the strength of the affinities of the different elementary parts of animal bodies for each other, but because we do not even know the manner in which these elements are combined, and consequently we cannot know by what particular forces these compounds are destroyed.

Putrefaction under ground.

In carcases buried in the earth, putrefication takes place much more slowly; but it is scarcely possible to observe its progress with accuracy. The abdomen is gradually dilated with elastic fluids, which make their appearance in it, and at last it bursts and discharges a horribly fetid and noxious gas; at the same time a dark coloured liquid flows out. If the earth be very dry, and the heat considerable, the moisture is often absorbed so rapidly, that the carcase, instead of putrifying, dries, and is transformed into what is called a *mummy*.

Putrefaction of accumulated animal matter.

Such are the phenomena when dead bodies are left to putrify separately: but when great numbers of carcases are crowded together in one place, and are so abundant as to exclude the action of external air and other foreign agents, their decomposition is entirely the consequence of the reciprocal action of their ingredients themselves upon each other, and the result is very different. The body is not entirely dissipated or reduced to mould, but all the soft parts are found diminished remarkably in size, and converted into a peculiar *saponaceous matter*. This singular change was first accurately observed in the year 1786.

Converted into a saponaceous matter.

The burial ground of the Innocents in Paris having become noxious to those who lived in its neighbourhood, on account of the disagreeable and hurtful odour which it exhaled, it was found necessary to remove the carcases to another place. It had been usual to dig very large pits in that burial ground, and to fill them with the carcases of the poorer sort of people, each in its proper bier; and when they were quite full, to cover them with about a foot deep of earth, and to dig another similar pit, and fill it in the same manner. Each pit held between 1000 and 1500 dead bodies. It was in removing the bodies from these pits that this saponaceous substance was found. The grave-diggers had ascertained, by long experience, that about thirty years were required before all the bodies had undergone this change in its full extent.* Every

* Fourcroy, Ann. de Chim. v. 154.

part of the body acquired the properties of this substance. The intestines and viscera of the thorax had completely disappeared; but what is singular enough, the brain had lost but little of its size or appearance, though it was also converted into the same substance.

Chap. IV.

This saponaceous matter was of a white colour, soft and unctuous to the touch, and melted, when heated, like tallow. It exhibited all the properties of a *soap*, containing, however, an excess of fatty matter. Fourcroy, who analysed it, found that it was composed of a fatty matter combined with ammonia, and that it contained also some phosphate of lime and ammonia. Diluted acids decomposed it, and separated the fatty matter: alkalies and lime, on the other hand, drove off the ammonia. When exposed to the air it gradually lost its white colour; the ammonia, in a great measure, evaporated; and what remained had something of the appearance of wax. It absorbed water with great avidity, and did not part with it readily. Its white colour was owing to the presence of that liquid. The oily matter, when separated by means of a diluted acid, was concrete, and of a white colour, owing to the mixture of a quantity of water. When dried, it acquires a greyish-brown colour, a lamellar and crystalline texture, like that of spermaceti; but if it has been rapidly dried, it assumes the appearance of wax.* It melts when heated to 126° ; when properly purified, by passing it through a linen cloth while fluid, it has scarcely any smell. Alcohol does not act upon it while cold, but at the temperature of 120° it dissolves it: when the solution cools, the fatty matter precipitates, and forms a gritty mass. With alkalies it forms a soap; and when set on fire it burns precisely like oil or fat, only that it exhales a more unpleasant odour.†

Its properties.

Dr. Smith Gibbes found the same substance in the pit into which animal matters are thrown at Oxford after dissection. A small stream of water constantly passes through this pit; a circumstance which induced him to try whether animal muscle exposed to the action of a running stream underwent the same change. The experiment succeeded completely: he attempted, in consequence, to render this substance, to which

* Chevreul has shown that it is a compound of two substances, one having nearly the properties of margaric acid, the other of fluid fat. Ann. de Chim. xcvi. 25.

† Fourcroy, Ann. de Chim. viii. 17. A set of experiments on a similar substance have been published by Dr. Eimke. See Gehlen's Jour. iv. 139.

Book V. he gave the name of *spermaceti*, useful in those manufactures which require tallow; but the fetid odour which it constantly exhales was an unsurmountable objection.* Attempts were indeed made to get over it; and a manufacture of Dr. Smith Gibbes's spermaceti was even established near Bristol.

Substances
which re-
tard putre-
faction.

Many attempts have been made to retard the destructive progress of putrefaction, in order to preserve animal bodies either as food or for other useful purposes; and several methods have been ascertained which prevent it from operating for a considerable time.

1. The freezing temperature is a complete preservative from putrefaction, as long as the animal substance is exposed to it. Hence the common practice of keeping meat in snow in the frozen climates of the north: and of packing fish in ice, and sending them in that state from Scotland to the London market.

2. Almost all bodies which have a strong affinity for water retard putrefaction for a longer or shorter time, doubtless by depriving the animal substances of their water, or preventing that liquid from acting upon these bodies in its usual manner. In this way the acids, sugar, alcohol, &c. seem to prevent or retard putrefaction.

3. It is well known that common salt is a powerful antiseptic. Hence the practice of salting meat, and the length of time which meat that has undergone this operation may be kept. Several other salts, especially nitre, possess the same property. In what manner these bodies act has not been ascertained; but they undoubtedly produce some chemical change upon the meat; for they alter its taste, its colour, and other sensible properties.

4. Many aromatics, such as camphor, resins, volatile oils, bitumens, and other similar bodies, act with considerable efficacy in preserving animal bodies from putrefaction. Hence their utility in embalming. In what the action of these substances consists has not been ascertained. Part of their efficacy is doubtless owing to the rapidity with which the animal substances to which they are applied lose their moisture; and something may be ascribed likewise to their odour, which keeps insects at a distance, and thus prevents the lodging of excrementitious matter, which always acts powerfully as a putrefactive ferment.

* Phil. Trans. 1794, 1795

APPENDIX.

*Tables exhibiting the weight of the atom of all substances
hitherto accurately examined.*

In the first table, the bodies are placed according to the weight of the atoms; in the second, in the alphabetical order.

TABLE I.

	Weight of atom, hy- drogen = 1.	Weight of atom, oxy- gen = 1.
Hydrogen	1	0·125
Carbon	6	0·75
Boron	7	0·875
Oxygen	8	1
Silicon	8	1
Carburetted hydrogen	8	1
Aluminum	9	1·125
Water	9	1·125
Lithium	10	1·25
Magnesium	12	1·5
Phosphorus	12	1·5
Hydroguret of phosphorus	13	1·625
Azote	14	1·75
Carbonic oxide	14	1·75
Olefiant gas	14	1·75
Bihydroguret of phosphorus ..	14	1·75
Sulphur	16	2
Silica	16	2
Fluorine	16	2
Ammonia	17	2·125
Alumina	17	2·125
Sulphuretted hydrogen	17	2·125
Deutoxide of hydrogen	17	2·125
Glucinum	18	2·25

	Weight of atom, hy- drogen=1.	Weight of atom, oxy- gen=1.
Lithia	18	2.25
Phosphuret of carbon	18	2.25
Magnesia	20	2.5
Phosphorous acid	20	2.5
Calcium	20	2.5
Carbonic acid	22	2.75
Protoxide of azote	22	2.75
Boracic acid	23	2.875
Fluoboric acid	23	2.875
Alcohol	23	2.875
Sodium	24	3
Hyposulphurous acid	24	3
Fluosilicic acid	24	3
Glucina	26	3.25
Cyanogen	26	3.25
Nickel	26	3.251
Cobalt	26	3.25
Hydrocyanic acid	27	3.375
Iron	28	3.5
Manganese	28	3.5
Chromium	28	3.5
Phosphoric acid	28	3.5
Phosphuret of sulphur	28	3.5
Lime	28	3.5
Deutoxide of azote	30	3.75
Nephrin	30	3.75
Soda	32	4
Sulphurous acid	32	4
Tellurium	32	4
Yttrium	32	4
Zinc	34	4.25
Protoxide of nickel	34	4.25
Protoxide of cobalt	34	4.25
Chlorine	36	4.5
Protoxide of iron	36	4.5
Protoxide of manganese	36	4.5
Oxalic acid	36	4.5
Muriatic acid	37	4.625
Formic acid	37	4.625
Zirconium?	37	4.625
Sulphuric ether	37	4.625

	Weight of atom, hy- drogen = 1.	Weight of atom, oxy- gen = 1.
Arsenic	38	4.75
Hyponitrous acid	38	4.75
Sulphuret of carbon	38	4.75
Potassium	40	5
Yttria	40	5
Sulphuric acid	40	5
Oxide of tellurium	42	5
Sulphuret of sodium	40	5
Selenium	41	5.125
Oxide of zinc	41	5.25
Seleniuretted hydrogen	42	5.25
Sulphuret of cobalt	42	5.28
Strontium	44	5.5
Peroxide of manganese	44	5.5
Protoxide of chlorine	44	5.5
Sulphuret of iron	44	5.5
Purpuric acid?	44	5.5
Antimony	45	5.625
Zirconia?	45	5.625
Uric acid?	45	5.625
Cerium	46	5.75
Nitrous acid	46	5.75
Lactic acid	46	5.75
Chloride of lithium	46	5.75
Molybdenum	48	6
Potash	48	6
Iridium	48	6
Protochloride of phosphorus ..	48	6
Chloride of magnesium	48	6
Oxide of selenium	49	6.125
Mellitic acid?	49	6.125
Sulphuret of zinc	50	6.25
Succinic acid	50	6.25
Chlorocarbonic acid	50	6.25
Chloric ether	50	6.25
Acetic acid	51	6.375
Strontian	52	6.5
Chromic acid	52	6.5
Manganesic acid	52	6.5
Chloride of sulphur	52	6.5
Nitric acid	54	6.75

	Weight of atom, hy- drogen=1.	Weight of atom, oxy- gen=1.
Protoxide of cerium	54	6.75
Arsenious acid	54	6.75
Cadmium	56	7
Palladium	56	7
Protoxide of molybdenum	56	7
Sulphuret of potassium	56	7
Chloride of calcium	56	7
Selenic acid	57	7.125
Tin	59	7.375
Citric acid	59	7.375
Sulpho-chyazic acid?	59	7.375
Chloride of sodium	60	7.5
Persulphuret of iron	60	7.5
Antimonic acid	61	7.625
Arsenic acid	62	7.75
Chlorocyanic acid	62	7.75
Gallic acid	63	7.875
Copper	64	8
Oxide of cadmium	64	8
Oxide of palladium	64	8
Peroxide of potassium	64	8
Chloride of iron	64	8
Molybdous acid	64	8
Muriatic ether	65	8.125
Sesquisulphuret of selenium	65	8.125
Tartaric acid	67	8.375
Protoxide of tin	67	8.375
Deutoxide of chlorine	68	8.5
Sub-bichloride of sulphur	68	8.5
Chloride of zinc	70	8.75
Barium	70	8.75
Bismuth	71	8.875
Tannin	71	8.875
Molybdic acid	72	9
Protoxide of copper	72	9
Peroxide of sodium	72	9
Sulphuret of cadmium	72	9
Sulphuret of palladium	72	9
Hyposulphuric acid	72	9
Peroxide of tin	75	9.375
Protosulphuret of tin	75	9.375

	Weight of atom, hy- drogen=1.	Weight of atom, oxy- gen=1.
Chloric acid	76	9.5
Chloride of potassium	76	9.5
Peroxide of cobalt	76	9.5
Peroxide of nickel	76	9.5
Barytes	78	9.75
Oxide of bismuth	79	9.875
Acetic ether	79	9.875
Peroxide of copper	80	10
Peroxide of iron	80	10
Sulphuret of copper	80	10
Sugar	81	10.125
Perchloride of phosphorus	84	10.5
Peroxide of barium	86	10.75
Sulphuret of bismuth	87	10.875
Chloride of strontium	88	11
Gum	90	11.25
Persulphuret of tin	91	11.375
Perchloric acid	92	11.5
Chloride of cadmium	92	11.5
Protochloride of tin	95	11.875
Tungsten	96	12
Perchloride of iron	100	12.5
Protochloride of copper	100	12.5
Persulphuret of bismuth	103	12.875
Lead	104	13
Saccharic acid	105	13.125
Chloride of barium	106	13.25
Chloride of bismuth	107	13.375
Realgar	108	13.5
Silver	110	13.75
Protoxide of lead	112	14
Brown oxide of tungsten	112	14
Bichloride of selenium	113	14.125
Peroxide of cerium	116	14.5
Seleniuret of arsenic	117	14.625
Oxide of silver	118	14.75
Rhodium	120	15
Peroxide of lead	120	15
Sulphuret of lead	120	15
Tungstic acid	120	15

	Weight of atom, hy- drogen=1.	Weight of atom, oxy- gen=1.
Benzoic acid	120	15
Orpiment	124	15.5
Iodine	125	15.625
Uranium	125	15.625
Hydriodic acid	126	15.75
Sulphuret of silver	126	15.75
Protoxide of rhodium	128	16
Perchloride of tin	131	16.375
Protoxide of uranium	133	16.625
Sulphuret of tin	134	16.75
Persulphuret of lead	136	17
Perchloride of copper	136	17
Deutoxide of rhodium	136	17
Iodide of phosphorus	137	17.125
Chloride of lead	140	17.5
Iodide of sulphur	141	17.625
Starch	142	17.75
Columbium	144	18
Titanium?	144	18
Peroxide of rhodium	144	18
Iodide of calcium	145	18.125
Chloride of silver	146	18.25
Cubic pyrites	146	18.25
Iodide of sodium	149	18.625
Deutosulphuret of tin	150	18.75
Protoxide of titanium	152	19
Columbic acid	152	19
Iodide of iron	153	19.125
Hydriodic ether	154	19.25
Iodide of zinc	158	19.75
Chloride of azote	158	19.75
Peroxide of titanium	160	20
Iodic acid	165	20.625
Iodide of potassium	165	20.625
Iodide of strontium	169	21.125
Iodide of cadmium	181	22.625
Platinum	181	22.625
Iodide of copper	189	23.625
Protoxide of platinum	189	23.625
Iodide of barium	195	24.375

	Weight of atom, hy- drogen = 1.	Weight of atom, oxy- gen = 1.
Iodide of bismuth	196	24.5
Chloriodic acid	197	24.625
Gold	199	24.875
Mercury	200	25
Peroxide of platinum	205	25.625
Protoxide of gold	207	25.875
Protoxide of mercury	208	26
Peroxide of mercury	216	27
Sulphuret of mercury	216	27
Peroxide of gold	223	27.875
Iodide of lead	229	28.625
Red oxide of lead	232	29
Iodide of silver	235	29.375
Periodide of phosphorus ...	262	32.75
Margaric acid	264	33
Peroxide of uranium	274	34.25
Oleic acid	288	36
Morphia	322	40.25
Pierotoxia	360	45
Strychnia	381	47.625
Brucia	412	51.5

TABLE II.

	Weight of atom.		Weight of atom.
A			
Acetic acid	6·375	Chloric acid	9·500
Acetic ether	9·875	Chloride of azote . . .	19·750
Alcohol	2·875	————— barium ..	13·250
Alumina	2·125	————— bismuth .	13·375
Aluminum	1·125	————— cadmium	11·500
Ammonia	2·125	————— calcium .	7·000
Antimonic acid	7·625	————— iron	8·000
Antimony	5·625	————— lead	17·500
Arsenic	4·750	————— lithium .	5·750
Arsenic acid	7·750	————— magnesium	6·000
Arsenious acid	6·750	————— potassium	9·500
Azote	1·750	————— silver . . .	18·250
B			
Barium	8·750	————— sodium .	7·500
Barytes	9·750	————— strontium	11·000
Benzoic acid	15·000	————— sulphur .	6·500
Bichloride of selenium	14·125	————— zinc	8·75
Bihydroguret of } phosphorus }	1·750	Chlorocarbonic acid .	6·250
Bismuth	8·875	Chlorocyanic acid ..	7·750
Boracic acid	2·875	Chlorine	4·500
Boron	0·875	Chloriodic acid	24·625
Brown oxide of } tungsten }	14·000	Chromic acid	6·500
Brucia	51·500	Chromium	3·500
C			
Cadmium	7·000	Citric acid	7·375
Calcium	2·625	Cobalt	3·250
Carbon	0·750	Columbic acid	19·000
Carbonic acid	2·750	Columbium	18·000
Carbonic oxide	1·750	Copper	8·000
Carburetted hydrogen	1·000	Cubic pyrites	18·250
Cerium	5·750	Cyanogen	3·250
Chloric ether	6·250	D	
		Deutoxide of azote ..	3·750
		————— chlorine .	8·500
		————— hydrogen	2·125
		————— rhodium	17·000
		————— tin	18·750

	Weight of atom.		Weight of atom.
E		Iodide of copper	23·625
Ether, acetic	9·875	———— iron	19·125
———— chloric	6·250	———— lead	28·625
———— hydriodic	19·250	———— phosphorus	17·125
———— muriatic	8·125	———— potassium	20·625
———— nitric	10·25?	———— silver	29·375
———— sulphuric	4·625	———— strontium	21·125
		———— sulphur	17·625
F		Iodine	15·625
Fluoboric acid	2·875	Iridium	6·000
Fluorine	2·000	Iron	3·500
Fluosilicic acid	3·000		
Formic acid	4·625	L	
		Lactic acid	5·750
G		Lead	13·000
Gallic acid	7·875	Lime	3·500
Glucina	3·250	Lithia	2·250
Glucinum	2·250	Lithium	1·250
Gold	24·875		
Gum	11·250	M	
		Magnesia	2·500
H		Magnesium	1·500
Hydriodic acid	15·750	Manganese	3·500
———— ether	19·250	Manganetic acid	6·500
Hydrocyanic acid	3·375	Margaric acid	33·000
Hydrogen	0·125	Mellitic acid	6·125?
Hydroguret of phos- } phorus }	1·625	Mercury	24·000
Hyponitrous acid	4·750	Molybdenum	6·000
Hyposulphuric acid	9·000	Molybdic acid	9·000
Hyposulphurous acid	3·000	Molybdous acid	8·000
		Morphia	40·250
I		Muriatic acid	4·625
Iodic acid	20·625	Muriatic ether	8·125
Iodide of barium	24·375		
———— bismuth	24·500	N	
———— cadmium	22·625	Nephrin	3·750
———— calcium	18·125	Nickel	3·250
		Nitrous acid	5·750
		Nitric acid	6·750

	Weight of atom.		Weight of atom.
O			
Olefiant gas	1·750	Peroxide of uranium	34·250
Oleic acid	36·000	Persulphuret of bis- muth	} 12·875
Orpiment	15·500	———— iron	7·500
Oxalic acid	4·500	———— lead	17·000
Oxide of bismuth	9·875	———— tin	11·375
———— cadmium	8·000	Phosphoric acid	3·500
———— palladium	8·000	Phosphorous acid	2·500
———— selenium	6·125	Phosphorus	1·500
———— silver	14·750	Phosphuret of carbon	2·250
———— tellurium	5·000	———— sulphur	3·500
———— zinc	5·250	Picrotoxia	45·000
Oxygen	1·000	Platinum	22·625
P			
Palladium	7·000	Potash	6·000
Perchloric acid	11·500	Potassium	5·000
Perchloride of copper	17·000	Protochloride of copper	} 12·500
———— iron	12·500	———— phosphorus	6·000
———— phosphorus	10·500	———— tin	11·875
———— tin	16·375	Protosulphuret of tin	9·375
Periodide of phos- phorus	} 32·750	Protoxide of azote	2·750
Peroxide of barium	10·750	———— cerium	6·750
———— cerium	14·500	———— chlorine	5·500
———— cobalt	9·500	———— cobalt	4·250
———— copper	10·000	———— copper	9·000
———— gold	27·875	———— gold	25·875
———— iron	10·000	———— iron	4·500
———— lead	15·000	———— lead	14·000
———— manganese	5·500	———— manganese	4·500
———— mercury	27·000	———— mercury	26·000
———— nickel	9·500	———— molybdenum	7·000
———— platinum	25·625	———— nickel	4·250
———— potassium	8·000	———— platinum	23·625
———— rhodium	18·000	———— rhodium	16·000
———— sodium	9·000	———— titanium	19·000
———— tin	9·375	———— tin	8·375
———— titanium	20·000	———— uranium	16·625
		Purpuric acid	—

	Weight of atom.		Weight of atom.
R			
Realgar	13·500	Sulphuret of palladium	9·000
Red oxide of lead ...	29·000	————— potassium .	7·000
Rhodium	15·000	————— silver	15·750
		————— sodium ...	5·000
		————— tin	16·750
		————— zinc	6·250
S			
Sacclactic acid	13·125	Sulphuretted hydro- } gen	2·125
Selenic acid	7·125	Sulphuric acid	5·000
Selenium	5·125	————— ether	4·625
Seleniuret of arsenic	14·625	Sulphurous acid	4·000
Seleniuretted hydrogen	5·250		
Sesquisulphuret of } selenium	8·125	T	
Silica	2·000	Tannin	8·875
Silicon	1·000	Tartaric acid	8·375
Silver	13·750	Tellurium	—————
Soda	4·000	Tin	7·375
Sodium	3·000	Titanium	18·000
Starch	17·750	Tungsten	12·000
Strontian	6·500	Tungstic acid	15·000
Strontium	5·500		
Strychnia	17·625	U	
Sub-bichloride of } sulphur	8·500	Uranium	15·625
Succinic acid	6·250	Uric acid	5·625
Sugar	10·125		
Sulphochyazic acid ..	7·375	W	
Sulphur	2·000	Water	1·125
Sulphuret of bismuth	10·875		
————— cadmium .	9·000	Y	
————— carbon ...	4·750	Ytria	5·000
————— cobalt	5·250	Yttrium	4·000
————— copper ...	10·000		
————— iron	5·500	Z	
————— lead	15·000	Zinc	4·125
————— mercury ..	27·000	Zirconia	5·625
		Zirconium	4·625

TABLE III.

Specific Gravity of Gases.

GASES.	Specific gravity.	Weight of 100 cubic inches, in grains.
Air	1·0000	30·5000
Vapour of iodine	8·6805	264·7550
Hydriodic ether	5·4750	166·9870
Oil of turpentine vapour	5·0130	152·8960
Hydriodic acid	4·3750	133·4340
Flusilicic acid	3·5735	108·0020
Chlorocarbonic acid	3·4722	105·9020
Chloric ether	3·4750	105·2570
Sulphuric acid vapour	2·7777	84·6980
Sulphuret of carbon vapour ..	2·6447	80·6630
Sulphuric ether vapour	2·5666	78·1833
Chlorine	2·5000	76·2500
Protoxide of chlorine	2·4444	74·5555
Fluoboric acid	2·3709	72·3120
Sulphurous acid	2·2222	67·7777
Muriatic ether	2·2190	67·6790
Chlorocyanic acid vapour	2·1520	65·6360
Cyanogen	1·8055	55·0694
Alcohol vapour	1·6133	49·2060
Protoxide of azote	1·5277	46·5972
Carbonic acid	1·5277	46·5972
Muriatic acid	1·28472	39·1839
Sulphuretted hydrogen	1·1805	36·0069
Sulphur vapour	1·1111	33·8888
Oxygen	1·1111	33·8888
Deutoxide of azote	1·0416	31·7708
Hydrocarbonic oxide	0·9930	30·3470
Olefiant gas	0·9722	29·6527
Azote	0·9722	29·6527
Carbonic oxide	0·9722	29·6527
Bihydroguret of phosphorus ..	0·9722	29·6527
Hydrocyanic acid vapour	0·9368	28·5720

GASES.	Specific gravity.	Weight of 100 cubic inches, in grains.
Hydrouret of phosphorus ...	0·90277	27·5376
Phosphorus vapour	0·8333	25·4166
Steam	0·6250	19·0620
Ammonia	0·59027	18·0035
Carburetted hydrogen	0·5555	16·9444
Carbon vapour	0·4166	12·6083
Hydrogen	0·0694	2·1180



INDEX.

A

- ACANTICONE**, vol. iii. p. 274
Aceric acid, iv. 211
Acetates, ii. 142
Acetic acid, ii. 135
 in animals, iv. 442
 in plants, iv. 3
Acetous acid, ii. 138
Acid principle, opinions respecting, ii. 81
Acids, ii. 80
 how distinguished, ii. 226
 with a simple basis, ii. 81
 combustible, ii. 134
 composition of, ii. 225
 freezing of, i. 49
 in plants, iv. 3, and 186
 in animals, iv. 439
Actinolite, iii. 368
Adamantine spar, iii. 254
Adepts, i. 4
Adipocire, iv. 566
Adularia, iii. 330
Aerated alkaline water, ii. 97
Aerial acid, ii. 93
Æs, i. 471
Ætites, iii. 481
Affinity, iii. 1
 homogeneous, iii. 6
 heterogeneous, iii. 7
 whether a definite force, iii. 13
 whether elective, iii. 15
 tables of, iii. 8
Agalmatolite, iii. 359
Agaric mineral, iii. 385
Agaricus acris, iv. 290
 bulbosus, iv. 290
 campestris, iv. 289
 juglandis, iv. 291
 muscarius, *ibid.*
 piperatus, iv. 290
 stypticus, *ibid.*
 theogalus, iv. 291
 volvaccens, iv. 239
Air, iii. 162
 how reduced to a given density, iii. 172
 in the swimming bladders of fishes, iv. 533
 inflammable, i. 217
Alalite, iii. 377
Albumen, iv. 398
 uncoagulated, iv. 401
 coagulated, iv. 405
Alburnum, iv. 312
Alchymists, i. 4
Alcionium, iv. 460
Alcohol, ii. 318
 its composition, ii. 337
 action, on solids, iii. 104
 of sulphur, i. 293
Ale, iv. 373
Algaroth, powder of, i. 552
Alkali, affinities of, iii. 141
 aluminated, ii. 74
 fixed, ii. 49
 fossil, i. 336
 mineral, i. 336
 phlogisticated, ii. 294
 prussian, *ibid.*
 silicated, ii. 103
 vegetable, i. 328
 volatile, ii. 27
Alkalies, ii. 26, and 49
 in animals, iv. 443
 in plants, iv. 187
Alkalinity, i. 323
Allagite, iii. 497
Allanite, iii. 501
Allochroite, iii. 282
Allophanite, iii. 303
Alloys, table of, i. 547
Alluvial formations, iii. 579
Almonds, iv. 270
Aloes, iv. 163
Alum, ii. 537, and iii. 418
 burnt, ii. 538
 cubic, ii. 540
 earth, iii. 434
 slate, iii. 316

Alun stone, iii. 340
 Alumina, i. 373, and ii. 74
 salts of, ii. 535
 acetate of, ii. 542
 arsenate, *ibid.*
 benzoate, ii. 543
 borate, ii. 536
 camphorate, ii. 543
 carbonate of, ii. 536
 fluante, ii. 283
 gallate, ii. 544
 gelatinous, i. 373
 hyposulphite, ii. 542
 lactate, ii. 544
 malate, *ibid.*
 mellate, *ibid.* and iii. 419
 muriate, ii. 255
 nitrate, ii. 535
 oxalate, ii. 543
 phosphate, ii. 536, and
 iii. 420
 sac lactate, ii. 544
 seleniate, ii. 542
 silicate, ii. 536, and iii.
 422
 soap, ii. 716
 soluble in alkalies, ii. 74
 spongy, i. 373
 suberate, ii. 543
 succinate, *ibid.*
 sulphate, ii. 536, 540
 subsulphate, iii. 418
 sulphite, ii. 541
 supersulphate, ii. 537
 tartrate, ii. 544
 tungstate, ii. 542
 urate, ii. 544
 Aluminite, iii. 418
 Aluminum, i. 372
 Amalgam, i. 491
 of silver, iii. 443
 Amber, iii. 428, and iv. 144
 salt of, ii. 150
 varnish, iv. 146
 Ambergris, iv. 435
 Amblygonite, iii. 335
 Amethyst, iii. 291
 Amianthus, iii. 364
 Amiatiti, iii. 300
 Ammonia, i. 224, and ii. 26
 acetate of, ii. 419
 action of potassium on,
 ii. 36

Ammonia antimoniate, ii. 418
 arseniate, ii. 417
 arsenite, ii. 418
 benzoate, ii. 420
 bicarbonate, ii. 414
 borate, ii. 414
 camphorate, ii. 420
 carbonate, ii. 413
 chlorate, ii. 230
 chromate, ii. 418
 citrate, ii. 422
 composition of, ii. 36
 dissolves oxides, ii. 32
 fluante, ii. 281
 fluosilicate, ii. 287
 formate, ii. 423
 hydriodate, ii. 276
 hydrosulphuret, ii. 698
 hypophosphite, ii. 415
 hyposulphite, ii. 416
 iodate, ii. 272
 lactate, ii. 422
 liquid, ii. 27
 malate, ii. 422
 mellate, *ibid.*
 molybdate, ii. 418
 moroxylate, ii. 420
 muriate, ii. 248
 nitrate, ii. 411
 oxalate, ii. 421
 phosphate, ii. 414
 phosphite, ii. 415
 sac lactate, ii. 422
 salts of, ii. 411
 seleniate, ii. 417
 soap, ii. 715
 suberate, ii. 421
 succinate, ii. 420
 sulphate, ii. 415
 sulphite, ii. 416
 sulphuret, ii. 30
 tartrate, ii. 422
 tungstate, ii. 418
 urate, ii. 422
 Ammoniac, iv. 162
 fixed, i. 350
 vitriolated, ii. 415
 Ammonium, ii. 36
 Amnios, liquor of, iv. 527
 human, iv. 528
 of cows, iv. 530
 Amniotic acid, iv. 441
 Amphibole, iii. 366

- Amphigene, iii. 277
 Amygdaloid, iii. 569
 Analcime, iii. 320
 Anatase, iii. 518
 Anatto, iv. 272
 Andalusite, iii. 326
 Andreolite, iii. 321
 Andropogon schænanthus, iv. 225
 Anhydrite, iii. 406
 Animal poisons, iv. 531
 heat, iv. 617
 substances, iv. 390, 443
 Animals, iv. 388
 decomposing of, iv. 631
 functions of, iv. 588
 parts of, iv. 445
 saccharine matters of,
 iv. 423
 Anime, iv. 134
 Anthrazothion, ii. 302
 Anthracolite, iii. 389
 Antimonial silver ore, iii. 448
 sulphuret of silver, iii.
 450
 Antimonic acid, i. 553, and ii. 128
 Antimonious acid, i. 552, and ii.
 129
 Antimony, i. 549
 acetate of, ii. 692
 alloys of, i. 556
 argentine flowers of,
 i. 551
 salts of, ii. 689
 benzoate, ii. 692
 butter of, i. 554
 chloride, i. 554
 expansion of, i. 44
 grey ore of, iii. 544
 hydrosulphuret, ii.
 707
 native, iii. 543
 nickeliferous sulphu-
 ret, iii. 545
 ochre, iii. 546
 ores of, iii. 543
 ores, analysis of, iii. 613
 oxalate, ii. 692
 oxides of, i. 551
 phosphuret, i. 555
 red ore of, iii. 546
 seleniuret, i. 556
 succinate, ii. 692
 Antimony, sulphuret, i. 555
 tartrate, ii. 692
 white ore of, iii. 546
 Ants, acid of, ii. 195
 oil of, iv. 434
 resin of, iv. 439
 Apatite, iii. 400
 Aphrite, iii. 385
 Aplome, iii. 286
 Apophyllite, iii. 319
 Aqua fortis, ii. 82
 Aqueous humour of the eye, iv.
 519
 Aquila alba, i. 485
 mitigata, *ibid.*
 Arbor Dianæ, i. 497
 Arcanum duplicatum, ii. 435
 tartari, ii. 442
 Archil, iv. 285
 Arctizite, iii. 336
 Arendate, iii. 274
 Argentine, iii. 384
 flowers of antimony, i.
 552
 Argentum vivum, i. 479
 Argil, i. 373
 Argillite, iii. 347
 Arnica montana, iv. 249
 Arragonite, iii. 390
 Arseniates, ii. 125
 Arsenic, i. 304
 acid, i. 306
 alloys, i. 335, 342, 402,
 409, 436, 449, 459,
 469, 491, 500, 508,
 525, 556, 566
 bloom, iii. 411
 butter, i. 308
 chloride of, *ibid.*
 hydrosulphuret, ii. 710
 iodide of, i. 309
 native, iii. 548
 native oxide, iii. 550
 ores, iii. 547, 614
 oxides of, i. 305
 phosphuret of, i. 311
 sulphuret of, *ibid.*
 white, i. 305
 Arsenical neutral salt of Macquer,
 ii. 438
 pyrites, iii. 548
 Arsenious acid, i. 305, and ii. 125

Arsenites, ii. 126
 Arseniuretted hydrogen, i. 309
 Asafœtida, iv. 165
 Asbestos, iii. 364
 Asparagin, iv. 63
 Asparagus, iv. 241
 stone, iii. 401
 Asphalt, ii. 398
 Assimilation, iv. 625
 Atmosphere, iii. 160
 as food of plants, iv.
 315
 composition of, iii.
 165
 stones from, iii. 184
 unknown bodies in,
 iii. 181
 Atomic theory, account of, iii. 17
 Atropa belladonna, iv. 242
 Attraction explained, iii. 2
 Aubier, iv. 312
 Augite, iii. 376
 Auriferous silver, iii. 447
 Auripigmentum, iii. 550
 Aurum graphicum, iii. 441
 musivum, musicum, or
 mosaicum, i. 458
 paradoxicum, iii. 441
 Automalite, iii. 251
 Axestone, iii. 361
 Axinite, iii. 276
 Azote, i. 202
 chloride of, i. 210
 deutoxide of, i. 207, and
 i. 7.
 iodide of, 213
 oxides of, ii. 3
 protoxide of, i. 206, and ii. 3
 given out by plants, iv. 345
 Azotic gas, i. 203
 Azotites, ii. 7
 Azure stone, iii. 421
 Azure de cuivre, iii. 464

B

Baikalite, iii. 375
 Balass ruby, iii. 252
 Balm of Gilead, iv. 152
 Balsam of Canada, iv. 131
 of Copaiva, iv. 152
 of Peru, iv. 154
 of sulphur, ii. 367

Balsam of Tolu, iv. 154
 iv. 151
 liquid, *ibid.*
 solid, iv. 157
 Baltic, iii. 205
 Balneum regale, i. 558
 Barilla, i. 336, and ii. 451
 Barium, i. 352
 chloride of, i. 356
 iodide of, i. 358
 peroxide of, i. 356
 Barley, iv. 259
 malting of, iv. 366
 Baroselenite, iii. 413
 Barote, i. 353
 Barras, iv. 131
 Barytes, i. 352, and ii. 66
 salts of, ii. 496
 acetate of, ii. 503
 antimoniate, *ibid.*
 antimonite, *ibid.*
 arsenate, ii. 502
 arsenite, *ibid.*
 benzoate, ii. 504
 boletate, ii. 505
 borate, ii. 498
 camphorate, ii. 504
 carbonate, ii. 497, and
 iii. 412
 chlorate, ii. 234
 chromate, ii. 503
 citrate, ii. 506
 fluante, ii. 282
 fluosilicate, ii. 288
 formate, ii. 506
 gallate, ii. 507
 hydrates of, ii. 67
 hydriodate, ii. 277
 hydrosulphuret, ii. 700
 hypophosphite, ii. 499
 hyposulphite, ii. 501
 hyposulphate, ii. 502
 iodate, ii. 274
 lactate, ii. 507
 malate, ii. 506
 mellate, ii. 505
 muriate, i. 357
 nitrate, ii. 496
 oxalate, ii. 505
 phosphate, ii. 498
 phosphite, ii. 499
 phosphuret, ii. 68

- Barytes, sacclactate, ii. 506
 seleniate, ii. 502
 silicate, ii. 498
 suberate, ii. 505
 succinate, ii. 504
 sulphate, ii. 500, and iii.
 413
 sulphite, ii. 501
 sulphuret, ii. 68
 tannate, ii. 507
 tartrate, ii. 506
 tungstate, ii. 503
 urate, ii. 506
 water, ii. 64
 Basalt, iii. 380
 Basaltine, octahedral, iii. 376
 Basanos, iii. 296
 Bdellium, iv. 170
 Bean, iv. 266
 Bee, venom of, iv. 533
 Beer, iv. 366
 Bell metal, i. 473
 Benzoates, ii. 148
 Benzoic acid, ii. 146
 plants containing,
 iv. 7
 in animals, iv. 440
 Benzoin, iv. 157
 flowers of, ii. 146
 Bergmannite, iii. 336
 Bergman's doctrine of affinity, iii.
 11
 Berthollet's doctrine of affinity, iii.
 11
 Beryl, iii. 265
 shortous, iii. 261
 Bezoardic acid, ii. 184
 Bezoars, iv. 562
 Bihydroguret of carbon, i. 235
 of phosphorus, i. 272
 Bildstein, iii. 359
 Bile, iv. 509, 598
 of ateles arachnoides, iv. 512
 fishes, *ibid.*
 fowls, iv. 511
 human, iv. 512
 the sow, iv. 511
 use of, iv. 596
 Biliary calculi, iv. 565
 Birdlime, iv. 118
 Bismuth, i. 475
 alloys of, i. 492, 501, 511,
 526, 532, 567, 573
 salts, ii. 652
 Bismuth, acetate of, ii. 655
 arseniate, *ibid.*
 benzoate, *ibid.*
 borate, ii. 654
 butter, i. 477
 carbonate, ii. 654
 chloride, i. 477
 cupreous sulphuret, iii.
 528
 expansion of, i. 44
 glance, iii. 527
 iodide, i. 478
 magistery, ii. 653
 molybdate, ii. 655
 muriate, ii. 262
 native, iii. 527
 needle ore of, iii. 528
 nitrate, ii. 652
 ochre, iii. 529
 ores, iii. 526
 analysis of, iii. 614
 oxalate, ii. 656
 oxides of, i. 476
 phosphate, ii. 654
 seleniuret, i. 479
 succinate, ii. 655
 sulphate, ii. 654
 sulphite of, *ibid.*
 sulphuret, iii. 478
 tartrate, ii. 656
 Bitter principles, iv. 49
 artificial, iv. 34
 Bittern, ii. 521
 Bitter spar, iii. 394
 Bitumens, ii. 393, and iii. 430
 proper, ii. 398
 Bituminous oils, ii. 393
 marl slate, iii. 399
 Black, Dr. his theory of latent
 heat, i. 51
 Black cobalt ore, iii. 508
 copper ore, iii. 459
 jack, iii. 522
 lead, i. 227
 lead ore, iii. 535
 ore of manganese, iii. 494
 ore of tellurium, iii. 442
 poplar resin, iv. 136
 Bleaching powder, ii. 239
 Bleischimmer, iii. 533
 Bleiniere, iii. 541
 Blende, iii. 522
 Blight, iv. 302
 Blisters, liquor of, iv. 587

- Blood, iv. 489
 colouring matter of, iv. 412
 effect of different gases on,
 iv. 608
 how altered by respiration,
 iv. 613
 diseased, iv. 493
 of the fœtus, *ibid.*
- Blowpipe, iii. 242
- Blue carbonate of copper, iii. 464
 calx of copper, *ibid.*
 iron earth, iii. 485
 lead ore, iii. 531
 flowers, iv. 245
 spar, iii. 325
 vesuvian limestone, iii. 388
- Bog-iron ore, iii. 482
- Boiling points, table of, i. 58
 explained, *ibid.*
- Bole, iii. 357
- Boletates, ii. 156
- Boletic acid, ii. 155
- Boletus igniarius, iv. 293
 laricis, iv. 292
 pseudo igniarius, iv. 293
 viscidus, iv. 291
- Bologna stone, ii. 500
- Bones, iv. 446
 fossil, iv. 452
 earth of, iv. 448
- Boracic acid, i. 241, and ii. 100
 composition of, i.
 245
- Boracite, iii. 417
- Borax, i. 241, ii. 454, and iii. 426
- Boron, i. 241
 chloride of, i. 245
- Botany Bay resin, iv. 134
- Botryolite, iii. 409
- Bournonite, iii. 531
- Boyle, fuming liquor of, ii. 30,
 698
- Brain, iv. 474
- Brandy, ii. 318
- Brass, i. 470
- Brassica oleracea, iv. 244
- Brazil wood, iv. 227
- Brewing, iv. 369
- Brimstone, i. 277
- Brionia alba, iv. 218
- Brittle silver ore, iii. 449
- Bronze, i. 473
- Bronzite, iii. 373
- Brown cobalt ochre, iii. 508
- Brown hematite, iii. 480
 iron ore, iii. 480
 lead ore, iii. 537
 ore of titanium, iii. 521
 spar, iii. 397
- Brucia, ii. 45
 muriate of, ii. 253
 salts of, ii. 475
- Bucholzite, iii. 422
- Buffy coat of blood, iv. 494
- Bulbs, iv. 276
- Buntkupfererz, iii. 457
- Butter, iv. 496
 of bismuth, i. 477
 of wax, iv. 104
- C
- Cadmia, i. 429
- Cadmium, i. 437
 chloride, i. 440
 iodide, i. 441
 oxide, i. 439
 phosphuret, i. 442
 sulphuret, i. 449
 muriate, ii. 261
 salts of, ii. 600
 acetate, ii. 602
 carbonate, ii. 600
 citrate, ii. 602
 nitrate, ii. 601
 oxalate, ii. 602
 phosphate, *ibid.*
 sulphate, *ibid.*
 tartrate, *ibid.*
- Caffein, iv. 52
- Calaguala, iv. 219
- Calamine, iii. 524
- Calamite, iii. 368
- Calcareous acid, ii. 93
 spar, iii. 387
- Calcedony, iii. 297
- Calchantum, ii. 553
- Calcium, i. 347
 chloride of, i. 350
 iodide of, i. 351
 oxide of, i. 349
- Calcsinter, iii. 388
- Calctuff, *ibid.*
- Calculi, biliary, iv. 565
 urinary, iv. 568
 arranged, iv. 575
 solvents for, iv. 579
 of inferior animals, iv. 580
- Calomel, i. 485

- Caloric, i. 26
 and electricity, analogy
 between, i. 155
 escape from surfaces, i. 75
 how conducted, i. 37
 latent, i. 51
 nature of, i. 118
 of evaporation, i. 71
 of fluidity, i. 51
 radiation of, i. 75
 sources of, i. 124
 specific, i. 96
 table of, i. 101
 Calorimeter, i. 98
 Camphor, iv. 109
 artificial, iv. 115
 oil of, iv. 112
 Camphorates, ii. 155
 Camphoric acid, ii. 154
 Cancer, matter of, iv. 586
 Cannon metal, i. 473
 Cantharidin, iv. 426
 Caseic acid, iv. 499
 Caoutchouc, iv. 170
 mineral, ii. 400, and
 iii. 430
 Capacity for heat, i. 97
 Caranna, iv. 170
 Carbon, i. 225
 absorption of gases by, i.
 227
 combustion of, i. 230
 phosphuret of, i. 276
 sulphuret of, i. 293
 Carbonates, ii. 96
 Carbonic acid, i. 230, and ii. 93
 in the atmosphere,
 iii. 179
 in animals, iv. 440
 absorbed by plants,
 iv. 340
 emitted by plants,
 iv. 350
 oxide, i. 231, and ii.
 21
 Carburet of manganese, i. 420
 Carburetted hydrogen gas, i. 238
 Carica papaya, iv. 98
 Carinthine, iii. 367
 Carnelian, iii. 298
 Caromel, iv. 21
 Carrying power, what, i. 91
 Carthamus, iv. 247
 Cartilage, iv. 449
 Cascic acid, iv. 499
 Cassava, iv. 78
 Cassius, precipitate of, ii. 268
 Cast iron, i. 388
 Castor, iv. 437
 Cat's eye, iii. 293
 Catechu, ii. 213
 extractive of, iv. 59
 Caustic, i. 328
 lunar, ii. 666
 Cautery, potential, *ibid.*
 Celestine, iii. 414
 Cementation, i. 390
 Cerasin, iv. 64
 Cerate, iv. 105
 Cerin, iv. 104
 Cerite, iii. 500
 Cerium, i. 422
 salts of, ii. 584
 acetate, ii. 586
 arsenate, *ibid.*
 benzoate, ii. 587
 carbonate, ii. 585
 carburet of, i. 424
 citrate, ii. 587
 fluat, iii. 503
 molybdate, ii. 586
 muriate, ii. 260
 nitrate, ii. 584
 oxalate, ii. 587
 oxides of, i. 424
 phosphate, ii. 585
 phosphuret of, i. 424
 seleniate, ii. 587
 succinate, ii. 587
 sulphate, ii. 585
 sulphite, ii. 586
 sulphuret of, i. 425
 tartrate, ii. 587
 ores of, iii. 500
 Cerumen of the ear, iv. 513
 Ceruse, ii. 608
 Cetic acid, ii. 391
 Ceylanite, iii. 250
 Chabasite, iii. 320
 Chalcolite, iv. 499
 Chalk, iii. 385, 574
 black, iii. 346
 silvery, iii. 385
 stones, iv. 582
 Charcoal, i. 223
 action on guni, iv. 37
 proportion yielded by
 different trees, iv. 182

- Chemical combination, heat from,
i. 156
- Chemistry, definition of, i. 2
origin of, i. 3
as a science, i. 8
object, i. 11
- Chenopodium olidum, iv. 244
- Cheese, iv. 499
- Chert, iii. 298
- Chiastolite, iii. 327
- Chlorates, ii. 230
- Chloric acid, i. 186, and ii. 229
- Chloride, i. 182
of azote, i. 210
sulphur, i. 287
manganese, i. 420
zinc, i. 433
cadmium, i. 440
lead, i. 448
tin, i. 455
copper, i. 466
bismuth, i. 477
mercury, i. 483
selenium, i. 300
silver, i. 496
phosphorus, i. 268
platinum, i. 521
antimony, i. 554
- Chlorides, ii. 238
table of, i. 543
- Chlorine, i. 180
protoxide of, i. 183
deutoxide of, i. 185
compounds of, ii. 228
- Chloriodic acid, i. 193
- Chlorite, iii. 351
- Chlorocarbonic acid, i. 234
- Chlorocyanic acid, ii. 289
- Chloruret of iodine, i. 195
- Chromic acid, i. 561, and ii. 130
- Chromium, i. 558
ores of, iii. 618
- Churning, iv. 496
- Chyle, iv. 487, 597
- Chrysoberyl, iii. 258
- Chrysocolla, ii. 454
- Chrysolite, iii. 378
- Chrysophrase, iii. 298
- Chyme, iv. 590
- Cimolite, iii. 355
- Cinchona, iv. 231
- Cinnabar, i. 489
native, iii. 444
- Cinnamon, iv. 238
stone, iii. 289
- Cistic oxide, iv. 573
- Citric acid, ii. 176
plants containing, iv. 5
formed from gum, iv.
38
- Civet, iv. 438
- Clay, iii. 341
common, iii. 341
iron-stone, iii. 481, 484
potters, iii. 342
slate, iii. 342, 347, 559
stone, iii. 343
variegated, iii. 342
- Clink-stone, iii. 380
- Club-moss, iv. 273
- Clyssus, ii. 427
- Coal, ii. 401
brown, iii. 431
black, iii. 432
glance, iii. 435
cannel, iii. 433
earth, iii. 431
formation, iii. 571
Kilkenny, iii. 435
paper, iii. 432
moor, *ibid.*
slate, iii. 433
- Cobalt, i. 409
alloys of, i. 413, 436, 450,
460, 501
salts of, ii. 572
acetate, ii. 575
ammonionitrate, ii. 576
antimoniate, ii. 575
antimonite, *ibid.*
arseniate, ii. 574
borate, ii. 573
carbonate, *ibid.*
chloride, i. 413
fluante, ii. 283
lactate, ii. 576
muriate, ii. 258
nitrate, ii. 573
ores, iii. 505, 615
oxalate, ii. 575
oxides of, i. 411
phosphate, ii. 573
phosphuret, i. 413
pyrites, iii. 507
seleniuret, i. 413
seleniate, ii. 574

- Cobalt, soap, ii. 717
 sulphate, ii. 573
 sulphuret, i. 413
 tartrate of, ii. 575
 ore, black, iii. 508
 brown, *ibid.*
 glance, iii. 507
 grey, iii. 506
 red, iii. 509
 white, iii. 505
 yellow, iii. 509
- Cobalus, i. 409
- Coccolite, iii. 376
- Cocculus indicus, ii. 47
- Cochenilin, iv. 427
- Coco, iv. 270
- Coffee, iv. 268
- Cohesion, iii. 107
- Coin, gold, i. 514
 silver, i. 501
- Colcothar of vitriol, ii. 554
- Cold, i. 112
 reflection of, i. 113
- Collyrite, iii. 355
- Colophonite, iii. 283
- Colour, iii. 234
- Colouring matters of plants, iv. 47
- Columbic acid, ii. 132
- Columbite, iii. 504
- Columbium, i. 574
- Combination, iii. 133
- Combinations by heat, iii. 40
 spontaneous, iii. 39
- Combustible bases, ii. 26
- Combustibles, iii. 427
 heat given out by,
 i. 141
 acidifiable, i. 216
 alkalifiable, i. 323
 simple, i. 215
 compound, ii. 317
 analysis of, iii. 600
 without inflammation,
 i. 147
- Combustion, i. 128
 opinions relative to,
ibid.
 Berzelius' theory of,
 i. 161
 Lavoisier's theory of,
 i. 133
 explanation of, i. 135
 supporters of, i. 137,
 173
- Compounds, ii. 1.
 primary, ii. 2
 secondary, ii. 405
- Concretions, morbid, iv. 559
- Condensation, heat evolved by, i.
 147
- Conductors of caloric, i. 87
 electricity, i. 166
- Conium maculatum, iv. 244
- Contagion, iii. 182
- Cooling, i. 108
- Cooling powers of gases, i. 94
- Copaiva, iv. 152
- Copal, iv. 137
- Copper, i. 462
 acetate of, ii. 643
 alloys of, i. 469, 493, 501,
 514, 527, 532, 536,
 557, 568, 573
 salts of, ii. 630
 ammonio-oxalate, ii. 648
 ammonio-sulphate, ii. 637
 anhydrous carbonate, iii.
 463
 antimoniate, ii. 643
 antimonite, *ibid.*
 arseniate, ii. 640, and iii.
 468
 arsenite, ii. 642
 azure, iii. 464
 benzoate, ii. 645
 borate, ii. 632
 carbosilicate, iii. 467
 carbonate, ii. 632
 anhydrous, iii.
 463
 bismuthis, iii.
 466
 citrate, ii. 651
 chlorate, ii. 237
 chlorides, i. 466
 chromate, ii. 643
 emerald, iii. 467
 fluuate, ii. 284
 formate, ii. 651
 glance, iii. 456
 hyposulphite, ii. 639
 iodide, i. 467
 lactate, ii. 651
 martial arseniate, iii. 469
 mellate, ii. 650
 mica, iii. 468
 molybdate, ii. 643

- Copper, muriate, ii. 263, and iii. 469
 native, iii. 456
 nickel, iii. 511
 nitrate, ii. 630
 nitrite, ii. 632
 ores of, iii. 455
 ores, analysis of, iii. 607
 oxalate, ii. 645
 oxides, i. 463
 phosphate, ii. 633, and iii. 470
 phosphuret, i. 467
 pyrites, iii. 457
 sacclactate, ii. 651
 seleniate, ii. 639
 seleniuret, i. 468, and iii. 459
 silicate, ii. 633, and iii. 467
 soap, ii. 718
 soda oxalate, ii. 647
 suberate, ii. 645
 succinate, *ibid.*
 sulphate, ii. 634
 sulphite, ii. 636
 sulphuret, i. 468, and iii. 456
 tartrate, ii. 650
 tinned, i. 474
 tungstate, ii. 643
 white, iv. 458
- Copperas, ii. 553
 blue, ii. 634
- Coral, iv. 459
- Cordierite, iii. 266
- Corivindum, iii. 255
- Cork, ii. 157
 rock, iii. 364
- Corneous mercury, iii. 446
 silver ore, iii. 454
- Corrosive sublimate, i. 483
- Cortical layers, iv. 230
- Corundum, iii. 255
- Cotton, iv. 176
- Cotyledons, iv. 304
- Couch, iv. 367
- Cream, iv. 495
- Crocus metallorum, ii. 690
- Cross-stone, iii. 321
- Croton eleutheria, iv. 236
- Cruor of blood, iv. 492
- Crusts, composition of, iv. 455
- Cryolite, ii. 283, and iii. 423
- Crystals, iii. 113
- Crystallization, iii. 111
 expansion during, i. 43
 nature of, iii. 113
- Cube ore, iii. 486
 spar, iii. 406
- Cubizite, iii. 320
- Cucumber, iv. 275
- Cupellation, i. 447
- Cuprium, i. 471
- Curd, iv. 497
- Cutis, iv. 469
- Cuttlefish, bone of, iv. 455
- Cyanite, iii. 372
- Cyanogen, i. 235, and ii. 288
- Cymophane, iii. 258
- D
- Dalton, Mr. John, atomic theory
 of, iii. 17
- Daourite, iii. 271
- Daphnia, iv. 53
- Date-tree, pollen of, iv. 251
- Datholite, ii. 482, and iii. 409
- Dead Sea, iii. 205
- Decomposition, iii. 135
 tables of, iii. 137, 156
 of animal bodies, iv. 631
- Decrement, iii. 121
- Decrepitation, ii. 500
- Delphia, ii. 49
 salts of, ii. 478
- Delphinite, iii. 274
- Deoxidizing rays, i. 21
- Dephlogisticated air, i. 177
- Deutoxide of hydrogen, i. 220
- Diallage, iii. 373
 rock, iii. 569
- Diamond, i. 228, and iii. 247
- Dichroite, iii. 266
- Digestion, iv. 588
- Diopside, iii. 377
- Diopbase, iii. 467
- Dippel's animal oil, iv. 434
- Dipyre, iii. 322
- Disthene, iii. 372
- Distinct concretions, iii. 240
- Dolomite, iii. 394
- Draco mitigatus, i. 485
- Dragon's blood, iv. 159

Drawing slate, iii. 346
 Dropsy, liquor of, iv. 586
 Ductility, i. 325

E

Eagle-stone, iii. 481
 Earths, ii. 56
 in plants, iv. 190
 whether formed in plants,
 iv. 317
 as the food of plants, iv.
 316
 alkaline, ii. 58
 proper, ii. 73
 in animals, iv. 443
 Earthy smell, i. 374
 Egg shells, iv. 504
 white of, *ibid.*
 yolk of, iv. 505
 Eggs, iv. 504
 Eisenram, iii. 478
 Ekebergite, iii. 332
 Elaïn, ii. 384
 Elaolite, iii. 337
 Elecampane, iv. 224
 Electricity, i. 159, 162
 analogy of, with heat,
 i. 155
 distribution of, i. 166
 as a source of heat, i.
 159
 resinous, i. 164
 vitreous, *ibid.*
 Electrics, i. 165
 Electrum, i. 517, and iii. 439
 Elements, weight of, i. 319, 540
 Elemi, iv. 133
 Ellazic acid, ii. 223
 Emerald, iii. 265
 copper ore, iii. 467
 Emery, iii. 256
 Emetin, iv. 61
 Empyrean air, i. 177
 Emulsion, ii. 377
 Enamel of the teeth, iv. 451
 Epidermis of animals, iv. 468
 of trees, iv. 229
 Epidote, iii. 275
 Epsom salt, ii. 520
 Ether, ii. 340
 acetic, ii. 360
 chloric, i. 237, and ii. 355

Ether, formic, ii. 362
 hydriodic, ii. 359
 muriatic, ii. 356
 nitric, ii. 348
 sulphuric, ii. 340
 composition, ii. 344
 Ethiop's mineral, i. 488
 martial, i. 383
 per se, i. 481
 Euchlorine, i. 184
 Euclase, iii. 264
 Eudiometers, iii. 166
 Eudyalite, iii. 288
 Eukairite, iii. 450
 Euphorbium, iv. 169
 Excitement, i. 165
 Excrementitious matter, iv. 553
 Expansion, i. 26
 exceptions to it, i. 39
 Expectorated matter, iv. 516
 Extract, iv. 56
 Goulard's, ii. 617
 Extractive, iv. 56
 species of, iv. 59
 Eye, humours of, iv. 518

F

Fahl ore, iii. 458
 Fassaite, iii. 375
 Fat, iv. 431
 Feathers, iv. 483
 Feces, iv. 553
 human, *ibid.*
 of cattle, iv. 556
 of fowls, iv. 557
 Fecula, green, contains gluten, iv.
 96
 Felspar, iii. 330
 Labradore, *ibid.*
 compact, iii. 331
 Fer oligiste, iii. 476
 oxydulé, iii. 474
 Fermentation, iv. 364
 acetous, iv. 385
 vinous, iv. 365
 Ferrochyzates, ii. 312
 Ferrochyzatic acid, ii. 308
 Ferruretted chyzic acid, *ibid.*
 Fettstein, iii. 337
 Fibrin, iv. 98, 408
 Fibrolite, iii. 273
 Figure-stone, iii. 359
 Finery, i. 389

- Fiorite, iii. 300
 Fire-damp, i. 240
 Fish-scales, iv. 463
 Fixed, what, ii. 49
 air, ii. 93
 Flake, white, i. 477
 Flesh, iv. 463
 Flexibility, iii. 242
 Flint, iii. 296
 iron, iii. 294
 slate, iii. 296
 Float-stone, iii. 344
 Flowers, iv. 245
 of sulphur, i. 278
 Fluates, ii. 280
 Fluidity produced by heat, i. 46
 Fluids conduct caloric, i. 90
 Fluoborate of ammonia, ii. 286
 Fluoboric acid, i. 245, and ii. 285
 Fluor, iii. 403
 Fluoric acid, i. 197, 223, and ii. 280
 Fluorine, i. 196, and ii. 279
 Fluosilicates, ii. 287
 Fluosilicic acid, i. 252, and ii. 286
 Flux, black, ii. 429
 white, *ibid.*
 Foliated earth of tartar, ii. 442
 crystallized, ii. 464
 ore of tellurium, iii. 442
 Food of plants, iv. 311
 animals, iv. 588
 Formations, iii. 555
 alluvial, iii. 579
 floetz, iii. 569
 independent coal, iii.
 570
 new Paris, iii. 577
 primitive, iii. 556
 transition, iii. 565
 volcanic, iii. 580
 Formates, ii. 197
 Formic acid, ii. 195
 Fracture, iii. 239
 Franchipan, iv. 502
 Frangibility, iii. 242
 Frankincense, iv. 131
 Freezing, expansion by, i. 43
 contraction by, i. 45
 mixtures, i. 115
 of bodies, i. 48
 point, i. 47
 Friction, i. 151
 Frigorific particles, i. 113
 Fruits, iv. 274
 Fuci, iv. 297
 Fulminating powder, ii. 429
 Fuller's earth, iii. 358
 Functions of animals, iv. 588
 Fungates, iv. 288
 Fungic acid, *ibid.*
 Fungin, iv. 184
 Fustic, iv. 228

G

 Gadolinite, i. 369, and iii. 501
 Gahnite, iii. 280
 Galbanum, iv. 162
 Galena, i. 449, and iii. 530
 Galipot, iv. 131
 Gall-stone, iv. 565
 Gallic acid, ii. 203
 plants containing, iv. 8
 Galvanic pile, i. 168
 trough, *ibid.*
 Galvanism, i. 167
 Gamboge, iv. 167
 Garlic, iv. 279
 Garnet, iii. 284
 black, iii. 281
 common, iii. 285
 precious, *ibid.*
 white, iii. 277
 Gas, what, i. 175
 Gases, iii. 19
 absorption by water, iii. 54
 combination of, iii. 39
 ditto with liquids, iii. 53
 ditto with solids, iii. 69
 as conductors of heat, i. 94
 constitution of, iii. 20
 elasticity of, iii. 21, 25
 expansion of, i. 27.
 list of, iii. 22
 mixture of, iii. 29
 in the intestines, iv. 599
 in plants, iv. 216
 respirable, iv. 602
 specific gravity of, iii. 23,
 and iv. 648
 unrespirable, iv. 601
 water which they contain,
 iii. 26
 weight of their atoms, iii. 24
 Gastric juice, iv. 592
 Gehlenite, iii. 325
 Gelatin, iv. 390
 oxygenized, iv. 392

- Gelberz, iii. 442
 Gellert's table of decompositions, iii. 9
 Gentian, iv. 221
 Geoffroy's table of affinity, iii. 9
 Geognosy, iii. 551
 Germination, iv. 304
 Gieseckite, iii. 333
 Gin, ii. 318.
 Glacies Mariæ, iii. 349
 Glance cobalt ore, iii. 507
 nickel, iii. 511
 Glands, iv. 474
 Glass, ii. 103
 expansion of, i. 34
 phosphoric, ii. 107
 Glauberite, iii. 408
 Glauber's salt, ii. 458
 Gleadin, iv. 91
 Glucina, i. 371, and ii. 74
 salts of, ii. 531
 acetate, ii. 534
 carbonate, ii. 532
 chromate, ii. 534
 chromo-sulphate, *ibid.*
 fluante, ii. 233
 hydrosulphuret, ii. 700
 muriate, ii. 256
 nitrate, ii. 532
 oxalate, ii. 534
 phosphate, ii. 532
 seleniate, ii. 534
 succinate, ii. 534
 sulphate, ii. 533
 Glucinum, i. 370
 Glue, iv. 396
 Gluten, iv. 90
 fermented, iv. 93
 plants containing it, iv. 96
 Gneiss, iii. 558
 Gold, i. 503
 alloys, i. 508, 528, 532, 535, 539, 557, 568, 573
 coin, i. 514
 native, iii. 438
 ores of, iii. 437
 ores, analysis of, iii. 602
 oxides of, i. 505
 salts of, ii. 676
 muriate, ii. 266
 nitrate, ii. 677
 fulminating, ii. 32
 Gold, phosphuret, i. 508
 sulphate, ii. 678
 sulphuret, i. 508
 soap, ii. 718
 Gorgonia, iv. 459
 Goulard's extract, ii. 617
 Gouty concretions, iv. 582
 Grainer, iv. 471
 Grammatite, iii. 370
 Granatite, iii. 286
 Granite, iii. 557
 Grapes, iv. 275
 Graphic gold ore, iii. 441
 Graphite, i. 227, and iii. 435
 Graugiltigerz, iii. 459
 Gratiola officinalis, iv. 243
 Gravitation, iii. 5
 Graywacka, iii. 567
 slate, *ibid.*
 Greek fire, ii. 399
 Green, Brunswick, ii. 651
 Scheele's, ii. 642
 colour of plants, iv. 347
 earth, iii. 352
 lead ore, iii. 537
 porphyry, iii. 562
 resin, iv. 137
 sand of Peru, iii. 469
 stone, iii. 561
 Grey copper ore, iii. 458
 cobalt ore, iii. 506
 ore of antimony, iii. 544
 ore of manganese, iii. 493
 Grossularia, iii. 281
 Guaiacum, iv. 147
 Gum, iv. 34
 arabic, iv. 41
 cherry tree, iv. 66
 congo, *ibid.*
 gut, iv. 167
 kuteera, iv. 42
 resins, iv. 160
 Senegal, iv. 41
 tragacanth, iv. 64
 West Indian, iv. 67
 Gummihederæ, iv. 170
 Gun metal, i. 473
 Gunpowder, ii. 427
 Gurhofite, iii. 397
 Gypsum, ii. 485, and iii. 404, 563

H

- Hæmatites, brown, iii. 480
 red, iii. 479
 Hair, iv. 480
 Hairy concretions, iv. 562
 Hardness, iii. 241
 Harmotome, iii. 321
 Hart, horn of, iv. 462
 Hartshorn, ii. 27
 Hauyne, iii. 324
 Hazle-tree, pollen of, iv. 253
 Heat, i. 26
 animal, iv. 617
 conduction of, i. 87
 during combustion, i. 141
 escape of, from surfaces,
 i. 75
 latent, i. 51
 nature of, i. 118
 radiation of, i. 75
 red, i. 23
 specific, i. 96
 sources of, i. 124
 Hedenbergite, iii. 489
 Heliotrope, iii. 299
 Hellebore, iv. 218
 Helvine, iii. 283
 Hematin, iv. 48
 Hempseeds, iv. 268
 Hepar sulphuris, ii. 701
 Hepatic gas, i. 289
 concretions, iv. 562
 mercurial ore, iii. 445
 Highgateresin, iv. 140
 Ho-ang-lien, iv. 220
 Hohlspath, iii. 327
 Hollow spar, iii. 327
 Honey, iv. 425
 stone, iii. 419
 Hooke, Dr. theory of combustion
 by, i. 129
 Hops, iv. 370
 Hornblende, iii. 366
 basaltic, iii. 367
 Labradore, iii. 374
 slate, iii. 366
 Hornmangan, iii. 497
 Horns, iv. 461
 of the hart and buck, iv.
 462
 Hornsilver, i. 496, and iii. 454
 Hornstone, iii. 295
 Horsechestnut, iv. 237
 Horseradish, iv. 224
 Hospital sore, matter of, iv. 585
 Humours of the eye, iv. 518
 Hyacinth, iii. 249
 blanche cruciform, iii.
 321
 white, iii. 339
 yields mucus, iv. 43
 Hyacinthine, iii. 279
 Hyalite, iii. 301
 Hydnum repandum, iv. 293
 hybridum, iv. 294
 Hydrargillite, iii. 420
 Hydrargyrum, i. 479
 Hydrates, ii. 16, 405
 Hydriodates, ii. 276
 Hydriodic acid, i. 222, and ii. 276
 Hydrocarbonic oxide, i. 232
 Hydrochloric acid, i. 222, and ii.
 242
 Hydrocyanic acid, ii. 291, 298
 plants containing,
 iv. 7
 Hydrogen, i. 216
 arsenuretted, i. 309
 bisulphuretted, ii. 694
 boruretted, i. 247
 phosphuretted, i. 272
 combustion of, iii. 40
 deutoxide of, i. 220
 gas, i. 216
 absorbed by charcoal,
 i. 227
 sulphuretted, ii. 694
 telluretted, i. 315
 Hydroguret of carbon, i. 235
 phosphorus, i. 272
 Hydroguretted sulphur, ii. 696
 sulphurets, ii. 701
 metallic, ii. 704
 sulphuret of am-
 monia, ii. 703
 sulphuret of pot-
 ash, ii. 702.
 Hydrometer, Clarke's, ii. 331
 Hydropete, iii. 497
 Hydrophosphoric gas, i. 275
 Hydrosiderum, i. 398
 Hydrosulphurets, ii. 696
 metallic, ii. 704

- Hydrosulphuric acid, i. 289, and
ii. 694
- Hydrosulphurous acid, i. 291
- Hydrothionic acid, ii. 694
- Hyperoxymuriates, ii. 230
- Hyperstene, iii. 374
- Hyponitrous acid, i. 210
- Hypophosphorous acid, i. 265,
and ii. 112
- Hyposulphuric acid, i. 286, and
ii. 123
- Hyposulphurous acid, i. 284, and
ii. 122
- I.
- Jade, iii. 360
- Jalap, iv. 220
- Jargon, i. 373
- Jasper, iii. 305
agate, iii. 307
basalt, iii. 307
Egyptian, iii. 305
porcelain, iii. 306
- Ice, ii. 12
expansion of, i. 43
spar, iii. 339
- Ichthyophthalmite, iii. 319
- Idocrase, iii. 279
- Jelly, iv. 44
- Jet, iii. 433
- Igasuric acid, ii. 194
- Incombustibles, simple, i. 202
- Indian rubber, iv. 170
- Indianite, iii. 328
- Indigo, iv. 79
- Indigofera tinctoria, *ibid.*
- Indigofera anil, iv. 243
- Inflammable air, i. 217
heavy, i. 235
- Ink, ii. 219
printers', ii. 375
sympathetic, ii. 259
- Intestinal concretions, iv. 562
- Intestines, gases in, iv. 599
- Inulin, iv. 67
- Iodates, ii. 272
- Iodic acid, i. 192, and ii. 271
union with acids, ii. 271
- Iodide, i. 191
of arsenic, i. 309
azote, i. 213
cadmium, i. 441
iron, i. 388
- Iodide of zinc, i. 434
lead, i. 448
tin, i. 456
copper, i. 467
bismuth, i. 478
mercury, i. 487
phosphorus, i. 271
potassium, i. 333
silver, i. 498
sodium, i. 341
sulphur, i. 288
antimony, i. 554
tellurium, i. 315
- Iodides, table of, i. 544
- Iodine, i. 189
how prepared, i. 190
properties of, *ibid.*
combines with starch, i.
195
compounds of, ii. 270
- Iolite, iii. 266
- Ipecacuan, iv. 220
- Iridium, i. 536
salts of, ii. 686
ores of, iii. 440
- Iron, i. 372
affinities of, i. 403
alloys, i. 402, 409, 413, 421,
436, 450, 460, 469, 491,
500, 509, 525, 532, 556,
566, 573.
acetate, ii. 560
antimoniate, *ibid.*
arseniate, ii. 559, and iii.
486
benzoate, ii. 561
boletate, ii. 562
borate, ii. 551
boruret, i. 396
carbonate, i. 550, and iii.
483
cast, expansion of, i. 44
cast, i. 388
chlorate, ii. 236
chlorides of, i. 386
chromate, ii. 560, and iii.
488
citrate, ii. 563
clay, iii. 381
cold short, i. 397
flint, iii. 294
fluated, ii. 283
froth, iii. 478

Iron, gallate, ii. 564
 glancce, iii. 476
 hydrate, iii. 479
 hydrosulphuret, ii. 716
 hyposulphite, ii. 557
 iodide of, i. 388
 lactate, ii. 563
 malate, ii. 563
 mellate, ii. 562
 mica, iii. 477
 molybdate, ii. 560
 muriated, ii. 257
 native, iii. 472
 nitrate, ii. 549
 ores of, iii. 472
 ores, analysis of, iii. 608
 oxalate, ii. 562
 oxides of, i. 382
 phosphate, ii. 552, and iii.
 485
 phosphuret, i. 397
 pyrites, iii. 473
 saclactate, ii. 563
 salts of, ii. 548
 sand, iii. 475
 seleniate, ii. 558
 silicate, iii. 489
 silicuret, i. 394
 sinter, iii. 487
 soap, ii. 718
 subcarburet, i. 396
 suberate, ii. 562
 succinate, ii. 561
 sulpharseniate, iii. 487
 sulphate, ii. 553
 sulphite, ii. 557
 sulphurets, i. 398
 tartrate, ii. 562
 tungstate, ii. 560, and iii.
 492
 Isatis tinctoria, iv. 243
 Iserine, iii. 520
 Isinglass, iv. 397
 Juice, peculiar, of plants, iv. 211,
 358
 Jupiter, ii. 625

K

Kali, i. 328
 Kaolin, iii. 341
 Karpholite, iii. 316
 Kelp, i. 336
 Kermes mineral, ii. 707

Kidney bean, iv. 266
 Kidneys, action of, iv. 620
 calculi of, iv. 577
 Kilkenny coal, iii. 435
 Kinate of lime, ii. 493
 Kinates, ii. 181
 Kinic acid, ii. 179
 Kiffekill, iii. 356
 Kino yields a species of tannin, iv. 9
 Klebshiefer, iii. 343
 Klinkstone, iii. 380
 Knebelite, iii. 491
 Konia, i. 327
 Koumiss, iv. 501
 Krameriac acid, ii. 202
 Kupfer lazur, iii. 464
 nickel, iii. 511
 glas, iii. 456

L

Labdanum, iv. 134
 Lac, iv. 141
 sulphuris, i. 287
 Laccic acid, ii. 189
 Laccin, iv. 143
 Lackers, iv. 144
 Lactates, ii. 200
 Lactic acid, ii. 198
 Ladanum, iv. 134
 Lampic acid, ii. 160
 Lana philosophica, i. 432
 Lapis infernalis, ii. 666
 nephriticus, iii. 360
 lazuli, iii. 323
 philosophorum, i. 4
 Latent heat explained, i. 51
 tables of, i. 56, 73
 Lathyrus tuberosus, iv. 282
 Latialite, iii. 324
 Lazulite, iii. 323
 Lazurstein, *ibid.*
 Lead, i. 443
 alloys, i. 449, 461, 472, 493,
 501, 512, 526, 532, 557,
 567, 573
 salts of, ii. 603
 acetate, ii. 615
 aluminate, iii. 542
 antimonial sulphuret of, iii.
 533
 antimoniate, ii. 613
 antimonite, ii. 614

- Lead, arseniate, ii. 613, and iii. 541
 arseniophosphate, iii. 538
 arseniosulphuret, iii. 533
 benzoate, ii. 618
 black, i. 227
 borate, ii. 609
 carbonate, ii. 608, and iii. 534
 chlorate, ii. 236
 chloride, i. 448
 citrate, ii. 620
 chromate, i. 614, and iii. 538
 fluuate, ii. 284
 gallate, ii. 622
 hyposulphite, ii. 612
 iodide, i. 448
 lactate, ii. 621
 malate, ii. 621
 mellate, ii. 620
 molybdate, ii. 614, and iii. 540
 muriate, ii. 262
 murio-carbonate, iii. 536
 nitrate, ii. 603
 nitrite, ii. 605
 nitrophosphate, ii. 610
 ores, iii. 529
 ore, analysis of, iii. 610
 oxalate, ii. 620
 oxides, i. 444
 phosphate, ii. 609, and iii. 537
 phosphite, ii. 610
 red, i. 446
 refining of, i. 447
 saccharate, ii. 621
 seleniate, ii. 612
 seleniuret, i. 449
 soap, ii. 718
 suberate, ii. 619
 submuriate, ii. 262
 succinate, ii. 619
 sugar of, ii. 615
 sulphate, ii. 610, and iii. 539
 sulphite, ii. 611
 sulphuret, i. 449
 tannate, ii. 623
 tartrate, ii. 620
 tungstate, ii. 615, and iii. 542
 vinegar of, ii. 617
 white, i. 445, and ii. 608
- Leather, iv. 470
 Leaves, iv. 239
- Leaves, functions of, iv. 338
 why they fall off, iv. 341
 the digesting organs of
 plants, iv. 357
 how produced, *ibid.*
 seminal, iv. 310
 absorb carbonic acid, iv. 341
 absorb water, iv. 353
 emit oxygen, iv. 342
- Lemnian earth, iii. 357
 Lemons, essential salt of, ii. 444, 562
 Lens, crystalline, iv. 520
 Lentils, iv. 267
 Lepidolite, iii. 348
 Leucite, iii. 277
 Leucolite, iii. 261
 Libavius, smoking liquor of, i. 455
 Lichens, iv. 283
 Lievrite, iii. 490
 Ligaments, iv. 473
 Light, i. 12
 its peculiarities, i. 22
 sources of, i. 23
- Lignin, iv. 180
 Lilalite, iii. 348
 Limbourg's table of affinity, iii. 10
 Lime, i. 347, and ii. 59
 acetate of, ii. 490
 antimoniate, ii. 489
 antimonite, *ibid.*
 arseniate, ii. 488
 arsenite, ii. 489
 benzoate, ii. 490
 biphosphate, ii. 484
 boletate, ii. 157
 borate, ii. 482
 borosilicate, *ibid.*
 camphorate, ii. 491
 carbonate, ii. 480
 chlorate, ii. 234
 chromate, ii. 489
 citrate, ii. 493
 effects of, on mould, iv. 325
 fluuate of, ii. 281
 fluosilicate, ii. 288
 formate, ii. 495
 gallate, ii. 495
 hydrate, ii. 59
 hydriodate, ii. 277
 hydrosulphuret, ii. 700
 hypophosphite, ii. 485
 hyposulphate, ii. 488

- Lime, hyposulphite, ii. 487
 iodate, ii. 273
 kinate, ii. 493
 lactate, ii. 495
 malate, ii. 494
 mellate, ii. 492
 molybdate, ii. 489
 moroxylate, ii. 491
 muriate, i. 350
 nitrate, ii. 479
 oxalate, ii. 492
 oxymuriate of, ii. 239
 phosphate, ii. 483
 phosphuret, ii. 62
 quadriphosphate, ii. 484
 sacclactate, ii. 494
 salts of, ii. 479
 silicate, ii. 482
 slacking of, ii. 59
 soap, ii. 716
 subbichloride, ii. 239
 suberate, ii. 491
 subphosphate, ii. 484
 succinate, ii. 490
 sulphate, ii. 485
 —anhydrous, ii. 486
 sulphite, ii. 487
 sulphuret, ii. 63
 supermalate, ii. 494
 tannate, ii. 495
 tartrate, ii. 492
 tungstate, ii. 489
 urate, ii. 494
 water, ii. 58
- Limestone, iii. 385
 blue vesuvian, iii. 388
 magnesian, iii. 395
 primitive, iii. 562
 floetz, iii. 572
 transition, iii. 568
- Liquefaction explained, i. 51
- Liquids, iii. 75
 constitution of, iii. 76
 table of, iii. 78
 specific gravity of, iii. 79
 expansion of, i. 30
 action on each other, iii.
 80
 mixture of, *ibid.*
 combination with solids,
 iii. 85
- Liquor silicum, ii. 103
- Liquorice, iv. 224
- Litharge, i. 447
- Lithia, i. 345, and ii. 57
 borate, ii. 469
 carbonate, *ibid.*
 muriate of, ii. 253
 nitrate, ii. 469
 salts of, ii. 468
 phosphate, ii. 470
 sulphate, *ibid.*
- Lithic acid, ii. 184
- Lithium, i. 343
 chloride of, i. 346
 oxide of, i. 345
- Lithomarge, iii. 353
- Liver of sulphur, ii. 701
- Lixivia, i. 328
- Loam, iii. 341
- Loboite, iii. 280
- Lobster crust, iv. 456
- Logwood, iv. 226
- Lomonite, iii. 322
- Lucullite, iii. 389
- Luna cornea, i. 496
- Lunar caustic, ii. 666
- Lupine, white, iv. 268
- Lustre, iii. 239
- Lute, what, i. 176
- Lycopërdon, tuber, iv. 295
- Lycopodium clavatum, iv. 273
- Lydian stone, iii. 296
- Lymph of plants, iv. 328
 animals, iv. 598
- M
- Macle, iii. 327
- Madder, iv. 222
- Madrapores, iv. 457
- Madriporite, iii. 389
- Magistery of bismuth, i. 477
- Magnesia, i. 364, and ii. 71
 salts of, ii. 515
 acetate of, ii. 525
 ammonio-nitrate, ii.
 516
 ammonio-phosphate, ii.
 520
 ammonio-sulphate, ii.
 522
 ammonio-sulphite, ii.
 524
 arseniate, ii. 525
 benzoate of, ii. 526

- Magnesia, borate, ii. 518, and iii. 417
 camphorate, ii. 526
 carbonate, ii. 517, and iii. 416
 chlorate, ii. 235
 chromate, ii. 525
 citrate, ii. 527
 fluante, ii. 282
 gallate, ii. 528
 hydriodate, ii. 278
 hydrosulphuret, ii. 700
 hyposulphite, ii. 524
 lactate, ii. 528
 malate, *ibid.*
 molybdate, ii. 525
 muriate, i. 365
 native, iii. 356
 nitrate, ii. 516
 oxalate, ii. 526
 phosphate, ii. 519
 saccharate, ii. 528
 seleniate, ii. 524
 silicate, ii. 519
 soap of, ii. 716
 soda-sulphate, ii. 522
 suberate, ii. 526
 succinate, *ibid.*
 sulphate, ii. 520, and iii. 416
 sulphite, ii. 523
 tartrate, ii. 527
 tungstate, ii. 525
 urate, ii. 528
 Magnesian lime-stone, iii. 395
 Magnesite, iii. 416
 Magnesium, i. 364
 chloride of, i. 366
 iodide of, i. 367
 oxide of, i. 365
 Magnetic iron-stone, iii. 474
 sand, iii. 475
 Magnets, i. 400
 Maize, iv. 263
 Malachite, iii. 465
 Malaccolite, iii. 376
 Malates, ii. 193
 Malic acid, ii. 190
 plants containing it, iv. 6
 in animals, iv. 442
 Malleability, i. 325
 Malt, iv. 366
 Maltha, ii. 397
 Manganblende, iii. 495
 Manganese, i. 415
 alloys, i. 421, 510, 567
 salts of, ii. 577
 acetate, ii. 582
 antimoniate, ii. 581
 arseniate, *ibid.*
 benzoate, ii. 582
 carbo-silicate, iii. 497
 chromate, ii. 581
 citrate, ii. 583
 carbonate, ii. 578
 carburet, i. 420
 chloride, *ibid.*
 chromate, ii. 581
 fluante, ii. 284
 hydrosulphuret, ii. 709
 hyposulphite, ii. 580
 hyposulphate, ii. 581
 muriate, ii. 259
 nitrate, ii. 577
 ores, iii. 493, 616
 oxalate, ii. 583
 oxides of, i. 416
 phosphate, ii. 579, and iii. 496
 phosphuret, i. 421
 seleniate, ii. 581
 silicate, ii. 578 and iii. 496
 soap, ii. 718
 succinate, ii. 582
 sulphate, ii. 579
 sulphuretted oxide, i. 421
 tartrate, ii. 583
 tungstate, ii. 581
 black ore of, iii. 494
 grey ore of, iii. 493
 Manganetic acid, i. 419
 Manganspath, iii. 496
 Manna, iv. 28
 metallorum, i. 485
 Manure, iv. 322
 Marble, iii. 386
 Portsoy, iii. 361
 Marcasite of gold, i. 430
 Margarates, ii. 388
 Margaric acid, ii. 387
 Marine acid, ii. 242

- Marl, iii. 398
 Marrow, iv. 477
 Mars, ii. 548
 Stahl's tincture of, ii. 551
 tartarised tincture of, ii. 565
 tincture of, *ibid.*
 Marsh bean, iv. 266
 Martial æthiops, i. 383
 salts, ii. 548
 Massicot, i. 445
 Mastich, iv. 132
 Mayow, Dr. Theory of Combustion by, i. 129
 Meadow ore, iii. 482
 Meconic acid, ii. 158
 Medicine, universal, i. 7
 Medullin, iv. 180
 Meershaum, iii. 356
 Meionite, iii. 338
 Melanite, iii. 281
 Melasses acid, iv. 16
 Mellate of alumina, iii. 419
 Mellates, ii. 171
 Mellite, iii. 419
 Mellitic acid, ii. 170
 Melt of the carp, iv. 526
 Melting explained, i. 51
 Melting points, i. 47
 Membrane, iv. 473
 Menachanite, iii. 519
 Menachine, i. 577
 Menianthis trifoliata, iv. 244
 Menilite, iii. 304
 Mephitic acid, ii. 93
 Mercurius dulcis, i. 485
 Mercury, i. 479
 action on metals, iii. 106
 amalgams of, i. 491, 503, 516, 527
 salts of, ii. 656
 acetate, ii. 662
 arseniate, ii. 661
 benzoate, ii. 662
 boiling point, i. 480
 borate, ii. 659
 carbonate, ii. 658
 carbosulphuret, iii. 445
 chlorate, ii. 237
 chloride, i. 483
 chromate, ii. 661
 citrate, ii. 663
 congelation of, i. 480
 cornceous, iii. 446
 Mercury, fluate, ii. 284
 fulminating. ii. 34, 354
 hepatic ore of, iii. 445
 Howard's fulminating, ii. 354
 hydrosulphuret of, ii. 706
 iodides, i. 487
 lactate, ii. 664
 malate of, ii. 664
 mellate, ii. 663
 muriate, i. 483
 corrosive muriate of, i. 483
 molybdate, ii. 661
 native, iii. 443
 nitrate, ii. 657
 ores of, iii. 442
 ores, analysis of, iii. 606
 oxalate, ii. 663
 oxides of, i. 481
 phosphate, ii. 659
 phosphite, *ibid.*
 phosphuret, i. 488
 sac lactate, ii. 664
 seleniate, ii. 661
 soap of, ii. 717
 succinate, ii. 663
 sulphate, ii. 659
 sulphurets, i. 488
 tartrate, ii. 663
 tungstate, ii. 661
 Merulius cantharellus, iv. 294
 Mesembryanthemum crystallinum, iv. 245
 Mesolyte, iii. 316
 Mesotype, iii. 314
 Metallic hydrogurets, ii. 704
 lustre, i. 323
 Metals, i. 323
 as conductors of electricity, i. 160
 tenacity of, i. 325
 in plants, iv. 198
 how obtained pure, iii. 619
 conducting power of, i. 88
 in animals, iv. 443
 Meteoric stones, iii. 184
 Miasmata, putrid, destroyed by muriatic acid, ii. 270
 Mica, iii. 349
 green, iii. 499
 iron, iii. 477

- Mica slate, iii. 558
 Micaceous iron ore, iii. 477
 Micarell, iii. 350
 Miemite, iii. 397
 Milk, iv. 494
 acid of, ii. 198
 woman's, iv. 502
 ass's, iv. 503
 goat's, *ibid.*
 ewe's, iv. 504
 mare's, *ibid.*
 mountain, iii. 385
 rock, *ibid.*
 Millepores, iv. 458
 Milt of the carp, iv. 526
 Mineral caoutchouc, ii. 400
 crystal, ii. 426
 pitch, iii. 430
 tallow, ii. 397
 tar, ii. 399
 Minerals, iii. 232
 characters of, iii. 233
 compound, iii. 551
 simple, iii. 243
 analysis of, iii. 590
 Minium, i. 446
 native, iii. 534
 Mirrors of telescopes, i. 474
 Mispickel, i. 402
 Misy, ii. 553
 Mixture, as a source of heat, i. 156
 Molybdate of lead, iii. 540
 Molybdates, ii. 131
 Molybdena, iii. 513
 Molybdenum, i. 562
 ores of, iii. 513, 617
 oxides of, i. 563
 Molybdic acid, i. 564, and ii. 130
 Molybdous acid, i. 564
 Moonstone, iii. 330
 Moorcoal, iii. 432
 Morass ore, iii. 482
 Moroxylates, ii. 153
 Moroxylic acid, ii. 152
 Morphia, ii. 37
 muriate of, ii. 253
 salts of, ii. 472
 Mortar, ii. 64
 water, ii. 65
 Mosaic gold, i. 458
 Mother ley, i. 364
 Mother of pearl, iv. 454
 Mould, vegetable, iv. 323
 acid, iv. 325
 Mountain blue, iii. 464
 green, iii. 465
 iron-shot, iii. 466
 Mucic acid, ii. 181
 Mucilage, iv. 34
 Mucous acid, ii. 181
 Mucus, iv. 42, 415
 of the nose, iv. 515
 Muriacite, iii. 406
 Muriatic acid, i. 221, and ii. 242
 how prepared, ii. 243
 destroys putrid miasmata, ii. 270
 dephlogisticated, i. 180
 in animals, iv. 439
 Muriates, ii. 248
 Muscles of animals, iv. 463
 Muscovy glass, iii. 349
 Mushrooms, iv. 287
 Musk, iv. 438
 Mussite, iii. 377
 Must, iv. 379
 fermentation of, iv. 380
 Mustard, iv. 272
 Myricin, iv. 104
 Myrrh, iv. 168

N

 Nacrite, iii. 363
 Nadelstein, iii. 517
 Nails, iv. 462
 Nanceic acid, ii. 202
 Naphtha, ii. 394
 Narcotic salt, i. 241
 Natrolite, iii. 315
 Natron, i. 336
 Nature, chemical examination of, iii. 159
 Needle-ore, iii. 528
 Needle-stone, iii. 528
 Nepheline, iii. 339
 Nephryn, iv. 419
 Nephrite, iii. 360
 Nerium tinctorium yields indig o, iv. 80
 Nerves, iv. 474
 Neutralization, iii. 130
 Nickel, i. 403

- Nickel, alloys, i. 409, 436, 450, 470, 510, 566
 salts of, ii. 565
 acetate, ii. 571
 ammonio-nitrate, ii. 567
 ammonio-sulphate, ii. 569
 arseniate, ii. 570, and iii. 513
 arseniuretted, iii. 511
 borate, ii. 568
 carbonate, ii. 567
 carburet, i. 408
 chloride, *ibid.*
 chromate, ii. 570
 molybdate, ii. 571
 muriate, ii. 258
 native, iii. 510
 nitrate, ii. 566
 ochre, iii. 512
 ores, iii. 510
 ores, analysis of, iii. 612
 oxalate, ii. 571
 oxides of, i. 406
 phosphate, ii. 568
 phosphuret, i. 408
 seleniate, ii. 570
 sulphate, ii. 568
 sulphuret, i. 409
 tartrate, ii. 571
- Nicotin, iv. 55
- Nightshade, deadly, iv. 242
- Nigrine, iii. 520
- Nihil album, i. 432
- Nitratès, ii. 87
- Nitre, ii. 424
 cubic, ii. 450
 fixed, ii. 430
 fixed by charcoal, ii. 427
- Nitric acid, i. 207, and ii. 82
 phlogisticated, ii. 87
- Nitricum, i. 214
- Nitrogen, i. 202
- Nitrous acid, i. 205, 208, and ii. 90
 vapour, ii. 92
 etherised gas, ii. 353
 gas, i. 206, and ii. 7
 gas, dephlogisticated, ii. 3
 oxide, i. 207
 gas, eudiometer, iii. 168
- Nitroxi, ii. 7
- Nitrum, i. 336
 fixum, ii. 435
- Nitrum, flammans, ii. 411
 semivolatile, *ibid.*
- Novaculite, iii. 347
- Nutgalls, ii. 203
- Nutmeg, iv. 271
- Nuts, husk of, iv. 275
- Nux vomica, ii. 42
- O
- Oats, iv. 259
- Obsidian, iii. 308
- Ochre, iii. 478
- Ochroite, i. 422
- Octahedrite, iii. 518
- Oetites, iii. 481
- Oil of vitriol, ii. 111
- Oils, ii. 363
 animal, iv. 431
 bituminous, ii. 393
 fixed, ii. 371
 list of, iv. 100
 volatile, ii. 363
 list of, iv. 101
 drying, ii. 374
 essential, ii. 363
 fat, ii. 376
 philosophical, ii. 374
 poisonous, ii. 392
 sweet, principle of, ii. 379
- Oleates, ii. 390
- Olefiant gas, i. 235
- Oleïc acid, ii. 390
- Oleo-saccharum, ii. 369
- Oleum sulphuris, ii. 111
 vitrioli dulce, ii. 341
- Olibanum, iv. 164
- Oliven ore, iii. 469
- Olivile, iv. 121
- Olivine, iii. 379
- Onion, iv. 281
- Opacity, i. 324
- Opal, iii. 301
 jasper, iii. 307
- Opium, ii. 37
- Opobalsamum, iv. 152
- Opoponax, iv. 166
- Ores, iii. 436
 analysis of, iii. 602
- Orichalcum, i. 471
- Orpiment, i. 312, and iii. 549
- Orthite, iii. 503
- Osmazome, iv. 416
- Osmium, i. 317

- Ossifications, iv. 560
 Ouretic acid, ii. 456
 Ourmia, lake, iii. 206
 Oxalates, ii. 168
 Oxalic acid, ii. 165
 plants containing, iv. 3
 in animals, iv. 442
 Oxide of antimony, i. 551
 azote, i. 204, and ii. 3
 bismuth, i. 476
 cerium, i. 424
 chromium, i. 560
 cobalt, i. 411
 copper, i. 463
 gold, i. 505
 hydrogen, i. 218
 iridium, i. 538
 iron, i. 382
 lead, i. 444
 manganese, i. 416
 mercury, i. 481
 molybdenum, i. 563
 nickel, i. 406
 osmium, i. 318
 palladium, i. 531
 phosphorus, i. 267
 platinum, i. 520
 rhodium, i. 534
 selenium, i. 299
 silver, i. 495
 tantalum, i. 575
 tellurium, i. 314
 tin, i. 452
 titanium, i. 579
 tungsten, i. 570
 uranium, i. 427
 zinc, i. 432
 Oxides, ii. 3
 metallic, i. 541, and ii. 79
 table of, i. 541
 division of, ii. 3
 unsalifiable, *ibid.*
 Oxycarburetted hydrogen, i. 240
 Oxygen, i. 174
 gas, how prepared, *ibid.*
 combination with nitrous
 gas, ii. 9
 combination with simple
 combustibles, ii. 3
 emitted by plants, iv. 343
 absorbed by plants, iv.
 349
 Oxymuriatic acid, i. 180
- Oxyprussic acid, ii. 289
- P
- Pakfong, i. 436
 Palladium, i. 529
 alloys, i. 532
 salts of, ii. 682
 ferrocyanate, ii. 684
 muriate of, ii. 683
 nitrate, ii. 682
 oxides of, i. 531
 seleniuret, i. 531
 sulphate, ii. 683
 sulphuret, i. 531
 Panacea duplicata, ii. 435
 holsatica, *ibid.*
 Panchymagogum minerale, i. 485
 Panchymagogus quercitanus, *ibid.*
 Pancreatic juice, iv. 508
 concretions, iv. 561
 Papaw tree contains fibrin, iv. 98
 Paranthine, iii. 335
 Parenchyma, iv. 230
 Patellæ, iv. 461
 Paulite, iii. 374
 Pea stone, iii. 388
 Pearl, iv. 455
 ash, i. 326
 white, ii. 653
 stone, iii. 309
 sinter, iii. 300
 Peas, iv. 263
 Pechblende, iii. 498
 Peliom, iii. 268
 Pepper, iv. 271
 Perchloric acid, i. 187
 Perchloride of phosphorus, i. 270
 Percussion, i. 147
 Pericardium, liquor of, iv. 518
 Peridot, iii. 378
 Periodide of phosphorus, i. 271
 Perlated acid, ii. 456
 Perspiration, iv. 536, 621
 Peru, balsam of, iv. 154
 Peruvian bark, iv. 233
 Petalite, i. 343, and iii. 328
 Petroleum, ii. 393
 Petro-silex, iii. 331
 Pewter, i. 462
 Phallus impudicus, iv. 294
 Pharmacolite, iii. 411
 Phenix dactylefera, iv. 251
 Phlogisticated air, i. 202

- Phlogisticated alkali, ii. 294
 Phlogiston, i. 217
 Phosgene gas, i. 233
 Phosphate of iron and manganese,
 iii. 496
 Phosphatic acid, ii. 110
 Phosphorana, i. 268
 Phosphorane, *ibid.*
 Phosphoric acid, i. 259, and ii. 109
 plants contain-
 ing, iv. 186
 in animals, iv. 440
 Phosphorite, iii. 401
 Phosphorous acid, i. 262, and ii.
 111.
 Phosphorus, i. 254
 chlorides of, i. 268
 hydrogurets of, i. 272
 iodides of, i. 270
 liquid, ii. 368
 oxide of, i. 267
 of Homberg, i. 351
 of Balduin, ii. 480
 in animals, iv. 439
 Phosphuret of antimony, i. 573
 arsenic, i. 311
 barytes, ii. 66
 carbon, i. 276
 cerium, i. 424
 cobalt, i. 413
 copper, i. 467
 gold, i. 508
 iron, i. 397
 lead, i. 448
 lime, ii. 60
 manganese, i. 421
 mercury, i. 488
 molybdenum, i. 566
 nickel, i. 408
 platinum, i. 522
 selenium, i. 303
 silver, i. 499
 strontian, ii. 69
 sulphur, i. 295
 tin, i. 457
 titanium, i. 579
 tungsten, i. 572
 zinc, i. 434
 Phosphuretted alcohol, ii. 334
 azotic gas, i. 272
 hydrogen gas, *ibid.*
 Photocite, iii. 497
 Phytolacca berries, iv. 273
 Picnite, iii. 261
 Picromel, iv. 417
 Picrotonia, muriate of, ii. 254
 salts of, ii. 476
 Picrotoxia, ii. 47
 Pierre à fusil, iii. 296
 de croix, iii. 286
 Pig iron, i. 388
 Pimelite, iii. 353
 Pinchbeck, i. 471
 Pineal concretions, iv. 560
 Pinite, iii. 350
 Pisiform iron ore, iii. 481
 Pitch, mineral, iii. 430
 ore, iii. 498
 stone, iii. 309
 Pistasite, iii. 274
 Pitchy iron ore, iii. 487, 496
 Pitcoal, ii. 401
 caking, ii. 402
 cannel, ii. 403
 cherry, ii. 403
 Kilkenny, ii. 402
 splint, ii. 403
 gas from, i. 240
 Pittizite, iii. 487
 Piztazite, iii. 274
 Plants, death of, iv. 362
 decay of, iv. 361
 diseases of, iv. 300
 food of, iv. 311
 foreign substances in, iv.
 186
 ingredients of, iv. 198
 parts of, iv. 205
 peculiar juices of, iv. 211
 sap of, iv. 206
 Plasma, iii. 298
 Plaster of Paris, ii. 486
 Plasters, ii. 718
 Platina, iii. 440
 Platinum, i. 517
 alloys of, i. 525, 532,
 558, 569
 chlorides, i. 521
 fulminating, ii. 33
 muriate, ii. 269
 ores of, iii. 439
 analysis of, iii. 603
 oxides of, i. 520
 phosphurets, i. 522
 salts of, ii. 678
 sulphate, *ibid.*

- Platinum, sulphurets, i. 523
triple salts of, ii. 679
- Pleonaste, iii. 250
- Plombgumme, iii. 542
- Plumbago, i. 227, and iii. 435
- Plumbum corneum, i. 428
- Plumula, iv. 304
- Poisons, animal, iv. 531
- Polarization, i. 15
- Pollen, iv. 251
- Pollenin, iv. 97
- Polychroite, iv. 47
- Polyhalite, iii. 408
- Pompholyx, i. 432
- Ponderable bodies, i. 173
- Ponderous spar, i. 352, and iii. 413
- Porcelain earth, iii. 341
- Porphyry, iii. 560
- Potash, i. 326, 328, and ii. 51
dissolves oxides, ii. 51
acetate of, ii. 442
ammonio-sulphate, ii. 436
antimoniate, ii. 439
antimonite, ii. 440
arseniate, ii. 438
arsenite, ii. 439
benzoate, ii. 443
binarseniate, ii. 426
boletate, ii. 142
borate, ii. 432
camphorate, ii. 443
carbonate, ii. 430
chlorate, ii. 230
chromate, ii. 440
citrate, ii. 448
columbate, ii. 442
ferrosulphate, ii. 564
ferrotartrate, ii. 565
fluuate, ii. 281
fluosilicate, ii. 287
hydrate of, ii. 53
hydriodate, ii. 276
hydrosulphuret, ii. 698
hypophosphite, ii. 434
hyposulphite, ii. 437
iodate, ii. 272
lactate, ii. 449
malate, *ibid.*
mellate, ii. 446
molybdate, ii. 441
muriate, i. 331
nitrate, ii. 424, and iii. 424
nitrite, ii. 430
- Potash, oxalate, ii. 444
phosphate, ii. 433
quadroxalate, ii. 445
sac lactate, ii. 449
salts of, ii. 423
seleniate, ii. 438
silicate, ii. 433
soap of, ii. 718
succinate, ii. 561
suberate, ii. 444
succinate, ii. 443
sulphate, ii. 435
sulphite, ii. 436
sulphuret, ii. 54
tartrate, ii. 446
tungstate, ii. 441
nitro, *ibid.*
urate, ii. 449
- Potassium, i. 326
alloys of, i. 342, 402,
436, 450, 460, 469,
491, 509
peroxide of, i. 330
chloride of, i. 331
iodide of, i. 333
hydruret of, *ibid.*
phosphuret of, i. 334
seleniuret, i. 335
sulphuret of, i. 334
arseniuret, i. 335
action of, on ammonia,
ii. 31
- Potatoe, iv. 276
- Potential cautery, i. 318
- Potstone, iii. 362
- Powder of algaroth, i. 552
- Prase, iii. 293
- Precipitants, table of, iii. 141
- Precipitate *per se*, i. 482
- Precipitation, iii. 150
- Prehite, foliated, iii. 312
fibrous, iii. 313
- Prehnite, iii. 312
- Prince's metal, i. 471
- Principle, vegetative, iv. 362
animal, iv. 631
- Products of combustion, i. 138
- Propolis, iv. 436
- Prostate concretions, iv. 562
- Protiodide of phosphorus, i. 271
- Protochloride of phosphorus, i.
268
- Prussian alkali, ii. 294
2 x 2

Prussian blue, ii. 292
 native, iii. 485
 Prussiates, ii. 312
 Prussic acid, ii. 291, 298
 plants containing it,
 iv. 7
 Prussous acid, ii. 302
 Puddling, i. 390
 Pulmonary concretions, iv. 561
 Pumice, iii. 310
 Purple of amorgos, iv. 285
 Purpurates, ii. 189
 Purpuric acid, ii. 187
 Pus, iv. 583
 how distinguished from mu-
 cus, iv. 584
 Putrefaction, iv. 387, 632
 in air, iv. 632
 under ground, iv.
 634
 of accumulated mat-
 ters, *ibid.*
 preventatives from,
 iv. 636
 Pycnite, iii, 261
 Pyreneite, iii. 278
 Pyrites, i. 398, and iii. 473
 magnetic, i. 398, and iii.
 473
 arsenical, iii, 548
 cobalt, iii. 507
 copper, iii. 457
 hepatic, iii. 474
 Pyro-acetic spirit, ii. 142
 Pyromalite, iii. 491
 Pyrolignous acid, ii. 163
 Pyromachus, iii. 296
 Pyrometer of Wedgewood, ii. 75
 Pyromucic acid, ii. 164
 Pyrop, iii. 287
 Pyrophorus, Homberg's, ii. 541
 Canton's, i. 18
 Pyrophysalite, iii. 261
 Pyrotartaric acid, ii. 162
 Pyrotartarous acid, ii. 163
 Pyroxen, iii. 376

Q

Quartz, iii. 290
 milk, iii. 292
 Quassin, iv. 50
 Quercitron, iv. 237
 Quercus nigra, *ibid.*

Quicklime, i. 347
 Quicksilver, i. 479
 Quinquina, iv. 231

R.

Radiant heat, i. 75
 Radiating power of bodies, *ibid.*
 Radical vinegar, ii. 138
 Radicle, iv. 304
 Rælizite, iii. 372
 Rancidity, ii. 379
 Rarification, effect of, on the com-
 bustion of bodies, i. 145
 Rays of light, i. 13
 caloric, i. 75
 Realgar, i. 311, and iii. 549
 Red antimonial ore, iii. 546
 cobalt ochre, iii. 509
 copper ore, iii. 462
 flowers, iv. 246
 hematite, iii. 478
 iron ore, *ibid.*
 lead, i. 446
 lead ore of Siberia, iii. 538
 ochre, iii. 478
 ore of antimony, iii. 546
 ore of manganese, iii. 496
 precipitate, i. 482
 sanders, iv. 228
 silver ore, iii. 451
 zinc ore, iii. 524
 Reflection of heat, i. 121
 light, i. 14
 Refraction of light, *ibid.*
 Regulus of antimony, i. 549
 martial, i. 556
 Venus, i. 557
 Reissite, iii. 425
 Repulsion, iii. 21
 Resin, iv. 135, and iii. 428
 Botany Bay, iv. 134
 from bitumen, iv. 146
 green, iv. 137
 of bile, iv. 417
 Highgate, iii. 424, and iv.
 140
 Resinous electricity, i. 163, 164
 Resins, iv. 122
 animal, iv. 435
 Respiration, iv. 601
 of fishes, iv. 611
 Priestley's theory of,
 iv. 615

- Respiration, Lavoisier's theory of, iv. 616
 Lagrange's theory of, *ibid.*
 changes on air by, iv. 606
 changes on blood by, iv. 614
- Respirations, number of, iv. 603
- Rete mucosum, iv. 472
- Reticularia portensis, iv. 296
- Retinasphaltum, ii. 401, and iii. 429
- Rhodium, i. 533
 salts of, ii. 684
- Rhodonite, iii. 497
- Rhodorite, iii. 496
- Rhomb spar, iii. 394
- Rhubarb, iv. 221, 243
- Ricc, iv. 262
- Richter's law of double decomposition, iii. 15
- Rock cork, iii. 364
 crystal, ii. 292
 soap, iii. 354
 wood, iii. 364
- Rocks, iii. 551
 position of, iii. 554
 structure of, iii. 552
- Roestone, iii. 386
- Roots, iv. 217
 absorb oxygen, iv. 353
- Rosa mallos, iv. 156
- Rosacic acid, iv. 441
- Rosin, iv. 131
- Rouge, iv. 247
- Rubellite, iii. 271
- Ruby, iii. 252
 balass, *ibid.*
 family, iii. 250
 occidental, iii. 2
 octahedral, iii. 252
 oriental, iii. 251
 spinell, iii. 252
- Rum, ii. 318
- Rust, ii. 551
- Rutile, iii. 517
- Rutilite, iii. 521
- Rye, iv. 257
- S.
- Sacar-nambu, iv. 30
- Saccharine acid, ii. 165
- Sacclactates, ii. 183
- Sacclactic acid, ii. 181
- Saffron, iv. 250
 of Mars, i. 384
- Sagapenum, iv. 165
- Sagenite, iii. 517
- Sago, iv. 77
- Sahlite, iii. 376
- Sal ammoniac, ii. 248
 secret, ii. 415
 catharticus amarus, ii. 520
 de duobos, ii. 438
 gem, ii. 249
 mirabile, ii. 458
 polychrest Glaseri, ii. 435
 perlatum, ii. 455
- Salifiable bases, ii. 26
- Saline solutions, action of salt on, iii. 101
 table of, iii. 94
 freezing of, i. 48
- Saliva, iv. 306
 of the horse, iv. 508
- Salivary concretions, iv. 560
- Salop, iv. 77
- Salsola, i. 336
- Salt, i. 339, and iii. 426
 arsenical, neutral, ii. 438
 its action on saline solutions, iii. 101
 digestive, of Sylvius, ii. 442
 diuretic, *ibid.*
 Epsom, ii. 520
 febrifuge, i. 331
 essential of lemons, ii. 444
 essential of wine, ii. 442
 Glauber's, ii. 458
 microcosmic, ii. 457
 narcotic, i. 241
 of amber, ii. 150
 of Saturn, ii. 615
 of Seignette, ii. 466
 of Sylvius, i. 331
 of tartar, ii. 436
 of wood-sorrel, ii. 444
 perlated, ii. 455
 Rochelle, ii. 466
 rock, i. 339
 sea, *ibid.*
 sedative, i. 241
 Seydler, ii. 520
 petre, ii. 424

- Salt, rock, iii. 426
 sulphureous, of Stahl, ii. 437
 Salts, ii. 407, and iii. 424
 calcareous, iii. 383
 compound, ii. 408
 neutral, *ibid.*
 nomenclature, ii. 409
 simple, ii. 408
 solubility in alcohol, ii. 335
 solubility in water, iii. 92
 triple, ii. 409
 food of plants, iv. 319
 Sandarach, iv. 133
 Sandstone, iii. 570
 Sap of plants, iv. 206
 motion of, iv. 327
 transpiration of, iv. 338
 Saponaceous matter of putrefac-
 tion, iv. 634
 Saponules, ii. 368
 Sappare, iii. 372
 Sapphyr, iii. 253
 oriental, *ibid.*
 occidental, iii. 260
 Sarcocoll, iv. 31
 Saturation, iii. 89
 Saturn, i. 443
 salt of, ii. 615
 sugar of, *ibid.*
 extract of, ii. 617
 Saussurite, iii. 327
 Savonule, ii. 368
 Scales of fish, iv. 463
 serpents, *ibid.*
 Scammony, iv. 166
 Scapolite, iii. 335
 Schalstone, iii. 410
 Shaumearth, iii. 385
 Scheelium, i. 570
 Schiefer spar, iii. 384
 Schillerstone, iii. 373
 Schillitin, iv. 52
 Schmelzstein, iii. 322
 Schorl, iii. 269
 common, iii. 270
 red, iii. 271
 white, iii. 339
 Schorlous beryl, iii. 261
 Schorlite, iii. 261
 Schriftez, iii. 441
 Schwarzgiltigerz, iii. 459
 Scorodite, iii. 487
 Scorpion, venom of, iv. 533
 Sea-froth, iii. 356
 salt, ii. 249
 salt, regenerated, *ibid.*
 wax, ii. 397
 Sebacic acid, ii. 149
 Secretions, morbid, iv. 583
 Sedative salt, i. 241
 Seeds of plants, iv. 255
 Selenite, ii. 485, and iii. 405
 Seleniates, ii. 125
 Selenic acid, i. 299, and ii. 124
 Selenium, i. 297
 chloride of, i. 300
 oxide of, i. 299
 phosphuret of, i. 303
 sulphuret of, *ibid.*
 Seleniuretted hydrogen gas, i. 301
 Semen, iv. 534
 Semiopal, iii. 302
 Senna, iv. 239
 Serosity, iv. 490
 Serpentine, iii. 361, 563
 Serum of blood, iv. 490
 Seydler's salt, ii. 520
 Shale, iii. 346
 Shalstein, ii. 482
 Shells, iv. 453
 porcellaneous, iv. 453
 mother-of-pearl, iv. 454
 of eggs, iv. 456
 Shistus, argillaceous, iii. 347
 Siberite, iii. 271
 Siderocalcite, iii. 397
 Siderum, i. 398
 Sienite, iii. 565
 Silberkupperglanz, iii. 450
 Silicious stalactite, iii. 300
 Silica, i. 248, and ii. 102
 composition of, i. 250
 Silicated potash precipitates gum,
 iv. 37
 Silicon, i. 248
 Silk, iv. 483
 Silver, i. 494
 alloys of, i. 500, 516, 528,
 532, 535, 539, 557, 568,
 573
 salts of, ii. 664
 acetate, ii. 674
 amalgam of, iii. 443

- Silver, antimonial, iii. 448
 antimonial sulphuret, iii. 448
 arseniate, ii. 672
 arsenical silver ore, iii. 448
 arsenite, ii. 673
 auriferous, iii. 447
 benzoate, ii. 674
 bismuthic, iii. 448
 borate, ii. 668
 carbonate, ii. 668, and iii. 455
 chlorate, ii. 237
 chloride, i. 496
 chromate, ii. 673
 citrate, ii. 675
 coins, purity of, i. 502
 fluuate, ii. 285
 fulminating, ii. 33
 glance, iii. 449
 horn, i. 496, and iii. 454
 hyposulphite, ii. 670
 iodate, ii. 274
 iodide, i. 498
 lactate, ii. 675
 malate, *ibid.*
 mellate, *ibid.*
 molybdate, ii. 674
 muriate, i. 496
 native, iii. 447
 nitrate, ii. 665
 nitrite, ii. 667
 ores of, iii. 447
 ores, analysis of, iii. 603
 oxalate, ii. 674
 oxides, i. 495
 phosphate, ii. 669
 phosphuret, i. 499
 red ore, iii. 451
 sacclactate of, ii. 675
 seleniate, ii. 672
 seleniuret, i. 500
 soap, ii. 718
 succinate, ii. 674
 sulphate, ii. 669
 sulphite, ii. 670
 sulphuret, i. 499, and iii. 449
 tartrate, ii. 674
 white ore, iii. 450
- Silvery chalk, iii. 385
 Simple combustibles, i. 215
 Simple incombustibles, i. 202
 substances, i. 11
 supporters, i. 173
 Sinovia, iv. 532
 Size, iv. 397
 Skin, iv. 469
 whether it absorbs, iv. 623
 Skolezite, iii. 316
 Slate, adhesive, iii. 343
 drawing, iii. 346
 polishing, iii. 343
 whet, iii. 347
 clay, iii. 342
 spar, iii. 384
 Smaragdite, iii. 373
 Soap, ii. 711
 hard, ii. 712
 soft, ii. 714
 Starkey's, ii. 368
 Soaps, alkaline, ii. 711
 earthy, ii. 716
 metallic, ii. 717
 Soda, i. 336
 composition of, i. 338, and ii. 56
 salts of, ii. 450
 acetate of, ii. 464
 ammonio-phosphate, ii. 457
 arseniate, ii. 462
 arsenite, ii. 463
 benzoate, ii. 464
 bisulphate, ii. 460
 borate, ii. 453
 camphorate, ii. 465
 carbonate, ii. 451, and iii. 424
 citrate, ii. 467
 chlorate, ii. 233
 chromate, ii. 463
 fluuate of, ii. 281
 fluosilicate, ii. 288
 hydriodate, ii. 277
 hydrosulphuret, ii. 699
 hypophosphite, ii. 458
 hyposulphite, ii. 451
 iodate, ii. 273
 lactate, ii. 468
 malate, *ibid.*
 mellate, ii. 466
 molybdate, ii. 463
 muriate, i. 339, and ii. 249, and iii. 426

- Soda, nitrate, ii. 45
 oxalate, ii. 465
 phosphate, ii. 455
 phosphite, ii. 458
 sac lactate, ii. 467
 seleniate, ii. 462
 silicate, ii. 455
 soap of, ii. 712
 sub-borate, ii. 454
 suberate, ii. 465
 succinate, *ibid.*
 sulphate, ii. 458, and iii. 425
 sulphite, ii. 461
 tartrate, ii. 466
 tungstate, ii. 463
 urate, ii. 468
 how extracted from common salt, ii. 249
- Sodalite, iii. 337
- Sodium, i. 336
 alloys of, i. 342, 402, 436, 450, 460, 469, 491, 509, 525, 556
 peroxide of, i. 338
 chloride of, i. 339
 iodide of, i. 341
 phosphuret of, i. 342
 sulphuret of, *ibid.*
 arsenuret of, *ibid.*
 alloy with potassum, *ibid.*
 action of on ammonia, ii. 30
- Soil, iv. 316
- Solids, iii. 106
 table of, iii. 126
 expansion of, i. 34
 conversion into liquids, i. 46
 specific gravity of, iii. 127
- Solution, iii. 92
- Sommite, iii. 339
- Sorbic acid, ii. 191
- Sory, ii. 553
- Sowans, iv. 78
- Spar, ponderous, iii. 413
 fluor, iii. 403
 calcareous, iii. 387
 rhomb, iii. 394
- Sparry iron ore, iii. 484
- Spath perlé, iii. 397
- Specific caloric, i. 96
- Specificum purgans, ii. 435
- Spectrum, prismatic, i. 16
- Specular iron ore, iii. 476
- Spelter, i. 430
- Spermaceti, ii. 386
 oil, iv. 434
- Sphene, iii. 521
- Spherulite, iii. 308
- Sphragid, iii. 357
- Spiders' webs, iv. 486
 venom, iv. 533
- Spinell, iii. 252
- Spirit of nitre, ii. 82
 mindererus, ii. 419
 sal ammoniac, ii. 27
 salt, ii. 242, 521
 urine, ii. 27
 wine, ii. 318
- Spirit, proof, ii. 331
- Spirits, ardent, ii. 318
 rectified, *ibid.*
- Spleen, calculi from, iv. 562
- Spodumene, iii. 334
- Sponges, iv. 460
- Spreustein, iii. 336
- Squills, iv. 282
- Stalactite, siliceous, iii. 300
- Starch, iv. 69
 plants containing it, *ibid.*
 potatoe, iv. 77
- Starchy fibrous matter, iv. 264
- Staurolithe, iii. 286
- Staurotide, *ibid.*
- Steam, ii. 12
 elasticity of, i. 61
 latent heat of, i. 71
- Stearin, ii. 381
- Steatite, iii. 359
- Steel, i. 390
 natural, i. 394
 of cementation, i. 394
 cast, *ibid.*
 tempering, i. 385
- Steinheilite, iii. 267
- Stibium, i. 549
- Stilbite, iii. 318
- Stilpnosiderite, iii. 480
- Stinking sulphureous air, i. 289
- Stones, analysis of, iii. 592
 saline, iii. 381
 earthy, iii. 244
 from the atmosphere, iii. 184
 vitrifiable, i. 248

- Storax, iv. 159
 Strahlstein, iii. 368
 Strahlerz, iii. 469
 Strontian, i. 359, and ii. 69
 salts of, ii. 508
 acetate of, ii. 513
 arsenate, ii. 512
 arsenite, *ibid.*
 benzoate, ii. 513
 borate, ii. 510
 carbonate, ii. 509, and
 iii. 414
 chlorate, ii. 235
 chromate, ii. 512
 citrate, ii. 514
 fluante, ii. 282
 hydriodate, ii. 278
 hydrosulphuret, ii. 700
 hypophosphite, ii. 511
 hyposulphate, ii. 512
 hyposulphite, ii. 511
 iodate, ii. 274
 lactate, ii. 514
 malate, *ibid.*
 muriate, i. 362
 nitrate, ii. 508
 oxalate, ii. 513
 phosphate, ii. 510
 seleniate, ii. 512
 silicate, ii. 510
 succinate, ii. 513
 sulphate, ii. 511, and iii.
 414
 sulphuret, ii. 71
 tartrate, ii. 514
 urate, *ibid.*
 water, ii. 68
 Strontianite, iii. 414
 Strontites, i. 359
 Strontium, *ibid.*
 oxide of, i. 361
 chloride of, i. 362
 iodide of, i. 363
 Strychnia, ii. 42
 muriate of, ii. 253
 salts of, ii. 174
 Styraç, iv. 156
 Suber, iv. 178
 Suberic acid, ii. 157
 Suberin, iv. 179
 Subsalts, ii. 409
 Subsoap, ii. 368
 Succinates, ii. 151
 Succinic acid, ii. 150
 Sugar, iv. 81
 acid of, ii. 165
 alcohol from, iv. 376
 of Botany Bay, iv. 27
 cane, juice of, iv. 23
 liquid, *ibid.*
 of beet, iv. 23
 of diabetic urine, iv. 426
 of figs, iv. 24
 of grapes, *ibid.*
 of lead, ii. 615
 of milk, iv. 423
 of mushrooms, iv. 27
 of starch, iv. 25
 of Saturn, ii. 615
 plants containing, iv. 29
 Sulphate of cobalt, iii. 510
 indigo, iv. 85
 lead, iii. 539
 zinc, iii. 526
 Sulphates, ii. 117
 Sulphites, ii. 120
 Sulphochyazates, ii. 306
 Sulphochyazic acid, ii. 302
 Sulphur, i. 277, and iii. 427
 auratum, ii. 708
 chloride of, i. 287
 iodide of, i. 288
 flowers of, i. 278
 combustion of, i. 279
 in animals, iv. 439
 native, iii. 427
 Sulphureous salt of Stahl, ii. 437
 Sulphuret of ammonia, ii. 30
 antimony, i. 558
 arsenic, i. 311
 barytes, ii. 66
 bismuth, i. 478
 carbon, i. 293
 cerium, i. 425
 cobalt, i. 413
 copper, i. 468, and
 iii. 449
 gold, i. 508
 iron, i. 398
 lead, i. 449
 lime, ii. 61
 magnesia, ii. 69
 manganese, i. 421
 mercury, i. 488
 molybdenum, i. 566,
 and iii. 515

Sulphuret of nickel, i. 409
 palladium, i. 531
 phosphorus, i. 295
 platinum, i. 523
 potash, ii. 52
 selenium, i. 303
 silver, i. 499
 soda, ii. 54
 strontian, ii. 69
 tin, i. 457
 tungsten, i. 572
 uranium, i. 428
 zinc, i. 435

Sulphuretted chyzic acid, ii. 302
 hydrogen gas, i. 289
 alcohol, ii. 334

Sulphurets, table of, i. 549
 metallic, *ibid.*

Sulphuric acid, i. 281, and ii. 113
 phlogisticated, ii. 117
 in animals, iv. 440

Sulphurous acid, i. 279, and ii. 119

Sulphurs, i. 286

Sumach, iv. 229

Sun, nature of, i. 125
 a source of heat, *ibid.*
 heat of its rays, i. 127

Supersalts, ii. 409

Supersulphuret of iron, i. 399
 lead, i. 449

Supporters of combustion, i. 173

Surface of minerals, iii. 238

Surturbrand, iii. 431

Swamp ore, iii. 482

Sweat, iv. 530

Sweet-flag, iv. 225

Swimming bladders of fishes, air contained in, iv. 533

Sylvanite, i. 313

Syrup, iv. 18

T

Tablespar, ii. 482, and iii. 410

Tacamahac, iv. 134

Talc, iii. 363
 Venetian, *ibid.*

Talcite, iii. 363

Tallow, iv. 431
 mineral, ii. 397

Tamarinds, iv. 275

Tannin, ii. 207
 artificial, ii. 220
 natural, ii. 209
 species of, iv. 9

Tanning, iv. 470

Tantalite, iii. 504

Tantalum, i. 574
 alloys of, i. 577
 ores of, iii. 503
 oxide, i. 575

Tapioca, iv. 78

Tar, mineral, ii. 398

Tartar, ii. 446
 cream of, *ibid.*
 vitriolated, ii. 435
 chalybeated, ii. 565
 emetic, ii. 690
 regenerated, ii. 442
 of the teeth, iv. 508
 secret foliated earth of, ii. 442
 salt of, i. 328
 soluble, ii. 446

Tartaric acid, ii. 173
 plants containing it, iv. 4

Tartarin, i. 328

Tartrates, ii. 175

Tears, iv. 514

Tectum argenti, i. 475

Teeth, iv. 452

Telesia, iii. 254

Tellurates, ii. 687

Telluretted hydrogen gas, i. 315

Tellurium, i. 313
 chloride of, i. 315
 iodide, i. 315
 salts of, ii. 686
 muriate of, ii. 688
 native, iii. 441
 nitrate, ii. 688
 ores of, iii. 441, 614
 oxides of, i. 314
 seleniuret, i. 316
 sulphate, ii. 689
 sulphuret of, i. 316

Temperature at which bodies burn, i. 145

Tempering of steel, i. 385

Tenacity, iii. 242

Tendons, iv. 473

Tennantite, iii. 459

Terra ponderosa, i. 352

- Thallite, iii. 274
 Thermometer explained, i. 37
 differential, i. 77.
 varieties of, i. 38
 Thermoxygen, i. 137
 Thorina, i. 377, and ii. 79
 Thorinum, i. 377
 Thummerstone, iii. 276
 Thus, iv. 131
 Tile ore, iii. 463
 Tin, i. 452
 alloys, i. 459, 472, 493, 502,
 513, 526, 532, 557, 567,
 573
 salts of, ii. 624
 acetate, ii. 628
 arseniate, *ibid.*
 benzoate, ii. 629
 borate, ii. 626
 carbonate, ii. 626
 chlorites, i. 455
 fluat, ii. 284
 hydrosulphurets, ii. 707
 hyposulphite, ii. 628
 iodide, i. 456
 muriate, ii. 263
 nitrate, ii. 625
 ores, iii. 514
 ores, analysis of, iii. 609
 oxalate, ii. 629
 oxides of, i. 452
 phosphate, ii. 627
 phosphuret, i. 457
 pyrites, iii. 515
 seleniate, ii. 628
 seleniuret, i. 459
 stone, iii. 515
 succinate, ii. 629
 soap, ii. 717
 sulphate, ii. 627
 sulphite, ii. 628
 sulphuret, i. 457
 tartrate, ii. 629
 Tincal, ii. 454
 Tinfoil, i. 452 λ
 Tinning, i. 474
 Tinplate, i. 460
 Tinstone, iii. 515
 Titaneous iron ore, iii. 477
 Titanite, iii. 517, 521
 Titanium, i. 577
 brown ore of, iii. 521
 carbonate, ii. 694
 Titanium, muriate, ii. 693
 nitrate, *ibid.*
 ores of, iii. 517, 617
 oxides of, i. 579
 oxalate, ii. 694
 salts of, ii. 692
 silicate, iii. 521
 sulphate, ii. 693
 tartrate, *ibid.*
 Tobacco, iv. 241
 Tolu, balsam of, iv. 154
 Tombac, white, i. 469
 Topaz, iii. 260, 378
 occidental, *ibid.*
 oriental, iii. 253
 Topaz rock, iii. 563
 Saxon, iii. 260
 Tortoise-shell, iv. 462
 Touchstone, iii. 296
 Tourmaline, iii. 269
 Tourquois, iii. 259
 Tracheæ, iv. 332
 Tragacanth, iv. 64
 Train oil, iv. 432
 Transparency, iii. 241
 Transpiration of plants, iv. 338
 Traps, primitive, iii. 561
 floetz, iii. 575
 transition, iii. 568
 Tremella nostoc, iv. 297
 Tremolite, iii. 370
 Triphane, iii. 334
 Tripoli, iii. 344
 Trona, ii. 452
 Truffles, iv. 295
 Tube of safety, ii. 82
 Tungstate of lime, iii. 411
 Tungstates, ii. 132
 Tungsten, i. 569, and iii. 411
 alloys, i. 573
 ores of, iii. 616
 oxides of, i. 570
 Tungstic acid, i. 571, and ii. 132
 Turmeric, iv. 223
 Turpentine, iv. 131
 Turpeth mineral, ii. 660
 Type metal, i. 557
- U
- Ulcers in plants, iv. 301
 Ulmin, iv. 45
 Umber, iii. 483
 Uran mica, iii. 499

- Uran ochre, iii. 499
 Uranitic ochre, *ibid.*
 Uranium, i. 425
 ores of, iii. 498, 617
 oxides of, i. 427
 salts of, ii. 588
 acetate, ii. 591
 arsenate, *ibid.*
 muriate, ii. 261
 nitrate, ii. 588
 oxides of, i. 427
 phosphate, ii. 591
 seleniate, ii. 591
 sulphate, ii. 590
 sulphuret of, i. 428
 tartrate, ii. 592
 Urates, ii. 186
 Urea, iv. 419
 Uric acid, ii. 184
 sublimate from, ii. 186
 Urine, iv. 539
 fusible salt of, ii. 457
 spirit of, ii. 27
 changes produced on, by
 diseases, iv. 547
 of the ass, iv. 549
 camel, iv. 550
 castor, iv. 552
 cow, iv. 549
 fowls, iv. 552
 guinea-pig, iv. 551
 horse, iv. 548
 insects, iv. 552
 lion, iv. 551
 rabbit, *ibid.*
 sow, *ibid.*
 tiger, *ibid.*
 saponaceous extract of, iv.
 419
 Urinary calculi, iv. 568
 of inferior animals, iv. 580

 V
 Valerian, iv. 223
 Vapour, explained, i. 57
 elasticity of, i. 61
 state of, in the atmo-
 sphere, iii. 175
 Vapours, nature of, iii. 27
 specific gravity of, i. 67
 Varec, i. 336
 Variegated copper ore, iii. 457
 Variolite, iii. 327
 Varnish of silk, iv. 484
 Varnishing, iv. 138
 Vegetable substances, iv. 1
 Vegetables, *ibid.*
 ingredients of, iv. 2
 decomposition of, iv.
 362
 gas from, i. 240
 Vegetation, iv. 303
 Veins, iii. 580
 Venus, ii. 630
 salts of, *ibid.*
 Verdigris, ii. 643
 Verjuice, iv. 25
 Vermillion, i. 489
 Vessels of plants, iv. 331
 Vesuvian, iii. 277, 279
 Vinegar, ii. 136
 distilled, ii. 138
 radical, ii. 136
 of Venus, ii. 138
 Viper, poison of, iv. 531
 Vital air, i. 177
 Vitreous copper ore, iii. 456
 electricity, i. 164
 humour of the eye, iv.
 520
 silver ore, iii. 449
 Vitriol, green, ii. 553
 blue, ii. 634
 white, ii. 596
 Vitriolic acid, ii. 111
 Volatile and volatilization, i. 278,
 and iii. 155
 liniment, ii. 715
 Volcanite, iii. 376
 Volta's eudiometer, iii. 170
 Vulpinite, iii. 407

 W
 Wacka, iii. 380
 Walton mineral, ii. 88
 Wash, iv. 374
 Wasp, venom of, iv. 533
 Wasserblei, iii. 513
 Water, i. 218, and ii. 12
 boiling point of, as affect-
 ed by salts, ii. 13
 as the food of plants, iv.
 312
 composition of, i. 218, and
 ii. 17

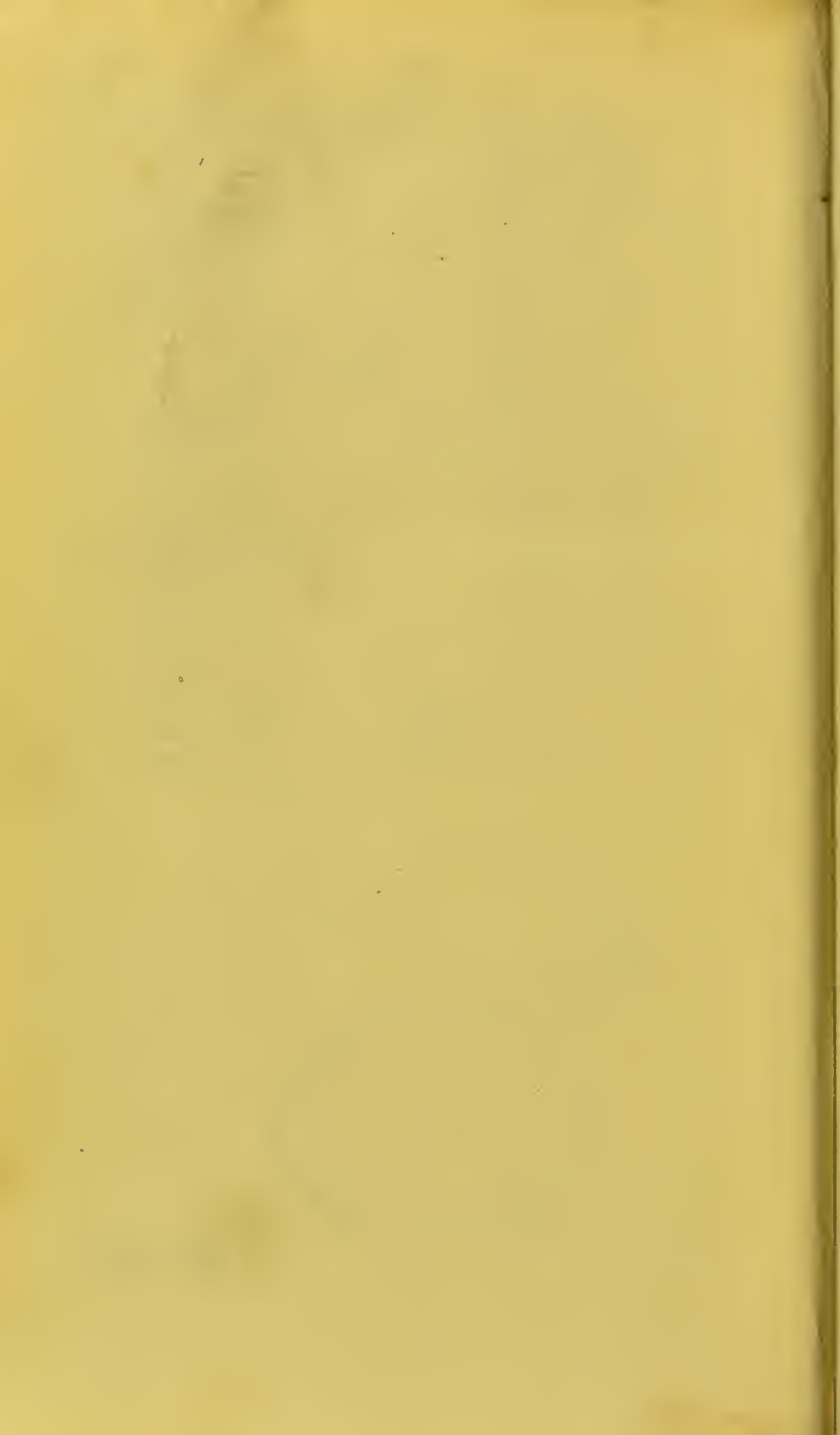
- Water, decomposed by iron, i. 383
 of nitre, ii. 82
 expansion by cold, i. 40
 sea, iii. 195
 specific gravity of, iii. 200
 in the atmosphere, iii. 174
- Waters, iii. 190
 acidulous, iii. 212
 analysis of, iii. 216
 chalybeate, iii. 213
 common, iii. 191
 hepatic, iii. 213
 mineral, iii. 207
 table of, iii. 215
 rain, iii. 191
 river, iii. 193
 saline, iii. 213
 snow, iii. 192
 spring, iii. 192
- Wavellite, iii. 420
- Wax, iv. 103
 Brazil, iv. 108
 myrtle, iv. 107
 punic, iv. 105
 from rose and lavender oil,
 iv. 108
- Wedgewood's pyrometer, ii. 75
- Weiserz, iii. 549
- Wernerite, iii. 336
- Wheat flour, iv. 256
- Whet slate, iii. 347
- Whey, iv. 500
- Whisky, ii. 318
- White cobalt ore, iii. 505
 copper ore, iii. 458
 flowers, iv. 246
 gold ore, iii. 442
 lead, i. 445
 lead ore, iii. 534
 ore of antimony, iii. 546
 silver ore, iii. 450
 tombac, i. 469
 of eggs, iv. 504
- Willow, white, bark of, iv. 236
- Wine, iv. 379
 essential salt of, ii. 442
 its fermentation, iv. 380
 component parts, iv. 381
- Wines, table of their component
 parts, iv. 383
- Wire gauze screen, i. 146
- Witherite, iii. 412
- Woad, iv. 243
- Wodan pyrites, i. 414
- Wodanium, *ibid.*
- Wolfram, i. 569, and iii. 492
- Wolframium, i. 570
- Wood, iv. 180, 217, 225
 opal, iii. 303
 rock, iii. 365
 tin, iii. 516
- Woods, conducting power of, i. 89
- Woodstone, iii. 295
- Wool, iv. 483
- Wort, iv. 369
- Wormwood, iv. 240
- Y
- Yanolite, iii. 276
- Yeast, iv. 371
- Yenite, iii. 490
- Yellow earth, iii. 354
 cobalt, ochre, iii. 509
 acid, iv. 466
 flowers, iv. 246
 ore, iii. 442
- Yolk of egg, iv. 504
- Yttria, i. 368, and ii. 73
 salts of, ii. 528
 acetate of, ii. 531
 arseniate, ii. 536
 carbonate, iv. 529
 chromate, ii. 530
 fluuate, ii. 282
 hydrosulphuret, ii. 700
 muriate, ii. 254
 nitrate, ii. 529
 oxalate, ii. 531
 phosphate, ii. 529
 seleniate, ii. 530
 succinate, ii. 531
 sulphate, ii. 530
 tartrate, ii. 531
- Yttrium, i. 368
- Yttrocerite, iii. 502
- Yttrotantalite, iii. 504
- Z
- Zeolite, iii. 314
 foliated, iii. 318
- Zimome, iv. 92
- Zinc, i. 429

- Zinc, alloys, i. 436, 451, 460, 470,
492, 501, 511, 525, 556,
567
salts of, ii. 592
acetate, ii. 598
antimoniate, *ibid.*
arseniate, ii. 598
benzoate, ii. 599
borate, ii. 594
butter, i. 433
carbonate, ii. 593
 anhydrous, iii. 525
 hydrous, iii. 526
chlorate, ii. 236
chloride, i. 433
chromate, ii. 598
citrate, ii. 599
flowers of, i. 432
fluante, ii. 284
hydriodate, ii. 278
hydrosulphuret, ii. 707
hyposulphite, ii. 597
iodate, ii. 274
iodide, i. 434
lactate, ii. 599
malate, *ibid.*
molybdate, ii. 598
muriate, ii. 261
nitrate, ii. 592
ores of, iii. 521
ores, analysis of, iii. 612
- Zinc, oxalate, ii. 599
oxide of, i. 432
phosphate, ii. 594
phosphuret, i. 434
red ore, iii. 515
seleniate, ii. 597
silicate, ii. 594
soap of, ii. 717
succinate, ii. 599
sulphate, ii. 595
sulphite, ii. 597
sulphuret, i. 435
tartrate, ii. 599
tungstate, ii. 598
- Zircon, i. 375, and iii. 248
- Zirconia, i. 376, and ii. 78, 545
acetate, ii. 547
borate, ii. 546
carbonate, ii. 546
fluante, ii. 283
muriate, ii. 255
nitrate, ii. 546
phosphate, ii. 547
seleniate, *ibid.*
sulphate, *ibid.*
sulphite, *ibid.*
- Zirconium, i. 375
- Zoisite, iii. 274
- Zoophites, iv. 457
- Zumic acid, ii. 202

THE END.











COMPLIMENT (11)

TITLE PAGE

